



Redox dynamics of sulphur with Ni/GDC anode during SOFC operation at mid- and low-range temperatures: An *operando* S K-edge XANES study



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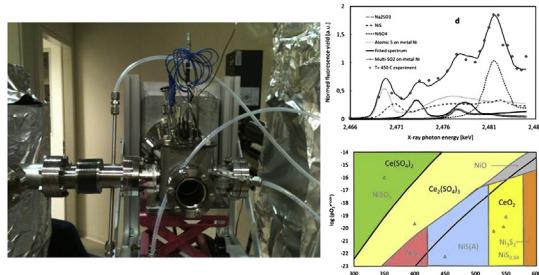
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HIGHLIGHTS

- An *operando* SOFC anode S-poisoning XAS experiment at 550 °C to 250 °C was performed.
- S K-edge XANES spectral information at Ni/GDC outer surface was collected.
- Intermediates with different sulphur oxidation states (6+, 4+, 0, 2-) were observed.
- Proportion of oxidation states changed as a function of temperature.
- Differences between TD calculations and XAS information were observed and discussed.

GRAPHICAL ABSTRACT



ARTICLE INFO

Article history:

Received 12 July 2012

Received in revised form

12 February 2013

Accepted 28 March 2013

Available online 18 April 2013

Keywords:

Sulphur poisoning

Ni-GDC

SOFC

XANES

S K-edge

In situ

ABSTRACT

Sulphur poisoning of nickel-based solid oxide fuel cell (SOFC) anode catalysts is a well-documented shortcoming, but not yet fully understood. Here, a novel experiment is demonstrated to obtain spectroscopic information at *operando* conditions, in particular the molecular structure of sulphur species in the sulphur K-shell X-ray absorption near edge structure (XANES) region for a SOFC anode under realistic *operando* conditions, thus, with the flux of O²⁻ from cathode to anode. Cooling from $T = 550\text{ }^{\circ}\text{C}$ stepwise down to 250 °C, 5 ppm H₂S/H₂ reacting with Ni-gadolinium doped ceria (GDC) anode resulted in several sulphur species in different oxidation states (6+, 4+, 0, -2) and in amounts being at a minimum at high temperature. According to sulphur speciation analysis, the species could either relate to SO_4^{2-} or $\text{SO}_3(\text{g})$, $\text{SO}_2(\text{g})$, $\text{S}_2(\text{g})$ or surface-adsorbed S atoms, and Ni or Ce sulphides, respectively. The coexistence of different sulphur oxidation states as a function of temperature was analysed in the context of thermodynamic equilibrium calculations. Deviations between experimental results and calculations are most likely due to limitations in the speed of some intermediate oxidation steps as well as due to differences between stoichiometric CeO_2 used in calculations and partially reduced $\text{Ce}_{0.9}\text{Gd}_{0.1}\text{O}_{2-\delta}$.

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1. Introduction

Solid oxide fuel cells (SOFC) are promising energy converters with high electrical efficiency (from 45 to 65% based on lower heating value), fuel flexibility (can oxidise virtually any fuel from hydrogen to hydrocarbons up to carbon) and environmental friendliness. To reduce thermal degradation of system components, system costs and radiative heat transfer and to increase the speed of start-up and shut-down cycles, attempts have been made to decrease the operating temperature of SOFC from high temperature at $T = 900\text{ }^{\circ}\text{C}$ to intermediate temperature (IT) between $T = 500\text{--}750\text{ }^{\circ}\text{C}$ [1,2] and even to temperatures as low as $T = 350\text{ }^{\circ}\text{C}$ [3,4]. In addition to systems doped with noble metals such as Pt and Pd, until now the best electrocatalytic performance for IT-SOFC application was achieved with Ni-based catalysts. But Ni-based SOFC anode catalysts have some weaknesses when operating at intermediate temperature region such as coke formation (when operating on hydrocarbon fuels) and intolerance against some inorganic impurities [5–8]. Particular attention has been paid to sulphur poisoning, because most fuels contain sulphur and the complete removal of it is a costly process. The majority of sulphur poisoning studies have been carried out on Ni–YSZ (yttrium stabilized zirconia) SOFC anode at elevated working temperatures ($T = 700\text{--}900\text{ }^{\circ}\text{C}$) [9–18]. Only few studies examine sulphur poisoning of Ni-gadolinium doped ceria (GDC) anodes at intermediate temperature region [13,19–23] and the influence of chemical composition of ceramic phase on the poisoning process of Ni-cermet electrodes [24–26]. In good agreement with thermodynamics [9,20] and calculations based on density functional theory (DFT), [27] it has been demonstrated in several electrochemical studies that the extent of sulphur poisoning is a function of temperature, oxygen partial pressure and sulphur concentration as well as chemical nature of fuel and anode material [28–30]. Several studies have demonstrated that there is an initial reversible (or partially reversible) degradation step during the very first minutes when the anode is exposed to H_2S [9,12,31], followed by a slow, but continuous, non-reversible decrease of electrode performance [9,12,31]. First, rapid decrease of electrochemical performance is likely caused by adsorption of sulphur species at electrode and concomitant blocking of the so-called three-phase-boundary (TPB) for hydrogen oxidation [9,12]. There is some consensus about the first step of poisoning, but there are several significantly different if not controversial views about the second, slower step. Some authors are supporting the idea that slow degradation is caused by the formation of volatile Ni_xS_y phase, which causes restructuring and degradation of electrode structure [14,15,27]. Sasaki et al. proposed that blocking of hydrogen diffusion to the TPB caused by sulphur adsorption leads to the oxidation of Ni at TPB interface, causing depletion of active sites [9]. Slow decline in performance might also be caused by adsorption of sulphur on nickel surface grains which are more difficult to access, such as pores with bottlenecks [14], or, because of surface reconstruction or recrystallization of nickel grains from catalytically more active into catalytically less active crystallographic form [14], or, because of bulk phase diffusion of sulphur penetrating into Ni grains [14]. One possible reason of slow degradation could be the formation of $\text{Ce}_2\text{O}_2\text{S}$ as well as $\text{Ce}(\text{SO}_4)_2$ at relatively low H_2S concentrations in reaction with non-stoichiometric CeO_n ($n < 2$), because reduced or doped ceria exhibits significantly larger affinity for H_2S compared to CeO_2 used in theoretical calculations [20,32]. Thus, while there is some information about changes in electrical performance and degradation of electrode structure caused by sulphur at different conditions, there is no comprehensive explanation of the physico-chemical mechanisms which ultimately constitute sulphur poisoning.

