

# CHEM 20100 (Inorganic Chemistry I) Problem Sets

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## 0 Course Prep Problems

9/13: **2.8** The details of several steps in the particle-in-a-box model in this chapter have been omitted. Work out the details of the following steps:

- a. Show that if  $\Psi = A \sin rx + B \cos sx$  ( $A$ ,  $B$ ,  $r$ , and  $s$  are constants) is a solution to the wave equation for the one-dimensional box, then

$$r = s = \sqrt{2mE} \left( \frac{2\pi}{h} \right)$$

*Solution.*

$$\begin{aligned} \frac{-h^2}{8\pi^2m} \cdot \frac{\partial^2 \Psi(x)}{\partial x^2} &= E\Psi(x) \\ \frac{-h^2}{8\pi^2m} \cdot \frac{\partial^2}{\partial x^2} (A \sin rx + B \cos sx) &= E(A \sin rx + B \cos sx) \\ \frac{-h^2}{8\pi^2m} \cdot \frac{\partial}{\partial x} (Ar \cos rx - Bs \sin sx) &= E(A \sin rx + B \cos sx) \\ \frac{-h^2}{8\pi^2m} \cdot (-Ar^2 \sin rx - Bs^2 \cos sx) &= E(A \sin rx + B \cos sx) \\ \frac{Ar^2 h^2}{8\pi^2m} \sin rx + \frac{Bs^2 h^2}{8\pi^2m} \cos sx &= AE \sin rx + BE \cos sx \\ 0 &= \left( \frac{Ar^2 h^2}{8\pi^2m} - AE \right) \sin rx + \left( \frac{Bs^2 h^2}{8\pi^2m} - BE \right) \cos sx \end{aligned}$$

Choose  $x = 0$ .

$$\begin{aligned} &= \frac{Bs^2 h^2}{8\pi^2m} - BE \\ E &= \frac{s^2 h^2}{8\pi^2m} \\ \frac{8\pi^2mE}{h^2} &= s^2 \\ s &= \sqrt{\frac{8\pi^2mE}{h^2}} \\ \boxed{s = \sqrt{2mE} \frac{2\pi}{h}} \end{aligned}$$

With this result ...

$$\begin{aligned} 0 &= \left( \frac{Ar^2 h^2}{8\pi^2m} - AE \right) \sin rx + \left( \frac{Bs^2 h^2}{8\pi^2m} - BE \right) \cos sx \\ &= \left( \frac{Ar^2 h^2}{8\pi^2m} - AE \right) \sin rx + \left( B \left( \frac{s^2 h^2}{8\pi^2m} \right) - BE \right) \cos sx \\ &= \left( \frac{Ar^2 h^2}{8\pi^2m} - AE \right) \sin rx + (BE - BE) \cos sx \\ &= \left( \frac{Ar^2 h^2}{8\pi^2m} - AE \right) \sin rx \end{aligned}$$

Choose  $x = \frac{\pi}{2r}$ .

$$= \frac{Ar^2 h^2}{8\pi^2m} - AE$$

$$r = \sqrt{2mE} \frac{2\pi}{h}$$

□

- d. Show that substituting the value of  $r$  given in part c into  $\Psi = A \sin rx$  and applying the normalizing requirement gives  $A = \sqrt{2/a}$ .

*Solution.*

$$\begin{aligned} 1 &= \int_{\text{all space}} \Psi \Psi^* d\tau \\ &= \int_0^a \left( A \sin \frac{n\pi x}{a} \right) \left( A \sin \frac{n\pi x}{a} \right) dx \\ &= \int_0^a A^2 \sin^2 \frac{n\pi x}{a} dx \end{aligned}$$

Use  $\sin^2 u = \frac{1 - \cos 2u}{2}$ .

$$\begin{aligned} &= A^2 \int_0^a \frac{1 - \cos \frac{2n\pi x}{a}}{2} dx \\ &= \frac{A^2}{2} \left( \int_0^a dx - \int_0^a \cos \frac{2n\pi x}{a} dx \right) \\ &= \frac{A^2}{2} \left( [x]_0^a - \left[ \frac{a}{2n\pi} \sin \frac{2n\pi x}{a} \right]_0^a \right) \\ &= \frac{A^2}{2} \left( (a - 0) - \left( \frac{a}{2n\pi} \sin 2n\pi - \frac{a}{2n\pi} \sin 0 \right) \right) \\ &= \frac{A^2}{2} \left( a - \left( \frac{a}{2n\pi} \sin 2n\pi \right) \right) \end{aligned}$$

Since  $n$  is an integer,  $\sin 2n\pi = 0$ .

$$\begin{aligned} &= \frac{aA^2}{2} \\ \frac{2}{a} &= A^2 \\ A &= \sqrt{\frac{2}{a}} \end{aligned}$$

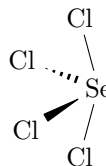
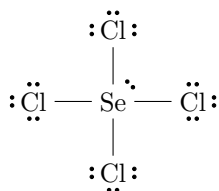
□

# 1 VSEPR and Point Groups

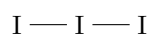
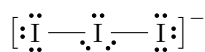
1/21: I) Do the following (VSEPR) problems from your text (Miessler et al. (2014)): Chapter 3: #8, 9f-i, 20, 29.

