

Separation and Quantitative Analysis of An Unknown Pharmaceutical mixture using RP-HPLC

CHEM 4303

Analytical Separations

Robby RENZ

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Partner: Jaya Roe
Lab Instructor: Kevin Stroski

Abstract

High-performance liquid chromatography utilizing the reverse-phase... is a separation technique blah blah blah ...

1 Introduction

High performance liquid chromatography (HPLC) is a type of chromatographic technique which utilizes high pressure to push the solvent through the column [1]. In an HPLC, the stationary phase is usually inorganic particles that are made up of silica particles that are microporous [1]. The type of chromatographic techniques that utilizes this type of stationary phase are normal phase, ion pair, and reversed phase [2]. Furthermore, one would choose liquid chromatography over gas chromatography for the sole purpose of the analyte in question being not volatile enough to actually undergo gas chromatography [1].

In normal phase chromatography, the stationary phase is more polar than the solvent [1]. However, in reversed phase chromatography, it's the exact opposite; the solvent is more polar than the stationary phase [1]. In reverse phase HPLC, or RP-HPLC, the stationary phase is silica coated with C_{18} that acts as a surface [3]. One main advantage of reverse phase over normal phase is that the separation is more reproducible and robust [3].

In terms of detectors, there are numerous ones that are available for the HPLC. The most commonly used detectors are the ultraviolet-visible detectors and mass spectrometers [4]. However, in this experiment, the diode array detector (DAD) is utilized, which is actually a type of UV-visible detector [4]. This detector consists of two lamps (tungsten-halogen lamp and a deuterium lamp) that is able to produce electromagnetic radiation, which shine onto a detector array, and it's continuously recorded [4]. That way, if it's disturbed by the flow of the analyte, the disturbance gets recorded due to the difference in absorption properties [4]. This is how it is able to detect analytes.

Acetaminophen, an antipyretic and an analgesic for humans, is known to cause poisons in cats and dogs [5].

There were three main objectives in this experiment; in the first week, the purpose was to test the efficiency of separations in RP-HPLC by varying the composition of the mobile phase (methanol). In the qualitative part of the experiment, the purpose was to identify the types of compounds present in an unknown pharmaceutical mixture. Finally, in the quantitative part of the experiment, the purpose was to find out the mass of one of the drugs present in the aforementioned unknown pharmaceutical mixture.

2 Chemicals, Methods and Instrumentation

2.1 Chemicals

Benzaldehyde (Sigma Aldrich, Lot: 41696 PKV), nitrobenzene (Alfa Aesar, Lot: A06V040, 99%, CAS: 98-95-3), methanol (HPLC grade, Caledon, Lot: 103267, CAS: 108-90-7), salicylamide (Sigma Aldrich, Lot: S51885-279, CAS: 65-45-2), phenacetin (Sigma Aldrich, Lot: 78C-0014), acetaminophen (Sigma Aldrich, Lot: 32F-0073), caffeine (Sigma Aldrich, Lot: 0316B4), and an unknown pharmaceutical mixture was used in this experiment.

2.2 Instrumentation

The separation and analysis of the unknown pharmaceutical mixtures, and the standards, were performed on an Agilent 1100 Series, with a 1260 Infinity degasser (both by Agilent Technologies), fitted with a diode array detector (DAD). The type of column used was a Symmetry® C₁₈ (particle size of 5 μ m, diameter of 4.6 mm, and a length of 150 mm). The flow rate for each analysis was kept at a constant value of 1 mL/min, and the injection volume was 5 μ L for each analysis.

2.3 Methods

In “testing the retention behavior in RP-HPLC” part of the experiment, benzaldehyde, nitrobenzene and chlorobenzene, with volumes of 30 μ L, 30 μ L and 4 mL, respectively, were mixed together. 20 μ L of this mixture was diluted with 5 mL of HPLC-grade methanol. This diluted mixture was analysed in the HPLC utilizing the RP technique, and the mobile phase was varied until the separations were deemed satisfied. Standards were also prepared for benzaldehyde, nitrobenzene and chlorobenzene. For nitrobenzene, 30 μ L of it were diluted with 4.3 mL of methanol. And from this mixture, 20 μ L were diluted with 5 mL of methanol. The same procedure was done for making the benzaldehyde standard. However, for the chlorobenzene standard, 4 mL of it were added to 60 μ L of methanol, and then 20 μ L of this mixture was transferred to 5 mL of methanol. The UV detection was set at 254 nm.