1.1. Motivation and goal

Different studies have proposed and supported experimentally the idea that the release of sulphur from SOFC electrode might take place through the oxidation of sulphur [10,18,22,23,29], which premises the presence of oxidised particles or intermediates at the anode surface. But, up to now, spectroscopic measurements with X-ray photoelectron spectroscopy (XPS), X-ray absorption spectroscopy (XAS), or, Raman microspectroscopy either obtain information from SOFC anode surface in the presence of H_2S , H_2 , and, possibly even heat, or *ex-situ* and under *post-mortem* conditions [14,15,22,33]. The molecular structure of sulphur compounds has been studied with sulphur K-shell edge X-ray absorption spectroscopy in various fields of application, particularly in relation to fossil fuels, such as sulphur in coal [34], sulphur in residual oil fly ash [35] and heterogeneous catalysis [36]. High temperature oxidation and reduction studies [34] and catalysis studies [36] have also been made recently with *in-situ* or *operando* sulphur XAS, for example. Sulphur as the ligand ion is also of interest in connection with metal ions [37] and could in future studies open up new opportunities for the understanding and quantification of electronic transport processes in SOFC anodes, where newly formed Ni–S compounds are subject to exchange interactions, in analogy to 3d metal oxides at SOFC cathodes, for example [38,39]. Latter would constitute an extension of sulphur XAS from molecular structure and chemical speciation of sulphur motifs towards electronic structure, valence band and electronic transport properties of aged SOFC anodes. However, so far, there are no spectroscopic experiments made *operando* under realistic SOFC anode operation conditions with oxide ion flux from cathode to anode or under electrochemical polarization of the anode. And, when it comes to study of the nickel–sulphur chemistry in detail, we realize that Ni K-edge XAS is less informative than S K-edge XAS [40], particularly because the sulphur K-edge spectra provide a very detailed picture about the molecular structure of the various sulphur species being formed on the anode during typical SOFC operation. Therefore, the present study demonstrates a novel powerful approach with sulphur-sensitive XAS under *operando* SOFC conditions for studying processes during sulphur poisoning of an IT-SOFC dedicated electrochemical element.

2. Experiments and methods

2.1. Preparation of SOFC single cell assembly

The SOFC single cell assemblies tested in this study were typical electrolyte (GDC) supported planar fuel cells with rectangular shape, which were made befitting with the XAS cell compartments. Supporting electrolyte pellets were made from $\text{Ce}_{0.9}\text{Gd}_{0.1}\text{O}_{2-\delta}$ (GDC) powder with $S_{\text{BET}} = 19.9\text{ m}^2\text{ g}^{-1}$ (NexTech Materials) pressed at 20 MPa cm^{-2} and sintered at 1773 K for 10 h . The electrolyte supported cell configuration was used because of better compatibility with the gas sealing concept used. The anode layer was screen printed onto the sintered electrolyte pellet by applying commercial NiO-GDC ink (Fuel Cell Materials) through 200 mesh screen and fired at $1350\text{ }^{\circ}\text{C}$ for 3 h with heating and cooling rates of 2.5 K min^{-1} . The choice of sintering temperature was based on the results of our previous studies [41]. Cathode powder was prepared by nitrate solution thermal combustion method using $\text{La}(\text{N-O}_3)_3 \cdot 6\text{H}_2\text{O}$, $\text{Sr}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ (both from Aldrich, 99.9%) and $\text{Co}(\text{N-O}_3)_2 \cdot 6\text{H}_2\text{O}$ (98%, Riedel de Haen) as precursors for $\text{La}_{0.6}\text{Sr}_{0.4}\text{CoO}_{3-\delta}$ cathode powder and glycine (99.9%, Sigma–Aldrich) as a reducing agent. A detailed description of the experimental procedure can be found from Lust et al. [41]. Finally the cathode paste was

deposited onto a GDC electrolyte pellet by screen printing followed by 5 h sintering at 1110 °C with heating and cooling rates of 2.5 K min⁻¹ [42].

2.2. Design and electrochemical testing of operando SOFC-XAS cell

An *in situ* reaction cell (for use with hard X-rays), which had been constructed at Paul Scherrer Institute, followed the design of Grunwaldt (KIT, Germany) [43], was adapted by us into a SOFC-dedicated XAS cell suitable for performing X-ray absorption experiments (Fig. 1) in the fluorescence mode of detection at the S K-edge ($E \sim 2.5$ keV). For this purpose, the cell was furnished with two gas manifolds (2 and 3 in Fig. 1) to supply the anodic and cathodic sides of the IT-SOFC pellet with H₂ and synthetic air, respectively. Windows (1 in Fig. 1) being permeable to soft X-rays were installed. Cell components (7 in Fig. 1) in direct contact with the GDC electrolyte pellet were made from Crofer 22 APU steel to achieve compatible thermal expansion coefficient. Electrical contact was taken with Pt wire and paste (MaTecK). Electrical sealing of steel cell compartments (7 in Fig. 1) from current collectors (4 and 5 in Fig. 1) was achieved using alumina tubing and glass paste (Schott). X-ray spectroscopy cell was mounted on the heating element (6 in Fig. 1) with temperature precision of 0.1 °C. A security cell with Kapton as X-ray windows was mounted over the SOFC-XAS cell to protect the measurement chamber against sulphur contamination. Amel 7050 Potentiostat with Junior Assist software for data acquisition were used for electrochemical tests. Gas flows were controlled with digital mass flow controllers (Vögtlin Instruments AG, Switzerland). High purity gases ($\geq 99.995\%$) were used for testing.

2.3. Sulphur K-edge spectra

X-ray absorption near edge structure (XANES) spectra were recorded at the Phoenix beamline of the Swiss Light Source (SLS) at PSI, Switzerland. The source of the X-rays is an elliptical undulator (APPLE II). The beamline optics consist of a focussing mirror (radius ~ 1.4 km), a filter to suppress higher harmonics contributions, a double-crystal monochromator and a Kirkpatrick–Baez mirror system to focus the beam. For the experiments in this study we used the unfocussed beam with a spot size of about 1 × 1 mm. A double-crystal monochromator equipped with a silicon [111] crystal, and available Bragg angles ranging from 14 to 75° allow generation monochromatic light in the energy range from 2050 to 8000 eV. The end station of the Phoenix beamline is specially designed to host *in situ* and *operando* experiments. For the experiments the standard end station was removed and replaced by the aforementioned custom-made *operando* XAS cell. The cell was surrounded by a safety compartment with Kapton as X-ray transmitting windows and purged constantly with He during operation.

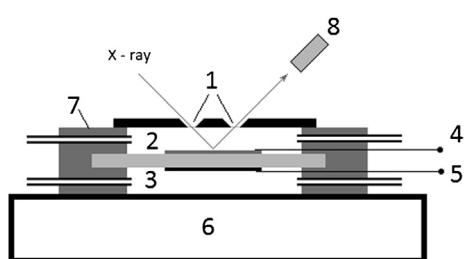


Fig. 1. X-ray absorption spectroscopy cell for *operando* measurements of IT-SOFC: (1) windows for soft X-ray; (2) anode gas compartment; (3) cathode gas compartment; (4) current collector for anode; (5) current collector for cathode; (6) heating element; (7) cell body with tubing; and (8) fluorescence detector.

The spectra were collected in fluorescence mode using a dispersive Si-drift chamber solid-state detector. The beam energy was calibrated to the S K-edge (2.472 keV). Spectra were recorded at the S K-edge within energy range from 2.40 to 3.20 keV, while scanning monochromator and the gap of the undulator synchronously. With the *operando* XANES spectra, we only recorded a few energy positions around noticeable white-line peaks to obtain discernible peak features between $E = 2.466$ –2.485 keV (comprising all known peak positions) within reasonable data accumulation times (30 min at low T and 60–120 min at high T).

