25 What are major difference between metal & non metal

Metals	Non metals
1. Have strong tendency to lose Electrons to form cation	1. Non metals have a strong tendency to accept Electrons to form anions
2. Metals are strong reducing agents	2. Non metal are strong Reducing agent.
3. Metals have low Ionization Enthalpies	3. Non metals have high Ionization Enthalpies
4. Metals form basic oxides and ionic compound	4. Non metals form acidic oxides and covalent compond.

26 Use periodic table and find out
(i) Identify the Element with 5 electrons in outer subshell
(ii) Identify the Element that would tend to lose two Electrons
(iii) Identify the Element that belong tend to gain two Electrons.
(i) Elements belonging to Nitrogen family (group 15) nitrogen
(ii) Elements belonging to Alkaline Earth family (group 2) Magnesium
(iii) Element belonging to Oxygen family (group 16) Oxygen

28 Write the general Electronic configuration of s-p-d- and f-block Elements?
ans (i) s-Block Elements: ns^{1-2} where $n = 2-7$
(ii) p-Block Elements: ns^2np^{1-6} where $n = 2-6$
(iii) d-Block element $(n-1)d^{1-10}ns^{0-2}$ where $n = 4-7$
(iv) f-block elements $(n-2)f^{0-14}(n-1)d^{0-1}ns^2$ where $n = 6-7$

29 The increasing order of reactivity among group I elements is Li < Na < K < Rb < Cs whereas that of group II is F > Cl > Br > I Explain?
ans The Elements of Group I have only one Electron in their respective valence shells and thus have a strong tendency to lose this Electron. The tendency to lose Electron in turn, depends upon the ionization Enthalpy. It is linked with Electronegativity. Since both of them decrease down the group, the reactivity decreases.

30 Assign the position of elements having outer Electronic Configuration,
(i) ns^2np^4 for $n = 3$
(ii) ns^2 for $n = 4$
(iii) $(n-2)f^7(n-1)d^1ns^2$ for $n = 6$ in periodic table?
ans (i) p-block, group-16, 3-rd period } Sulphur
(ii) 4th period, group 4 } Titanium
(iii) 6th period, f block, group 3 } Gadolinium

31 Choose the Incorrect option
(b) ans (b) The d-block has 8 columns, because a maximum of 8 Electrons can occupy all the orbitals in a d-subshell.