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(54) **SYSTEM FOR FORMING JANUS PARTICLES**

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(52) **U.S. Cl.** **516/22; 428/403**

(58) **Field of Classification Search** **516/22**
See application file for complete search history.

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(57) **ABSTRACT**

The invention is a method of forming Janus particles, that includes forming an emulsion that contains initial particles, a first liquid, and a second liquid; solidifying the first liquid to form a solid that contains at least a portion of the initial particles on a surface of the solid; and treating the exposed particle sides with a first surface modifying agent, to form the Janus particles. Each of the initial particles on the surface has an exposed particle side and a blocked particle side.

14 Claims, 11 Drawing Sheets

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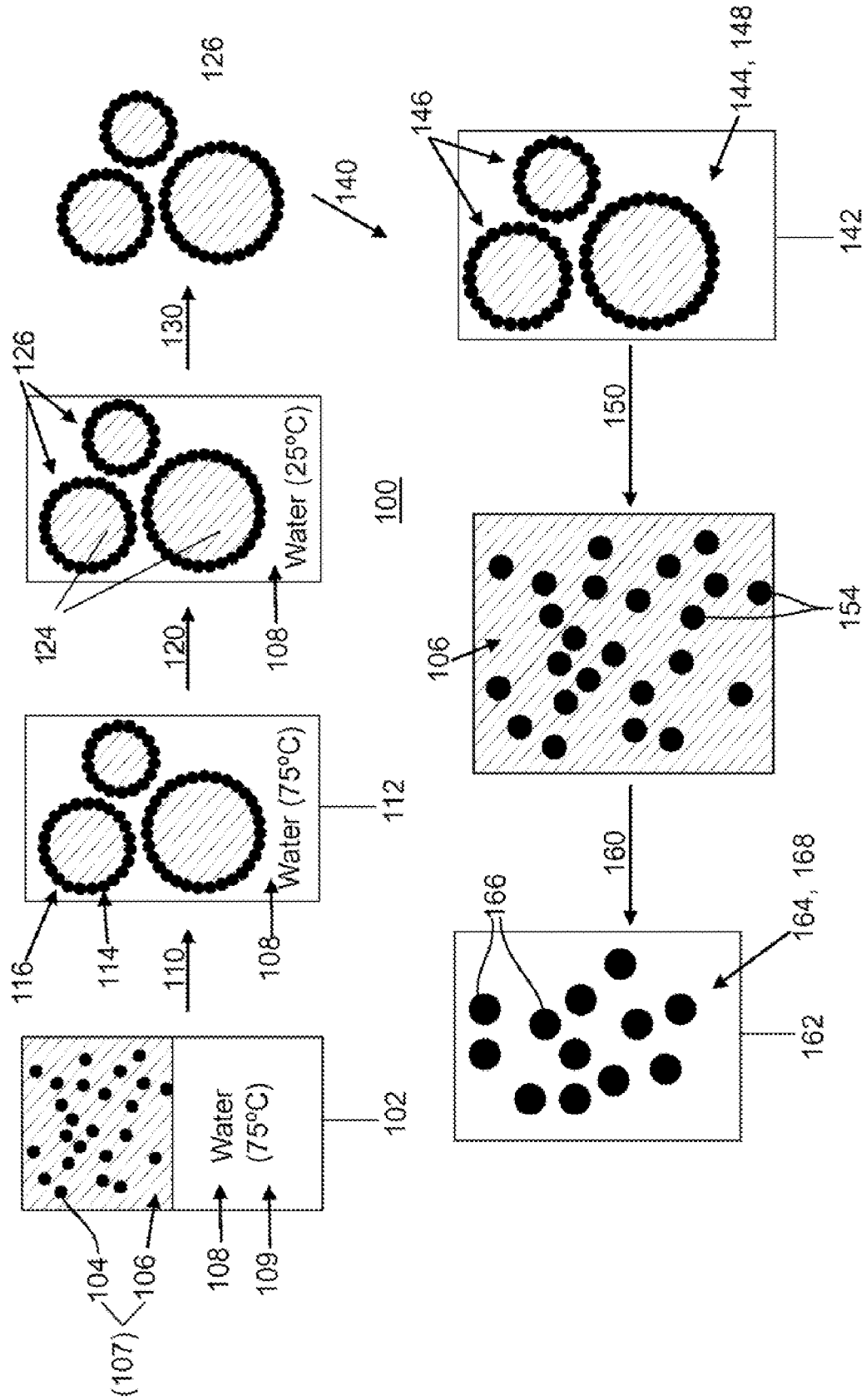


