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The electron density around each nucleus is, for a moment, greater in one region of each cloud. Each molecule forms a temporary dipole. 394 Chapter 13 States of Matter. When temporary dipoles are close together, a weak dispersion force exists between oppositely charged regions of the dipoles, as shown in Figure 13-8.

Chapter 13: States of Matter - Jayne Heier

Chapter 13 States of Matter 137. SECTION 13.1 THE NATURE OF GASES (pages 385–389) This section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the particles of a substance. Kinetic Theory and a Model for Gases (pages 385–386)

Name Date Class STATES OF MATTER 13

13.1 The Fluid States 300 States of Matter FIGURE 13–1 The ice cube, a solid, has a definite shape. But water, a fluid, takes the shape of its container.

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Chapter 13 States of Matter - Chapter 13 "States of... Section 13.2 The Nature of Liquids OBJECTIVES: Describe the equilibrium between a liquid and its vapor. Section 13.2 The Nature of Liquids OBJECTIVES: Identify the conditions at which boiling occurs. Section 13.2 The Nature of Liquids Liquid particles are also in motion.

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Chapter 13 States of Matter notes. So $F/A = F'/A'$. A small force applied over a small area can give a bigger force over a bigger area like in a hydraulic system. The pressure you feel under water is due to the weight of the water over you. For a uniform fluid this depends on the density and the depth. $P = \rho hg$ where is ρ the density of the fluid.

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Name CHAPTER 13 Section 13.1 continued Date Class STUDY GUIDE FOR CONTENT MASTERY In your textbook, read about gas pressure. Circle the letter of the choice that best completes the statement or answers the question. 13. Pressure is defined as force per unit c. time. a. area. b. mass. d. volume. 14.

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Chapter 13- The States of Matter Gases- indefinite volume and shape, low density. Liquids- definite volume, indefinite shape, and high density. Solids- definite volume and shape, high density Solids and liquids have high densities because their molecules are close together.

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Chapter 13 Concept Map: ... Most of the states of matter are pretty steady, but solids have two different type of solids. Notice how above, the graph says a solid is packed orderly? This is recognizing the crystal structure of a solid. Most solids are crystal, which means the particles are arranged in a repeating, 3D pattern.

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Prentice Hall Chemistry Chapter 13: States of Matter ...

Chapter 13 - States of Matter - 13.4 Changes of State - 13.4 Lesson Check - Page 439: 27 Answer One useful application of sublimation is the use of dry ice as a coolant in the transportation of frozen goods such as ice cream.

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