

Lab4

Objectives

- BSSE
- Understanding basis sets
- Thermochemistry

Basis set superposition error (BSSE)

Gaussian input file for water dimer:

```
%chk=water.chk
```

```
%nprocshared=2
```

```
# hf/3-21g counterpoise=2
```

Title Card Required

```
0 1 0 1 0 1
```

| | | | |
|---------------|-------------|-------------|------------|
| O(Fragment=1) | -1.97619047 | -1.11904760 | 0.00000000 |
| H(Fragment=1) | -1.01619047 | -1.11904760 | 0.00000000 |
| H(Fragment=1) | -2.29664505 | -0.21411177 | 0.00000000 |
| O(Fragment=2) | 0.80952381 | -0.38095238 | 0.00000000 |
| H(Fragment=2) | 1.76952381 | -0.38095238 | 0.00000000 |
| H(Fragment=2) | 0.48906922 | 0.52398346 | 0.00000000 |

Basis set superposition error (BSSE)

```
$ grep A.U. water.log
```

| | | | | |
|-----------|----------|----------------|------------|-----------|
| SCF Done: | E(RHF) = | -151.183376551 | A.U. after | 10 cycles |
| SCF Done: | E(RHF) = | -75.5871648459 | A.U. after | 9 cycles |
| SCF Done: | E(RHF) = | -75.5908152208 | A.U. after | 9 cycles |
| SCF Done: | E(RHF) = | -75.5858099772 | A.U. after | 10 cycles |
| SCF Done: | E(RHF) = | -75.5858099773 | A.U. after | 10 cycles |

```
grep -A 4 'Counterpoise ' water.log
```

| | |
|---------------------------------|-----------------------------|
| Counterpoise corrected energy = | -151.177016439265 |
| BSSE energy = | 0.006360112084 |
| sum of fragments = | -151.171619954583 |
| complexation energy = | -7.38 kcal/mole (raw) |
| complexation energy = | -3.39 kcal/mole (corrected) |

Hands-on

1. Compute the BSSE for the water dimer
2. Determine the BSSE for the water dimer for several distances between the water molecules
3. Determine BSSE for the water dimer as a function of basis set size?
(Use HF method in all cases)

Hands-on

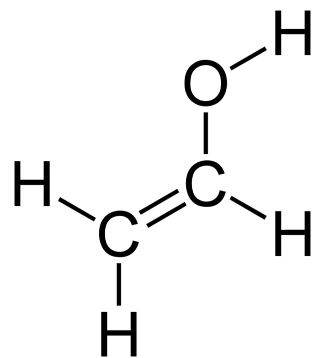
1. Compute the number of primitive gaussians for H₂O and CH₄ monomer for the following basis sets
 - a. STO-3g
 - b. 3-21g
 - c. 6-31g
 - d. 6-31g(d)
 - e. 6-31g(d,p)
 - f. 6-111g(d,p)

Vibrational spectroscopy

- Number of frequencies
- Visualization of normal modes using Jmol
- Number of imaginary frequencies
- IR intensities

Hands-on

- Find two isomers of vinyl alcohol and the transition state between them.



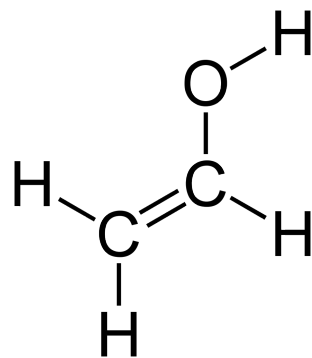
Thermochemistry

- Enthalpy and free energy

<https://gaussian.com/wp-content/uploads/dl/thermo.pdf>

Hands-on

- Check the stability of vinyl alcohol isomers as a function of temperature



References

1. <https://avogadro.cc/>
2. <https://gaussian.com/gaussian16/>
3. <https://gaussian.com/wp-content/uploads/dl/thermo.pdf>