Fig. 1

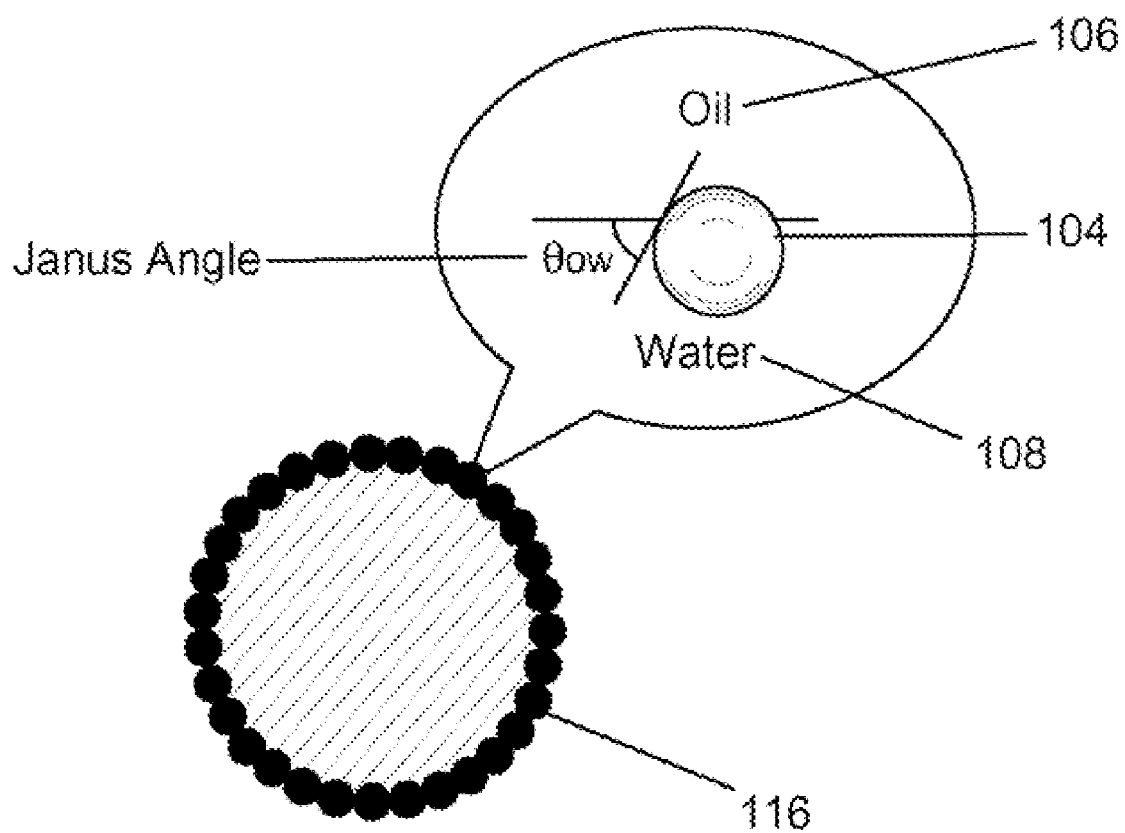


Fig. 2

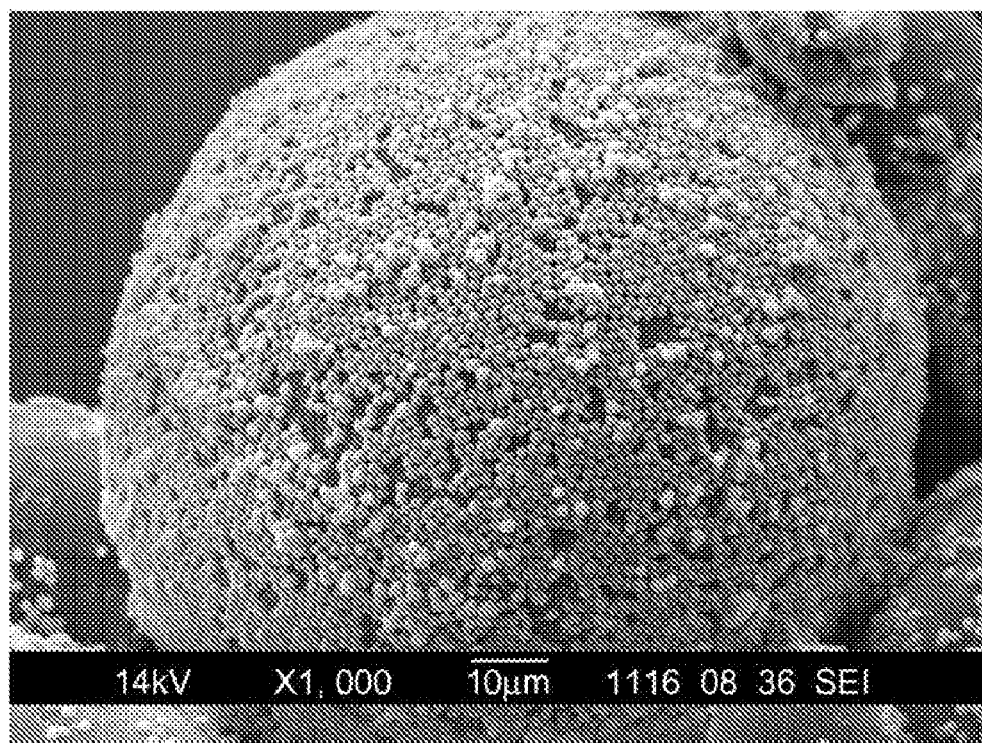


Fig. 3A

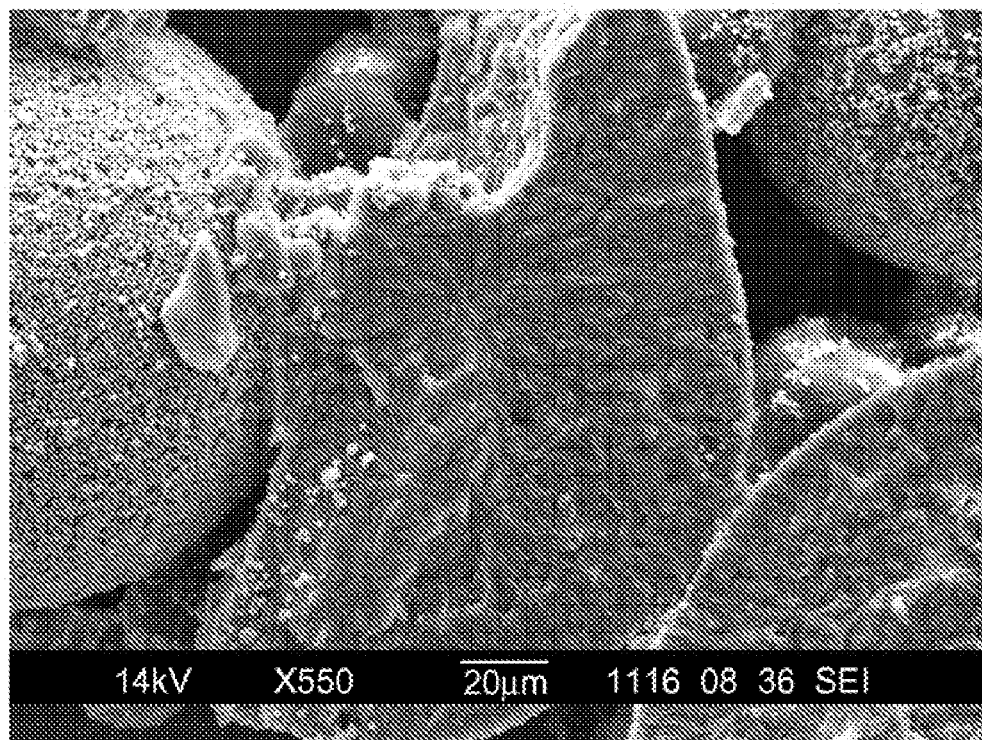


Fig. 3B

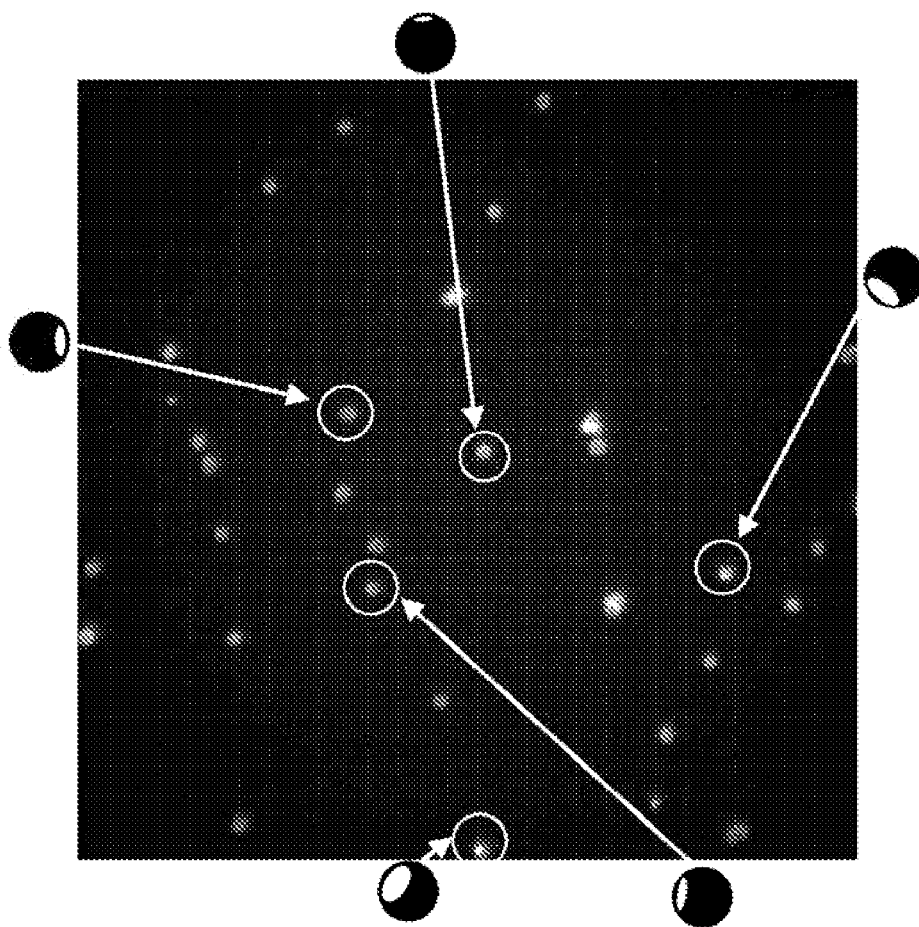
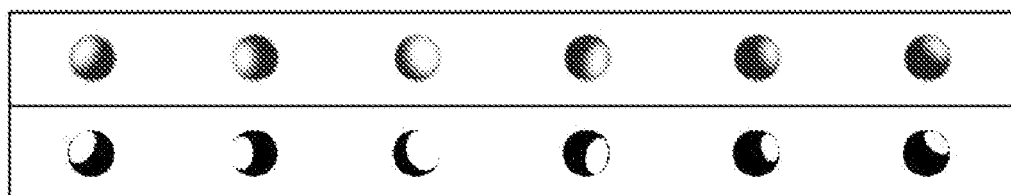


Fig. 4A



Time (s)

