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## THERMODYNAMIC RESPONSE FUNCTIONS IN A CELL FLUID MODEL

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*Thermodynamic response functions, namely the isothermal compressibility, the thermal pressure coefficient, and the thermal expansion coefficient, are calculated for a many-particle system interacting through a modified Morse potential. These calculations are based on an equation of state previously derived for a cell fluid model in the grand canonical ensemble. The calculated quantities are presented graphically as functions of density and chemical potential.*

**Keywords:** Cell model, Morse potential, Thermodynamic response functions

### 1. Introduction

In this work we continue our study of the thermodynamic behavior of a cell fluid model, which was defined in [1, 2]. In works [3–6] this model was used with a modified Morse potential describing the particle interaction. In particular, the equation of state was obtained in [3] in the zero-mode approximation. This equation of state is used in the current work to calculate thermodynamic response functions, namely the isothermal compressibility, the thermal pressure coefficient, and the thermal expansion coefficient.

As the problem has been considered in the framework of the grand canonical ensemble, the initial equation of state is formulated in terms of pressure  $P$ , temperature  $T$  and the chemical potential  $\mu$ ,  $P = P(T, \mu)$ . To leverage this form of the equation of state for calculation of the response functions, the corresponding definitions should be transformed to proper derivatives with respect to temperature and chemical potential [7]. For each response function considered in this paper, we present such transformation in a dedicated subsection. On the other hand, within the approach applied here, the chemical poten-

tial can be explicitly expressed via the number particle density  $\rho$  and temperature. This gives rise to the equation of state in the form  $P = P(T, \rho)$ , in which case it is suitable to re-express the response functions in terms of derivatives with respect to temperature and density.

In Section 2 we present the modified Morse potential used in our work, and briefly compare it with other possible modifications. Section 3 is dedicated to the explicit expressions for the equation of state. In Section 4 the expressions for response functions are derived in terms of thermodynamic derivatives suitable for different forms of the equation of state.

### 2. The interaction potential

The potential of interaction between particles is taken in a form of a modified Morse potential

$$U(r) = \varepsilon C_H \left[ A e^{-n_0(r-R_0)/\alpha} + e^{-\gamma(r-R_0)/\alpha} - 2e^{-(r-R_0)/\alpha} \right], \quad (1)$$

where  $R_0$  is the coordinate of the potential minimum,  $\alpha$  is an effective range of interaction,  $\gamma$  and  $n_0$  are

parameters of the model. Other two constants  $C_H$  and  $A$  are expressed via  $\gamma$  and  $n_0$  as follows

$$C_H = \frac{n_0}{n_0 + \gamma - 2}, \quad A = \frac{2 - \gamma}{n_0}, \quad (2)$$

where  $\varepsilon$  is the depth of the potential well at  $r = R_0$ . This potential is reduced to the ordinary Morse potential [8] at  $\gamma = 2$ . For a more detailed discussion of such modified Morse potential, see Sections 1 and 2 in [3].

Modifications of the Morse potential have been used in other works as well. For example, in [9] a repulsive term in a form of a power of  $r^{-1}$  was added to the ordinary Morse potential, and the influence of the softness of such a term was investigated on the coordinates of the critical point. The generalized form of the Morse potential was suggested in [10]

$$U(r) = A_1 e^{-\lambda_1 r} + A_2 e^{-\lambda_2 r} \quad (3)$$

with application to silicon structural energies, and was also considered in [11] as the potential for Be-S and H-Na compounds.

Our modification contains an additional repulsive term, similarly to [9], as well as introduces parameter  $\gamma$ , which can vary as opposed to being strictly equal to 2 in the Morse potential. Including the repulsive term enables us to single out a reference system (in the reciprocal space) and apply the method of collective variables to calculating the grand partition function [12].

### 3. The equation of state

#### 3.1. Pressure as a function of temperature and chemical potential

The equation of state obtained in [3, see Eq.(42)] reads

$$Pv\beta = E_\mu(M, T) + M\bar{\rho}_0 + \frac{1}{2}\tilde{D}(0)\bar{\rho}_0^2 - \frac{a_4}{24}\bar{\rho}_0^4. \quad (4)$$

The quantities in the left-hand side of the equation are  $P$ , the pressure;  $\beta = (k_B T)^{-1}$ , the inverse temperature;  $k_B$ , the Boltzmann constant;  $T$ , the temperature;  $v$ , cell volume. The quantities in the right-hand side are, in general, functions of the temperature  $T$  and the chemical potential  $\mu$ . Let us present their expressions explicitly.

First, the quantity  $M$  depends linearly on the chemical potential

$$M = \frac{\tilde{\mu}}{W(0)} + g_1 - \frac{g_3}{g_4}\tilde{D}(0) - \frac{1}{6}\frac{g_3^3}{g_4^2}, \quad (5)$$

$$\tilde{\mu} = \mu - \mu_0(1 + \tau), \quad (6)$$

where  $\mu_0$  is some positive constant,  $\tau$  is the relative temperature  $\tau = (T - T_c)/T_c$ ,  $T_c$  is the critical temperature. We will call  $M$  the effective chemical potential.

The quantity  $W(0)$  is expressed via parameters of the potential (1) as follows

$$W(0) = \Phi^{(r)}(0) [B - 1 + \chi_0 + \tau(\chi_0 + A_\gamma)], \quad (7)$$

where

$$B = 2\gamma^3 e^{(1-\gamma)R_0/\alpha},$$

$$A_\gamma = A e^{(n_0 - \gamma)R_0/\alpha} (\gamma/n_0)^3,$$

and  $\Phi^{(r)}(0)$  is the Fourier transform of the repulsive part of the potential at  $|\mathbf{k}| = 0$

$$\Phi^{(r)}(0) = \varepsilon C_H 8\pi e^{\gamma R_0/\alpha} \left( \frac{\alpha}{\gamma R_0} \right)^3.$$

The parameter  $\chi_0$  is used in [3] to single out a contribution in the Fourier transform of the potential that is treated as a reference system defined in the reciprocal space, and is selected as  $\chi_0 = 0.07$  [3, see Eq.(24)].

