Chapter 11: Nuclear Physics

11.1. Basic Structure

- What is a nucleon?
- What do N, Z, and A signify?
- What varies between two isotopes of the same atom?
- What is the difference between stable and unstable nuclei?
- What is the atomic mass unit?
- What did Rutherford discover when he shot alpha particles at gold, and what did it prove?
- What is the rough formula for the size of a nucleus?
- How does the mass density of a nucleus depend on its atomic number?

• 11.2. Binding

- Why don't the protons inside a nucleus explode outward due to electrostatic repulsion?
- How does the strong force depend on distance between nucleons?
- How has the strong force's existence been proven?
- Why can a proton and a neutron bond together (into a deuteron), but two protons or two neutrons cannot?

Arbitrary Nucleon Number

- How does the strong force change as you increase the number of nucleons in the nucleus?
- What effect does Coulomb repulsion have on the stability of nuclei?
- Why must nuclei include protons? Why not neutrons only?
- Why do smaller nuclei have equal numbers of protons and neutrons, but larger nuclei have more neutrons?

• Stability: The Experimental Truth

- How do we calculate the binding energy of a nucleus"
- What is the curve of stability?
- Which element(s) has the largest binding energy per nucleon? Overall binding energy?

11.3 Nuclear Models

Liquid Drop Model

- Why is a water drop like an atomic nucleus?
- What quantity does this model predict?

- Why does this model include a *volume term*? How does it depend on A? (Why doesn't it go as A^2?) Why is it positive?
- Why does it include a surface term? What does it look like?
- Explain the *Coulomb* term.
- Why is there an asymmetry term? How does it depend on N and Z?
- Where do the coefficients come from?
- Is the liquid drop model a good fit to experiment?
- How is the shell model different from the liquid drop model?
- What are magic numbers?
- What is the pairing effect?

11.5 Radioactivity

- What happens during radioactive decay?
- Why does an atom decay? How does it "benefit"?

Alpha Decay

- What is an alpha particle?
- What happens to N, Z, and A during alpha decay?
- Why do nuclei like to emit helium nuclei? Why not single protons, or lithium nuclei?
- Why does the alpha particle end up with most of the kinetic energy created by the decay?

Beta Decay

- What is beta decay?
- What happens to N, Z, and A during beta decay?
- How can a *nucleus* emit an electron, if there are no electrons in the nucleus?
- What is a *neutrino*, and why is it necessary to explain beta decay?
- What happens in electron capture?
- Are neutrons stable?
- What kinds of nuclei prefer alpha decay? Beta- decay? Beta+ decay or electron capture?
- What is a decay series?
- What is gamma decay, and why does it occur?

• What is *spontaneous fission*? What are the end-products?

11.6 Radioactive Decay Law

- Why do we say that radioactive decay is probabilistic?
- Derive the radioactive decay law
- What is the *decay constant*, the *decay rate*, and the *half-life*?
- How does radioactive dating work?

11.7 Nuclear Reactions

• What is a nuclear reaction, and how does it differ from radioactive decay?

Fission

- What is nuclear fission?
- What kinds of nuclei will undergo it?
- Why is it energetically favorable?
- When does it occur?
- How does a *chain reaction* occur? Why doesn't it occur in nature (very often)?
- What is the difference between an uncontrolled and a controlled chain reaction? How is control
 maintained?
- Why do countries that want to create nuclear fission need centrifuges?
- What does a moderator do?
- What is a breeder reactor?

Fusion

- What is *nuclear fusion*?
- What kinds of nuclei will undergo it?
- What is the proton-proton cycle?
- What is the carbon cycle? Why is it faster than the proton-proton cycle?
- Why can't fusion occur spontaneously? (i.e. Why is it so freaking hard to create controlled nuclear fusion on Earth?)
- What two things does hydrogen need to fuse? How are these achieved?
- What two things make fusion preferable to fission?