0

2.0

4.0

6.0

8.0

10.0

Fig. 4B

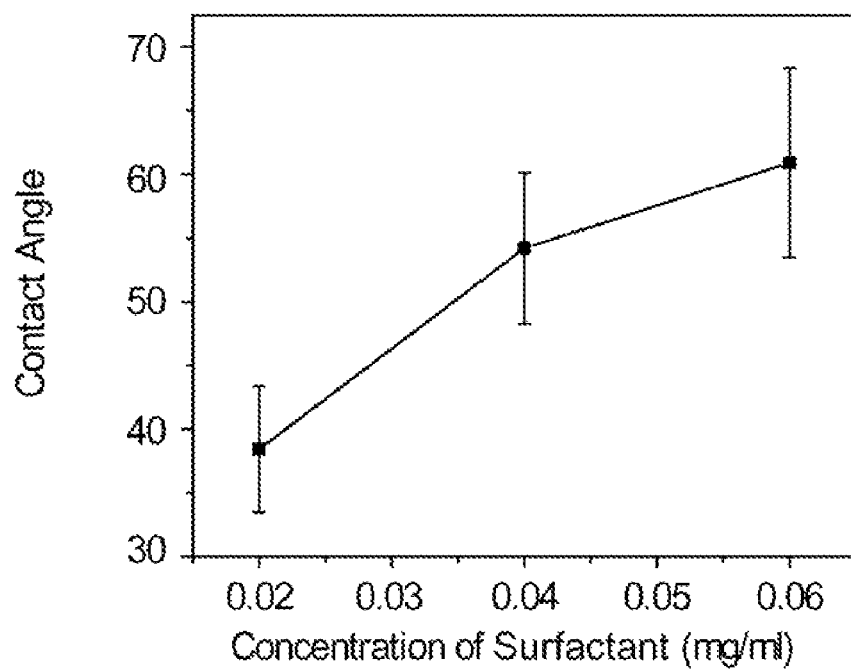


Fig. 5A

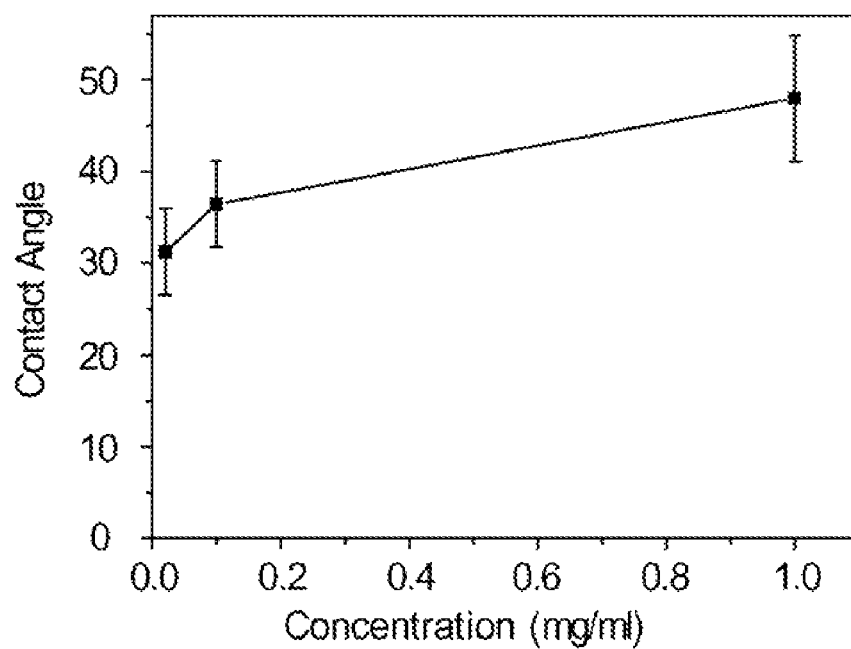


Fig. 5B

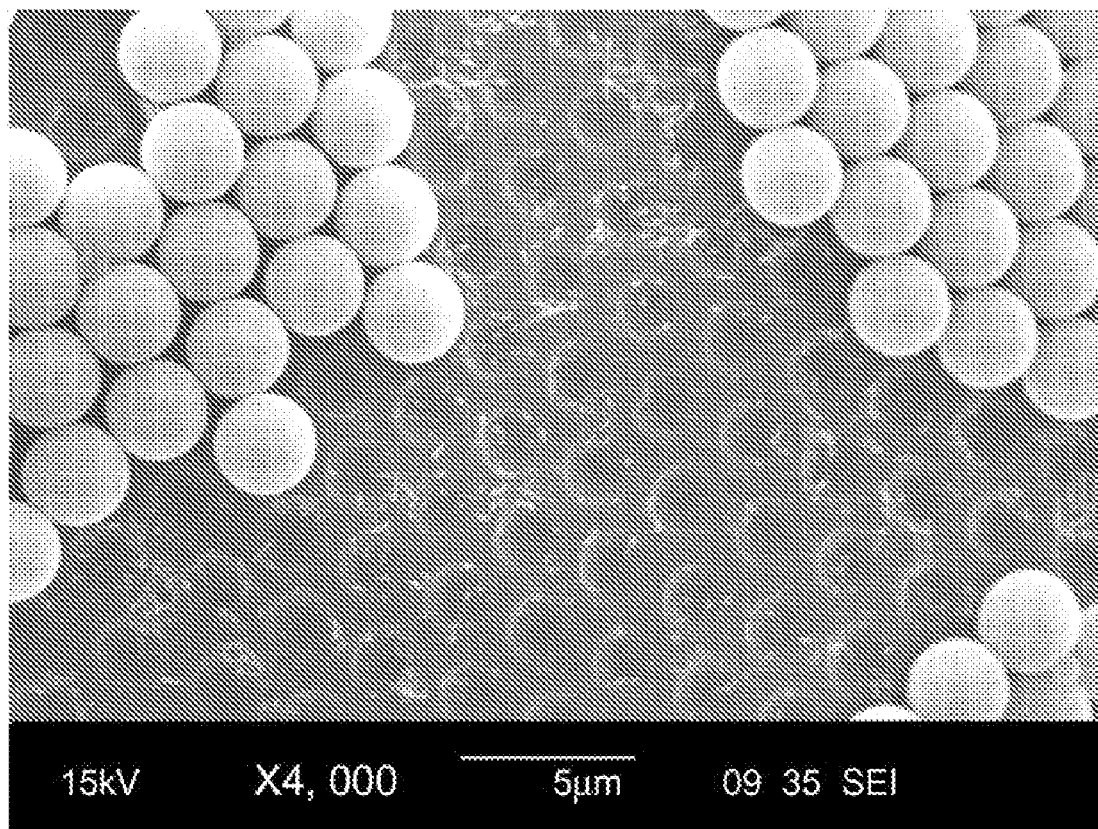


Fig. 6A

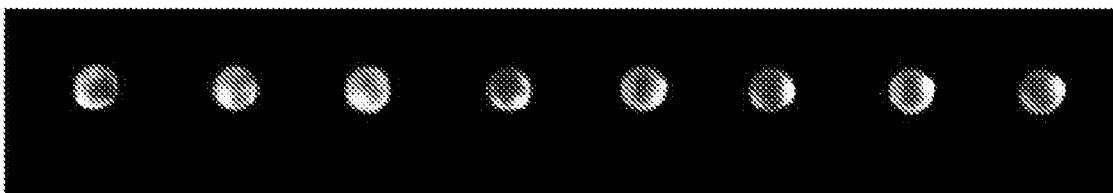


Fig. 6B

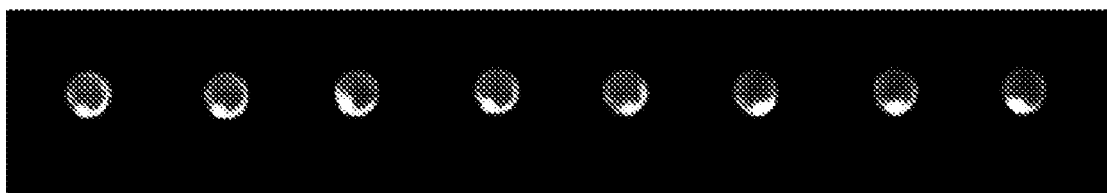


Fig. 6C

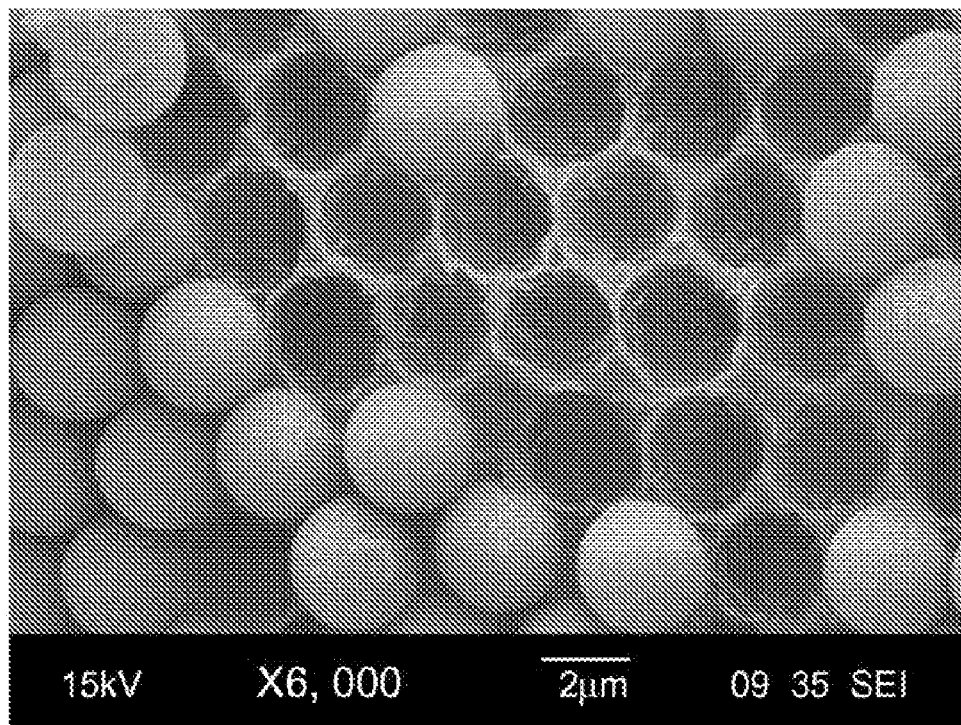


Fig. 7A

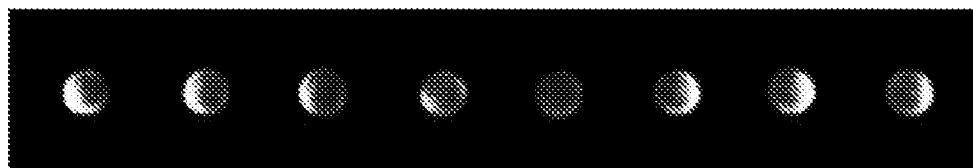


Fig. 7B

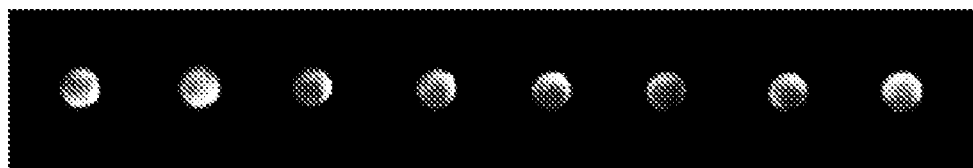


Fig. 7C

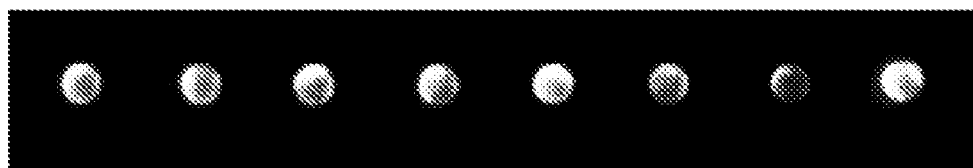


Fig. 7D

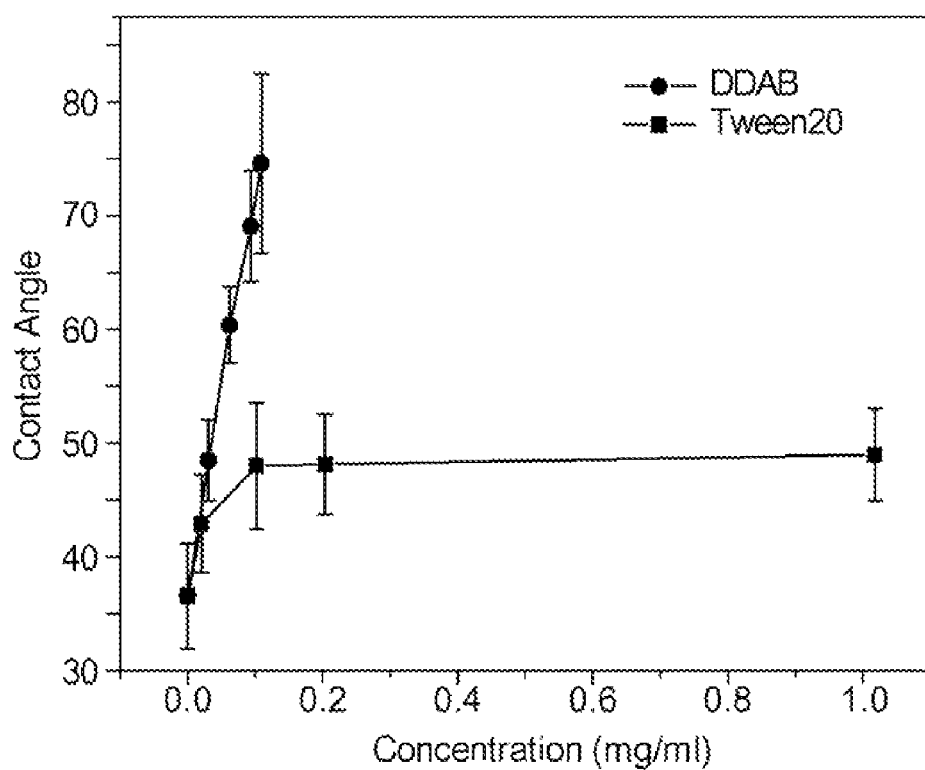


Fig. 8

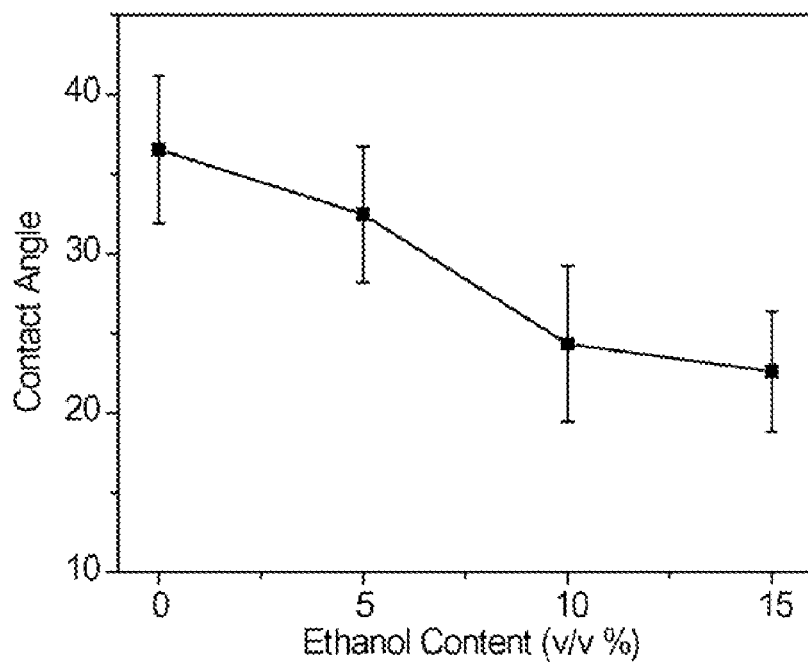


Fig. 9

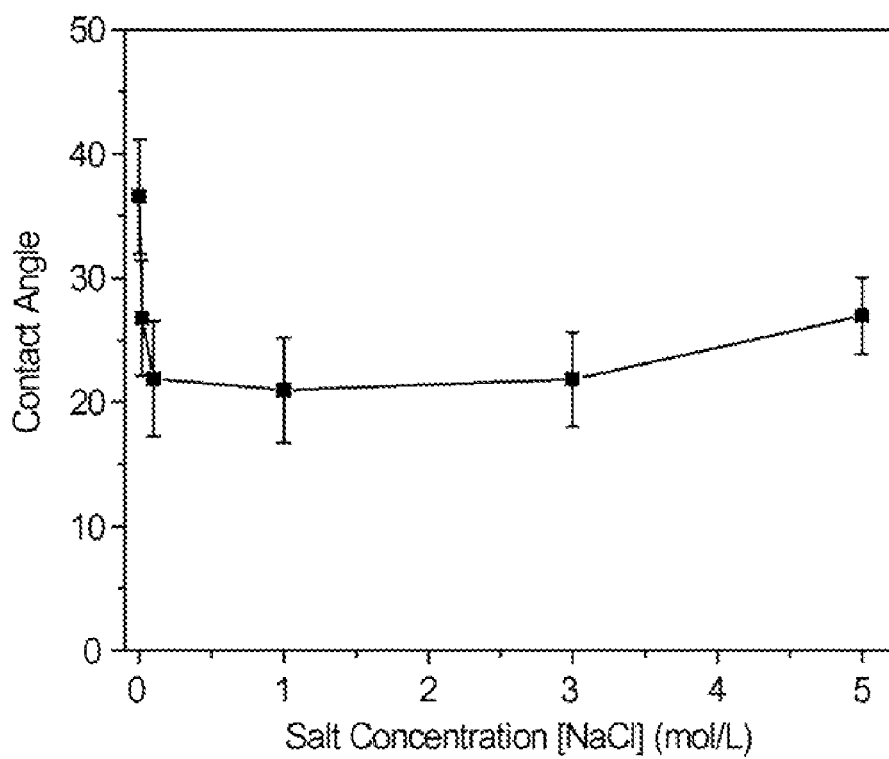


Fig. 10

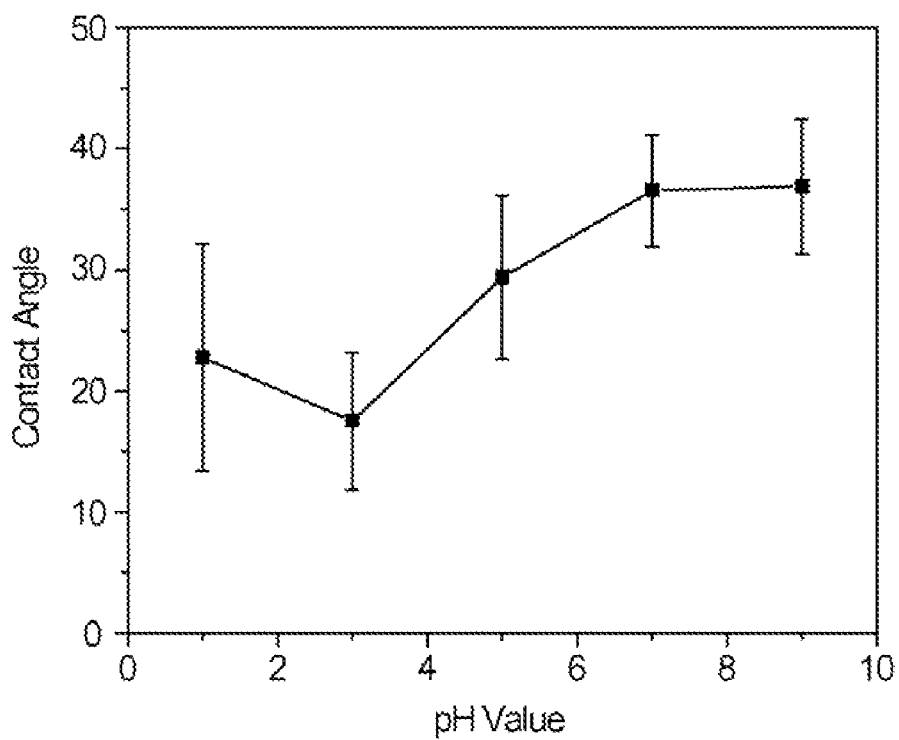


Fig. 11

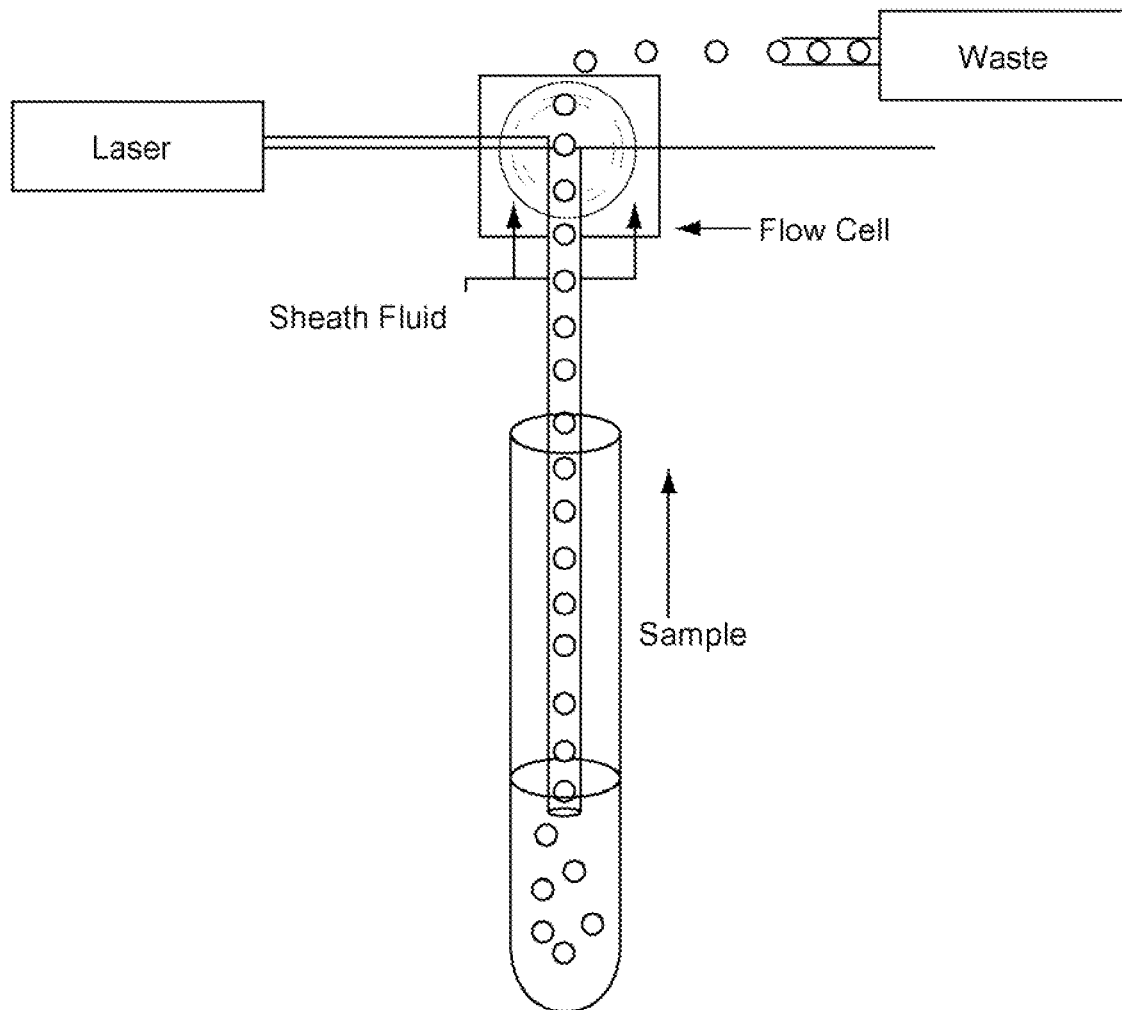


Fig. 12

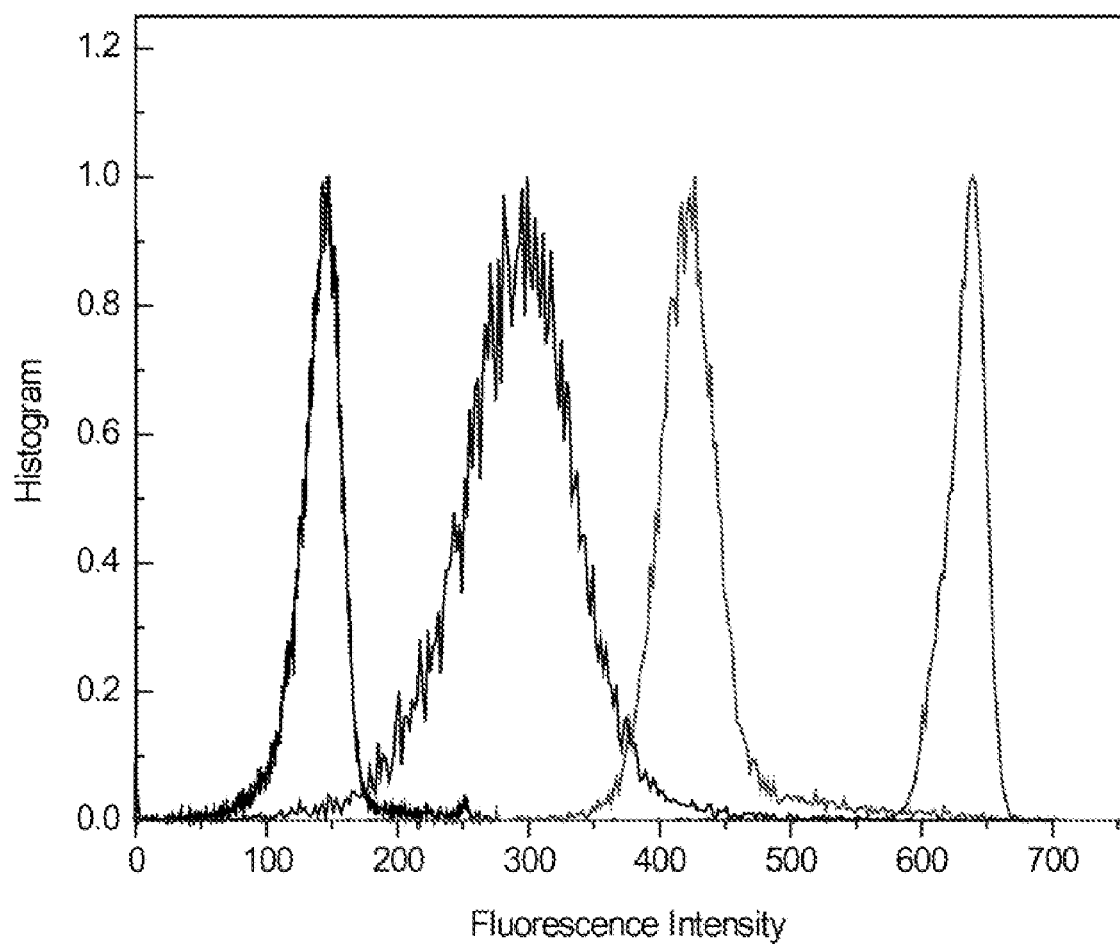


Fig.13

SYSTEM FOR FORMING JANUS PARTICLES

FEDERALLY SPONSORED RESEARCH OR
DEVELOPMENT

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BACKGROUND

Colloidal particles are fundamental in nature and technology. These particles are larger than molecules but small enough to sustain Brownian motion. A traditional goal in colloid and nanoparticle science is to obtain particles that are homogeneous in their chemical composition, and this attribute continues to have value in applications such as paints, ceramics, and photonic materials (Qi et al., 2004). Of emerging interest, however, is a heterogeneous material known as a Janus particle, in which the surface chemical composition differs on two faces of the particle (de Gennes, 1992)).

Janus particles have numerous applications. For example, an anisotropic surface chemical composition, superposed on a homogeneous shape, introduces asymmetric interactions that may induce particles to self-assemble (Hong et al., 2008). Spherical Janus particles can be used as microrheological probes (Behrend et al., 2004). With magnetic fields being used to exert force on one side of the particle, spherical Janus particles can be used to manipulate particles with controlled magnetic fields (Behrend et al., 2005). Janus particles can also be used as potential building blocks for new three-dimensional self-assembled structures.

