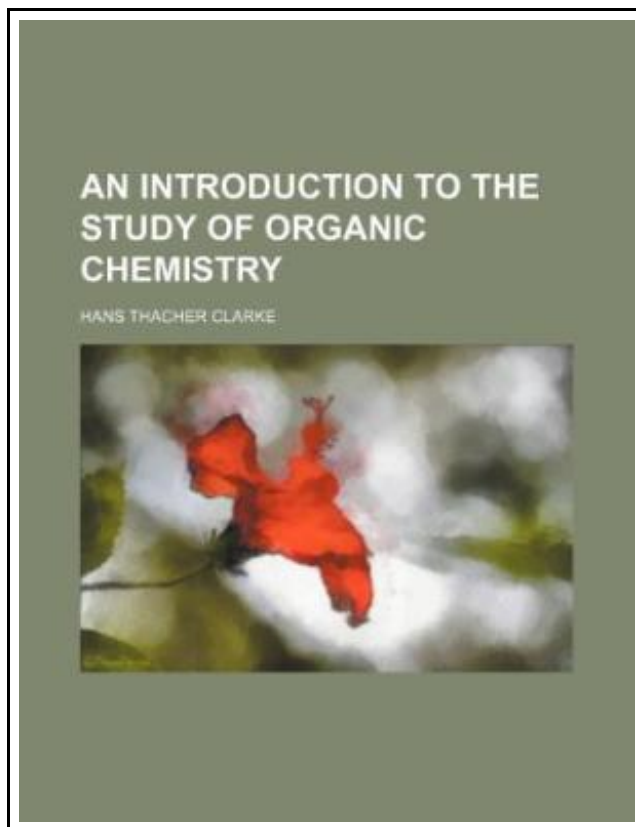


An Introduction to the Study of Organic Chemistry



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Reviews

Here is the best publication i have got go through until now. It is actually writter in simple phrases and never hard to understand. I realized this publication from my dad and i suggested this ebook to find out.

(Lorena White)

AN INTRODUCTION TO THE STUDY OF ORGANIC CHEMISTRY



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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1914 Excerpt: .of ferric salts, due to the formation of double salts such as $9 \text{ KSCN} \cdot 2 \text{ Fe(SCN)}_8$. The thiocyanates are more stable than the cyanates, and remain without decomposition in aqueous solution for an indefinite period. Ammonium thiocyanate $\text{NH}_4\text{S.C: N}$ is formed on treating a solution of hydrogen cyanide with yellow ammonium sulphide; a more convenient method of preparation is to heat carbon bisulphide with alcoholic ammonia: $\text{CS}_2 + 4\text{NH}_3 = \text{NH}_4\text{S.CN} + (\text{NH}_4)_2\text{S}$. It is a crystalline substance, soluble in water and in alcohol, which melts at about 1500. On heating to 1700-1800 it undergoes an intramolecular rearrangement similar to that of ammonium cyanate, yielding thiocarbamide (p. 183). Alkyl thiocyanates R.S.CiN in which an alkyl radicle replaces the metal, may be prepared by two different methods, which clearly show their constitution: firstly by the action of alkyl halides, or alkyl esters of inorganic salts, such as metallic alkylsulphates RO.SOsOK , upon alkali thiocyanates: $\text{C}_2\text{H}_5\text{I} + \text{K.S.CiN} = \text{C}_2\text{H}_5\text{S.C: N} + \text{KI}$ $\text{CH}_3\text{O.SOsOK} + \text{K.S.C-N} = \text{CH}_3\text{S.C: N} + \text{K}_2\text{SO}_4$; secondly by the interaction of cyanogen halides upon metallic mercaptides (p. 210): $\text{C}_2\text{H}_5\text{SNa} + \text{Cl.CN} = \text{C}_2\text{H}_5\text{S.ON} + \text{NaCl}$. They are oily liquids of unpleasant odour. Their constitution is established by treatment with nascent hydrogen--for instance, with zinc and dilute sulphuric acid--whereby they are decomposed into mercaptans and hydrogen cyanide: $\text{C}_2\text{H}_5\text{S.GN} + 2\text{H} = \text{C}_2\text{H}_5\text{SH} + \text{H.ON}$. From this it is clear that the alkyl group is directly united to the sulphur atom. On boiling with...



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