The coefficients  $g_n$  are given by the formulas:

$$\begin{aligned} g_0 &= \ln T_0, & g_1 &= T_1/T_0, & g_2 &= T_2/T_0 - g_1^2, \\ g_3 &= T_3/T_0 - g_1^3 - 3g_1g_2, \\ g_4 &= T_4/T_0 - g_1^4 - 6g_1^2g_2 - 4g_1g_3 - 3g_2^2, \\ a_4 &= -g_4, \end{aligned} \quad (8)$$

where  $T_n(p, \alpha^*)$  are the following special functions

$$T_n(p, \alpha^*) = \sum_{m=0}^{\infty} \frac{(\alpha^*)^m}{m!} m^n e^{-pm^2}. \quad (9)$$

Here  $\alpha^* = v e^{\beta_c \mu_0}$ , and the parameter  $p$  has the form

$$p = \frac{\beta_c}{2} \Phi^{(r)}(0) [\chi_0 + A_\gamma]. \quad (10)$$

The quantity  $\beta_c$  denotes the critical value of the inverse temperature. In [3, see Eq.(31)] it was found that

$$\varepsilon\beta_c = 0.200, \quad \frac{k_B T_c}{\varepsilon} = 4.995.$$

We also use the reduced temperature defined as  $T = k_B T/\varepsilon$ , and thus its critical value  $T_c^* = 4.995$ .

Since  $p$  is independent of temperature, the coefficients  $g_n$  are also independent of temperature. The numerical values for other coefficients used in this paper are the same as those in [3, see Eqs.(5), (23), and (24)]:

$$\begin{aligned} \chi_0 &= 0.07, & \gamma &= 1.65, \\ n_0 &= 1.521, & R_0/\alpha &= 2.9544, \\ \alpha^* &= 5.0 & p &= 1.0. \end{aligned} \quad (11)$$

The quantity  $\tilde{D}(0)$  entering equations (4) and (5) is a function of temperature

$$\tilde{D}(0) = g_2 - \frac{1}{2} \frac{g_3^2}{g_4} - \frac{1}{\beta W(0)}. \quad (12)$$

The condition  $\tilde{D}(0) = 0$  defines the critical temperature [3, see Eq.(31)]

$$k_B T_c = \left( g_2 - \frac{1}{2} \frac{g_3^2}{g_4} \right) (B - 1 + \chi_0) \Phi^{(r)}(0). \quad (13)$$

The function  $E_\mu(M, T)$  from the equation (4) is provided by

$$\begin{aligned} E_\mu(M, T) &= -\frac{\ln(2\pi\beta W(0))}{2N_v} \\ &+ g_0 - \frac{\beta W(0)}{2} \left( \frac{\tilde{\mu}}{W(0)} \right)^2 \\ &- \frac{g_3}{g_4} M - \frac{g_3^2}{2g_4^2} \tilde{D}(0) - \frac{1}{24} \frac{g_3^4}{g_4^3}. \end{aligned} \quad (14)$$

Here the quantity  $N_v$  defines the number of cubic cells in volume  $V$  for the initial model. In the thermodynamic limit,  $N_v \rightarrow \infty$ , and thus the first term can be neglected. The term  $\tilde{\mu}/W(0)$  can be expressed in terms of  $M$  using (5). The temperature and the inverse temperature can always be expressed in terms of the reduced temperature and a corresponding critical value:

$$T = T_c(1 + \tau), \quad \beta = \beta_c(1 + \tau)^{-1}$$

The quantity  $\bar{\rho}_0$  is a solution to the following cubic equation

$$M + \tilde{D}(0)\bar{\rho}_0 - \frac{a_4}{6}\bar{\rho}_0^3 = 0. \quad (15)$$

For any  $\tau > 0$ , the latter equation has one real root

$$\bar{\rho}_0 = \left( -\frac{3M}{g_4} + \sqrt{Q_t} \right)^{1/3} - \left( \frac{3M}{g_4} + \sqrt{Q_t} \right)^{1/3}, \quad (16)$$

where

$$Q_t = \left( \frac{2\tilde{D}(0)}{g_4} \right)^3 + \left( -\frac{3M}{g_4} \right)^2, \quad g_4 < 0. \quad (17)$$

Thus,  $\bar{\rho}_0$  is a function of the temperature and the chemical potential.

Let us introduce the reduced pressure

$$P^* \equiv \frac{Pv}{\varepsilon}. \quad (18)$$

Considering the equation of state (4),  $P^*$  is explicitly written as

$$\begin{aligned} P^* &= (1 + \tau)T_c^* \left[ E_\mu(M, T) \right. \\ &\quad \left. + M\bar{\rho}_0 + \frac{1}{2}\tilde{D}(0)\bar{\rho}_0^2 - \frac{a_4}{24}\bar{\rho}_0^4 \right]. \end{aligned} \quad (19)$$

This equation can be easily represented graphically. Figure 1 illustrates the relationship between the reduced pressure  $P^*$  and the effective chemical potential  $M$  for various values of the relative temperature  $\tau$ , at and above the critical temperature. At the critical point  $M = 0, \tau = 0$ ,  $P_c^* = 1.606$ .

