

Section 2.2 Homework

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16th April 2024

Notice of ADA Accommodation and Methods

I have an ADA accommodation to do my assignment on paper. This document is a utilization of that accommodation. This assignment will utilize questions from the textbook, *Chemistry: Atoms First, 2e*, to practice the skills and learning objectives for this class.

Learning Objective 2.2.1 & 2.2.2: Chapter 2.4

2.4.37: Write a sentence that describes how to determine the number of moles of a compound in a known mass of the compound if we know the molecular formula

If we know the mass and the molecular formula of a sample, we can determine its moles by deriving the molar mass via atomic mass units and dividing the mass of the sample by its molar mass.

2.4.39: Which contains the greatest mass of oxygen: 0.75 mol of ethanol, 0.60 mol of formic acid, or 1.0 mol of water?

0.60 mol of formic acid contains 1.20 mol of oxygen, which is more than any other chemical provided.

2.4.41: How are the molecular mass and molar mass of a compound related?

The molecular mass of a compound is the mass in amu of the molecule, while the molar mass is the mass of one mole of the compound in grams. Generally, the numbers of these are the same, with different units.

2.4.43: Calculate the molar mass of each of the following:

2.4.43.a: S_8

256.48 g

2.4.43.b: C_5H_{12}

72.146 g

2.4.43.c: $Sc_2(SO_4)_3$

378.10 g

2.4.43.d: CH_3COCH_3

58.078 g

2.4.43.e: $\text{C}_6\text{H}_{12}\text{O}_6$

180.156 g

2.4.47: Determine the mass for each of the following:

2.4.47.a: 0.0416 mol KOH

2.33 g

2.4.47.b: 10.2 mol C_2H_6

306 g

2.4.47.c: 1.6×10^{-3} mol Na_2SO_4

0.20 g

2.4.47.d: 6.854×10^3 mol $\text{C}_6\text{H}_{12}\text{O}_6$

$1,235 \times 10^6$ g

2.4.47.e: 2.86 mol $\text{Co}(\text{NH}_3)_6\text{Cl}_3$

765 g

Learning Objective 2.2.3: Chapter 6.1

6.1.1: What is the total mass in amu of carbon in each of the following?

6.1.1.a: CH_4

12.01 amu

6.1.1.b: CHCl_3

12.01 amu

6.1.1.c: $\text{C}_{12}\text{H}_{10}\text{O}_6$

144.12 amu

6.1.1.d: $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$

60.05 amu

6.1.3: Calculate the molecular or formula mass of each of the following:

6.1.3.a: P_4

123.88

6.1.3.b: H_2O

18.016

6.1.3.c: $Ca(NO_3)_2$

164.10

6.1.3.d: CH_3CO_2H

60.052

6.1.3.e: $C_{12}H_{22}O_{11}$

338.292

Learning Objective 2.2.4 & 2.2.5: Chapter 6.2

6.2.7: What information is needed to determine the molecular formula from the empirical formula?

The molecular formula can be obtained from the empirical formula if the molar mass is known, or the number of moles of sample and mass of sample.

6.2.9: Determine the following to four sigfigs:

6.2.9.a: Percent Composition of HN_3

H: 2.342% ; N: 97.66%

6.2.9.b: Percent Composition of $C_6H_2(CH_3)(NO_2)_3$

C: 37.01% ; H: 2.219% ; N: 18.50% ; O: 42.26%

6.2.9.c: Percent of SO_4^{2-} in $Al_2(SO_4)_3$

84.23%

6.2.11: Determine the percent water in $CuSO_4 \cdot 5H_2O$ to three sigfigs

36.1%

6.2.15: Dichloroethane, a compound that is often using for dry cleaning, contains carbon (24.3%), hydrogen (4.1%) and chlorine with a molar mass of 99 g/mol

$C_2H_4Cl_2$

Learning Objective 2.2.6: Chapter 6.3

6.3.19: Explain what changes and what stays the same when 1.00L of a solution of NaCl is diluted to 1.80L

The number of moles of salt stays the same, while the molarity and concentration changes.

6.3.21: A 200-mL sample and a 400-mL sample of a solution of salt have the same molarity. How are they related?

The samples both have the same concentration, but the 400 mL sample has twice the amount of salt as the 200 mL solution.

Learning Objective 2.2.7 - 2.2.9: Chapter 7.3

Learning Objective 2.2.10 - 2.2.12: Chapter 7.4

Learning Objective 2.2.13 - 2.2.14: Chapter 7.5

Learning Objective 2.3.1 - 2.2.5: Chapter 9.1

Learning Objective 2.3.6 - 2.3.7: Chapter 9.2

Learning Objective 2.3.8 - 2.3.12: Chapter 9.3

Learning Objective 2.3.13 - 2.3.14: Chapter 9.4