2.4. Sulphur speciation with the *operando*-recorded XANES spectra

This study focusses on the molecular speciation of sulphur motifs using XANES spectra recorded at the S K-edge under virtually complete steady state conditions at different SOFC temperatures. The establishment of steady state-like conditions was checked by monitoring the integrated fluorescence signal intensity between $E = 2.466$ –2.485 keV as a function of run time. Depending on the SOFC temperature under investigation, the establishment took several hours at start up to 15–30 min at lower temperatures. S K-edge X-ray absorption near edge structure spectroscopy (XANES) at the sulphur K-edge can be used to identify and quantify sulphur atoms exhibiting different oxidation states and molecular structure and those being part of different functional groups [44]. A powerful statistical tool facilitating the speciation analysis is principal component analysis (PCA), allowing estimating the minimal number of independent signal patterns (“principal components”, PCs) needed to reproduce a set of data, here the in-total eight *operando*-recorded S K-edge XANES spectra. The combined principal component analysis–target transformation–linear combination fitting (PCA–TT–LCF) approach has been applied and described by different authors (cf. [40,45]). The goal of PCA is to find a transformation matrix allowing reproducing a set of data by a minimal number of relevant, independent PCs using Eigenvalue analysis. The higher the eigenvalues, the more important are the associated PCs for reproducing the data set. Importantly, PCA does not relate to any compound spectra nor does it need any other input than a set of non-identical data files. The assignment of a sulphur compound (or sulphur form-representing) spectrum matching best with the characteristics of a given PC is done in a separate step, called target transformation (TT). The PC-representing spectra are then used to determine their amounts with each *operando*-recorded spectrum by means of least squares sum-based linear combination fitting (LCF). Here, PCA was performed with StatistiXL [46] and with WINXAS version 2.3 [47]. The best-suited sulphur compound spectra (“references”) were selected from a 12-fold database using WINXAS using TT. LCF was done using a hand-written program for commercial software (Berkeley Madonna from Kagi Shareware, Berkeley, CA, USA). For these purposes, the fluorescence signal with each *operando* spectrum was divided (normalised) by the value recorded at $E = 2.4845$ keV, thus, clearly after the highest-known white-line peak-energy position from sulphates [44]. Each normalised *operando* spectrum was then fitted between $E = 2.466$ and 2.485 keV to normalised reference spectra, so that their fitted amounts reflect fractions adding ideally up to unity. The peak-energy position of the sulphate-representing reference spectrum (NiSO₄) was also adjusted (fitted) in view of its disproportionately high impact on the accuracy with the other, minor, sulphur constituents even in case of a slight mismatch.

2.5. Thermodynamic phase-predominance diagrams

To understand the changes in the sulphur oxidation states with the Ni-GDC electrode, we have to take into account thermodynamic

as well as kinetic arguments. To estimate thermodynamically-defined phase transition conditions for the studied system, phase-predominance diagrams were calculated with the HSC software (Outotec Research Oy; Finland) for Ni–O–S and Ce–O–S, respectively, as functions of the partial pressure of O₂ at the anode, $\log[p\text{O}_2^{\text{anode}}]$, temperature, T, and, at the sulphur partial pressure used in XAS experiment (5 ppm H₂S in H₂). The sulphur partial pressure was introduced with the software as S₂ with $p\text{S}_2 = 1/2p\text{H}_2\text{S} = 2.5 \times 10^{-6}$, because phase-predominance calculations with HSC allows only usage of three different elements. Absence of H₂ in the calculations might introduce non-practise-real conditions compared with experiment. The oxygen partial pressure at the anode, characteristic with the *operando*-recorded XANES spectra, was calculated according to the Nernst equation using experimental temperature, cell voltage, and $p\text{O}_2^{\text{cathode}} = 0.2$ as input.

3. Results and discussion

3.1. Preliminary electrochemical tests and microstructural characterization

Preliminary electrochemical tests were carried out using the *operando* XAS electrochemical cell. Cyclic voltammograms characteristic to SOFC were measured and sulphur poisoning kinetics at constant current conditions and at different H₂S concentrations were recorded (Fig. 2). The current density at 600 °C with supply of hydrogen and synthetic air to the anode and cathode, respectively, at a cell potential 0.6 V was 50 mA cm⁻². Relatively low current densities of the studied cell were mainly caused by ohmic losses in supporting electrolyte with relatively high thickness (670 µm). The open circuit voltage (OCV) at conditions mentioned above was 0.95 V, which is slightly below theoretical value (1.138 V), but typical for ceria-based electrolytes. The OCV value did not depend on gas flow rate, which confirms that there was no gas leakage between cathode and anode compartment and potential drop was caused by typical slight electronic current leakage through the GDC membrane. It was electrochemically verified that SOFC single cell was gas-sealed and properly electrically connected. Preliminary sulphur poisoning experiments were carried out at three different H₂S concentrations – 0.25 ppm, 2.5 ppm, and, 5 ppm of H₂S in H₂. The speed of the first rapid poisoning step, which is most likely caused by blocking of TPB, increased

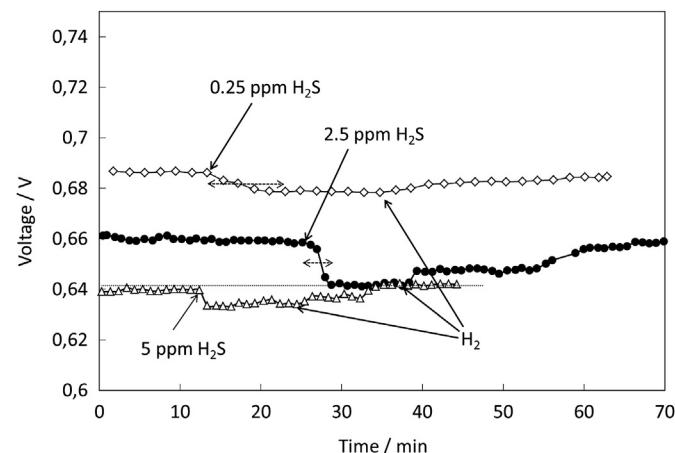


Fig. 2. Dependence of cell voltage on time during sulphur poisoning and regeneration process performed at $T = 600$ °C and at constant current condition (50 mA cm^{-2}) with H₂S concentrations of 0.25, 2.5 and 5 ppm in hydrogen (as indicated in the figure). Dotted horizontal arrows indicate durability of first poisoning step.

with the increase of H₂S concentration (Fig. 2). The recovery of electrochemical performance after closing of H₂S for 20–30 min was approximately 90%. A smaller drop of the potential and even better electrochemical performance after recovering in short range was achieved after closing of 5 ppm of H₂S in hydrogen stream. Similar behaviour of Ni/GDC electrode at low concentrations of H₂S and after short contamination with H₂S was documented in a comparative study of Ni/YSZ and Ni/GDC by Zhang et al. [21]. This phenomenon was most likely caused by small structural modifications of GDC close to TPB because of interaction with sulphur. In the case of Ni–YSZ, this effect was not observed. Comparison of micrographs for GDC operated in H₂ and H₂S in Fig. 3 shows microstructural differences expressed as etched interfaces of nickel and GDC in the TPB area after 20 h poisoning with 5 ppm of H₂S. We have highlighted three locations in Fig. 3B by the white boxes and seven locations with arrows where we suspect that the etched interfaces originate from the reaction of Ni with sulphur. Results of preliminary electrochemical experiments confirm that the SOFC single cell installed with the *operando* XAS cell was