### 3.2. Pressure as a function of temperature and density

In this work, by density we mean the particle number density  $\rho = \langle N \rangle / V$ . In the framework of the grand canonical ensemble the average number of particles  $\langle N \rangle$  is found by

$$\begin{aligned} \langle N \rangle &= \left( \frac{\partial \ln \Xi}{\partial \beta \mu} \right)_{T, V} \\ &= V \left( \frac{\partial P}{\partial \mu} \right)_{T, V}. \end{aligned}$$

From this equation it follows

$$\frac{\langle N \rangle}{V} = \frac{\langle N \rangle}{vN_v} = \left( \frac{\partial P}{\partial \mu} \right)_{T, V}$$

and thus

$$\rho^* \equiv \frac{\langle N \rangle}{V} v = v \left( \frac{\partial P}{\partial \mu} \right)_{T,V} = \left( \frac{\partial(Pv\beta)}{\partial(\beta\mu)} \right)_{T,V}. \quad (20)$$

The quantity  $\rho^*$ , on the one hand, is the reduced particle number density, which is the notation commonly used in the literature on simple fluids [13]. In the context of the cell model, on the other hand, it is the average number of particles per cell.<sup>1</sup> Taking explicit derivatives, we arrive at

$$\rho^* = n_g - M + \frac{\bar{\rho}_0}{\beta W(0)}. \quad (21)$$

The quantity  $n_g$  in the equation (21) is the critical density [3]

$$\begin{aligned} n_g &= g_1 - \frac{g_3}{g_4} \left( g_2 - \frac{1}{2} \frac{g_3^2}{g_4} \right) - \frac{1}{6} \frac{g_3^3}{g_4^2} \\ &= g_1 - \frac{g_2 g_3}{g_4} + \frac{g_3^3}{3g_4^2} \end{aligned} \quad (22)$$

Its numerical value for parameters (11) is

$$n_g = 0.978.$$

Equations (21) and (4) jointly define a parametric relationship between pressure and density, with  $M$  serving as the parameter. Figure 1 shows the dependency of the reduced pressure  $P^*$  on the density  $\rho^*$  for various values of the reduced temperature  $\tau$ , at and above the critical temperature.

This dependency can also be expressed explicitly. To achieve this, we combine the equations (21) and (15) to express the effective chemical potential  $M$  as a function of density and temperature:

$$\bar{M} = \frac{\rho_n}{\beta W(0)} - (\rho^* - n_g), \quad (23)$$

where

$$\begin{aligned} \rho_n &= -2 \left( \frac{g_3^2 - 2g_2 g_4}{g_4^2} \right)^{\frac{1}{2}} \cos \left( \frac{\alpha_n}{3} + \frac{\pi}{3} \right), \\ \alpha_n &= \arccos \left( \left( -\frac{9g_4^4}{(2g_2 g_4 - g_3^2)^3} \right)^{\frac{1}{2}} (n_g - \rho^*) \right). \end{aligned} \quad (24)$$

<sup>1</sup> In our previous works we denoted the reduced number density by  $\bar{n}$ . In the current work we switch to more common notation  $\rho^*$ .

The notation  $\bar{M}$  represents the effective chemical potential as a function of temperature and density  $\rho^*$ , while  $M$  represents the effective chemical potential as a function of temperature and chemical potential  $\mu$ .

At  $T > T_c$  the equation of state of a cell fluid model in terms of the density and the temperature has the following form

$$Pv\beta = E_\eta(\rho^*, T) + \bar{M}\rho_n + \frac{\tilde{D}(0)}{2}\rho_n^2 - \frac{a_4}{24}\rho_n^4. \quad (25)$$

The quantity  $E_\nu(\rho^*, T)$  in (24) is the function  $E_\mu(M, T)$  (14) rewritten in terms of density and temperature taking into account the expression (23)

$$\begin{aligned} E_\nu(\rho^*, T) &= -\frac{\ln(2\pi\beta W(0))}{2N_v} + g_0 \\ &\quad - \frac{\beta W(0)}{2} \left( \bar{M} - g_1 + \frac{g_3}{g_4} \tilde{D}(0) + \frac{g_3^3}{6g_4^2} \right)^2 \\ &\quad - \frac{g_3}{g_4} \bar{M} - \frac{g_3^2}{2g_4^2} \tilde{D}(0) - \frac{1}{24} \frac{g_3^4}{g_4^3}. \end{aligned} \quad (26)$$

For the reduced pressure we write

$$\begin{aligned} P^* &= (1 + \tau)T_c^* \left[ E_\eta(\rho^*, T) \right. \\ &\quad \left. + \bar{M}\rho_n + \frac{\tilde{D}(0)}{2}\rho_n^2 - \frac{a_4}{24}\rho_n^4 \right]. \end{aligned} \quad (27)$$

Figure 1 shows the isotherms for the pressure  $P^*$  versus density  $\rho^*$  (left-hand side) and versus the effective chemical potential  $M$  (right-hand side).

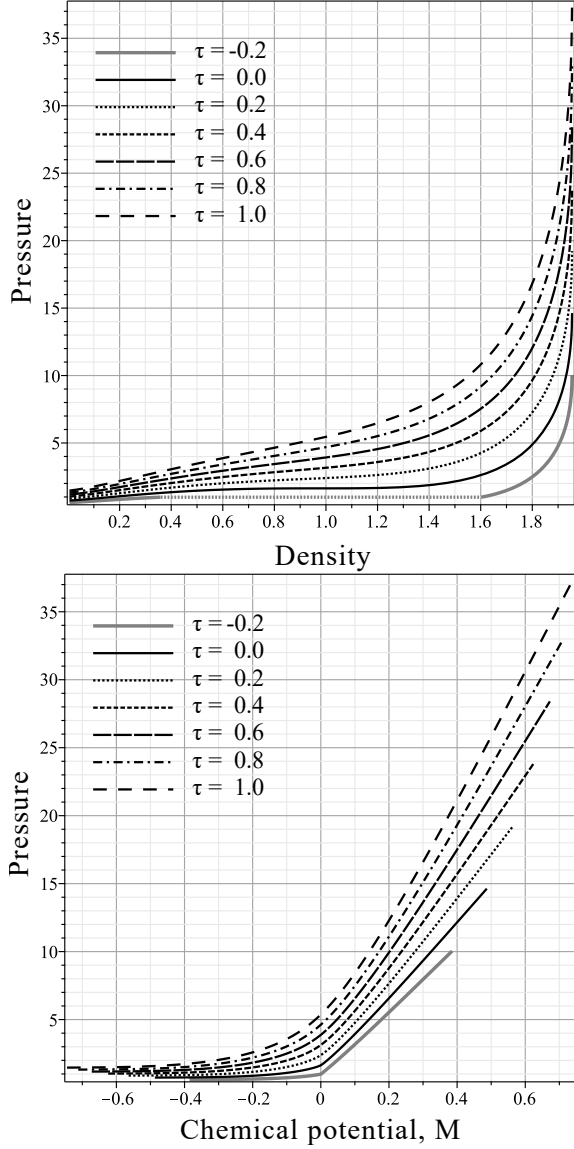
Thus, in this Section, we presented two forms of the equation of state. The first one expresses the pressure as a function of temperature and chemical potential,  $P = P(T, \mu)$ , and is represented by equivalent Eqs. (4) and (19). The second one expresses the pressure as a function of temperature and density,  $P = P(T, \rho^*)$ , and is represented by equivalent Eqs. (25) and (27). These equations, together with the explicit dependency of density on temperature and chemical potential,  $\rho^* = \rho^*(T, \mu)$ , Eq. (21), are the basis for the calculation of the thermodynamic response functions in the next Section 4.