**3.8** Give Lewis dot structures and sketch the shapes of the following:

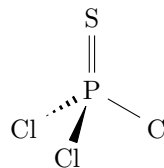
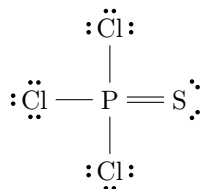
a.  $\text{SeCl}_4$



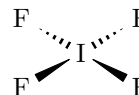
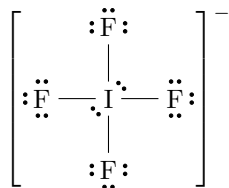
b.  $\text{I}_3^-$



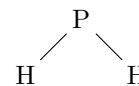
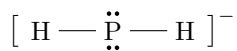
c.  $\text{PSCl}_3$



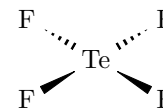
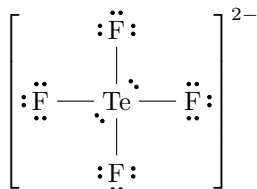
d.  $\text{IF}_4^-$



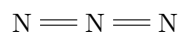
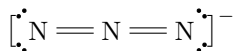
e.  $\text{PH}_2^-$



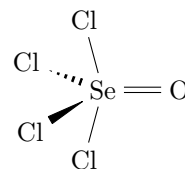
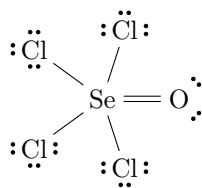
f.  $\text{TeF}_4^{2-}$



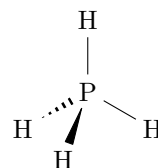
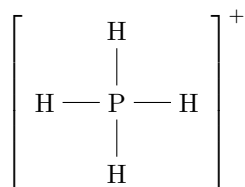
g.  $\text{N}_3^-$



h.  $\text{SeOCl}_4$

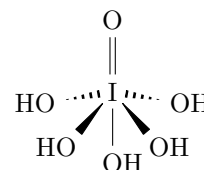
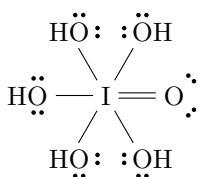


i.  $\text{PH}_4^+$

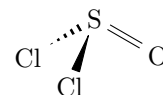
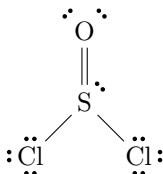


**3.9** Give Lewis dot structures and sketch the shapes of the following.

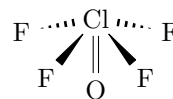
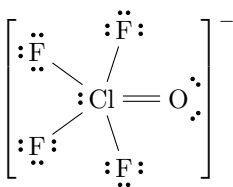
f.  $\text{IO}(\text{OH})_5$



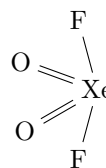
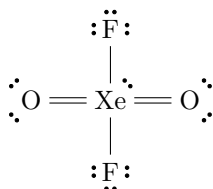
g.  $\text{SOCl}_2$



h.  $\text{ClOF}_4^{-[1]}$

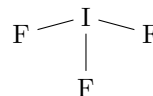
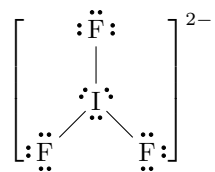


i.  $\text{XeO}_2\text{F}_2$

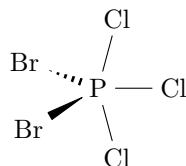


<sup>1</sup>Note that it is unclear whether the equatorial fluorines will be bent away from the lone pair and toward the oxygen, or the other way around. Hence, I arbitrarily chose to show them pointed away from the lone pair.

**3.20** Predict and sketch the structure of the (as yet) hypothetical ion  $\text{IF}_3^{2-}$ .



**3.29** Sketch the most likely structure of  $\text{PCl}_3\text{Br}_2$  and explain your reasoning.



*Answer.* Bromine is more electropositive than chlorine. Thus, by Bent's rule, the bromines will bond to the hybrid orbitals with greater  $s$ -character (the equatorial  $sp^2$  ones) first.  $\square$

- II) Assign the symmetry point group to the 13 ions and molecules in problems #8, 9f-i in Chapter 3 of your text.

**3.8****a.**  $\text{SeCl}_4$ 

*Answer.* Not low or high symmetry. Has a  $C_2$  axis. No perpendicular  $C_2$  axes. No  $\sigma_h$ . Has two perpendicular  $\sigma_v$  planes.

Therefore,  $\text{SeCl}_4$  is of the  $C_{2v}$  point group. ☐

**b.**  $\text{I}_3^-$ 

*Answer.*  $\text{I}_3^-$  is of the  $D_{\infty h}$  point group. ☐

**c.**  $\text{PSCl}_3$ 

*Answer.* Not low or high symmetry. Has a  $C_3$  axis. No perpendicular  $C_2$  axes. No  $\sigma_h$ . Has three  $\sigma_v$  planes all offset by  $60^\circ$ .

Therefore,  $\text{PSCl}_3$  is of the  $C_{3v}$  point group. ☐

**d.**  $\text{IF}_4^-$ 

*Answer.* Not low or high symmetry. Has a  $C_4$  axis. Has 4 perpendicular  $C_2$  axes. Has  $\sigma_h$ .

Therefore,  $\text{IF}_4^-$  is of the  $D_{4h}$  point group. ☐

**e.**  $\text{PH}_2^-$ 

*Answer.* Not low or high symmetry. Has a  $C_2$  axis. No perpendicular  $C_2$  axes. No  $\sigma_h$ . Has two perpendicular  $\sigma_v$  planes.

Therefore,  $\text{PH}_2^-$  is of the  $C_{2v}$  point group. ☐

**f.**  $\text{TeF}_4^{2-}$ 

*Answer.* Not low or high symmetry. Has a  $C_4$  axis. Has 4 perpendicular  $C_2$  axes. Has  $\sigma_h$ .

Therefore,  $\text{TeF}_4^{2-}$  is of the  $D_{4h}$  point group. ☐

**g.**  $\text{N}_3^-$ 

*Answer.*  $\text{N}_3^-$  is of the  $D_{\infty h}$  point group. ☐

**h.**  $\text{SeOCl}_4$ 

*Answer.* Not low or high symmetry. Has a  $C_2$  axis. No perpendicular  $C_2$  axes. No  $\sigma_h$ . Has two perpendicular  $\sigma_v$  planes.

