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(b) The molecular formula, CH<sub>4</sub>, gives the number of each kind of atom. Follow steps 1–3 as given

in the student textbook. Step 1 Carbon has the lower electronegativity in the molecule and will be the central atom. Step 2 Find the total number of valence electrons.  $(1 \text{ atom C} \times 4e-/C) + (4 \text{ atoms H} \times 1e-/H) = 8e-$  Determine the total number of electrons required for a noble gas

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