## 4. Thermodynamic response functions

### 4.1. Isothermal compressibility

The isothermal compressibility is defined by

$$K_T = -\frac{1}{V} \left( \frac{\partial V}{\partial P} \right)_{T,N}. \quad (28)$$



**Fig. 1.:** Isotherms of the reduced pressure  $P^*$  as a function of the density  $\rho^*$  (on the left-hand side) and a function of the effective chemical potential  $M$  (on the right-hand side) at  $T \geq T_c$  represented by black lines. Thick grey lines on both figures correspond to isotherms of pressure at  $T < T_c$  based on the results taken from [3].

Let us perform some transformations to rewrite  $K_T$  into a form that is more suitable for the equation of

state  $P = P(T, \mu)$ , see (4) and (19):

$$\begin{aligned} K_T &= \frac{1}{\rho} \left( \frac{\partial \rho}{\partial P} \right)_{T, N} \\ &= \frac{1}{\rho} \frac{(\partial \rho / \partial \mu)_T}{(\partial P / \partial \mu)_T}. \end{aligned}$$

We have omitted the condition of constant  $N$  in the last line of the above equation since we have explicit dependencies on temperature and chemical potential for both pressure  $P = P(T, \mu)$ , Eq. (4), and density  $\rho = \rho(T, \mu)$ , Eq. (21). Applying the Gibbs–Duhem equation

$$Nd\mu = -SdT + VdP, \quad (29)$$

at  $T = \text{const}$  one has

$$dP = \rho d\mu,$$

or

$$\rho = \left( \frac{\partial P}{\partial \mu} \right)_T. \quad (30)$$

Substituting this into the last expression for  $K_T$ , one gets

$$K_T = \frac{1}{\rho^2} \left( \frac{\partial \rho}{\partial \mu} \right)_T \quad (31)$$

Finally, from (30) it follows that

$$\left( \frac{\partial \rho}{\partial \mu} \right)_T = \left( \frac{\partial^2 P}{\partial \mu^2} \right)_T \quad (32)$$

and ultimately we arrive at the very useful expression for the isothermal compressibility

$$K_T = \frac{1}{\rho^2} \left( \frac{\partial^2 P}{\partial \mu^2} \right)_T \quad (33)$$

Let us introduce the reduced isothermal compressibility

$$K_T^* \equiv \frac{\varepsilon K_T}{v}. \quad (34)$$

The quantity  $K_T^*$  is dimensionless and is of order unity, except at the critical point itself, where it is divergent. It is expressed in terms of the reduced quantities  $P^*$  and  $\rho^*$  as follows

$$\begin{aligned} K_T^* &= \frac{\varepsilon}{\rho^{*2}} \left( \frac{\partial \rho^*}{\partial \mu} \right)_T \\ &= \frac{\varepsilon^2}{\rho^{*2}} \left( \frac{\partial^2 P^*}{\partial \mu^2} \right)_T. \end{aligned} \quad (35)$$

Now, either Eq. (19) or Eq. (21) can be used to explicitly calculate  $K_T^*$ , with the result expressed as a function of temperature and chemical potential.

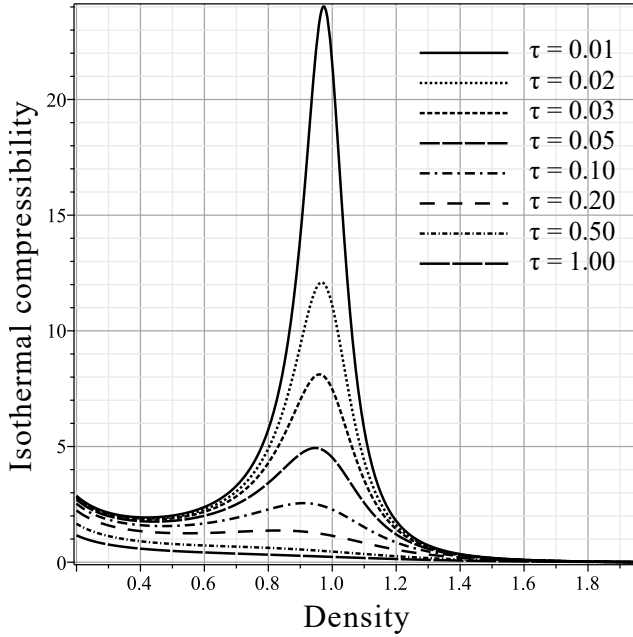
If it is preferable to use the equation of state in the form  $P = P(T, \rho^*)$ , see (25) and (27), then the most suitable expressions for the isothermal compressibility are

$$K_T = \frac{1}{\rho^*} \left( \frac{\partial P}{\partial \rho^*} \right)_T^{-1} \quad (36)$$

and

$$K_T^* = \frac{1}{\rho^*} \left( \frac{\partial P^*}{\partial \rho^*} \right)_T^{-1}. \quad (37)$$

Figure 2 illustrates the dependency of the compressibility  $K_T^*$  on the density  $\rho^*$  for various values of temperature above the critical one. The dependency of  $K_T^*$  on the effective chemical potential  $M$  is displayed in Fig 3.