Therefore,  $\text{SeOCl}_4$  is of the  $C_{2v}$  point group. ☐

**i.**  $\text{PH}_4^+$ 

*Answer.*  $\text{PH}_4^+$  is of the  $T_d$  point group. ☐

**3.9****f.**  $\text{IO}(\text{OH})_5$ 

*Answer.* Not low or high symmetry. Has a  $C_4$  axis. No perpendicular  $C_2$  axes. No  $\sigma_h$ . Has two perpendicular  $\sigma_v$  planes and two perpendicular  $\sigma_d$  planes.

Therefore,  $\text{IO}(\text{OH})_5$  is of the  $C_{4v}$  point group. ☐

**g.**  $\text{SOCl}_2$ 

*Answer.*  $\text{SOCl}_2$  is of the  $C_s$  point group. ☐

**h.**  $\text{ClOF}_4^-$ 

*Answer.* Not low or high symmetry. Has a  $C_4$  axis. No perpendicular  $C_2$  axes. No  $\sigma_h$ . Has two perpendicular  $\sigma_v$  planes and two perpendicular  $\sigma_d$  planes.

Therefore,  $\text{ClOF}_4^-$  is of the  $C_{4v}$  point group. ☐

**i.**  $\text{XeO}_2\text{F}_2$ 

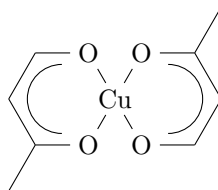
*Answer.* Not low or high symmetry. Has a  $C_2$  axis. No perpendicular  $C_2$  axes. No  $\sigma_h$ . Has two perpendicular  $\sigma_v$  planes.

Therefore,  $\text{XeO}_2\text{F}_2$  is of the  $C_{2v}$  point group. ☐



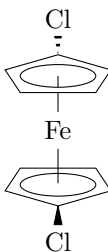
III) Assign the symmetry point group of the following molecules and objects. Ignore the H atoms in (a), (e), and (g). Note that (e) has pseudooctahedral geometry and (g) is square-planar.

a) The molecule pictured below.



*Answer.* Not low or high symmetry. Has a  $C_2$  axis. No perpendicular  $C_2$  axes. Has a  $\sigma_h$ .  
Therefore, the above molecule is of the  $C_{2h}$  point group. □

b) The molecule pictured below.



*Answer.* Not low or high symmetry. Has a  $C_2$  axis. No perpendicular  $C_2$  axes. Has a  $\sigma_h$ .  
Therefore, the above molecule is of the  $C_{2h}$  point group. □

c)  $\text{POCl}_3$

*Answer.* Not low or high symmetry. Has a  $C_3$  axis. No perpendicular  $C_2$  axes. No  $\sigma_h$ . Has three  $\sigma_v$  planes all offset by  $60^\circ$ .

Therefore,  $\text{POCl}_3$  is of the  $C_{3v}$  point group. □

d) Tennis ball (including the seam)

*Answer.* Not low or high symmetry. Has a  $C_2$  axis. Has 2 perpendicular  $C_2$  axes. No  $\sigma_h$ . Has two perpendicular  $\sigma_d$  planes.

Therefore, a tennis ball is of the  $D_{2d}$  point group. □

e)  $\text{trans}[\text{CrCl}_2(\text{H}_2\text{O})_4]^+$

*Answer.* Not low or high symmetry. Has a  $C_4$  axis. Has 4 perpendicular  $C_2$  axes. Has  $\sigma_h$ .

Therefore,  $\text{trans}[\text{CrCl}_2(\text{H}_2\text{O})_4]^+$  is of the  $D_{4h}$  point group. □

f) 1,3,5-trichlorobenzene.

*Answer.* Not low or high symmetry. Has a  $C_3$  axis. Has 3 perpendicular  $C_2$  axes. Has  $\sigma_h$ .

Therefore, 1,3,5-trichlorobenzene is of the  $D_{3h}$  point group. □

g)  $\text{trans-Pt}(\text{NH}_3)_2\text{Cl}_2$

*Answer.* Not low or high symmetry. Has a  $C_2$  axis. Has 2 perpendicular  $C_2$  axes. Has  $\sigma_h$ .

Therefore,  $\text{trans-Pt}(\text{NH}_3)_2\text{Cl}_2$  is of the  $D_{2h}$  point group. □

h)  $\text{SF}_5\text{Cl}$

*Answer.* Not low or high symmetry. Has a  $C_4$  axis. No perpendicular  $C_2$  axes. No  $\sigma_h$ . Has two perpendicular  $\sigma_v$  planes and two perpendicular  $\sigma_d$  planes.

Therefore,  $\text{SF}_5\text{Cl}$  is of the  $C_{4v}$  point group. □

i)  $\text{BFClBr}$

*Answer.*  $\text{BFClBr}$  is of the  $C_s$  point group. □

j)  $\text{PF}_2^+$

*Answer.* Not low or high symmetry. Has a  $C_2$  axis. No perpendicular  $C_2$  axes. No  $\sigma_h$ . Has two perpendicular  $\sigma_v$  planes.

Therefore,  $\text{PF}_2^+$  is of the  $C_{2v}$  point group. □

- IV) In the octahedral ion  $\text{FeF}_6^{3-}$ , what symmetry elements are destroyed if two *trans* F ions are moved away from the  $\text{Fe}^{3+}$  center in an equidistant fashion?

