1819-108-C1-ExamPartTwo

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We consider information measures from statistical information theory in the next section.

3.2. Information and Entropy

The good news: Computers allow us to work 100% faster. The bad news: They generate 300% more work.

The term entropy and its denotation by the letter S were introduced by Rudolf Julius Emanuel Clausius (1822–1888) in 1864 in his work on the foundations of classical thermodynamics, where he formulated the second law of thermodynamics in the form that the entropy of a closed system cannot decrease. Later the concept of thermodynamic entropy was clarified and propuded by Lydrig Edward Polymann (1844–1804). was clarified and grounded by Ludwig Eduard Boltzmann (1844-1906). The famous formula for thermodynamic entropy is

$$S = k \cdot \ln W \tag{3.2.1}$$

where W is the thermodynamic probability that the system is in the state

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The answer of Hartley (1928) to this question is based on the number n of possible outcomes of the experiment H or on the number n of

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    \documentclass[10pt]{article}
\usepackage[utf8]{inputenc}
\usepackage{geometry}
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%\usepackage{hyperref}
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\usepackage{ragged2e}
\usepackage{rotating}
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\date{}
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  \textit{Theory of Information}\hfill\lap{}\par
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\begin{flushright}\textit{The good news: Computers allows
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 $\label{thm:condition} $$ \operatorname{W}$ \int_{\Omega_{0}} \mathbb{S}^{(3.2.1)} \operatorname{where } W \text{ is the thermodynamic probability that the system is in the state with the entropy } S.$

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