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Group - E

Q. 3: Compare the properties of covalent and ionic compounds after defining the two bonds involved

Covalent:

Chemical compounds formed by sharing of electron between two atoms are called covalent compound. The bond is ~~called~~ that works between shared electron are called covalent bond. The properties of covalent compounds are stated below.

- ① Boiling point and melting points are relatively low.
- ② Poor Hardness
- ③ Does not dissolve in water.
- ④ Poor electrical conductivity.

Ionic:

Chemical compounds formed by transferring electron from one atom to another are called ~~the~~ ionic compounds. The bond which forms when transferring electron between two atoms are called ionic bond. Properties of ionic compounds are stated below.

- (i) High melting and boiling point
- (ii) Relatively High Hardness
- (iii) can dissolve in water
- (iv) conducts electricity and heat very well.
- (v) Bond energy is High, so Hard to ~~separate~~ break the compound.