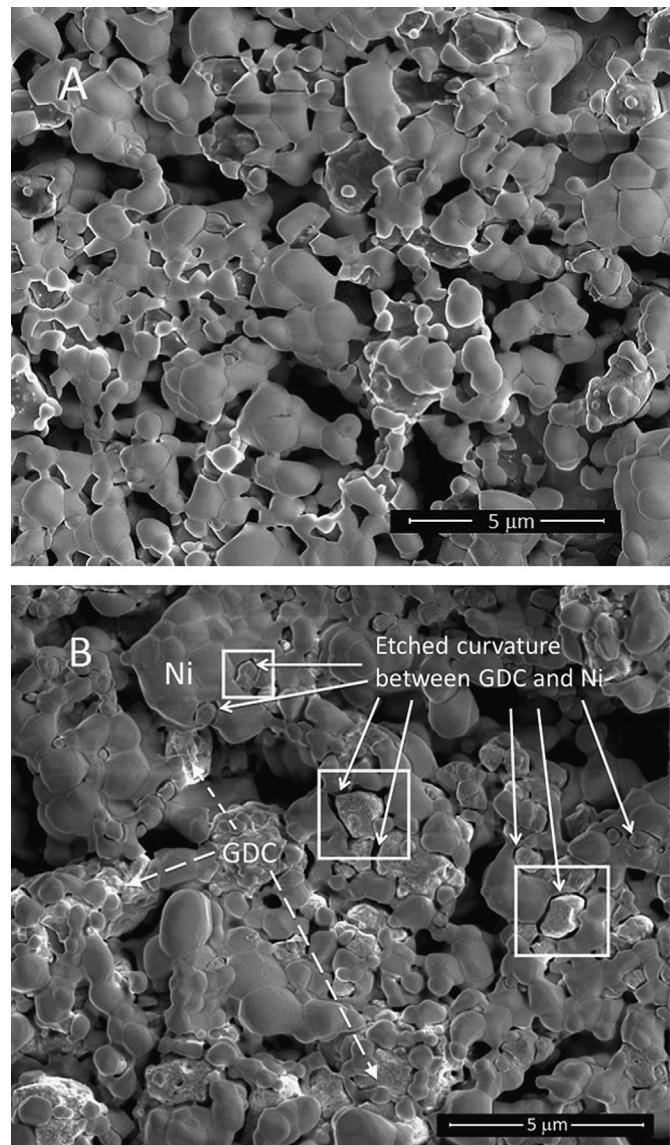


Fig. 3. Post reaction SEM micrograph of Ni-GDC anode after 20 h electrochemical tests at 600 °C: (a) with hydrogen as a fuel and (b) with hydrogen and 5 ppm H₂S as a fuel. Magnification of picture is 7500 times (5 µm scale bar in the images).

indeed working as a fuel cell. Results of sulphur poisoning experiments are in good accordance with previous studies and verify that H₂S interacts with Ni-GDC catalyst.

3.2. Preliminary operando XANES experiments at the S K-edge

Spectroscopic measurements were started from higher temperatures at $T = 600$ °C with direction to lower temperatures and from lower H₂S concentrations (0.25 ppm) to higher (5 ppm) H₂S concentrations. With lower H₂S concentrations at $T = 600$ °C, the fluorescence signal from the anode surface was too weak. Also the 5 ppm H₂S/H₂ gas mixture itself was not detectable with S K-edge XAS. But, in line with the present lower detection level of the beamline end station, the first detectable signal was recorded with 5 ppm H₂S/H₂ and at temperature of the studied electrode $T = 550$ °C, as soon as new sulphur species were formed in concentrations resembling that of (sub)monolayer levels with the anode surface. Spectroscopic data were then collected (Fig. 4), starting from $T = 550$ °C with 50 K-steps down to 350 °C. After the scan at $T = 350$ °C, temperature was raised again without interruption to $T = 550$ °C for making the final scan. It was found that the intensity of fluorescence signal increased enormously at lower SOFC temperatures and it revealed significant changes in the shape of the XANES spectra as a function of the irradiated photon energy (Fig. 4). At lower temperatures, the total amount of sulphur compounds residing at (or leaving) the anode surface increased as the absolute intensity of signal increased. This fact is also in good accordance with density functional theory calculations by Wang et al. [27] regarding sulphur surface coverage levels versus bulk sulphur compound formation as a function of temperature and H₂S/H₂ ratio for the Ni–S system, and with the common knowledge on the lifetime of particles at a surface as a function of temperature.

3.3. Relevant number of sulphur constituents with the operando spectra

XANES spectra with sulphur form-specific compounds often show a marked first peak ("white-line") at low energy, which is followed by several peak-like features within a modest energy interval (50–100 eV) thereafter. Helpful with a pure sulphur compound or phase is that the white-line energy position increases systematically with increasing sulphur oxidation state (cf. discussion and references in Ref. [40]). But, for mixtures, the "after peak" and white-line positions induced by different sulphur forms may

coincide (overlap) with each other. The sulphur compound spectra of importance with the present study are shown with Fig. 5a–h. It is then not possible to estimate the number of different forms just by counting the number of visible peaks with the *operando* spectra. Using a purely mathematical analysis with all (8) *operando*-recorded XANES spectra simultaneously (thus without using any sulphur compound spectrum as reference), principal component analysis (PCA) with StatistiXL [46] indicated that one component (PC-1) was particularly prominent, as it covered (accounted for) about 93% of data variance (Table 1), but it was not very revealing on the possible number of other (minor) principal components, although the *operando* spectra revealed, at least, three different peaks at about 2470, 2478, and, 2482 eV (see arrows in Fig. 4), resembling closely that of atomically adsorbed elemental sulphur ($\sim S^0$), S^{4+} (as SO₂ and/or SO₃²⁻), and, S^{6+} (as SO₃ and/or SO₄²⁻), respectively (cf. [40] and references cited). Instead, performing PCA with WINXAS [47] was inconclusive, because none of the calculated statistical indicators passed through a minimum value for increasing number of PCs assumed. This outcome can have simple causes, such as too high noise amplitude levels with the spectra (here, $\pm 10\%$ at best). But, whereas PCA is often found suited for data files in which each file (here, each *operando*-recorded spectrum) comprises different fingerprints in varying intensities, PCA fails e.g. when one file exhibits a unique fingerprint not found with the other data files. This option is followed further in the next section.