*Answer.* If the described change is made, the point group changes from  $O_h$  to  $D_{4h}$ . In this change, every  $C_3$  and  $S_6$  axis, two of the three  $C_4$  axes, four  $C_2$  axes, and every  $\sigma_d$  that does not contain the axis along which the F ions are stretched are destroyed.  $\square$

## 2 Representations, Character Tables, and Vibrations

1/28: I) Do the following problem from your text: Chapter 4: #22.

**4.22** Using the  $D_{2d}$  character table,

a. Determine the order of the group.

*Answer.*  $h = 8$ ; count the number of symmetry elements.  $\square$

b. Verify that the  $E$  irreducible representation is orthogonal to each of the other irreducible representations.

*Answer.*

$$\sum_{R_c} g_c \chi_E(R_c) \chi_{A_1}(R_c) = (1)(2)(1) + (2)(0)(1) + (1)(-2)(1) + (2)(0)(1) + (2)(0)(1) = 0$$

$$\sum_{R_c} g_c \chi_E(R_c) \chi_{A_2}(R_c) = (1)(2)(1) + (2)(0)(1) + (1)(-2)(1) + (2)(0)(-1) + (2)(0)(-1) = 0$$

$$\sum_{R_c} g_c \chi_E(R_c) \chi_{B_1}(R_c) = (1)(2)(1) + (2)(0)(-1) + (1)(-2)(1) + (2)(0)(1) + (2)(0)(-1) = 0$$

$$\sum_{R_c} g_c \chi_E(R_c) \chi_{B_2}(R_c) = (1)(2)(1) + (2)(0)(-1) + (1)(-2)(1) + (2)(0)(-1) + (2)(0)(1) = 0$$

$\square$

c. For each of the irreducible representations, verify that the sum of the squares of the characters equals the order of the group.

*Answer.*

$$\sum_{R_c} g_c [\chi_{A_1}(R_c)]^2 = 1 \cdot 1^2 + 2 \cdot 1^2 + 1 \cdot 1^2 + 2 \cdot 1^2 + 2 \cdot 1^2 = 8$$

$$\sum_{R_c} g_c [\chi_{A_2}(R_c)]^2 = 1 \cdot 1^2 + 2 \cdot 1^2 + 1 \cdot 1^2 + 2 \cdot (-1)^2 + 2 \cdot (-1)^2 = 8$$

$$\sum_{R_c} g_c [\chi_{B_1}(R_c)]^2 = 1 \cdot 1^2 + 2 \cdot (-1)^2 + 1 \cdot 1^2 + 2 \cdot 1^2 + 2 \cdot (-1)^2 = 8$$

$$\sum_{R_c} g_c [\chi_{B_2}(R_c)]^2 = 1 \cdot 1^2 + 2 \cdot (-1)^2 + 1 \cdot 1^2 + 2 \cdot (-1)^2 + 2 \cdot 1^2 = 8$$

$$\sum_{R_c} g_c [\chi_E(R_c)]^2 = 1 \cdot 2^2 + 2 \cdot 0^2 + 1 \cdot (-2)^2 + 2 \cdot 0^2 + 2 \cdot 0^2 = 8$$

$\square$

d. Reduce the following representations to their component irreducible representations.

| $D_{2d}$   | $E$ | $2S_4$ | $C_2$ | $2C'_2$ | $2\sigma_d$ |
|------------|-----|--------|-------|---------|-------------|
| $\Gamma_1$ | 6   | 0      | 2     | 2       | 2           |
| $\Gamma_2$ | 6   | 4      | 6     | 2       | 0           |

*Answer.* For  $\Gamma_1$ :

$$a_{A_1} = \frac{1}{8} \sum_{R_c} g_c \chi_{\Gamma_1}(R_c) \chi_{A_1}(R_c) = \frac{1}{8} [(1)(6)(1) + (2)(0)(1) + (1)(2)(1) + (2)(2)(1) + (2)(2)(1)] = 2$$

$$a_{A_2} = \frac{1}{8} \sum_{R_c} g_c \chi_{\Gamma_1}(R_c) \chi_{A_2}(R_c) = \frac{1}{8} [(1)(6)(1) + (2)(0)(1) + (1)(2)(1) + (2)(2)(-1) + (2)(2)(-1)] = 0$$

$$a_{B_1} = \frac{1}{8} \sum_{R_c} g_c \chi_{\Gamma_1}(R_c) \chi_{B_1}(R_c) = \frac{1}{8} [(1)(6)(1) + (2)(0)(-1) + (1)(2)(1) + (2)(2)(1) + (2)(2)(-1)] = 1$$

$$a_{B_2} = \frac{1}{8} \sum_{R_c} g_c \chi_{\Gamma_1}(R_c) \chi_{B_2}(R_c) = \frac{1}{8} [(1)(6)(1) + (2)(0)(-1) + (1)(2)(1) + (2)(2)(-1) + (2)(2)(1)] = 1$$

$$a_E = \frac{1}{8} \sum_{R_c} g_c \chi_{\Gamma_1}(R_c) \chi_E(R_c) = \frac{1}{8} [(1)(6)(2) + (2)(0)(0) + (1)(2)(-2) + (2)(2)(0) + (2)(2)(0)] = 1$$

