## Muti-Electron Wavefunctions

For atoms with more than one electron, the wavefunction depends on the sportial and spin coordinates of each electron eg- for a two-electron system:

4(c, s, ; [2, s2; t) The wavefunction also depends

Electrons are recessarily indistinguishable. So it we swap spin and positions of two electrons, an measureable quantities should remain unchanged.

Consider an operator whose effect swaps two electrons when the operator ô acts on a wavefurction with two electrons, it can only produce the same wavefurction with an unknown phase change. Acting twice produces the without wavefurction with some within phase.

so 
$$\hat{O}\Psi = \lambda \Psi$$
  $\hat{O}\hat{O}\Psi = \hat{O}\lambda \Psi = \hat{\lambda}(\lambda \Psi) = \lambda^2 \Psi = \Psi$   
 $\Rightarrow \lambda^2 = 1$ 

I here is the eigenvalue of particle exchange.

It is +1 the particles are called bosons.

If it is -1, the particles are called fermions.

Relativistic QM tells us we get +1 for integer spir particles and -1 for hulf integer spin particles.

80 the electrons give -1 and one fermious,

Now let's think about two more operators. One operator that exchanges position of two electrons but not spin. And another that exchanges spin of two electrons but not position.

We find that to a good well of approximation, reither of these operators change the namiltanian, so they both commute with the namiltanian. We can thus write energy eigenstates which are also eigenstates of the operators. As with the particle exchange operator, applying either operator take must give the initial wavefunction, so eigenvalues must be ±1. We find it the position swap operator has +1, the spin swap operator must have -1 and vice versa.

It is convenient to use single-electron orbitals as a complete set of states, i.e quantum numbers N.M. So to our two electron atom, every eigenstates are:

these are just the 1 we had in angular man. section.

Using M. to second electron

4(C, S, ; C2, S2; t) = \( \int \alpha\_{\lambda,m,s,s} \Pa(Ci) \Pi\_m(Cz) \ls, \ls\_2 \)

hie said that with some approximation, these are also eigenstates of the spatial endage and spin exchange operato. The approximation we make is to ignore interaction between spins and positions. We can factorise the above into the eigenstates of the two operators:

VEE, S. ; Ez, Sz; E) = Zan, MY(E) 4M(Ez) > bs, sz 15, > 15)

i.e. anmisisz = anim bsisz

so lbs., sell is probability of electrons having spin s. Sz

lan, m/2 is probability of electrons being in orbitals n. M

## Spir-States of Multi-Electron Systems

For a two electron system, the conducte basis of spir states is:

However, these states we not all eigenstates of the spireschange operator.

A complete set which are expertented of the spin exchange operator is:  $|\uparrow\rangle|\uparrow\rangle$   $\frac{1}{\sqrt{2}!}(|\uparrow\rangle|\downarrow\rangle + |\downarrow\rangle|\uparrow\rangle$   $\frac{1}{\sqrt{2}!}(|\uparrow\rangle|\downarrow\rangle - |\downarrow\rangle|\uparrow\rangle\rangle$ 

(called ever) and the last one has -1 (called odd).

so, using our approximation from the previous page, the first three have -1 under spatial exchange operator and the last one has +1.

The total spin is the vector sum of the spins of each electron.  $\hat{S} = \hat{s}_1 + \hat{s}_2$ 

Is with spin itself, we const simultaneously know every compared of the total spin vector but we can know  $\hat{\mathbb{S}}^2$  and one component, say  $\hat{\mathbb{S}}_2$ 

The speciator for 2 component of total spin is the sum of the 2 components of the individual spin:

Sz = Siz + Siz

We can therefore work out a quantum number associated with the 2 component of the total spir to each of the four spir states we listed.

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 $\hat{S}_{\epsilon}$   $|\uparrow\rangle|\uparrow\rangle = (\hat{S}_{i\epsilon} + \hat{S}_{i\epsilon}) |\uparrow\rangle|\uparrow\rangle = \hat{S}_{i\epsilon}|\uparrow\rangle|\uparrow\rangle + |\uparrow\rangle\hat{S}_{2=}|\uparrow\rangle$   $= t_{i}|\uparrow\rangle|\uparrow\rangle$ 

We say that this is quantum number 1. The other states have quantum number 0, 0, -1 in that order