The traditional methods for production of Janus particles having a colloidal size of 1 micron (μm) or less in diameter are severely limited, owing to the limited amount of product obtained. For example, the yields are typically less than 0.5% (Perro et al., 2006). Microfluidic methods can be used to produce larger quantities of particles, but the particle size at present is limited to about 10 to 100 μm diameter (Nie et al., 2006; Nisisako et al., 2006; Dendukuri, 2006). One current approach to produce Janus particles combines two different materials (Erhardt et al., 2001; Roh et al., 2005); however, the shape and monodispersity characteristics of the resulting particles are difficult to control.

In another approach, homogeneous precursor particles may be modified selectively by a protect-and-release process. This process uses a planar solid substrate as a protecting surface onto which particles are initially placed as a two-dimensional monolayer. The side of the particle that faces the solid substrate is protected, while the other side of the particle is unprotected and can undergo chemical modification. Sputtering (Takei et al., 1997; Lu et al., 2003), stamp coating (Cayre et al., 2003; Paunov et al., 2004), and Langmuir-Blodgett-based methods (Fujimoto et al. 1999) have been used to chemically modify the unprotected side of the particle. This approach offers good control over the surface area that sustains chemical modification; however, only a few milligrams of particles are produced in one batch because the approach relies on modification of a two-dimensional monolayer.

The experimental challenges of producing large quantities of Janus particles of colloidal size have limited their use in technological applications and scientific studies. Attempts to

solve the synthetic problem have focused on solution-based methodologies. For example, discrete particles can be joined from two phases at a liquid-liquid interface to form heterodimers (Gu et al., 2005). In another example, the addition of a small amount of water to a particle-in-oil dispersion can cause the particles to aggregate by capillary forces to form clusters, the exterior of which is accessible for chemical modification (Takahara et al., 2005). The problem with this approach is that adequate control of surface coverage of the resultant clusters is difficult to achieve, because the chemical process is subject to kinetic control.