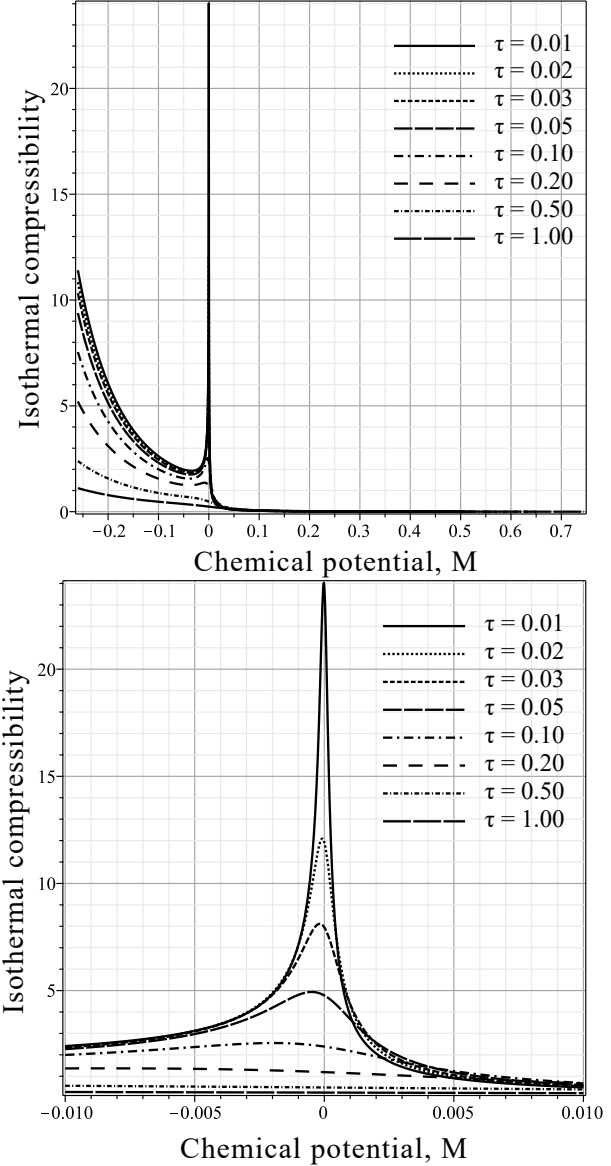


**Fig. 2.:** The reduced isothermal compressibility  $K_T^*$  as a function of the density  $\rho^*$  at different values of reduced temperature  $\tau > 0$  ( $T > T_c$ ).

#### 4.2. Thermal pressure coefficient

The thermal pressure coefficient is defined by

$$\beta_V = \left( \frac{\partial P}{\partial T} \right)_{V,N}. \quad (38)$$



**Fig. 3.:** The reduced isothermal compressibility  $K_T^*$  as a function of the effective chemical potential  $M$  for different temperatures  $\tau = (T - T_c)/T_c$  at  $T > T_c$ . The two figures differ in the scale of  $M$ . The left-hand-side figure covers a wider range of  $M$ , while the right-hand-side one focuses on a range of  $M$  around its critical value 0.

We rewrite  $\beta_V$  into a form that is suitable for the equation of state  $P = P(T, \mu)$ , see (4) and (19)

$$\beta_V = \left( \frac{\partial P}{\partial T} \right)_\mu + \left( \frac{\partial P}{\partial \mu} \right)_T \left( \frac{\partial \mu}{\partial T} \right)_{V,N}. \quad (39)$$

Applying the cyclic relation

$$\left( \frac{\partial \mu}{\partial T} \right)_{V,N} \left( \frac{\partial T}{\partial V} \right)_{\mu,N} \left( \frac{\partial V}{\partial \mu} \right)_{T,N} = -1$$

we obtain

$$\begin{aligned} \left( \frac{\partial \mu}{\partial T} \right)_{V,N} &= - \left( \frac{\partial V}{\partial T} \right)_{\mu,N} \left( \frac{\partial V}{\partial \mu} \right)_{T,N}^{-1} \\ &= - \left( \frac{\partial \rho^*}{\partial T} \right)_\mu \left( \frac{\partial \rho^*}{\partial \mu} \right)_T^{-1}. \end{aligned}$$

Substituting this result into the formula for  $\beta_V$ , we arrive at the final expression for the thermal pressure coefficient

$$\beta_V = \left( \frac{\partial P}{\partial T} \right)_\mu - \left( \frac{\partial P}{\partial \mu} \right)_T \left( \frac{\partial \rho^*}{\partial T} \right)_\mu \left( \frac{\partial \rho^*}{\partial \mu} \right)_T^{-1} \quad (40)$$

which<sup>2</sup> is easily calculated based on Eqs. (4) and (21). It is also worth noting that the first contribution to  $\beta_V$  is essentially the entropy per volume,  $S/V = (\partial P / \partial T)_\mu$ .