In the qualitative part of the experiment, 0.0615 g of an unknown pharmaceutical mixture was grinded with a mortar and pestle, and dissolved in 25 mL of methanol. 500 μ L of this mixture was added to 5 μ L of water and 5 μ L of methanol. This diluted mixture underwent RP-HPLC chromatography with an initial condition of 70 : 30 methanol:water. This composition would be varied if necessary in order to get an optimal separation of the mixture. Standards of concentration 2 mg/mL were made for salicylamide, phenacetin, caffeine and acetaminophen so that the compounds in the mixture can be recognized. Note that these standards were diluted 10-fold before being analyzed. The UV detection was set to 245 and 275 nm.

Finally, in the qualitative part of the experiment, a series of calibration solutions was prepared with caffeine acting as the internal standard (IS). 500 μ L was the volume of caffeine that was added to each calibration solutions. Salicylamide and phenacetin was found out to be the two compounds in the unknown mixture, and phenacetin was chosen to be quantitatively analyzed. Volumes of 100 μ L, 300 μ L, 500 μ L, 1500 μ L and 2500 μ L of phenacetin was added to each 10 mL of volumetric flask. A constant amount of IS was also added each flask (500 μ L), and it was filled with 50% methanol till the mark. These solutions was analyzed in the RP-HPLC.

3 Results and Discussion

3.1 Results

As stated earlier, in the first week of the experiment, the compositions of the mobile phase, methanol, was being varied with HPLC-grade water in order to notice the differences in separations in RP-HPLC. The corresponding chromatograms are given in the appendix (figures 1 to 4), and tables 1, 2 and 3 summarizes the aforementioned chromatograms in terms of its capacity factors. In addition, figure 1 illustrates the log of the capacity factors of the three compounds as a function of the polarity of methanol, and table 4 shows how the selectivity factor (α) varies with the polarity of methanol.

Table 1: Data that shows how the capacity factor (k') of benzaldehyde differ with the change in composition in methanol.

Methanol/%	k' of benzaldehyde	$\log k'$
90	0.218	-0.662
80	0.333	-0.478
70	0.497	-0.304

Table 2: Data that shows how the capacity factor (k') of nitrobenzene differ with the change in composition in methanol.

Methanol/%	k' of nitrobenzene	$\log k'$
90	0.274	-0.562
80	0.459	-0.338
70	0.737	-0.133

Table 3: Data that shows how the capacity factor (k') of chlorobenzene differ with the change in composition in methanol.

Methanol/%	k' of chlorobenzene	$\log k'$
90	0.629	-0.201
80	1.138	0.056
70	2.026	0.307

Table 4: Data that shows how the selectivity factors (α) change with the polarity of methanol.

Methanol/%	k'	$\log k'$
100		
90		
80		
70		

Figures 5, 6 and 7 show the chromatograms of a chlorobenzene, nitrobenzene and benzaldehyde standards, respectively.

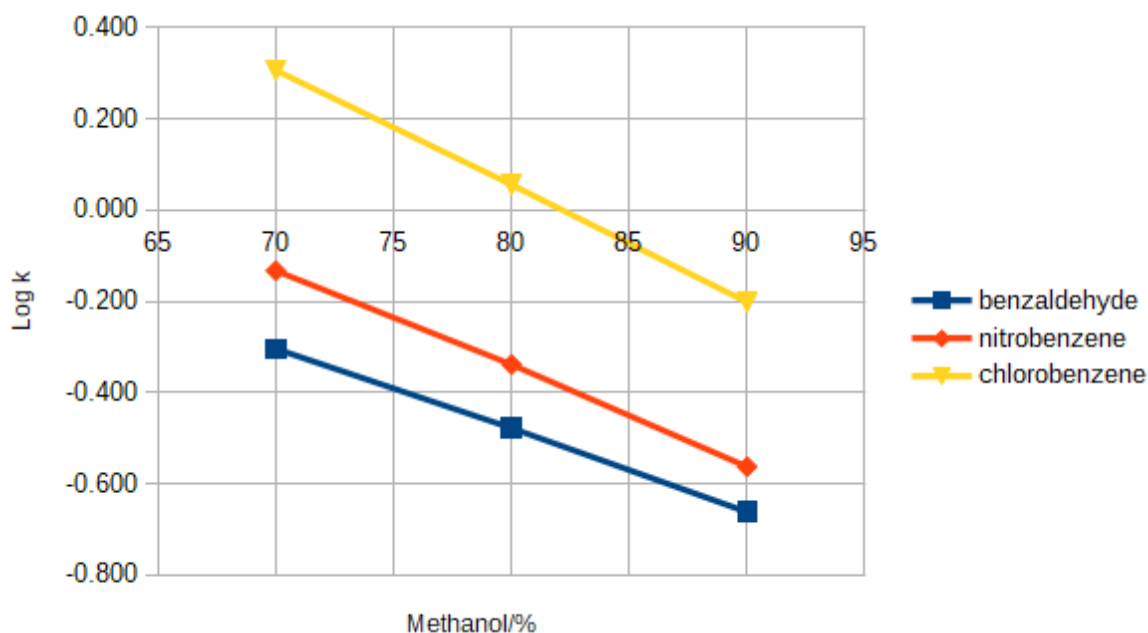


Figure 1: Plot of the Log k of benzaldehyde, nitrobenzene and chlorobenzene as a function of the polarity of methanol

In addition, figure 8 illustrates the chromatogram of the unknown pharmaceutical mixture. While figures 9, 10, 11 and 12 shows the chromatograms of the caffeine, phenacetin, salicylamide and the acetaminophen standard, respectively. Note that the optimal composition of the mobile phase was found out to be 50 : 50 methanol:water.