SUMMARY

In first aspect, the invention is a method of forming Janus particles that includes forming an emulsion that contains initial particles, a first liquid, and a second liquid; solidifying the first liquid to form a solid that contains at least a portion of the initial particles on a surface of the solid; and treating the exposed particle sides with a first surface modifying agent, to form the Janus particles. Each of the initial particles on the surface has an exposed particle side and a blocked particle side.

In a second aspect, the invention is a method of forming Janus particles that includes forming an emulsion that contains initial particles, a first liquid, a second liquid, and a surfactant; solidifying the first liquid to form a solid that contains at least a portion of the initial particles on a surface of the solid; and treating the exposed particle sides with a first surface modifying agent, to form the Janus particles. Each of the initial particles on the surface has an exposed particle side and a blocked particle side.

In a third aspect, the invention includes a particle. The particle is a Janus particle having a Janus balance of 10 percent to less than 50 percent.

DEFINITIONS

The term "emulsion" refers to a combination of at least two liquids, where one of the liquids is present in the form of droplets in the other liquid. IUPAC, *Compendium of Chemical Terminology: IUPAC Recommendations*, 2nd ed., compiled by A. D. McNaught and A. Wilkinson, Blackwell, Oxford (1997).

The term "colloidosome" refers to a colloidal-sized droplet of a first liquid in a second liquid having particles at the interface formed between the first liquid and a second liquid. A solidified colloidosome, in which the first liquid has solidified, is a colloidal-sized particle having particles on its surface.

The term "surface modifying agent" means any chemical reagent that can change the chemical functionality of a surface.

The term "Janus angle" refers to the three-phase contact angle formed between a particle and the liquid-liquid interface formed between a first liquid and a second liquid.

The term "Janus balance" refers to the percentage of the total contiguous surface area of a particle that contacts a first liquid in a liquid-liquid interface composed of a first liquid and a second liquid. For example, a Janus balance of 25% refers to a particle having 25% of the surface area of the particle in contact with a first liquid and the remaining 75% of the surface area of the particle exposed.

BRIEF DESCRIPTION OF THE DRAWINGS

FIG. 1 depicts a schematic of a procedure to create Janus particles;

FIG. 2 depicts the Janus angle (depicted as θ_{ow}) for a colloidosome **116** containing a particle **104**, a first liquid **106** (exemplified by oil) and a second liquid **108** (exemplified by water);

FIG. 3A depicts a scanning electron micrograph image of particles adsorbed onto the surface of a solidified colloidosome;

FIG. 3B depicts a scanning electron micrograph image of a broken solidified colloidosome;

FIG. 4A depicts an epifluorescent image of Janus particles modified with a rhodamine B isothiocyanate group;

FIG. 4B depicts a time-dependent rotation of epifluorescent images of the particles of FIG. 4A;

FIG. 5A depicts the relationship between the Janus angle and the concentration of a cationic surfactant (DDAB);

FIG. 5B depicts the relationship between the Janus angle and the concentration of a non-ionic surfactant (Tween 20);

FIG. 6A depicts a scanning electron micrograph image of a solidified colloidosome surface formed in the presence of a cationic surfactant (DDAB) at a concentration of 1 mg DDAB in 50 ml of water (representing the second liquid);

FIGS. 6B and 6C depict time-dependent rotation of epifluorescent images of the Janus particles of the colloidosome of FIG. 6A after modification with a rhodamine B isothiocyanate group;

FIG. 7A depicts a scanning electron micrograph image of a solidified colloidosome surface formed in the presence of a cationic surfactant (DDAB) at a concentration of 3 mg DDAB in 50 ml of water (representing the second liquid); and

FIGS. 7B, 7C, and 7D depict time-dependent rotation of epifluorescent images of Janus particles of the colloidosome of FIG. 7A after modification with a rhodamine B isothiocyanate group.

FIG. 8 depicts the relationship between the Janus angle and the concentration of surfactant (DDAB and Tween 20).

FIG. 9 depicts the relationship between the Janus angle and the concentration of ethanol relative to the water phase.

FIG. 10 depicts the relationship between the Janus angle and the concentration of salt relative to the water phase.

FIG. 11 depicts the relationship between the Janus angle and the pH value.

FIG. 12 illustrates a flow cytometry setup.

FIG. 13 depicts the distribution of particles, as analyzed by flow cytometry.

DETAILED DESCRIPTION

The present invention makes use of the discovery of an emulsion method to synthesize Janus particles in large quantity. Colloidosomes form in an emulsion containing particles, a first liquid, a second liquid, and optionally, a surfactant. The particles localize at the liquid-liquid interface of the colloidosome, and solidification of the first liquid forms a solidified colloidosome. The inclusion of a surfactant may permit control over the percentage of the surface area of the particles within the first liquid, thereby enabling control of the Janus balance of the final Janus particles obtained. The exposed surfaces of the immobilized particles are modified chemically by treating the solidified colloidosome with a surface modifying agent. The solid is removed from the solidified colloidosome to release the Janus particles. The previously blocked surfaces of the Janus particles are now accessible for optional subsequent chemical modification. Gram-sized quantities of Janus particles can be synthesized.

This method can achieve superior results over traditional methods of producing Janus particles. First, a dramatic increase in Janus particle yield may be realized with the

disclosed process (for example, by 2 orders of magnitude), as compared to Janus particles prepared from two-dimensional monolayer surfaces. Second, the mechanically rigid solidified colloidosomes may be stable in moderately aggressive chemical environments, thereby permitting a greater range of surface modification chemistries. Third, the surface coverage of the modified chemical composition of the particle surface is subject to thermodynamic control, thereby allowing improved monodispersity characteristics among particle batches. Finally, the Janus balance of the obtained particles may be controlled.

FIG. 1 represents an example of a process **100** for forming Janus particles. Process **100** includes processes **110**, **120**, and **140**. Process **100** further includes processes **130**, **150**, and optionally, **160**.

A mixture **102** is prepared that includes particles **104**, a first liquid **106**, a second liquid **108**, and optionally, a surfactant **109**. Optionally, the particles **104** and the first liquid **106** are combined initially to form mixture **107**, which is subsequently combined with the second liquid **108** to provide the mixture **102**.

The particles **104** include microparticles and nanoparticles. Preferred particles **104** include microparticles having a diameter greater than 100 nm, including 105 nm to 10 μ m, such as 105, 150, 200, 300, 500, 600, 700, 800, and 900 nm and 1, 1.5, 2.0, 2.5, 5, 7.5, and 10 μ m. An especially preferred particle **104** is a microparticle having a diameter of 800 nm to 5 μ m, including 1.0, 1.5, 2.0, 2.5, 3.0, 3.5, 4.0, and 4.5 μ m. Preferred particles **104** include nanoparticles having a diameter of 100 nm or less, including the range of 10 nm to 100 nm, such as 10, 20, 30, 40, 50, 60, 70, 80, 90, and 100 nm.

The particles **104** may have one of a variety of shapes, such as spherical, ellipsoid, rodlike, helical, and oblate, among others. The particles **104** may be solid or hollow. The particles **104** may be composed of any material, including polymers (for example, polystyrene, polypropylene, latex, polyacrylamide), polypeptides, proteins, nucleic acids, glass (for example, fused silica), ceramics (for example, TiO₂, Al₂O₃, ZrO₂), metals (for example, gold, silver, platinum, palladium or alloys), and chemical elements (for example, tungsten, titanium, carbon, etc.).

The particles **104** may be prepared by modifying precursor particles to possess modified surfaces bearing chemically reactive moieties. For example, precursor particles may be chemically modified to form particles **104** before their introduction into mixture **102** or optionally, mixture **107**. Examples of modification include alkylation, amidation, esterification, metallation, phosphorylation, sulfonation, oxidation, reduction, and covalent attachment of additional chemical moieties that impart further chemical functionality to a precursor particle, among others. Examples of modified particles include those particles described in Suslick et al. (2003) and Lu et al. (2006).

The first liquid **106** and the second liquid **108** of mixture **102** are selected such that the overall surface energy of the liquid-liquid interface **114** is lowered upon addition and subsequent adsorption of particles **104** in formation of the emulsion **112** in process **110**. Preferably, the first liquid **106** includes any organic polymer that undergoes that undergoes reversible melting.

Examples of preferred first liquids **106** include polymers and waxes. Waxes are water insoluble, organic materials that are solid or semi-solid at room temperature, usually of lower density than water, and typically can be melted above room temperature to form a liquid. Preferred waxes include any naturally occurring and synthetic waxes, wax esters, and greases that have a melting temperature of 30° C. or more,

with a melting range of less than 10° C., and that are non-reactive with the reagents or solvents to which they are exposed.

Examples of first liquids **106** include esters of various long-chain (fatty) alcohols and long-chain acids, preferably where at least one member of the ester has 10 or more carbon atoms, including various unsaturated and branched chain types and also those esters of glycerols and sterols. Also, certain free alcohols or acids have wax-like properties of melting temperature and inertness. Examples of saturated fatty acids include capric, lauric, myristic, palmitic, margaric, stearic, arachidic, behenic, tetracosanic, lignoceric, cerotic, and melissic. Examples of unsaturated fatty acids include tiglic, hypogaic, gaidic, phytetoleic, elaidic, oleic, isooleic, erudic, brassidic, and isorudic. Examples of fatty alcohols include octadecyl, carnaubyl, ceryl, melissyl, and phytol. Also included are various esters of these and other fatty acids with any suitable fatty alcohols, or sterols such as cholesterol, or glycerols.

Examples of first liquids **106** include natural or suitably modified waxes such as various plant derived waxes, greases and oils including carnauba wax, cranberry wax, ouricuri wax, candelilla wax, raphia wax, apple, cotton and cactus waxes; waxes (including greases) produced by bacteria (for example, cetyl stearate); fungi, protozoa and algae; various invertebrate waxes and greases including insect waxes such as beeswaxes (for example, triacontyl palmitate, palmatyl palmitate), and *Coccus* sp. derived waxes (for example, lac, cochineal and Chinese insect); and other animal fats (for example, triglycerides) and waxes including spermaceti (for example, cetyl palmitate), lanolin and wool grease. Also included are various derivatives, extracts, and combinations of these materials.

Examples of first liquids **106** include many natural or synthetic hydrocarbons such as white waxes, paraffins, ceresins, silicon greases and waxes, polychlorinated or polyfluorinated hydrocarbons, aromatic hydrocarbons (such as naphthalene and durene(1,2,4,5-tetramethylbenzene)), polyether waxes and polyester waxes. Waxes include waxy polymers, which are polymers that have wax-like chemical or physical properties alone or when combined with other waxes. Examples of waxy polymers include polyethylenes and polypropylenes. Examples of polymers that may be combined with waxes to produce waxy polymers include certain gums and rubbers, various kinds of latex, gutta-percha, balata, chicle and various derivatives. Also included are synthetic rubbers such as isoprene polymers, hydrogenated rubber, butadiene polymers, chloroprene polymers and butyl polymers.

Examples of first liquids **106** also include gelatin, guar gum, acacia (gum arabic), carob bean gum, carrageenan, xanthan gum, food starch, carboxymethyl cellulose, ethyl cellulose, methyl cellulose, cellulose acetate, cellulose nitrate, silicone rubber, butyl rubber, butadiene-styrene rubber, polyurethane, epoxy, polyvinyl alcohol, polyvinyl acetate, polydimethyl siloxane, urea formaldehyde, polyethylene, polyethylene glycol, polystyrene, polymethyl methacrylate, polypropylene, polyvinyl chloride, polyvinyl alcohol, polycarbonate, and polyamide. An especially preferred first liquid **106** is paraffin wax.