We introduce the reduced thermal pressure coefficient by

$$\beta_V^* = \frac{v}{k_B} \beta_V. \quad (41)$$

It is expressed in terms of reduced quantities as follows

$$\beta_V^* = \frac{1}{T_c^*} \left[ \left( \frac{\partial P^*}{\partial \tau} \right)_\mu - \left( \frac{\partial P^*}{\partial \mu} \right)_T \left( \frac{\partial \rho^*}{\partial \tau} \right)_\mu \left( \frac{\partial \rho^*}{\partial \mu} \right)_T^{-1} \right]. \quad (42)$$

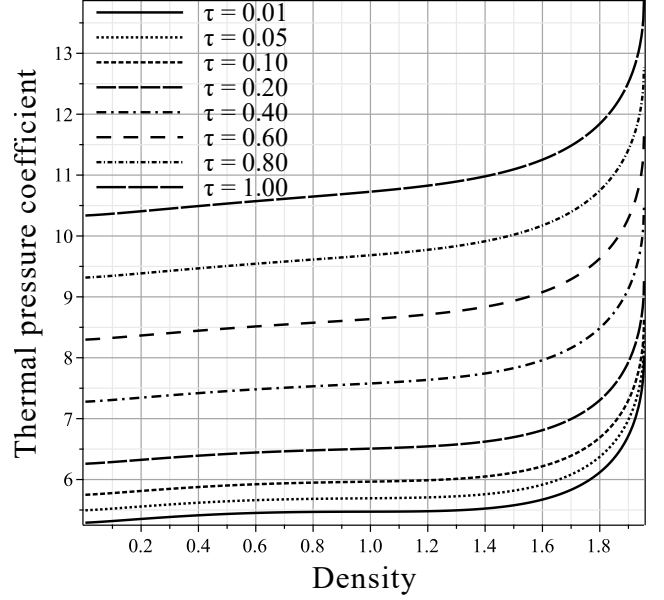
If it is preferable to use the equation of state in the form  $P = P(T, \rho^*)$ , see (25) and (27), then the most suitable expressions for the thermal pressure coefficient follow immediately from the definition (38)

$$\beta_V = \left( \frac{\partial P}{\partial T} \right)_\rho \quad (43)$$

<sup>2</sup> Compare this equation for  $\beta_V$  with Eq.(17) from [7]

and

$$\beta_V^* = \frac{1}{T_c^*} \left( \frac{\partial P^*}{\partial \tau} \right)_\rho. \quad (44)$$



**Fig. 4.:** The reduced thermal pressure coefficient  $\beta_V^*$  as a function of the density  $\rho^*$  at different values of relative temperature  $\tau > 0$  ( $T > T_c$ ).

Figure 4 shows the dependency of the pressure coefficient  $\beta_V^*$  on the density  $\rho^*$  for various values of temperature above the critical one. The dependency of  $\beta_V^*$  on the effective chemical potential  $M$  is displayed in Fig 5.

#### 4.3. Thermal expansion coefficient

The thermal expansion coefficient is defined by

$$\alpha_P = \frac{1}{V} \left( \frac{\partial V}{\partial T} \right)_{P,N}. \quad (45)$$

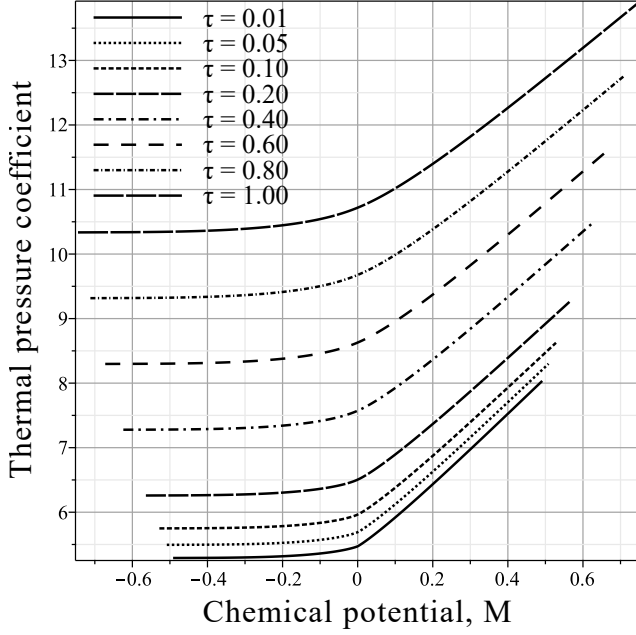
We rewrite  $\alpha_P$  into a form suitable for the equation of state  $P = P(T, \mu)$ , see (4) and (19)

$$\alpha_P = -\frac{1}{\rho} \left( \frac{\partial \rho}{\partial T} \right)_{P,N} \quad (46)$$

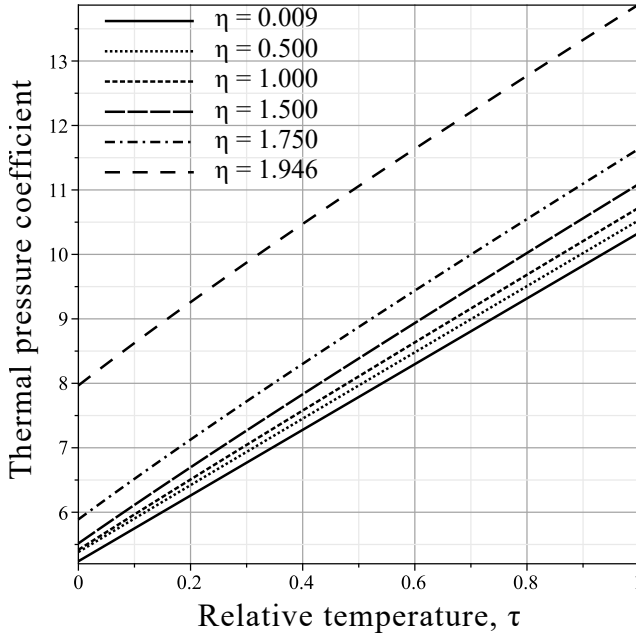
$$= -\frac{1}{\rho} \left( \frac{\partial \rho}{\partial T} \right)_\mu - \frac{1}{\rho} \left( \frac{\partial \rho}{\partial \mu} \right)_T \left( \frac{\partial \mu}{\partial T} \right)_{P,N}. \quad (47)$$

Applying the cyclic relation

$$\left( \frac{\partial \mu}{\partial T} \right)_{P,N} \left( \frac{\partial T}{\partial P} \right)_{\mu,N} \left( \frac{\partial P}{\partial \mu} \right)_{T,N} = -1$$



**Fig. 5.:** The reduced thermal pressure coefficient  $\beta_V^*$  as a function of the effective chemical potential  $M$  at different values of relative temperature  $\tau > 0$  ( $T > T_c$ ).



