

Ch-2 Acids, Bases and Salts

1. How will you test for a gas which is liberated when HCL reacts with an active metal?
2. What is baking powder? How does it make the cake soft and spongy?
3. When fresh milk is changed into curd will its pH value increase or decrease? Why?
4. Give Arrhenius definition of an acid and a base.
5. What happens chemically when quick lime is added to water?
6. Name the gas evolved when dilute HCL reacts with Sodium hydrogen carbonate. How is it recognized?
7. How does the flow of acid rain water into a river make the survival of aquatic life in the river difficult?
8. How is the pH of an acidic solution influenced when it is diluted?
9. How is the pH of a basic solution influenced when it is diluted?
10. Arrange these in increasing order of their pH values - NaOH, blood, lemon juice.
11. Two solutions of A and B have pH values of 5 and 8. Which solution will be basic in nature?
12. Why does tooth decay start when pH of mouth is lower than 5.5?
13. What would be the colour of litmus in a solution of sodium carbonate?
14. Name the products obtained when sodium hydrogen carbonate is heated. Write the chemical equation for the same.
15. Write the chemical formula of washing soda and baking soda. Which one of these two is an ingredient of antacids? How does it provide relief in stomachache?
16. What do you mean by 'water of crystallization' of a substance? Describe an activity to show that blue copper sulphate crystals contain water of crystallization.
17. How can washing soda be obtained from baking soda? Name an industrial use of washing soda other than washing clothes.
18. Why does 1M HCL solutions have a higher concentration of H^+ ions than 1M CH_3COOH solution?
19. Why is Plaster of Paris stored in a moisture proof container?
20. What do you mean by neutralization reaction? Give two examples.
21. Why does a milkman add a small amount of baking soda to fresh milk to shift the pH of fresh milk from 6 to slightly alkaline?
22. Why do acids not show acidic behavior in the absence of water?
23. Rain water conducts electricity but distilled water does not. Why?
24. Why don't we keep sour substances in brass and copper vessels?
25. What is the common name of $CaOCl_2$?
26. Name the compound used for softening hard water.
27. Give the properties and uses of bleaching powder.
28. Give a few uses of acids, bases and salts respectively.
29. What is universal indicator?
30. How is the concentration of hydroxide ions (OH^-) affected when excess base is dissolved in a solution of sodium hydroxide?
31. Name the gas evolved when dilute sulphuric acid acts on sodium carbonate. Write the chemical equation for the reaction involved.