Finally, in the quantitative analysis of the experiment, figures 13, 14, 15, 16 and 17 represent the chromatograms Figures 18 and 19 show the chromatograms of the analysis of the aforementioned pharmaceutical unknown. Finally, figures 20, 21 and 22 represents the chromatograms of the blank that was run (50% methanol).

The pressure of the column for each analysis is shown in table 5. Note that the figure headings correspond to the analysis of the chromatograms in the appendix.

Now draw a table for the data of the calibration curve, and include a figure.

3.2 Discussion

Give a reason as to why reverse-phase is used as a separation technique and not normal phase.

In the quantitative identification of the pharmaceutical unknown mixture, the initial concentration of the mobile phase was chosen as 50% methanol, and based on the chromatogram received, this composition was considered to be desirable for the analysis of the aforementioned unknown mixture. This was based on the resolution and the retention time for each peaks.

Tables 1, 2 and 3, and figure 1 clearly show that the capacity factor increases as the percentage of methanol decreases (that is, the polarity of methanol increases). Capacity factor, also

Table 5: Table that tabulates how the pressure of the column changes with each analysis. The figure number references the chromatogram in the appendix.

Figure number	Pressure/bar
Figure 1	72
Figure 2	98
Figure 3	122
Figure 4	143
Figure 5	143
Figure 6	143
Figure 7	143
Figure 8	178
Figure 9	178
Figure 10	178
Figure 11	178
Figure 12	178
Figure 13	174
Figure 14	174
Figure 15	174
Figure 16	174
Figure 17	174
Figure 18	174
Figure 19	174
Figure 20	174
Figure 21	174
Figure 22	174

known as retention factor [1], denotes the time required for the analyte in question to elute [1]. A greater capacity factor indicates that it is greatly attracted to the column, interacting with the stationary phase [1]. Therefore, one can conclude that, as it can be clearly seen in figure 4, benzaldehyde is the most polar while chlorobenzene is the least polar of the three compounds; as stated earlier, in RP chromatography, the solvent is more polar than the stationary phase [1].

As it can be seen in table 5, the pressure increases whenever the polarity of the mobile phase, in this case methanol, increases (by the addition of water). This is because as one increases the amount of water in the solvent, it increases the viscosity of the solvent, so a higher pressure is required to maintain the flow rate [1].

The reason for using internal standards and not standard addition for calculating the concentration of the unknown analyte is that internal standards are preferred when give an accurate answer when

In conclusion, reversed-phase HPLC gives a ...

4 References

References

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- (2) Moldoveanu, S.; David, V., *Essentials in modern HPLC separations*; Elsevier: Waltham, MA, 2013.
- (3) *Liquid chromatography: Fundamentals and Instrumentation*; Fanali, S., Haddad, P. R., Lloyd, D., Poole, C. F., Schoenmakers, P. J., Eds.; Elsevier: Amsterdam, 2013.
- (4) Vitha, M. F., *Chromatography: Principles and Instrumentation*; Wiley: Hoboken, New Jersey, 2017.
- (5) McConkey, S. E.; Grant, D. M.; Cribb, A. E. *Journal of Veterinary Pharmacology and Therapeutics*, 32, 585–595.

5 Appendix

5.1 Calculations

Calculating the response factor

$$\frac{\text{Area of Analyte Signal}}{\text{Concentration of Analyte}} = F \left(\frac{\text{Area of Standard Signal}}{\text{Concentration of Standard}} \right) \quad (1)$$
$$\frac{A_X}{[X]} = F \left(\frac{A_S}{[S]} \right)$$

Equation 1 was taken from [1], where $[X]$ and $[S]$ represent the concentrations of analyte and of the standard, respectively.

%RSD

Calculating the capacity factor of chlorobenzene for figure 2 of the chromatogram

$$K = \frac{t_r - t_m}{t_m}$$
$$K = \frac{2.343 - 1.438}{1.438} \quad (2)$$
$$K = \frac{0.905}{1.438}$$
$$K = 0.629346 \approx 0.629$$

Equation 2 was taken from [1].

5.2 Chromatograms

There are 22 sheets of HPLC chromatograms.