**Fig. 6.:** The reduced thermal pressure coefficient  $\beta_V^*$  as a function of the relative temperature  $\tau$  at different values of the density  $\rho^*$ .

we get

$$\left(\frac{\partial \mu}{\partial T}\right)_{P,N} = - \left(\frac{\partial P}{\partial T}\right)_\mu \left(\frac{\partial P}{\partial \mu}\right)_T^{-1}.$$

Substituting the result into the expression for  $\alpha_P$ , we arrive at the final formula for the thermal expansion coefficient

$$\alpha_P = -\frac{1}{\rho} \left(\frac{\partial \rho}{\partial T}\right)_\mu + \frac{1}{\rho} \left(\frac{\partial \rho}{\partial \mu}\right)_T \left(\frac{\partial P}{\partial T}\right)_\mu \left(\frac{\partial P}{\partial \mu}\right)_T^{-1} \quad (48)$$

which<sup>3</sup> is readily calculated based on Eqs. (4) and (21).

We introduce the reduced thermal expansion coefficient by

$$\alpha_P^* = \frac{\varepsilon}{k_B} \alpha_P. \quad (49)$$

It is expressed in terms of reduced quantities as follows

$$\alpha_P^* = \frac{1}{T_c^* \rho^*} \left[ - \left(\frac{\partial \rho^*}{\partial \tau}\right)_\mu + \left(\frac{\partial \rho^*}{\partial \mu}\right)_T \left(\frac{\partial P^*}{\partial \tau}\right)_\mu \left(\frac{\partial P^*}{\partial \mu}\right)_T^{-1} \right]. \quad (50)$$

If it is preferable to use the equation of state in the form  $P = P(T, \rho^*)$ , see (25) and (27), then we need to rewrite the definition for  $\alpha_P$  accordingly. We start with (46) and apply the cyclic relation

$$\left(\frac{\partial \rho}{\partial T}\right)_P = - \left(\frac{\partial P}{\partial T}\right)_\rho \left(\frac{\partial P}{\partial \rho}\right)_T^{-1} \quad (51)$$

to get for  $\alpha_P$

$$\alpha_P = \frac{1}{\rho} \left(\frac{\partial P}{\partial T}\right)_\rho \left(\frac{\partial P}{\partial \rho}\right)_T^{-1}. \quad (52)$$

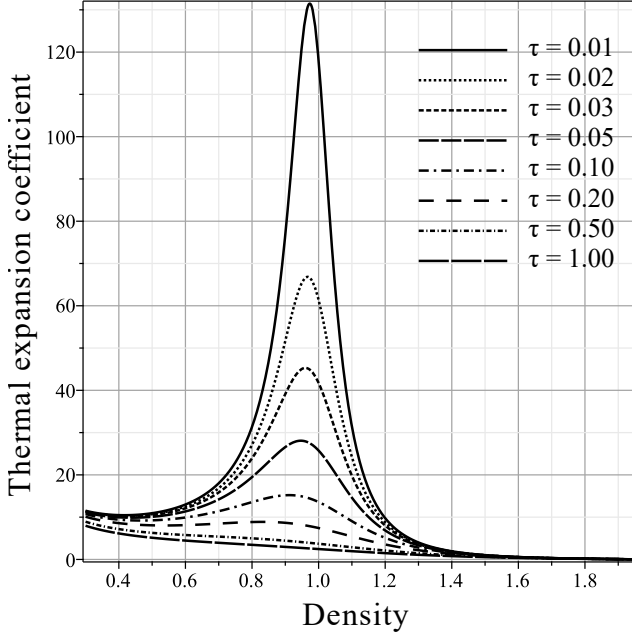
or for  $\alpha_P^*$

$$\alpha_P^* = \frac{1}{\rho^* T_c^*} \left(\frac{\partial P^*}{\partial \tau}\right)_\rho \left(\frac{\partial P^*}{\partial \rho^*}\right)_T^{-1}. \quad (53)$$

Figure 7 shows the dependency of the reduced thermal expansion coefficient  $\alpha_P^*$  on the density  $\rho^*$  for various values of temperature above the critical one. The dependency of  $\alpha_P^*$  on the effective chemical potential  $M$  is displayed in Fig. 8.

<sup>3</sup> Compare this equation for  $\alpha_P$  with Eqs.(57)-(58) from [7]





**Fig. 7.:** The reduced thermal expansion coefficient  $\alpha_P^*$  as a function of the density  $\rho^*$  at different values of relative temperature  $\tau > 0$  ( $T > T_c$ ).