3.4. Selection of form-representative S K-edge compound spectra

The assignment of suited sulphur form-representative compound spectra ("references") from a 12-fold, sulphur form-specific, spectral database was done by Target Transformation (TT) using the WINXAS software [47]. Some reference spectra had been measured in previous studies, other compound spectra had been digitised from literature, but corrected and normalised here [40]. The spectral database comprised inorganic (solid NiS, Ni₂S, Ni₃S₂, Na₂SO₃, NiSO₄), molecular (S₁, S₂ gases), elemental (S⁰), polymeric (S₈), and aromatic (thiophene) sulphur, and also that of S atoms and SO₂ molecules adsorbed at Ni⁰ surfaces. With our database, the emphasis lies with ionically bonded sulphur and less with sulphur in covalent bonds (often in connection with carbon atoms), because the presence of C in the form of CO had been excluded in the present SOFC experiments for simplicity. Most of the compounds are actually also well-known from other Ni-catalysed heterogeneous processes involving natural (biomass) product gases as feedstock. Using TT, we selected five reference spectra from our database. In decreasing order of importance, the TT procedure indicated NiSO₄ ≫ NiS ~ Molecular S₂ ≫ Na₂SO₃ ~ S-atoms-adsorbed-on-Ni⁰-surface ("S-atoms-on-Ni⁰") as the most-suitable sulphur form-representing compounds, capable of explaining at least 90% of data variance with the *operando*-recorded spectra, thus, well in line with the observed experimental noise within these spectra ($\pm 10\%$). For completeness, it must be mentioned that the S K-edge XANES spectrum of NiS is practically identical with that of Ni₃S₂, whereas they both (NiS, Ni₃S₂) differ clearly from that of Ni₂S. On the other hand, we have no reasons to believe that the XANES fingerprints e.g. with Ce₂O₂S or Ce₂S₃ will differ much from that of NiS or Ni₃S₂. Important here is that each reference compound stands for one specific sulphur form (oxidation state, functional group), but not necessarily for the chemical or physical state it had been measured in. For example: SO₂ gas and solid Na₂SO₃ are both S⁴⁺-containing compounds, whereas SO₃ (g) and NiSO₄(s) each comprise S⁶⁺. Linear combination fits reproduced each *operando* XANES spectrum fairly well, except the one recorded at $T = 350$ °C by featuring a sulphur form not present (or hardly seen) with the other spectra. Its fingerprint strikingly resembled that of SO₂ adsorbed at Ni⁰.

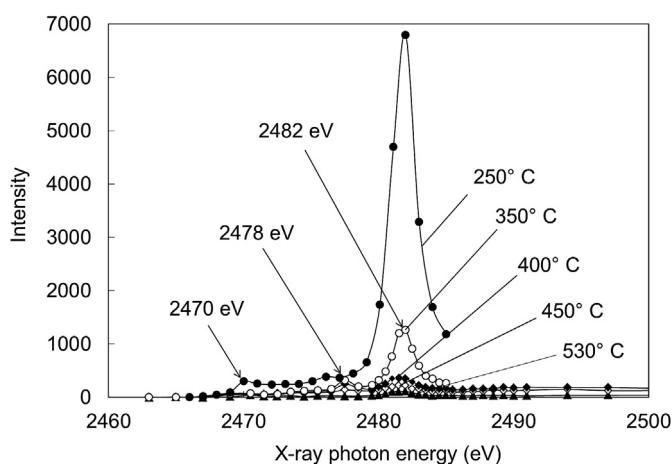


Fig. 4. Dependence of sulphur K-edge XANES spectral intensity on the X-ray photon energy at different measurement temperatures (250 °C–530 °C as shown in the figure) for the Ni-CGO anode in stream of hydrogen with 5 ppm of H₂S.

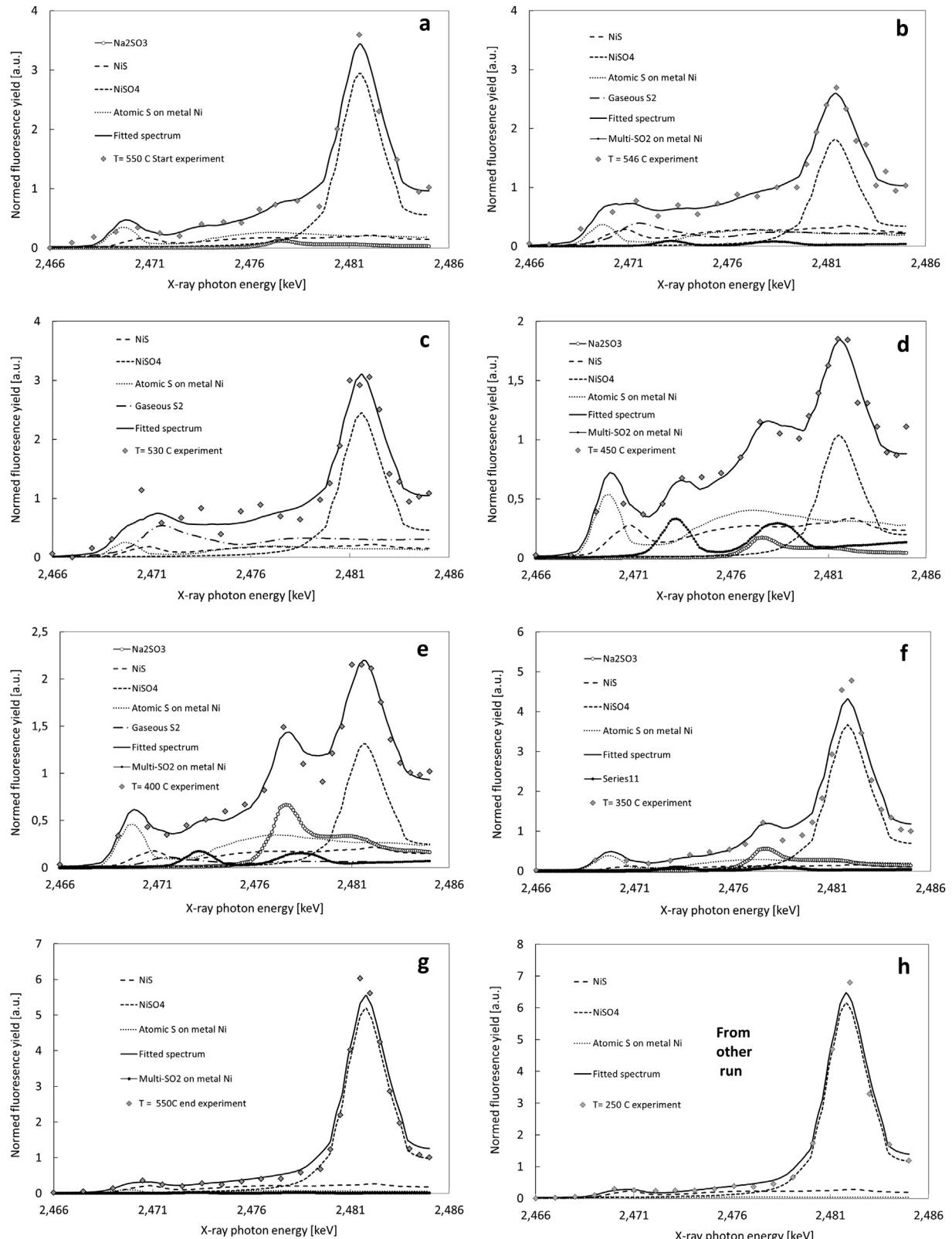


Fig. 5. Experimental sulphur K-edge XANES spectra at the Ni-GDC anode in presence of 5 ppm H₂S in H₂ in the working anode compartment at different temperatures (indicated in figure). Least square fitting results and model spectra used for fitting are shown as well.

surfaces (SO₂-molecules-on-Ni⁰⁺) [48,49] and it differs noticeably from that of S⁴⁺ probed with sulphite salt, such as Na₂SO₃. For consistency, the “SO₂-molecules-on-Ni⁰” reference spectrum was also fitted with the other operando-recorded spectra.