Therefore, we know that

$$\Gamma_1 = 2A_1 + B_1 + B_2 + E$$

For  $\Gamma_2$ :

$$a_{A_1} = \frac{1}{8} \sum_{R_c} g_c \chi_{\Gamma_2}(R_c) \chi_{A_1}(R_c) = \frac{1}{8} [(1)(6)(1) + (2)(4)(1) + (1)(6)(1) + (2)(2)(1) + (2)(0)(1)] = 3$$

$$a_{A_2} = \frac{1}{8} \sum_{R_c} g_c \chi_{\Gamma_2}(R_c) \chi_{A_2}(R_c) = \frac{1}{8} [(1)(6)(1) + (2)(4)(1) + (1)(6)(1) + (2)(2)(-1) + (2)(0)(-1)] = 2$$

$$a_{B_1} = \frac{1}{8} \sum_{R_c} g_c \chi_{\Gamma_2}(R_c) \chi_{B_1}(R_c) = \frac{1}{8} [(1)(6)(1) + (2)(4)(-1) + (1)(6)(1) + (2)(2)(1) + (2)(0)(-1)] = 1$$

$$a_{B_2} = \frac{1}{8} \sum_{R_c} g_c \chi_{\Gamma_2}(R_c) \chi_{B_2}(R_c) = \frac{1}{8} [(1)(6)(1) + (2)(4)(-1) + (1)(6)(1) + (2)(2)(-1) + (2)(0)(1)] = 0$$

$$a_E = \frac{1}{8} \sum_{R_c} g_c \chi_{\Gamma_2}(R_c) \chi_E(R_c) = \frac{1}{8} [(1)(6)(2) + (2)(4)(0) + (1)(6)(-2) + (2)(2)(0) + (2)(0)(0)] = 0$$

Therefore, we know that

$$\Gamma_2 = 3A_1 + 2A_2 + B_1$$

□

II) Decompose the following reducible representations into their irreducible components. Ordering of the classes is the same as in the character tables in Appendix C of your text.

a)  $D_{3h}$ : 5, 2, 1, 3, 0, 3

*Answer.*

$$a_{A'_1} = \frac{1}{12} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{A'_1}(R_c) = \frac{1}{12} [(1)(5)(1) + (2)(2)(1) + (3)(1)(1) + (1)(3)(1) + (2)(0)(1) + (3)(3)(1)] = 2$$

$$a_{A'_2} = \frac{1}{12} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{A'_2}(R_c) = \frac{1}{12} [(1)(5)(1) + (2)(2)(1) + (3)(1)(-1) + (1)(3)(1) + (2)(0)(1) + (3)(3)(-1)] = 0$$

$$a_{E'} = \frac{1}{12} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{E'}(R_c) = \frac{1}{12} [(1)(5)(2) + (2)(2)(-1) + (3)(1)(0) + (1)(3)(2) + (2)(0)(-1) + (3)(3)(0)] = 1$$

$$a_{A''_1} = \frac{1}{12} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{A''_1}(R_c) = \frac{1}{12} [(1)(5)(1) + (2)(2)(1) + (3)(1)(1) + (1)(3)(-1) + (2)(0)(-1) + (3)(3)(-1)] = 0$$

$$a_{A''_2} = \frac{1}{12} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{A''_2}(R_c) = \frac{1}{12} [(1)(5)(1) + (2)(2)(1) + (3)(1)(-1) + (1)(3)(-1) + (2)(0)(-1) + (3)(3)(1)] = 1$$

$$a_{E''} = \frac{1}{12} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{E''}(R_c) = \frac{1}{12} [(1)(5)(2) + (2)(2)(-1) + (3)(1)(0) + (1)(3)(-2) + (2)(0)(1) + (3)(3)(0)] = 0$$

Therefore, we know that

$$\Gamma = 2A'_1 + E' + A''_2$$

□

b)  $D_{3h}$ : 3, 0, -1, -3, 0, 1

*Answer.*

$$a_{A'_1} = \frac{1}{12} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{A'_1}(R_c) = \frac{1}{12} [(1)(3)(1) + (2)(0)(1) + (3)(-1)(1) + (1)(-3)(1) + (2)(0)(1) + (3)(1)(1)] = 0$$

$$a_{A'_2} = \frac{1}{12} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{A'_2}(R_c) = \frac{1}{12} [(1)(3)(1) + (2)(0)(1) + (3)(-1)(-1) + (1)(-3)(1) + (2)(0)(1) + (3)(1)(-1)] = 0$$

$$a_{E'} = \frac{1}{12} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{E'}(R_c) = \frac{1}{12} [(1)(3)(2) + (2)(0)(-1) + (3)(-1)(0) + (1)(-3)(2) + (2)(0)(-1) + (3)(1)(0)] = 0$$

$$a_{A''_1} = \frac{1}{12} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{A''_1}(R_c) = \frac{1}{12} [(1)(3)(1) + (2)(0)(1) + (3)(-1)(1) + (1)(-3)(-1) + (2)(0)(-1) + (3)(1)(-1)] = 0$$

$$a_{A''_2} = \frac{1}{12} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{A''_2}(R_c) = \frac{1}{12} [(1)(3)(1) + (2)(0)(1) + (3)(-1)(-1) + (1)(-3)(-1) + (2)(0)(-1) + (3)(1)(1)] = 1$$

$$a_{E''} = \frac{1}{12} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{E''}(R_c) = \frac{1}{12} [(1)(3)(2) + (2)(0)(-1) + (3)(-1)(0) + (1)(-3)(-2) + (2)(0)(1) + (3)(1)(0)] = 1$$

Therefore, we know that

$$\Gamma = A''_2 + E''$$

□

c)  $C_{2v}$ : 4, 0, 0, 0

*Answer.* We know the following by inspection.

$$\Gamma = A_1 + A_2 + B_1 + B_2$$

□

d)  $C_{2h}$ : 5, 1, 1, 1

*Answer.* We know the following by inspection.