#### 4.4. Relation between response functions

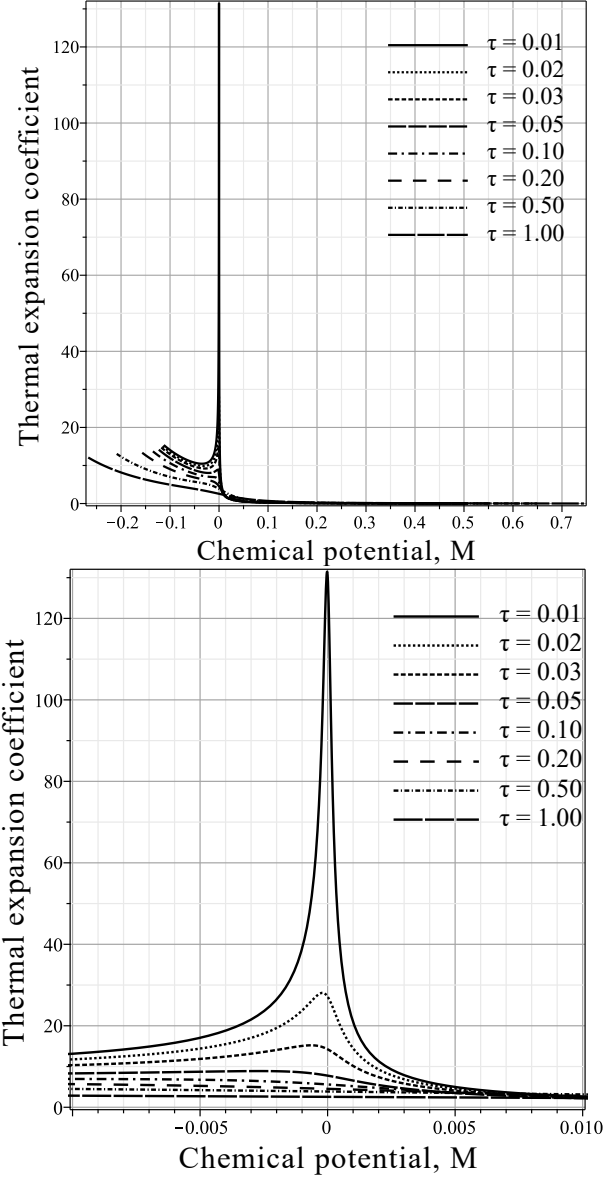
The following thermodynamic identity holds between the calculated response functions

$$\frac{\alpha_P}{K_T \beta_V} = \frac{\alpha_P^*}{K_T^* \beta_V^*} = 1, \quad (54)$$

which is easily derived from the cyclic relation between  $P$ ,  $V$  and  $T$ . In our case, this equality is exactly reproduced by substituting explicit expressions for  $K_T$ ,  $\alpha_P$ , and  $\beta_V$  obtained from the equation of state (4) (or (25)).

#### 5. Conclusions

Thermodynamic response functions, namely the isothermal compressibility, the thermal pressure coefficient, and the thermal expansion coefficient, are calculated for a many-particle system interacting through a modified Morse potential. The starting point for these calculations are the equation of state obtained for a cell fluid model within the framework of the grand canonical ensemble in our previous work [3]. The dependencies of the calculated response functions on the density and the effective chemical



**Fig. 8.:** The reduced thermal expansion coefficient  $\alpha_P^*$  as a function of the effective chemical potential  $M$  for different temperatures  $\tau = (T - T_c)/T_c$  at  $T > T_c$ . The two figures differ in the scale of  $M$ . The left-hand-side figure covers a wider range of  $M$ , while the right-hand-side one focuses on a range of  $M$  around its critical value 0.

potential are illustrated graphically. The thermodynamic identity (54) among these quantities is verified

to confirm the self-consistency of the performed calculations.

## Appendix A: Explicit expressions for derivatives

In this Appendix, we provide an explicit example of differentiating the equation of state, namely (27), by deriving the expression for the thermal pressure coefficient  $\beta_V^*$  from (44).

For the quantity  $\bar{M}$  from Eq. (23) one has

$$\left(\frac{\partial \bar{M}}{\partial \tau}\right)_{\rho^*} = \frac{k_B T_c \rho_n}{W(0)} \frac{1 - \omega_0}{1 + \tau \omega_0}, \quad (\text{A.1})$$

where

$$\omega_0 = \frac{\chi_0 + A_\gamma}{B - 1 + \chi_0}.$$

For  $\tilde{D}(0)$  from (12) one has

$$\frac{\partial \tilde{D}(0)}{\partial \tau} = -\frac{k_B T_c}{W(0)} \frac{1 - \omega_0}{1 + \tau \omega_0}. \quad (\text{A.2})$$

While calculating both derivatives, we take into account

$$\frac{\partial}{\partial \tau} [\beta W(0)] = -\frac{\beta_c \Phi^{(r)}(0)}{(1 + \tau)^2} (B - 1 - A_\gamma). \quad (\text{A.3})$$

The derivative of the quantity  $E_\nu$  from (26) with respect to temperature  $\tau$  is

$$\begin{aligned} \frac{\partial E_\nu}{\partial \tau} = & -\frac{1}{2} \left( \bar{M} - g_1 + \frac{g_3}{g_4} \tilde{D}(0) + \frac{g_3^3}{6g_4^2} \right)^2 \frac{\partial \beta W(0)}{\partial \tau} \\ & - \beta W(0) \left( \bar{M} - g_1 + \frac{g_3}{g_4} \tilde{D}(0) + \frac{g_3^3}{6g_4^2} \right) \\ & \times \left( \frac{\partial \bar{M}}{\partial \tau} + \frac{g_3}{g_4} \frac{\partial \tilde{D}(0)}{\partial \tau} \right) \\ & - \frac{g_3}{g_4} \frac{\partial \bar{M}}{\partial \tau} - \frac{g_3^2}{2g_4^2} \frac{\partial \tilde{D}(0)}{\partial \tau}. \end{aligned} \quad (\text{A.4})$$

Collecting these all formulas together, we can explicitly calculate  $\beta_V^*$  from (44)

$$\begin{aligned} \beta_V^* = & E_\eta(\rho^*, T) + \bar{M} \rho_n + \frac{\tilde{D}(0)}{2} \rho_n^2 - \frac{a_4}{24} \rho_n^4 \\ & + (1 + \tau) \left( \frac{\partial E_\nu}{\partial \tau} + \frac{\partial \bar{M}}{\partial \tau} \rho_n + \frac{\rho_n^2}{2} \frac{\partial \tilde{D}(0)}{\partial \tau} \right) \end{aligned} \quad (\text{A.5})$$

## Acknowledgement.

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ТЕРМОДИНАМІЧНІ ФУНКЦІЇ ВІДГУКУ КОМІРКОВОЇ МОДЕЛІ

Текст резюме укр. мовою.

Ключові слова: комірка модель, потенціал Морзе, термодинамічні функції відгуку.