33 In modern periodic Table, the period indicate
(c) principal quantum number

34 Predict the formulas of the stable Binary compound, that would formed by the combination
ans (a) Li_2O
(b) Mg_3N_2
(c) MI_3
(d) SiO_2
(e) PF_5
(f) LuF_3

35 Anything that influence the valence Electrons will affect the Chemistry of Element.
ans (c) Nuclear mass

36 The size of iso electronic species F^-, Ne and Na^+ is affected by
ans (a) Nuclear charge

37 Which of the following statement is incorrect in relation to Ionization Enthalpy?
(d) Removal of electron from orbitals bearing lower n value is easier than form orbital having higher n value.

20 Dec

1) BeCl is a polar compound yet $BeCl_2$ is a non polar compound,
2) Show σ and π bond formation in C_2H_6, C_2H_4, C_2H_2 Draw orbital diagrams and hybridisation
3) Account for discrepancy in bond angle
i) $\angle HOH$ in water
ii) $\angle HNH$ in NH_3
4) compare the bond order and relative stability of
i) O_2, O_2^-, O_2^{2-}, O_2^{2+}, O_2^+
ii) N_2, N_2^+
Predict the magnetic character
5) Define Resonance, draw the resonance structure of
i) CO_3^{-2} ii) CO_2 iii) N_3^-
6) Define Hybridisation i) SF_6
7) In PCl_5 are all the P-Cl bond same?
8) Define Hydrogen bond give condition for formation of H bond.
9) Discuss it types and give suitable eg.

1) $BeCl_2$ is non polar because of the symmetric structure it's structure is linear so the Dipole moment of the compound becomes 0. the force of attraction of electron of both chlorine atom neutralises.

Cl ⟵ Be ⟶ Cl
Electro negativity

So Be-Cl polar bond and overall it is non-polar bond.