Thus, in total, spectra for six different sulphur containing compounds have been used in the LCF of the *operando*-recorded XANES spectra. The significant number with the fit results is the sum of the sulphur form-specific fractions, which should be equal to one (1) at

Table 1

Cumulated variances and eigenvalues from PCA with StatistiXL in case of different number of components used in analysis.

Number of components	Eigenvalue	Cumulated variance (%)
PC-1	7.427	92.84
PC-2	0.370	97.46
PC-3	0.106	98.78
PC-4	0.050	99.41
PC-5	0.023	99.69
PC-6	0.016	99.88
PC-7	0.008	99.98
PC-8	0.001	100.00

ideal conditions (without noise) or 1 ± 0.1 within the presence of experimental noise. The latter estimate also implies that fitted fractions clearly falling below 0.1 are potentially disputable as being caused by an experimental noise.

3.5. Data reproduction quality with the operando-recorded XANES spectra

All operando-recorded S K-edge XANES spectra at Ni-GDC anode characteristic for different temperatures (indicated with the figures) in hydrogen with 5 ppm H₂S impurity are shown in Fig. 5, together with the fitted sulphur form-specific constituents and their sum ("fitted spectrum"). The fitted fractions and fitted sulphate peak positions (in terms of shift relative to 2482 eV) are summarized in Table 2, together with calculated "sum of fractions" and best-fit criterion values ("Residue (%)"). Note that in Fig. 5 and Table 2, certain sulphur forms (S^{2-} , S^{4+} , S^{6+}) are not specified regarding chemical or physical state. This was done for the reasons explained in Section 3.4. Focussing on the data reproduction quality, from Table 2 it can be seen that the Residue (%) -values varied not more than a factor of three (0.08–0.25%) between the, respectively, fitted XANES spectra. The sum calculated with the fitted fractions (Table 2) were all close to unity, except with the spectra recorded at $T = 350$ °C and $T = 250$ °C as well as after re-installing the SOFC temperature back to $T = 550$ °C, respectively. These three spectra rendered "sum of fractions" being at least ~25% larger than the averaged value of 0.96 ± 0.10 calculated with the other spectra, thus, they are clearly beyond experimental noise levels ($\pm 10\%$ maximally). Another oddity was that their fitted sulphate peak positions deviated systematically nearly 0.2 eV less compared with NiSO₄ as reference than the other spectra measured (Table 2).

3.6. First considerations to the sulphur speciation results

Fig. 6 shows sulphur form-specific fractions pertaining to the temperature-cycle performed with the functioning SOFC. The run started at $T = 550$ °C, followed by stepwise lowered residence at

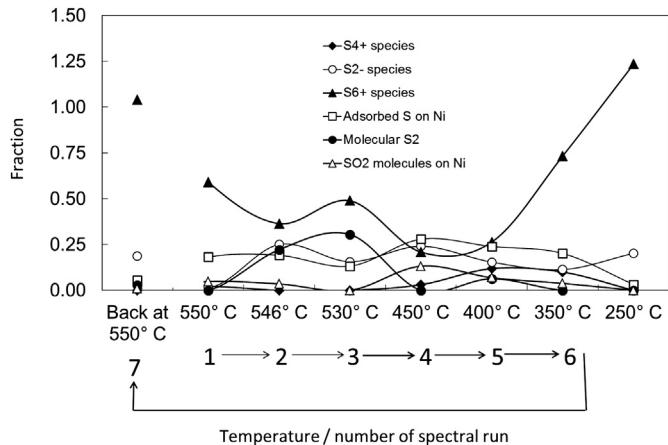


Fig. 6. Sulphur speciation results (fractional composition) as a function of the SOFC operating temperature, characteristic for the Ni-GDC anode, exposed with hydrogen and 5 ppm of H₂S. Sequence of spectroscopic runs (1→7) is indicated in the figure. Spectrum at 250 °C is measured in separate spectral run.

different temperatures down to $T = 350$ °C, and ended after returning back to $T = 550$ °C. In Fig. 6, these results are also indicated by their run time order (1→7) and it also includes speciation results for $T = 250$ °C recorded in a foregoing run ($T = 600 \rightarrow 250$ °C) with a new, freshly reduced, electrochemical element, labelled " $T = 250$ °C" (another run). From Fig. 6 (Table 2), it follows that only sulphur atoms on metallic Ni (Atom-S-on-Ni⁰) and S⁶⁺ species are noticeable at the beginning and may therefore be related intimately with the modest, reversible, drop in the SOFC performance occurring within minutes after the start of experiment (Section 3.1). Two sulphur forms are joined by the presence of S²⁻-species and that of molecular S₂ (g) at $T = 546$ °C. Their run time-delayed appearance indicates that it takes some time (nearly an hour) before reaching steady state-like reaction conditions for a new SOFC electrochemical element. Particularly noticeable thereafter is that the fraction values of S²⁻, S⁶⁺, and Atom-S-on-Ni⁰ do not change much between $T = 546 \rightarrow 350$ °C, except fraction of S⁶⁺ at $T = 350$ °C, where it started to increase drastically and grew even more at lower temperatures (" $T = 250$ °C (obtained from another run)"). In fact, similarly high values of S⁶⁺ fraction also persisted shortly after returning back from $T = 350 \rightarrow 550$ °C. Whatever the cause is, it does not explain the oddities with the sulphate-resembling peaks (Section 3.5). In fact, also the molecular S₂ (g) and the S⁴⁺ species are of special interest because they appear after each other (S₂ (g) within $T = 546\text{--}530$ °C; S⁴⁺ within $T = 400\text{--}350$ °C), with the rather unique occurrence of "SO₂-molecules-at-Ni⁰" at $T = 450$ °C. With respect to S₂ (g), pure thermal decomposition of H₂S into H₂ and S₂ is not very likely, because this process

Table 2

The fitted fractions of constituting sulphur forms, the sum-of-fractions, fitted sulphate peak positions (shift relative to 2482 eV), and best-fit criterion values (%Residue).

