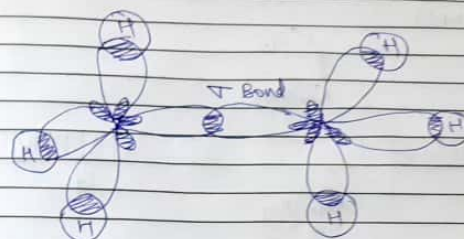
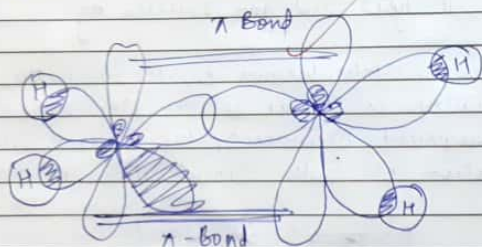


C₂H₄



- 3) (i) $\angle \text{HOM} = 104.45^\circ$
 (ii) $\angle \text{HNM} = 107.8$

As we can see the difference and As VSEPR Theory states lone pair lone pair > bond pair lone pair > bond pair - bond pair.

As NH_3 has only one lone pair so the Repulsion is comparatively less from H_2O as they have two lone pairs.