$$\Gamma = 2A_g + B_g + A_u + B_u$$

□

e)  $T_d$ : 13, 1, 5, -3, -3

*Answer.*

$$a_{A_1} = \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{A_1}(R_c) = \frac{1}{24} [(1)(13)(1) + (8)(1)(1) + (3)(5)(1) + (6)(-3)(1) + (6)(-3)(1)] = 0$$

$$a_{A_2} = \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{A_2}(R_c) = \frac{1}{24} [(1)(13)(1) + (8)(1)(1) + (3)(5)(1) + (6)(-3)(-1) + (6)(-3)(-1)] = 3$$

$$a_E = \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_E(R_c) = \frac{1}{24} [(1)(13)(2) + (8)(1)(-1) + (3)(5)(2) + (6)(-3)(0) + (6)(-3)(0)] = 2$$

$$a_{T_1} = \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{T_1}(R_c) = \frac{1}{24} [(1)(13)(3) + (8)(1)(0) + (3)(5)(-1) + (6)(-3)(1) + (6)(-3)(-1)] = 1$$

$$a_{T_2} = \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{T_2}(R_c) = \frac{1}{24} [(1)(13)(3) + (8)(1)(0) + (3)(5)(-1) + (6)(-3)(-1) + (6)(-3)(1)] = 1$$

Therefore, we know that

$$\Gamma = 3A_2 + 2E + T_1 + T_2$$

□

f)  $T_h$ : 8, -1, -1, 4, 8, -1, -1, 4

*Answer.* With respect to the two doubly degenerate groups, we must add the two parts together and also double the order that we are dividing out. Note that  $\varepsilon = e^{2\pi i/3} = \cos\left(\frac{2\pi}{3}\right) + i \sin\left(\frac{2\pi}{3}\right) = -0.5 + i\frac{\sqrt{3}}{2}$  and, thus,  $\varepsilon^* = -0.5 + i\frac{\sqrt{3}}{2}$ . It follows that  $\varepsilon + \varepsilon^* = -1$ .

$$\begin{aligned} a_{A_g} &= \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{A_g}(R_c) \\ &= \frac{1}{24} [(1)(8)(1) + (4)(-1)(1) + (4)(-1)(1) + (3)(4)(1) + (1)(8)(1) + (4)(-1)(1) + (4)(-1)(1) + (3)(4)(1)] \\ &= 1 \end{aligned}$$

$$\begin{aligned} a_{A_u} &= \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{A_u}(R_c) \\ &= \frac{1}{24} [(1)(8)(1) + (4)(-1)(1) + (4)(-1)(1) + (3)(4)(1) + (1)(8)(-1) + (4)(-1)(-1) + (4)(-1)(-1) + (3)(4)(-1)] \\ &= 0 \end{aligned}$$

$$\begin{aligned} 2a_{E_g} &= \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{E_g}(R_c) \\ a_{E_g} &= \frac{1}{48} [(1)(8)(2) + (4)(-1)(-1) + (4)(-1)(-1) + (3)(4)(2) + (1)(8)(2) + (4)(-1)(-1) + (4)(-1)(-1) + (3)(4)(2)] \\ &= 2 \end{aligned}$$

$$2a_{E_u} = \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{E_u}(R_c)$$

$$\begin{aligned} a_{E_u} &= \frac{1}{48} [(1)(8)(2) + (4)(-1)(-1) + (4)(-1)(-1) + (3)(4)(2) + (1)(8)(-2) + (4)(-1)(1) + (4)(-1)(1) + (3)(4)(-2)] \\ &= 0 \end{aligned}$$

$$a_{T_g} = \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{T_g}(R_c)$$

$$\begin{aligned} &= \frac{1}{24} [(1)(8)(3) + (4)(-1)(0) + (4)(-1)(0) + (3)(4)(-1) + (1)(8)(3) + (4)(-1)(0) + (4)(-1)(0) + (3)(4)(-1)] \\ &= 1 \end{aligned}$$

$$a_{T_u} = \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma}(R_c) \chi_{T_u}(R_c)$$

$$\begin{aligned} &= \frac{1}{24} [(1)(8)(3) + (4)(-1)(0) + (4)(-1)(0) + (3)(4)(-1) + (1)(8)(-3) + (4)(-1)(0) + (4)(-1)(0) + (3)(4)(1)] \\ &= 0 \end{aligned}$$

Therefore, we know that

$$\Gamma = A_g + 2\{E_g\} + T_g$$

□



- III) Draw the set of  $s$ ,  $p$ , and  $d$  orbitals, indicating the Cartesian axes and the proper phases of the orbitals. By noting how each orbital is affected by the symmetry operations in the  $C_{2h}$  point group ( $E$ ,  $C_2$ ,  $i$ ,  $\sigma_h$ ), write an irreducible representation for each orbital. Compare your results with the listing of the orbitals in the character table in Appendix C of the text.