The solidification of the first liquid **106** may be induced in a variety of ways. Examples include a temperature change (for example, heating or cooling about the melting temperature), a photochemical process (for example, photopolymerization or photodepolymerization), and a chemical process (for example, a catalyst, an initiator, or inclusion or removal of a co-solvent). When the first liquid **106** is a wax, it is preferred to induce the solidification of first liquid **106** by

lowering the temperature (for example, past a temperature about the melting temperature of the wax) or by removing a co-solvent (for example, a solvent that liquefies the wax).

The second liquid **108** includes any liquid that is immiscible with the first liquid **106** and does not solidify under the same conditions as the first liquid **106**. A preferred second liquid **108** is water, particularly where the first liquid **106** is a wax. An especially preferred combination of the first and second liquids **106** and **108** is paraffin wax and water, respectively.

The surfactant **109** includes any surfactant that is soluble in the second liquid **108**, such as an ionic surfactant or a non-ionic surfactant. Examples of ionic surfactants include cationic surfactants (for example, cetyltrimethyl-ammonium bromide (CTAB), hexadecyltrimethylammonium bromide (HTAB), dimethyldioctadecylammonium bromide (DDAB), and methylbenzethonium chloride (Hyamine™)) and anionic surfactants (for example, sodium dodecyl sulfate, sodium lauryl sulfate, lauryl dimethyl amine oxide, bile salts (such as sodium deoxycholate, sodium cholate)). Examples of non-ionic surfactants include polysorbates (Tween™), polyethoxylated alcohols, polyoxyethylene sorbitan, octoxynol (Triton X100™), N,N-dimethyldodecyl-amine-N-oxide, Polyoxyl 10 lauryl ether, Brij 721™, nonylphenol ethoxylate (Tergitol™), cyclodextrins, lecithin, among others. A preferred non-ionic surfactant is Tween 20. A preferred ionic surfactant is a cationic surfactant, such as DDAB.

The type of surfactant **109** selected for inclusion in process **100** depends upon the type of particle **104** that is selected for the mixture **102**. For particles **104** having an ionic surface or coating, selection of a surfactant **109** having an ionic character opposite that of the particles **104** is preferred. For example, a cationic surfactant, such as DDAB, is a preferred surfactant **109** for particles **104** having an anionic surface, such as fused silica microspheres. For particles that have a mixed ionic surface composition (that is, containing both positive and negative charges), a preferred surfactant may include a mixture of ionic surfactants, or alternatively, a non-ionic surfactant.

When surfactant **109** is an ionic surfactant, the preferred amount of the surfactant **109** in the mixture **102** corresponds to a mass of the surfactant **109** of 10 parts per million (ppm) to 90 ppm of the mass of the second liquid **108**, including 10, 15, 20, 25, 30, 35, 40, 45, 50, 55, 60, 65, 70, 75, 80, 85, and 90 ppm of the mass of the second liquid **108**, as well as other amounts within that range. An especially preferred amount of the surfactant **109** in mixture **102** corresponds to a mass of the surfactant **109** of 20 to 60 ppm of the mass of the second liquid **108**, such as 20 ppm. For example, if 10 g of the second liquid **108** is present in the mixture **102**, then an especially preferred amount of surfactant **109** in mixture **102** corresponds to 200 µg of surfactant **109**.

When surfactant **109** is a non-ionic surfactant, the preferred amount of surfactant **109** in the mixture **102** corresponds to a mass of the surfactant **109** of 20 parts per million (ppm) to 1000 ppm of the mass of the second liquid **108**, including 20, 50, 100, 200, 300, 500, 600, 700, 800, 900, and 1000 ppm of the mass of the second liquid **108**, as well as other amounts within that range. An especially preferred amount of the surfactant **109** in mixture **102** corresponds to a mass of the surfactant **109** of 100 to 1000 ppm of the mass of the second liquid **108**, such as 500 ppm. For example, if 10 g of the second liquid **108** is present in the mixture **102**, then an especially preferred amount of surfactant **109** in mixture **102** corresponds to 5 mg of surfactant **109**.

The preferred amounts of the particles **104** in the mixture **102** corresponds to a mass of the particles **104** of 0.25% to

35% of the mass of the first liquid **106**, including 0.25%, 0.5%, 1%, 5%, 10%, 15%, 20%, 25%, 30%, and 35% of the mass of the first liquid **106**, as well as other percentages within that range. An especially preferred amount of particles **104** in mixture **102** corresponds to 20% of the mass of the first liquid **106**. For example, if 1 g of the first liquid **106** is present in the mixture **102**, then an especially preferred amount of particles **104** in mixture **102** corresponds to 0.2 g of particles **104**.

The preferred mass ratios of the first liquid **106** to the second liquid **108** of mixture **102** is 1:1 to 1:20, including 1:1, 1:2, 1:3, 1:4, 1:5, 1:6, 1:7, 1:8, 1:9, 1:10, 1:12.5, 1:15, 1:17.5, and 1:20, as well as other ratios within that range. More preferably, the mass ratio of the first liquid **106** to the second liquid **108** in the mixture **102** is 1:5, 1:10, and 1:15. An especially preferred mass ratio of the first liquid **106** to the second liquid **108** is 1:10. For example, if 10 g of the second liquid **108** is present in the mixture **102**, then 1 g of the first liquid **106** is present in mixture **102**.

The preferred compositional amounts of the particles **104**, the first liquid **106**, the second liquid **108**, and optionally the surfactant **109** in the mixture **102** may be determined based upon factors such as the quantity of Janus particle yield desired, the type of modification chemistries to be used, and the desired Janus balance of the resultant Janus particles, among other considerations. Preferred compositional ratios of the first particles **104**, the first liquid **106**, and the second liquid **108** in mixture **102** include the following (particles **104**: first liquid **106**: second liquid **108**): 0.0025:1:1; 0.0025:1:2; 0.0025:1:3; 0.0025:1:4; 0.0025:1:5; 0.0025:1:6; 0.0025:1:7; 0.0025:1:8; 0.0025:1:9; 0.0025:1:10; 0.0025:1:12.5; 0.0025:1:15; 0.0025:1:17.5; 0.0025:1:20; 0.005:1:1; 0.005:1:2; 0.005:1:3; 0.005:1:4; 0.005:1:5; 0.005:1:6; 0.005:1:7; 0.005:1:8; 0.005:1:9; 0.005:1:10; 0.005:1:12.5; 0.005:1:15; 0.005:1:17.5; 0.005:1:20; 0.01:1:1; 0.01:1:2; 0.01:1:3; 0.01:1:4; 0.01:1:5; 0.01:1:6; 0.01:1:7; 0.01:1:8; 0.01:1:9; 0.01:1:10; 0.01:1:12.5; 0.01:1:15; 0.01:1:17.5; 0.01:1:20; 0.05:1:1; 0.05:1:2; 0.05:1:3; 0.05:1:4; 0.05:1:5; 0.05:1:6; 0.05:1:7; 0.05:1:8; 0.05:1:9; 0.05:1:10; 0.05:1:12.5; 0.05:1:15; 0.05:1:17.5; 0.05:1:20; 0.10:1:1; 0.10:1:2; 0.10:1:3; 0.10:1:4; 0.10:1:5; 0.10:1:6; 0.10:1:7; 0.10:1:8; 0.10:1:9; 0.10:1:10; 0.10:1:12.5; 0.10:1:15; 0.10:1:17.5; 0.10:1:20; 0.15:1:1; 0.15:1:2; 0.15:1:3; 0.15:1:4; 0.15:1:5; 0.15:1:6; 0.15:1:7; 0.15:1:8; 0.15:1:9; 0.15:1:10; 0.15:1:12.5; 0.15:1:15; 0.15:1:17.5; 0.15:1:20; 0.20:1:1; 0.20:1:2; 0.20:1:3; 0.20:1:4; 0.20:1:5; 0.20:1:6; 0.20:1:7; 0.20:1:8; 0.20:1:9; 0.20:1:10; 0.20:1:12.5; 0.20:1:15; 0.20:1:17.5; 0.20:1:20; 0.25:1:1; 0.25:1:2; 0.25:1:3; 0.25:1:4; 0.25:1:5; 0.25:1:6; 0.25:1:7; 0.25:1:8; 0.25:1:9; 0.25:1:10; 0.25:1:12.5; 0.25:1:15; 0.25:1:17.5; 0.25:1:20; 0.30:1:1; 0.30:1:2; 0.30:1:3; 0.30:1:4; 0.30:1:5; 0.30:1:6; 0.30:1:7; 0.30:1:8; 0.30:1:9; 0.30:1:10; 0.30:1:12.5; 0.30:1:15; 0.30:1:17.5; 0.30:1:20; 0.35:1:1; 0.35:1:2; 0.35:1:3; 0.35:1:4; 0.35:1:5; 0.35:1:6; 0.35:1:7; 0.35:1:8; 0.35:1:9; 0.35:1:10; 0.35:1:12.5; 0.35:1:15; 0.35:1:17.5; and 0.35:1:20. Thus, preferred compositions of mixture **102** include the particles **104** representing 1×10^{-4} to 1.5×10^{-1} of the total mass fraction of mixture **102**; the first liquid **106** representing 5×10^{-2} to 0.50 of the total mass fraction of mixture **102**; and the second liquid **108** representing 0.42 to 0.95 of the total mass fraction of mixture **102**. An especially preferred composition of mixture **102** includes 0.2 g of particles **104**, 1.0 g of first liquid **106**, and 10.0 g of second liquid **108**.

The amount of surfactant **109** in the mixture **102** can determine the desired Janus balance for the final Janus particles. Though the amount of surfactant **109** present in the mixture **102** contributes to the overall composition of the mixture **102**, the relative amount of surfactant **109** that represents the total

mass fraction of mixture **102** may be insignificantly small compared to the other components (for example, the mass fraction of a cationic surfactant would represent 4.2×10^{-6} to 8.8×10^{-5} of the total mass fraction of mixture **102**). Thus, the relative contribution of surfactant **109** may be ignored when calculating the amounts of the remaining components (that is, the particles **104**, the first liquid **106**, and the second liquid **108**) required to prepare the mixture **102**.

Referring again to FIG. 1, the mixture **102** is dispersed to form the emulsion **112** in process **110**. When the surfactant **109** is added to the mixture **102** before process **110**, the composition of the emulsion **112** is the same as mixture **102**. Optionally, the surfactant **109** may be added to the mixture **102** during process **110** to form the emulsion **112**. The emulsion **112** may be prepared by any means of dispersing the mixture **102**, including sonicating, high pressure jet homogenizing, vortexing, mechanical mixing, magnetic stirring, and the like. An especially preferred process **110** is forming the emulsion **112** by magnetic stirring of the mixture **102**.

Because emulsion **112** is formed by mixture **102**, it does not matter how mixture **102** is prepared initially. While it may be convenient to form a first mixture containing particles **104** with a first liquid **106** and to form a second mixture containing surfactant **109** and a second liquid **108** in route to preparing mixture **102**, this is not a requirement.

The emulsion **112** includes a liquid-liquid interface **114** between the first liquid **106** and the second liquid **108**, with particles **104** adsorbed at the liquid-liquid interface **114**. The liquid colloidosome **116** of emulsion **112** includes the first liquid **106** as an interior component and particles **104** as exterior components.

