Hydrogen Atom
$$\mathcal{L}_{nlm} = R(r) Y_{l}^{m}(\theta, \phi)$$

$$R(r) = \frac{u(r)}{r}$$

$$\rho = Kr \qquad K = \frac{J-2mE}{\hbar} \qquad E < 0$$

$$u(\rho) = \rho^{l+1} e^{-\rho} v(\rho)$$

$$u(\rho) = \sum_{j=0}^{l+1} C_{j} \rho^{j}$$

$$C_{j+1} = \frac{2(j+l+1-n)}{(j+l+2l+2)} C_{j} \qquad n=1,2,3,---$$

$$E_{n} = \frac{E_{2}}{n^{2}} \qquad E_{2} = -13.6eV$$

$$N = 1$$

$$\int_{max} \int_{a=0}^{n} \int_{a=0}^{n}$$

General

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$$V_{nelm} = \sqrt{\frac{2}{n\alpha}} \frac{3(n-l-1)!}{2n \left[(n+l)!\right]^3} e^{-r/n\alpha} \left(\frac{2r}{n\alpha}\right)^l \left(\frac{2l+1}{n\alpha}\right) \sqrt{\frac{ar}{n\alpha}} \left(\frac{ar}{n\alpha}\right) \sqrt{\frac{ar}{n\alpha}} \sqrt{\frac{ar}{n\alpha}} \left(\frac{ar}{n\alpha}\right) \sqrt{\frac{ar}{n\alpha}} \left(\frac{ar}{n\alpha}\right) \sqrt{\frac{ar}{n\alpha}} \sqrt{\frac{ar}{n$$

p has me-1,0,1

3 x 2 ° 6 states

1 has me-2,-1,1,2

5 x 2 = (0) states

$$\begin{bmatrix} x_1 x_1^2 = 0 \\ x_2 x_3 = 0 \end{bmatrix} = \vec{r} \times \vec{p}$$

$$\begin{bmatrix} x_1 x_1 = 0 \\ x_2 x_3 = 0 \end{bmatrix}$$

$$L_{x} = yp_{z} - 2p_{y} = -ih(yJz - 2J_{y})$$

$$L_{y} = zp_{x} - xp_{z} \qquad \text{etc}$$

$$L_{z} = xp_{y} - yp_{x} \qquad \text{etc}$$

$$[L_{x}, L_{y}] = i \hbar L_{z} [L_{y}, L_{z}] = i \hbar L_{x} [L_{z}, L_{x}] = i \hbar L_{y}$$

$$[L^{2}, L_{x}] = 0 = [L^{2}, L_{y}] = [L^{2}, L_{z}]$$
Suppose  $f$  is an eigenfunction of  $L^{2}$  &  $L_{z}$ 

$$L^{2}f = \lambda f \qquad L_{z}f = \mu f$$

Define 
$$L\pm = L_x \pm iL_y$$
  

$$[L\pm, L^2] = 0$$

$$[L_z, L\pm] = \pm \pm L\pm$$

Now  $L_{\pm}f$  is an eigenfunction of  $L^2$  with eigenvalue  $\lambda$   $L^2(L_{\pm}f) = L_{\pm}L^2f = L_{\pm}(\lambda f) = \lambda(L_{\pm}f)$ 

Lif is an eigenfunction of Le with eigenale juth.