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Chemistry Live! – Worked Solutions. 76. Workbook Chapter 9 – The Mole Concept. W9.1 (a) 1 mole of Li atoms = 7 g. (b) 1 mole of Na atoms = 23 g. (c) 1 mole of Ca atoms = 40 g. (d) 1 mole of Fe atoms = 56 g. (e) 1 mole of Ag atoms = 108 g.

Workbook - Worked Solutions Chapter 4 - The Periodic ...

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Chemistry Live! – Worked Solutions 2 Chapter 9 - The Mole Concept 9.1 Since a sulfur atom ($A_r = 32$) weighs twice as much as an oxygen atom ($A_r = 16$), 32 g of sulfur atoms and 16 g of oxygen atoms both contain the same number of atoms.

Worked Solutions The Periodic Table ... - Chemistry Live

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Chemistry Live! – Worked Solutions 2 Chapter 9 - The Mole Concept. 9.1. Since a sulfur atom ($A_r = 32$) weighs twice as much as an oxygen atom ($A_r = 16$), 32 g of sulfur atoms and 16 g of oxygen atoms both contain the same number of atoms. Answer: 32 g. 9.2 (a) 1 mole of H atoms = 1 g 1 mole of H_2 molecules = $2 \times 1 = 2$ g.

Textbook - Worked Solutions Chapter 4 - The Periodic Table ...

is called __ chemistry __. (e) An experiment is carefully planned so that it is a _ fair _ test with a single, clear result. (f) To ensure a fair comparison, a __ control __ experiment is carried out at the same time.

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