The Janus angle refers to the three-phase contact angle formed between the particles **104**, the first liquid **106**, and the second liquid **108** at the interface of the liquid colloidosome **116** with the second liquid **108** (FIG. 2). A Janus angle of 90 degrees would correspond to a particle having 50% of its surface in contact with the liquid **106**. A Janus angle of less than 90 degrees (for example, 50 degrees) would correspond to a particle having less than 50% of its surface in contact with the liquid **106**. Preferred Janus angles for liquid colloidosome **116** range from 40 degrees to 90 degrees, including 40, 45, 50, 55, 60, 65, 70, 75, 80, 85, and 90 degrees. Even more preferably, Janus angles for liquid colloidosome **116** range from 45 to 70 degrees, such as 50, 55, 60, and 65 degrees.

Referring again to FIG. 1, the first liquid **106** is solidified in process **120** to form the solid colloidosome **126**. The solid colloidosome **126** possesses a solid interior **124** composed of the solidified form of the first liquid **106** and particles **104** fixed in place on the surface of the solid. Preferably, the particles **104** that are associated with the solid colloidosome **126** are partially embedded in the first liquid **106** that forms the surface of solid colloidosome **126** during process **120**. Thus, particles **104** of solid colloidosome **126** have at least two different surfaces: (1) an embedded surface that is not exposed to solvent (that is, a blocked surface) and (2) a surface that is exposed (for example, to the second liquid **108** of emulsion **112**).

The Janus balance of the resultant particle is correlated with the Janus angle of colloidosome **116**, because the Janus angle reflects the extent to which the surface of the particle **104** contacts the solid interior **124** of the solid colloidosome **126**. Preferred Janus balances range from 10% to 50%, including 10, 15, 20, 25, 30, 35, 40, 45, and 50%, as well as other percentages within that range. Even more preferably, Janus balances range from 25%-45%, such as 30, 35, and 40%.

In process 130 of FIG. 1, the solid colloidosomes 126 are collected. Suitable methods of collection include any means of separating the solid colloidosomes 126 from the remaining liquefied components of the emulsion 112, including the second liquid 108 and surfactant 109. Preferred collection methods include filtration, centrifugation, sedimentation, and magnetic separation, with the latter method being useful for particles having magnetic properties. An especially preferred collection method is filtration. Optionally, the solid colloidosome 126 may be washed to remove any unbound or weakly-adsorbed particles 104 and surfactants 109. Preferred washing solutions include water, methanol and ethanol, among others. Methanol is an especially preferred washing solution, particularly for removing the surfactant 109 from the solid colloidosome 126. Additional purification steps may be performed to further separate free particles 104 and surfactants 109 from the solid colloidosome 126.

In process 140, the solid colloidosomes 126 are chemically modified on the exposed surfaces of particles 104 to form modified colloidosomes 146. The solid colloidosomes 126 may be suspended in solution 142 that includes one or more surface modifying agents 144. The solution 142 includes any liquid that preserves the integrity of the solid colloidosomes 126 and which solvates the surface modifying agents 144. Examples of the solution 142 include aqueous-based solvents like water or water-miscible solvents like polar protic solvents (for example, methanol, ethanol, etc.) or polar aprotic solvents (for example, dimethylformamide, dimethylsulfoxide, pyridine, acetone, etc.).

Optional additional components 148 may be included in the solution 142 of process 140, including buffering agents, acids, bases, salts, chaotropic agents, micellar agents, denaturants, renaturants, and the like.

The surface modifying agents 144 include any reagent suitable for adding chemical functionality to the exposed surface of particles 104 of the solid colloidosomes 126. Examples of the surface modifying agents 144 include alkylating agents, amidating agents, esterification agents, metalation agents, phosphorylating agents, sulfonating agents, oxidizing agents, reducing agents, or other chemistries that can add chemical functionality to the exposed surface of particles 104.

The exposed surface of particles 104 may also be modified directly using electromagnetic radiation, such as radiation in the infrared, ultraviolet, and visible wavelength ranges. Optionally, the surfaces of particles 104 may also contain reactive moieties that are responsive to electromagnetic radiation of specific wavelengths, thereby permitting further modification following irradiation.

Depending upon the complexity of modifications desired on particles 104 of the solid colloidosomes 126, a plurality of processes 140 may be performed sequentially, with each process 140 specific for adding a particular chemical functionality to the exposed surface of the particles 104. Preferably, the modified colloidosomes 146 are collected after each process 140 is completed. Any suitable collection process may be used for recovering the modified colloidosomes 146, such as those described for process 130. The obtained modified colloidosomes 146 may be extensively washed to remove solution 142, surface modifying agents 144, and any additional components 148.

In process 150, the modified colloidosomes 146 are treated to remove the solid interior 124, releasing the Janus particles 154. The colloidosomes 146 can be treated in a variety of ways to remove solid interior 124, such as by using a solvent to dissolve the solid interior 124 or by using heat to melt the solid interior 124. Preferably, the Janus particles 154 do not

dissolve or react with a solvent used to dissolve the solid interior 124. Preferably, the Janus particles 154 do not melt at the elevated temperature required to melt the solid interior 124. For colloidosomes 146 that include paraffin wax as the solid interior 124, process 150 includes treating the colloidosomes with a solvent, such as chloroform. The Janus particles 154 are subjected to collection using any separation and purification method available, including those described for process 130.

In process 160, the Janus particles 154 may be subjected to additional modification using chemistries to modify the second, previously blocked, side of the particles 104 to form modified Janus particles 166. The Janus particles 154 are suspended in a solution 162 that includes one or more surface modifying agents 164 and optional additional components 168. Additional components 168 may include those that are described as optional additional components 148 of process 140. Solution 162 includes any liquid that preserves the integrity of the Janus particles 154 and which solvates the surface modifying agents 164. The range of modification chemistries possible in process 160 may be greater than those possible in process 140, because preserving the integrity of the solid colloidosome 126 is no longer a consideration. Consequently, process 160 may include a broader range of solutions 162 than possible for solutions 142 of process 140, including both aqueous and organic solvents, as well as a broader range of surface modifying agents 164 than possible with surface modifying agents 144 of process 140.

Depending upon the complexity of modifications desired for the Janus particles 154, a plurality of processes 160 may be performed sequentially, with each process 160 specific for adding a particular chemical functionality to the surface of the Janus particles 154. Preferably, the modified Janus particles 166 are collected after each process 160 is completed. Any suitable collection process may be used for recovering the Janus particles 166, such as those described for process 130. The obtained Janus particles 166 may be extensively washed to remove solution 162, surface modifying agents 164, and any additional components 168.

The Janus angle of a colloidosome can be determined by removing at least a portion of the particles 104 from solid colloidosomes 126, imaging the resultant solid colloidosomes 126 with scanning electron microscopy (SEM), and measuring the three-phase contact angle of the voids left by the removed particles 104. Preferably, a plurality of individual three-phase contact angles is measured for one or more colloidosomes 126, and an averaged, three-phase contact angle determined by calculation.

The Janus balance may be determined by analyzing the Janus particles 154 or 166 themselves, such as by measuring the extent of surface modification achieved by a specific surface modifying agent during process 140.

One can infer the extent of particle coverage due to a surface modification on a particle having a known Janus balance. For example, a particle having a Janus balance of 35% that is subjected to surface modification with gold during process 140 refers to a particle having 65% of its surface covered with gold.

One possible explanation for the tunability of the Janus balance is that the surfactant 109 binds to the surface of the particles 104 and changes the surface hydrophobicity of the particles 104. With an apparent increase in surface hydrophobicity, the particles 104 can enter the first liquid 106 more deeply than if the surfactant 109 were omitted from the emulsion 112. Typically, the greater the hydrophobic character of the surfactant 109, the more deeply the particles 104 can enter the first liquid 106.

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For particles **104** having a hydrophilic surface (for example, an ionic surface), the desired Janus balance can be achieved by altering the hydrophilic-hydrophobic composition of the particle surface with a hydrophobic surface modifying agent other than the surfactant **109**. Janus particles of the desired Janus balance may be prepared by using pre-modified particles **104** in process **100**, thereby eliminating the use of the surfactant **109** in the emulsion **112**.

Besides the use of a surfactant **109** in process **100**, Janus particles having a desired Janus balance may be prepared by modifying the pH and ionic strength of emulsion **112**. For example, Janus particles having a desired Janus balance may be prepared with an emulsion **112** containing a liquid **108** (for example, water) having the pH adjusted to a specific pH (pH 1-11) or having a specific ionic strength (adjusted through inclusion of 0-5 M monovalent salt in liquid **108**, such as NaCl or KCl). FIGS. **10** and **11** illustrate the effects of ionic strength and pH, respectively. Additionally, inclusion of an alcohol, such as ethanol, to a volume fraction of 0-15% of liquid **108** (for example, water), including a volume fraction of 1%, 2.5%, 5%, 7.5%, 10%, 12.5%, and 15% (as well as other percentages within that range), may be used in process **100** to fine tune the Janus balance of the resultant Janus particles. FIG. **9** illustrates the effect of alcohol. One possible explanation for the tunability of the Janus balance with these different conditions (inclusion of surfactant **109** or alcohol, or adjustment of the pH or ionic strength) is that each condition varies the surface tension of the liquid **108** in emulsion **106**.

Process **100** can provide a dramatic improvement in the yield of Janus particles obtained when compared to traditional synthetic methods that use two-dimensional monolayers. For example, process **100** can provide Janus particle yields that are about two orders of magnitude greater than possible with traditional methods. Furthermore, the inclusion of a surfactant **109** in the emulsion **112** can result in improved synthetic yield of Janus particles. For example, inclusion of the cationic surfactant DDAB in an emulsion **112** containing fused silica microspheres resulted in a yield of 80% of Janus particles, as compared to a yield of 50% of Janus particles when the surfactant **109** was omitted.

One possible explanation for the improved yield is that the inclusion of a surfactant **109**, particularly an ionic surfactant having a charge that is opposite to that of particles **104**, may reduce electrostatic charge repulsion between particles **104**. The resultant particles **104** may pack more efficiently at the interface **114** of colloidosome **116** and on the surface of solid colloidosome **126**. A more efficient packing geometry for particles **104** may yield more robust colloidosomes **126** that contain or retain a greater number of particles **104** for subsequent modification.

In addition to improved Janus particle yield, the purity of the Janus particles produced with process **100** is increased relative to Janus particles obtained with traditional methods. For example, analysis of the Janus particles made according to process **100** by flow cytometry indicates that the particles are uniform in composition and possess the expected surface coverage. FIG. **12** illustrates the flow cytometry setup. FIG. **13** show the distribution of particles analyzed by flow cytometry. In FIG. **13**, dark area represents the area with lower fluorescence intensity; from left to right, the particles are: (1) homogeneous particles, which were modified by DCDMS first, then were labeled by APS-FITC; (2) Janus particles, the exposed area on the colloidosome was DCDMS modified first, and then inside was labeled by APS-FITC; (3) Janus particles, the exposed area on the colloidosome was labeled directly by APS-FITC; (4) homogeneous particles, which were labeled directly by APS-FITC.

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FIGS. **3A** and **3B** show SEM images of illustrative solid colloidosomes prepared according to process **120**. The particles **104** were adsorbed on the surfaces of the solid colloidosomes **126**. When emulsion **112** was cooled to a temperature below the melting temperature of the first liquid **106** (paraffin wax), the particle **104** monolayer surface coverage remained high (FIG. **3A**). When the first liquid **106** solidified as a wax, particles **104** were locked in place and ceased to rotate. An examination of broken wax emulsions showed that >99% of the particles were located at the emulsion surface (FIG. **3B**), having been adsorbed to this surface from prior dispersion in the wax phase.