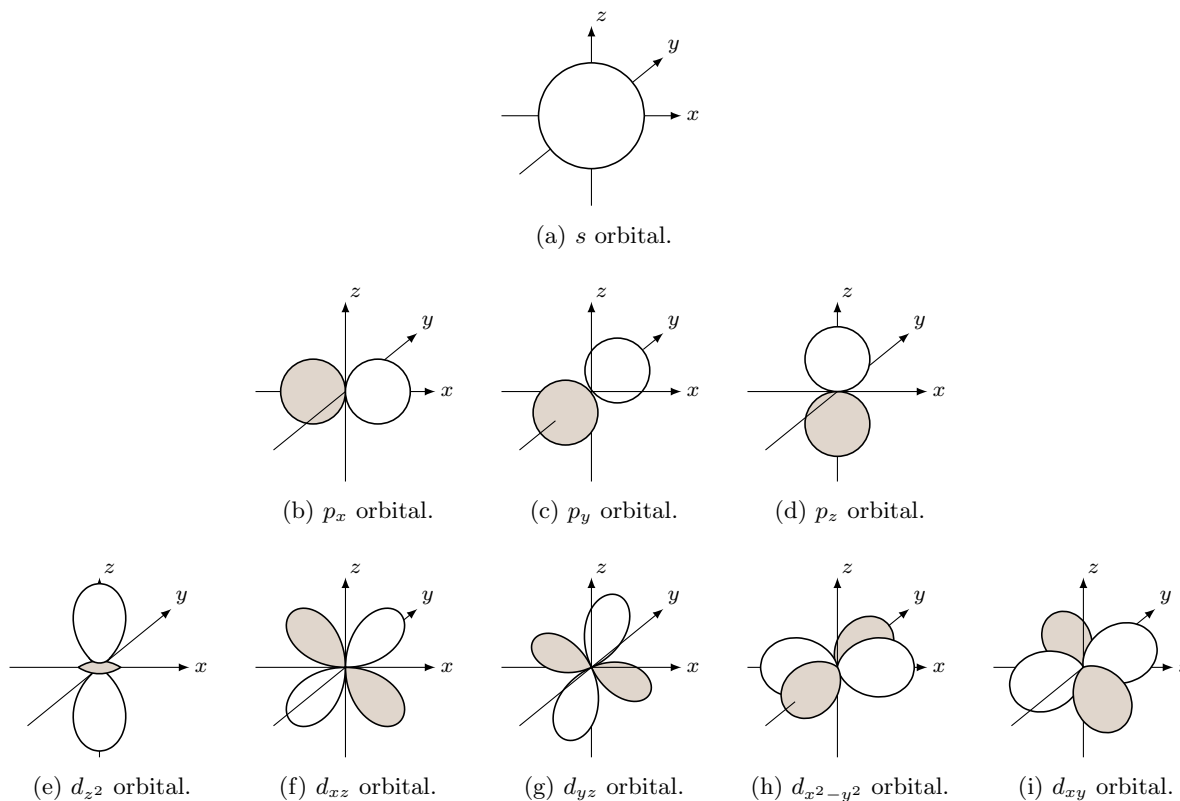


Figure 1:  $s$ ,  $p$ , and  $d$  orbitals.

*Answer.* In Figure 1, white shading means positive phase and orange means negative phase.

Here are irreducible representations for each orbital:

$$\begin{aligned}\Gamma_s &= (1, 1, 1, 1) = A_g \\ \Gamma_{p_x} &= (1, -1, -1, 1) = B_u \\ \Gamma_{p_y} &= (1, -1, -1, 1) = B_u \\ \Gamma_{p_z} &= (1, 1, -1, -1) = A_u \\ \Gamma_{d_{z^2}} &= (1, 1, 1, 1) = A_g \\ \Gamma_{d_{xz}} &= (1, -1, 1, -1) = B_g \\ \Gamma_{d_{yz}} &= (1, -1, 1, -1) = B_g \\ \Gamma_{d_{x^2-y^2}} &= (1, 1, 1, 1) = A_g \\ \Gamma_{d_{xy}} &= (1, 1, 1, 1) = A_g\end{aligned}$$

□

- IV) The molecule  $\text{Co}(\text{CO})_4(\text{SiMe}_3)$  has a structure based on a trigonal bipyramid. The infrared spectrum of  $\text{Co}(\text{CO})_4(\text{SiMe}_3)$  exhibits three  $\nu(\text{CO})$  stretching vibrations at 2100, 2041, and 2009  $\text{cm}^{-1}$ . Draw the two possible structures based on the TBP geometry, assign their proper point groups, and use the infrared data to determine which is the correct structure.

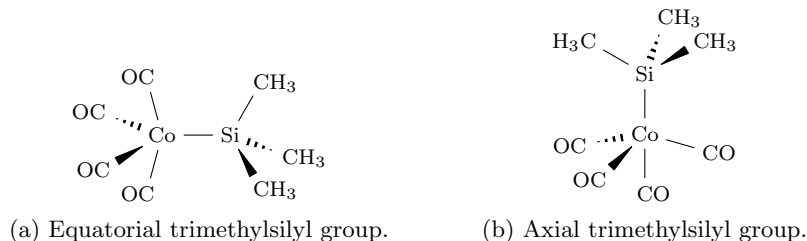


Figure 2: Structures of  $\text{Co}(\text{CO})_4(\text{SiMe}_3)$ .

*Answer.* Note that for the sake of point group assignments, the CO ligands will be treated as identical spheres. Additionally, the  $\text{SiMe}_3$  ligand will be treated as a sphere distinguishable from the CO “spheres.”

The structure in Figure 2a: Not high or low symmetry. Has a  $C_2$  axis. No perpendicular  $C_2$  axes. No  $\sigma_h$ . Has two perpendicular  $\sigma_v$  planes. Therefore, it is of the  $C_{2v}$  point group.

We can determine that  $\Gamma_\nu = (4, 0, 2, 2) = 2A_1 + B_1 + B_2$  by counting how many  $\overrightarrow{\text{Co}-\text{C}\ddot{\text{O}}}$  vectors stay the same under each symmetry operation and decomposing by inspection. With four stretching modes that are all IR active, we can expect there to be four peaks in the infrared spectrum of  $\text{Co}(\text{CO})_4(\text{SiMe}_3)$ . Therefore, this is not the correct structure.

The structure in Figure 2b: Not high or low symmetry. Has a  $C_3$  axis. No perpendicular  $C_2$  axes. No  $\sigma_h$ . Has three  $\sigma_v$  planes all offset by  $60^\circ$ . Therefore, it is of the  $C_{3v}$  point group.

We can determine that  $\Gamma_\nu = (4, 1, 2) = 2A_1 + E$  by counting how many  $\overrightarrow{\text{Co}-\text{C}\ddot{\text{O}}}$  vectors stay the same under each symmetry operation and decomposing by inspection. With four  $\nu(\text{CO})$  stretching modes that are all IR active (but two of the four being degenerate), we can expect there to be three peaks in the infrared spectrum of  $\text{Co}(\text{CO})_4(\text{SiMe}_3)$ . This confirms that this is the correct structure.  $\square$

V) Determine the number and symmetry types of normal vibrations in the following molecules:

- a)  $\text{H}_3\text{PBH}_3$ : with a staggered ethane-like geometry.

