

## ***Chapter 9 Stoichiometry Guided Reading And Study Workbook Answers***

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Chapter 9: Section 1 – Introduction to Stoichiometry Guided Reading 1. Define: Reaction Stoichiometry – 2. Reaction Stoichiometry is based on \_\_\_\_\_ and the Law of \_\_\_\_\_. What do all reaction stoichiometry calculations have to start with? 3.

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From above we can see that if we have 12.4 mol H<sub>2</sub> we need 4.13 mol N<sub>2</sub>. We don't have that much N<sub>2</sub> so the .892 mol of N<sub>2</sub> must be the limiting reagent. We can now determine how much

ammonia will be produced using the mole ratio in the balanced equation :

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