	550 °C start	546 °C	530 °C	450 °C	400 °C	350 °C	550 °C end	250 °C
S ⁶⁺ species	0.59	0.36	0.49	0.21	0.26	0.73	1.04	1.23
S ⁴⁺ species	0.02	0.00	0.00	0.03	0.12	0.10	0.00	0.00
Adsorbed S on Ni ⁰	0.18	0.19	0.13	0.28	0.24	0.20	0.05	0.03
Molecular S ₂	0.00	0.22	0.30	0.00	0.06	0.00	0.03	0.00
SO ₂ molecules-on Ni ⁰	0.05	0.04	0.00	0.13	0.07	0.04	0.01	0.00
S ²⁻ species	0.02	0.25	0.16	0.24	0.15	0.11	0.19	0.20
<i>Related fit results</i>								
Sum of fractions	0.86	1.06	1.08	0.89	0.91	1.19	1.32	1.47
Residue (%)	0.113	0.0989	0.214	0.064	0.097	0.262	0.250	0.192
Shift in sulphate peak position (eV) ^a	-0.45	-0.55	-0.41	-0.44	-0.35	-0.19	-0.22	-0.18

^a Relative to NiSO₄ reference spectrum used in the LCFs.

would require substantially higher temperatures ($T > 1000$ K) than applied here (cf. [50]). Recombination reactions between adsorbed sulphur atoms at the Ni^0 surface are also not known for high $\text{S}_2(\text{g})$ output. To better understand the changes in the sulphur oxidation states at the Ni-GDC electrode, phase-predominance diagrams were calculated for $\text{Ni}-\text{O}-\text{S}$ and $\text{Ce}-\text{O}-\text{S}$, respectively, which will be discussed first before presenting final considerations thereafter.

3.7. Agreement of spectroscopic results with thermodynamic predictions

At elevated temperatures, such as $T = 550$ °C and $T = 450$ °C, phase-predominance calculations at experimental $p\text{O}_2^\text{anode}$ values (Fig. 7) indicated that the formation of Ni and Ce sulphate-containing phases are not thermodynamically favourable, because for the $\text{Ni}-\text{O}-\text{S}$ system, they showed prevailing presence for nickel sulphides, particularly, that of amorphous NiS or $\text{NiS}_{0.84}$. For the $\text{Ce}-\text{O}-\text{S}$ system, calculations predicted CeO_2 to be the predominating phase. The speciation results based on our *operando* XANES spectra (Fig. 6) showed significant S^{6+} -related fractions prevailing with most of the intermediate temperatures investigated from $T = 550$ °C–450 °C. For the interpretation of these experiments, three important aspects should be taken into consideration as well: (i) with S K-edge XANES spectra, S^{6+} -containing species can either depict the presence of solid sulphate(s) or that of SO_3 gas. (ii) Besides Ni^0 , also ceria has significant affinity to adsorb SO_2 , leading to $\text{Ce}_2\text{O}_5\text{S}$ [51,52], or, to transform H_2S into SO_2 and/or SO_3 gas; cf [53]. (iii) The phase-predominance calculations are made for stoichiometric CeO_2 and do not include gadolinia doped ceria (GDC) or partially reduced GDC phases, but non-stoichiometric phases have very different oxygen chemical potentials in crystal lattice [54] and higher reactivity as well as adsorption activity compared with non-doped and non-reduced CeO_2 [32]. Another important detail is that there was no thermodynamic equilibrium in the studied system at working conditions, and only steady state conditions (flowing gas feeds and flow of oxide ions through the fuel cell membrane) have been established. Different studies proposed and supported experimentally the idea that the release of sulphur from SOFC electrode might take place through the oxidation of sulphur [10,18,22,23,29], which premises the presence of oxidised particles or intermediates on the anode surface.

The coexistence of different sulphur forms as evidenced with the *operando*-recorded XANES spectra at the S K-edge are unique and they are neither apparent with thermodynamic phase-

predominance calculations made here or by other studies [9,20], nor, with density functional theory calculations from Wang et al. [27]. Interestingly, the latter authors made calculations for the $\text{Ni}-\text{S}$ system as a function of T and $\text{H}_2\text{S}/\text{H}_2$ ratio, but without oxygen, showing the existence of an intermediate state between clean Ni^0 surfaces and nickel sulphides comprising solely sulphur atoms adsorbed on Ni surfaces. This is also supported by the speciation results presented here.

3.8. Final considerations

Considering all previously mentioned information, it is most likely that the S^{6+} signal from $T = 550$ –400 °C originates prevalently from SO_3 particles formed on the GDC surface through the oxidation of adsorbed sulphur particles. This interpretation is supported by the thermodynamic calculations as well as by the methanol synthesis studies using Pd/CeO_2 system with H_2S in reducing syngas, leading to formation of SO_3 as the highest oxidation product [53]. The S^{6+} fraction drops slightly when going from $T = 550$ –400 °C, most likely because the ion conductivity of GDC membrane and the supply of oxide ions to the TPB and GDC surfaces are decreasing. A relatively large energy shift fitted for the experimentally observed S^{6+} peak position at these temperatures (Table 2) compared with that for the reference spectrum (NiSO_4) also confirm that the S^{6+} signal might originate dominantly from SO_3 particles and not from sulphate groups.

At lower temperatures, such as $T = 400$ or 350 °C, the formation of sulphate groups is thermodynamically favoured with both, at Ni and GDC surfaces (Fig. 7). According to the *operando*-recorded XANES spectra (Fig. 6), we observed an increase of S^{6+} fraction and decrease of the fractions pertaining to less oxidised sulphur forms, which is in good accordance with thermodynamic calculations, because the majority of the S^{6+} signal probed at $T = 400$ °C and at $T = 350$ °C is most likely caused by formation of sulphate phases. This assumption is supported by the fact that the shift in the experimental sulphate peak position at these lower temperatures compared with NiSO_4 spectra is minor (Table 2).

After measuring the spectra, which started at $T = 550$ °C, and moving towards lower temperatures down to $T = 350$ °C (spectral run 1–6 with Fig. 6), the SOFC temperature was set back to $T = 550$ °C (spectral run 7) to compare the initial and final spectra at the same temperature. Much higher signal intensity in the final spectrum compared to the one taken at the beginning was detected, which indicates higher concentration of sulphur particles at the anode surface compared to the initial situation at the beginning of the experiment. From sulphur speciation analysis (Table 2), we observed higher sulphate and S^{2-} -related fractions for the final state spectrum, but much less Atom-S-on- Ni^0 compared with the initial one measured at $T = 550$ °C, whereas other sulphur forms were absent for both spectra. Strong similarities between fractional composition between initial and final spectrum indicate that the electrode surface, which had been modified at $T = 350$ °C, needs much longer time to achieve initial situation, except for non-reversible structural changes, as also suggested by recovery of the SOFC performance up to ~90% within 20–30 min (Section 3.1). However, even when the electrical performance of Ni-GDC is predominantly recovered to initial level with electrochemically-relevant TPB-domains, the recovery from chemical modification by sulphur at other, less active domains, could take much longer time.