The solid colloidosomes **126** of FIGS. **3A** and **3B** survived multiple washings, suggesting good mechanical stability. Nearly all particles **104** continued to remain at the surface, indicating fairly strong adsorption between particles **104** and the solid interior **124**. Strong particle-colloidosome surface adsorption provided the feasibility for performing sequential chemical modifications on these adsorbed particle surfaces according to process **140**.

The examples provide illustrations of the types of Janus particles that may be fabricated according to process **100**, including bipolar particles (cationic on one side and anionic on the other side) and colloidal surfactants (charged on one side and hydrophobic on the other). However, an expansive variety of Janus particles are possible, being limited only by the types of particles **104** that may be used to form the colloidosomes **116** and **126** (including use of pre-modified particles), the efficiency whereby solid colloidosomes **126** are formed, and the compatibility of the particle modification chemistries possible for both the solid colloidosomes **126** and released Janus particles **154**.

The efficacy of Janus particle surface chemical modification was confirmed by demonstrating that only one side of the particles was fluorescent following chemical modification of a solid colloidosome **126** with a fluorescent label. FIG. **4A** shows a time sequence of epifluorescence images that illustrates that the bipolar particles rotate in an aqueous suspension, consistent with the Janus particles having only one fluorescent side. The silica particles are opaque, and because of the asymmetric coating by the dye, the particles display fluorescence having an intensity that differs according to their orientation with respect to the illumination direction. FIG. **4B** illustrates a sequence of snapshots for one typical particle that tracks the particle's rotation, by Brownian motion, to different orientations.

The control of the Janus angle was confirmed by analyzing the three-phase angle of voids left by the escaped particles **104** in colloidosomes **126**. With the inclusion of a cationic surfactant in the emulsion **112** (for example, 20 mg/L to 60 mg/L of DDAB in water), the Janus angle was varied from 35 degrees to 70 degrees (FIG. **5A**). With the inclusion of a non-ionic surfactant in the emulsion **112** (for example, 20 mg/L to 1 mg/ml of Tween 20 in water), the Janus angle was varied from about 30 degrees to about 55 degrees (FIG. **5B**). Thus, the type of surfactant included in the emulsion **112** can influence the Janus angle of the formed colloidosomes and the Janus balance of the resultant Janus particles. FIG. **8** shows the data of FIGS. **5A** and **5B** on a single graph.

Applications

The Janus particles of the present invention find utility in numerous applications, depending upon their compositional nature and Janus balance. For example, bipolar particles that display charge asymmetry can undergo self-assembly wherein the surface charge asymmetry of the resultant clusters is preserved. Such model particle systems may be used to

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study the aggregation state of proteins that form aggregates where the surface charge distribution is similarly patchy.

Janus particles containing magnetically and fluorescently labeled surfaces have applications in new tools for the exploration of optical, chemical, and materials properties of surrounding environments. For example, such particles can be used in nanoinstruments, such as nanoviscosimeters, nanothermometers, nanobarometers, and nanochemical sensors. Modulation of these optical particles in magnetic fields allows for the local optical, chemical, and materials properties to be distinguished from the bulk properties.

Other applications include forming intracellular chemical images with improved signal over background noise, optical contrast agents for use in medical imaging methodologies like optical coherence tomography, and magnetic contrast agents for use in magnetic resonance imaging. Furthermore, high sensitivity immunoassays with simpler procedures are possible, as well as measurements of viscosities at chosen time and size scales, observation of microvorticity in fluids, and monitoring of microfluidic device performance. Finally, such particles should have applications in simultaneous intracellular physical and chemical imaging.

EXAMPLES

Fused silica particles used in these experiments were 800 nm and 1.5 μm in diameter, obtained from Alfa Aesa and Tokuyama, respectively. The particle surfaces were prepared for subsequent chemical modification in the following manner. The particles were initially treated with a piranha solution, followed by recovery of the particles by centrifugation and drying at 80° C. under vacuum. The paraffin wax used in these experiments had a melting point of about 55° C. Both (aminopropyl) triethoxysilane (APS) (99%) and n-octadecanetrichlorosilane (OTS) were purchased from Aldrich. To produce fluorescent-labeled APS, rhodamine B isothiocyanate was attached to APS using a procedure described in the literature. Before OTS was used in these experiments, it was freshly purified by distilling it under vacuum.

Example 1

Fabrication of Janus Particles

Fused silica particles (0.2 g) were dispersed homogeneously in paraffin wax (1 g) that was pre-heated at 75° C. The resultant suspension was then mixed with deionized water (10 g) that was pre-heated at 75° C. The emulsions were produced by subjecting the resultant mixture to magnetic stirring at 1600 rpm for 1 h. Upon cooling the emulsion to room temperature, the paraffin wax droplets of the emulsion became a solid.

Deionized water was used to wash the wax emulsions multiple times to remove particles in the aqueous solution as well as weakly attached particles. The resulting exposed surfaces of particles attached to the emulsions were allowed to react chemically with 2 mM labeled APS in methanol solution for 30 min. The reaction solutions were washed with methanol to remove excess silane.

The wax was dissolved in chloroform at room temperature to release the particles. The resulting particles were further modified chemically. For example, those particles that had been allowed to react with APS on one surface could be further allowed to react with OTS (on the second surface), producing particles that were cationic on the APS side and hydrophobic on the OTS side.

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Example 2

Characterization of the Janus Particles

Scanning electron microscopy, SEM (JEOL 6060 LV), was used to image the emulsions. Prior to imaging, a thin layer of gold was evaporated onto them to render them electrically conductive, to avoid surface charging in the electron beam.

A combination of optical imaging and epifluorescence imaging techniques were used to determine chemical composition of the particles. Optical imaging was performed with a Zeiss Axiovert 200 microscope fitted with a 40 \times objective. For epifluorescence imaging, 532 nm illumination (a Nd:YAG laser) was used and the laser beam was focused at the back focal point of a 63 \times air objective. The fluorescence images were collected using this same objective and recorded using an electron multiplying CCD camera (Andor Ixon) after filtering out light from the excitation laser. Images of particle rotation were recorded with an exposure time of 0.1 s for 800 frames.

Example 3

Fabrication of Colloidosomes Having Janus Angles of 38 \pm 2 Degrees and 63 \pm 4 Degrees

Janus particles were prepared according to Example 1 using mixtures that contain 1 mg/ml or 3 mg/ml of the cationic surfactant, DDAB, in the water phase. Following cooling of the emulsion to room temperature, the solid colloidosomes were collected by filtration, and washed extensively with methanol to remove the surfactant.

FIGS. 6A and 7A depict scanning electron microscopy images of the solid colloidosomes prior to chemical modification. Prior to imaging, a thin layer of gold was evaporated onto them to render them electrically conductive, to avoid surface charging in the electron beam.

The solid colloidosomes were chemically modified with APS and subsequently with rhodamine B isothiocyanate to produce fluorescent-labeled APS on one surface of the particles. The resultant Janus particles were recovered free of paraffin wax as described in Example 1. Time-dependent epifluorescence imaging was performed on the resultant Janus particles as described in Example 2. FIGS. 6B and 6C depict Janus particles released from colloidosomes having a Janus angle of 35 degrees and 36 degrees, respectively. FIGS. 7B, 7C, and 7D depict Janus particles released from colloidosomes having a Janus angle of 60 degrees, 59 degrees and 71 degrees, respectively.

Example 4

Fabrication of a Janus Protein Microparticle With a Gold Coating on One Surface (Prophetic Example)

BSA interacts very strongly with gold nanoparticles via its amine and thiol residues. Because of the presence of thiol and amine groups, the microparticle can be used as a template for the adhesion of gold nanoparticles. The gold colloid can adsorb directly onto the protein shell via interaction with thiol and amine groups. Thus, a Janus protein microparticle will be prepared having a gold coating on one side of the particle.

Air-filled microparticles 0.5 to 5 μm in diameter with a 50 Å thick protein shell will be initially prepared. The microparticles will be produced by sonicating a 5% weight per volume solution of bovine serum albumin (BSA) in water. The high-intensity ultrasound necessary for the reaction will be gener-

ated by a titanium horn with tip diameter of 0.5 inches, driven at 20 kHz. The BSA solution will be sonicated for 3 min at an acoustic power of 76 W/cm². The microparticle sizes will be generally under 5 μ m. This size distribution should enable them to pass readily through the microcirculation.

The obtained protein microparticles (0.2 g) will be mixed with paraffin wax (1 g) (pre-heated at 75° C.) and water (10 g) (pre-heated at 75° C.) to form a mixture. An emulsion will be prepared by magnetically stirring the mixture at 1600×rpm for 1 hour. The emulsion will be cooled to room temperature to form a suspension of colloidosomes containing solidified wax and BSA microparticles embedded on the surface of the wax. The colloidosomes will be collected and washed, followed by incubation in an aqueous solution containing reduced chloroauric acid. The mixture will be shaken gently for 25 min, followed by collection of the colloidosomes. The colloidosomes may take on a reddish color; which is indicative that the gold particles had been transferred to the exposed microparticle surface of the colloidosomes. These colloidosomes will be collected by centrifugation, washed with water, and suspended in chloroform to dissolve the paraffin wax component of the colloidosome and to release the Janus particles. The Janus particles will be purified and analyzed for the presence of a gold coating on one surface.

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What is claimed is:

1. A method of forming Janus particles, comprising: forming an emulsion comprising initial particles, a first liquid, and a second liquid; solidifying the first liquid to form solid droplets, each solid droplet of the first liquid comprising at least a portion of multiple initial particles on a surface of the solid droplet, wherein the initial particles on the surface each has an exposed particle side and a blocked particle side; and treating the exposed particle sides with a first surface modifying agent, to form the Janus particles.
2. The method of claim 1, wherein at least a portion of the initial particles is at the interface between the first and second liquids in the emulsion.
3. The method of claim 1, wherein the initial particles comprise from 1×10^{-4} to 1.5×10^{-1} of the total mass fraction of the emulsion; the first liquid comprises from 5×10^{-2} to 0.50 of the total mass fraction of the emulsion; and the second liquid comprises from 0.42 to 0.95 of the total mass fraction of the emulsion.
4. The method of claim 1, wherein the solidifying comprises lowering the temperature of the first liquid.
5. The method of claim 1, wherein the solidifying comprises polymerizing the first liquid.
6. The method of claim 1, further comprising removing the second liquid after solidifying the first liquid.
7. The method of claim 6, further comprising isolating the solid.
8. The method of claim 6, wherein the treating comprises contacting the exposed particle sides with a solution comprising the first surface modifying agent.
9. The method of claim 6, wherein the treating comprises exposing the exposed particle sides to electromagnetic radiation.
10. The method of claim 1, further comprising forming the initial particles by modifying precursor particles.
11. The method of claim 10, wherein the initial particles comprise microparticles or nanoparticles; the first liquid comprises paraffin wax; and the second liquid comprises water.
12. A method of forming Janus particles, comprising: forming an emulsion comprising initial particles, a first liquid, a second liquid, and a surfactant; solidifying the first liquid to form solid droplets, each solid droplet of the first liquid comprising at least a portion of multiple initial particles on a surface of the solid droplet, wherein the initial particles on the surface each has an exposed particle side and a blocked particle side; and treating the exposed particle sides with a first surface modifying agent, to form the Janus particles.
13. The method of claim 12, wherein the Janus particles have a Janus balance of 10 percent to 50 percent.
14. The method of claim 12, wherein the initial particles comprise microparticles or nanoparticles; the first liquid comprises paraffin wax; and the second liquid comprises water.

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