Finally, the appearance of “ SO_2 -molecules-on- Ni^0 ” might be related with stronger interaction of sulphur with Ni being oxidised further in the TPB region at $p\text{O}_2^\text{anode}$ values prevailing with the working SOFC. However, NiSO_3 under less complicated conditions is already unstable at room temperatures ($\text{NiSO}_3(\text{s}) \rightarrow \text{NiO}(\text{s}) +$

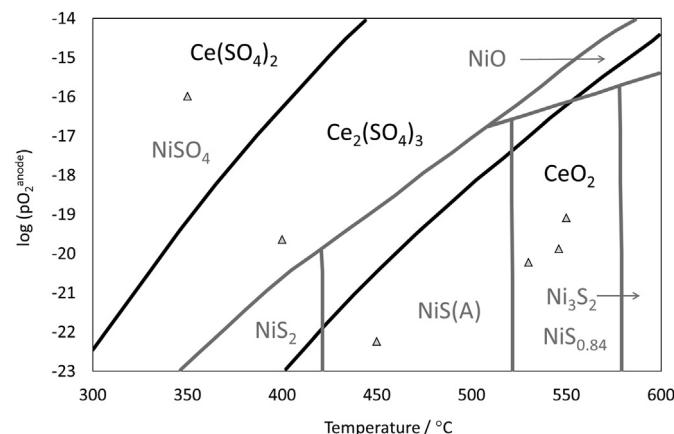


Fig. 7. Overlay of $\text{Ce}-\text{O}-\text{S}$ (black) and $\text{Ni}-\text{O}-\text{S}$ (grey) phase diagrams as a function of temperature and oxygen partial pressure at fixed sulphur partial pressure, $p_{\text{S}2} = 2.5 \times 10^{-6}$. Triangular points (Δ) are characteristic for the experimental conditions, where the XAS spectra were measured.

$\text{SO}_2(\text{g})$ [55]), so this could also explain the co-presence of S^{4+} -containing species in the form of $\text{SO}_2(\text{g})$ species at temperatures below which the “ SO_2 -molecules-on- Ni^0 ” appeared for the first time ($T = 450^\circ\text{C}$), unless suppressed (thermodynamically disfavoured) by the formation of NiSO_4 occurring at $T \leq 350^\circ\text{C}$. In parallel, some S^{4+} species can also be related to other processes taking place at the GDC electrolyte. As mentioned in Section 3.7, ceria can convert H_2S into SO_x . The less oxidised form is SO_2 and further oxidation gives SO_3 . One condition of the continuation of such reactions is that the oxygen removed from the ceria (as SO_x) is compensated. According to Yuchun Ma et al. [53], this can be done, e.g. by adding O_2 to the reaction gas, or, more applicable with the SOFC anode, by oxygen ions moving from cathode to the anode surface via the ceria electrolyte lying in between. However, at the lower temperatures investigated here, the O^{2-} -supply is much lower than at temperatures at $T \geq 450^\circ\text{C}$ (Fig. 7). Thus, the SO_3 formed at high temperatures is restricted to that of SO_2 at $T = 400$ and 350°C along with the formation of $\text{Ce}_2(\text{SO}_4)_3$, whereas the formation of SO_2 vanishes at $T = 250^\circ\text{C}$, because the formation of $\text{Ce}(\text{SO}_4)_2$ prevails thermodynamically. Another explanation for the absence of any SO_2 -related sulphur form within $T = 550 \rightarrow 450^\circ\text{C}$ and sudden presence within $T = 400 \rightarrow 450^\circ\text{C}$ could also be the discontinuation (or disablement) of the $\text{Ce}_2\text{O}_2\text{S}$ regeneration reaction with SO_2 due to insufficient O^{2-} supply, because the otherwise co-produced elemental sulphur (and here likely that of molecular S_2) was no longer detected as soon as the SO_2 -related forms appeared.

4. Conclusions

These first *operando* SOFC-XAS experiments at the S K-edge revealed formation of several sulphur containing compounds with different sulphur oxidation states (6+, 4+, 0, 2-) in parallel by reactions of H_2S with the Ni-GDC cermet anode. Here, these oxidation states are characteristic for $-\text{SO}_4^{2-}$ or SO_3 , $-\text{SO}_3^{2-}$ or $\text{SO}_2(\text{g})$, molecular S_2 or surface-adsorbed S atoms, and, Ni and/or Ce sulphides, respectively. The dynamics of sulphur oxidation states with the subsequently lowered temperatures down $T = 550 \rightarrow 350^\circ\text{C}$ were monitored and discussed. The “ S^{6+} ” fraction was nearly 60% at the beginning of experiment and decreased to approximately 20–25% till $T = 400^\circ\text{C}$, before it increased drastically at lower temperatures, reaching ~90% at $T = 250^\circ\text{C}$ (established use in a separate run), where NiSO_4 and $\text{Ce}(\text{SO}_4)_2$ are thermodynamically favoured. The “ S^0 ” species are indicating atomically adsorbed sulphur atoms at the Ni^0 surface and that of molecular $\text{S}_2(\text{g})$ formed by redox ($\text{Ce}^{3+}, \text{Ce}^{4+}$) reactions of ceria with H_2S with $\text{Ce}_2\text{O}_2\text{S}$ (“ S^{2-} ”) and SO_x (“ $\text{S}^{4+}, \text{S}^{6+}$ ”) as intermediates. Deviations from thermodynamic phase calculations were detected and are likely caused by the differences between assumed equilibrium and real kinetic-controlled steady state conditions prevailing for the SOFC conditions applied, as well as by different reactivity of stoichiometric CeO_2 assumed in calculations, and that of partially reduced ceria (GDC) inhibiting redox system based ($\text{Ce}^{3+}, \text{Ce}^{4+}$) reactivity with the experiment. From the present experiences, also kinetic information regarding individual sulphur oxidation steps can be accessed, because the time for reaching stable fluorescence signal after changing the SOFC temperature took longer than the shortest data accumulation period practised here (~30 min). Developed an *operando* SOFC-XAS cell is versatile as it also allows studying contamination effects at the air-fed cathode side, including Cr [56] poisoning and alkaline elements if appropriate X-ray windows for higher oxygen partial pressures are applied. The presented cell allows also to obtain the metal (in the present case Ni and Ce) absorption edges, which may facilitate the identification of Ni- and CeO_2 -related reaction products with sulphur. We are confident

that the outcome of our initial work will stimulate further *operando* experiments on SOFC assemblies under real exploitation conditions.

Acknowledgements

This work was funded by Sciex-NMS program under award number: 10.063 (*In-Situ-SOFC-XAS*). Partial funding was received from Swiss National Science Foundation, Estonian Science Foundation (Grant number 7791) and Estonian Center of Excellence (Materials for sustainable development). A.B. is grateful for the support from Peter Holtappels and Thomas Graule (Empa), and for financial support from the European Union under Marie Curie project # 042095 “X-ray and Electrochemical Studies on Solid Oxide Fuel Cells and Related Materials” and FCH JU project # 256885 “Solid Oxide Fuel Cells – Integrating Degradation Effects into Lifetime Prediction Models”, and the Swiss Competence Center for Energy and Mobility (CCEM), contract # CO 350260. Authors thank Markus Janousch (SLS) for fruitful discussion and other staff at PSI involved at different stages of the project, in particular, Lukas Oberer, Markus Mähr, Daniel Peitz, Peter Hottinger, Eric de Boni, Reto Wetter Maarten Nachtegaal, and Maria Casapu. Authors also thank Thyssen Krupp for providing the high performance steel used to construct the test cell.

The S K-edge spectroscopic measurements were performed on the Phoenix beamline at the Swiss Light Source, Paul Scherrer Institut, Villigen, Switzerland.

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