*Answer.*  $\text{H}_3\text{PBH}_3$  has  $3(8) - 6 = 18$  normal vibrations.

Not high or low symmetry. Has a  $C_3$  axis. No perpendicular  $C_2$  axes. No  $\sigma_h$ . Has three  $\sigma_v$  planes all offset by  $60^\circ$ . Therefore,  $\text{H}_3\text{PBH}_3$  is of the  $C_{3v}$  point group.

We can determine that  $\Gamma_{x,y,z} = (3, 0, 1)$ . We can also figure out that the number of atoms unmoved after applying each symmetry operation is  $(8, 2, 4)$ . Thus,  $\Gamma_{3N} = (24, 0, 4)$ . We can reduce this by inspection to  $\Gamma_{3N} = 6A_1 + 2A_2 + 8E$ .

Since  $\Gamma_{\text{trans}} = A_1 + E$  and  $\Gamma_{\text{rot}} = A_2 + E$ , we have by subtraction that  $\Gamma_{\text{vibs}} = 5A_1 + A_2 + 6E$ . Thus, of the 18 normal vibrations, 5 have symmetry  $A_1$ , 1 has symmetry  $A_2$ , and 12 have symmetry  $E$  (note that these 12 modes pair up into 6 pairs of vibration modes of the same type).  $\square$

- b)  $\text{Zr}_2\text{F}_{13}^{5-}$ : Each Zr is seven coordinate with monocapped trigonal prismatic geometry. The cap is a bridging F (linear Zr-F-Zr linkage) on the unique square face of the prism. The anionic complex has an eclipsed geometry about the bridging F.

*Answer.*  $\text{Zr}_2\text{F}_{13}^{5-}$  has  $3(15) - 6 = 39$  normal vibrations.

Not high or low symmetry. Has a  $C_2$  axis. Has 2 perpendicular  $C_2$  axes. Has a  $\sigma_h$  plane. Therefore,  $\text{Zr}_2\text{F}_{13}^{5-}$  is of the  $D_{2h}$  point group.

We can determine that  $\Gamma_{x,y,z} = (3, -1, -1, -1, -3, 1, 1, 1)$ . We can also figure out that the number of atoms unmoved after applying each symmetry operation is  $(15, 3, 1, 1, 1, 1, 7, 3)$ . Thus,  $\Gamma_{3N} = (45, -3, -1, -1, -3, 1, 7, 3)$ . We can reduce this as follows.

$$\begin{aligned} a_{A_g} &= \frac{1}{8} \sum_{R_c} g_c \chi_{\Gamma_{3N}}(R_c) \chi_{A_g}(R_c) \\ &= \frac{1}{8} [(1)(45)(1) + (1)(-3)(1) + (1)(-1)(1) + (1)(-1)(1) + (1)(-3)(1) + (1)(1)(1) + (1)(7)(1) + (1)(3)(1)] \\ &= 6 \end{aligned}$$

$$\begin{aligned} a_{B_{1g}} &= \frac{1}{8} \sum_{R_c} g_c \chi_{\Gamma_{3N}}(R_c) \chi_{B_{1g}}(R_c) \\ &= \frac{1}{8} [(1)(45)(1) + (1)(-3)(1) + (1)(-1)(-1) + (1)(-1)(-1) + (1)(-3)(1) + (1)(1)(1) + (1)(7)(-1) + (1)(3)(-1)] \\ &= 4 \end{aligned}$$

$$\begin{aligned} a_{B_{2g}} &= \frac{1}{8} \sum_{R_c} g_c \chi_{\Gamma_{3N}}(R_c) \chi_{B_{2g}}(R_c) \\ &= \frac{1}{8} [(1)(45)(1) + (1)(-3)(-1) + (1)(-1)(1) + (1)(-1)(-1) + (1)(-3)(1) + (1)(1)(-1) + (1)(7)(1) + (1)(3)(-1)] \\ &= 6 \end{aligned}$$

$$\begin{aligned} a_{B_{3g}} &= \frac{1}{8} \sum_{R_c} g_c \chi_{\Gamma_{3N}}(R_c) \chi_{B_{3g}}(R_c) \\ &= \frac{1}{8} [(1)(45)(1) + (1)(-3)(-1) + (1)(-1)(-1) + (1)(-1)(1) + (1)(-3)(1) + (1)(1)(-1) + (1)(7)(-1) + (1)(3)(1)] \\ &= 5 \end{aligned}$$

$$\begin{aligned}
a_{A_u} &= \frac{1}{8} \sum_{R_c} g_c \chi_{\Gamma_{3N}}(R_c) \chi_{A_u}(R_c) \\
&= \frac{1}{8} [(1)(45)(1) + (1)(-3)(1) + (1)(-1)(1) + (1)(-1)(1) + (1)(-3)(-1) + (1)(1)(-1) + (1)(7)(-1) + (1)(3)(-1)] \\
&= 4
\end{aligned}$$

$$\begin{aligned}
a_{B_{1u}} &= \frac{1}{8} \sum_{R_c} g_c \chi_{\Gamma_{3N}}(R_c) \chi_{B_{1u}}(R_c) \\
&= \frac{1}{8} [(1)(45)(1) + (1)(-3)(1) + (1)(-1)(-1) + (1)(-1)(-1) + (1)(-3)(-1) + (1)(1)(-1) + (1)(7)(1) + (1)(3)(1)] \\
&= 7
\end{aligned}$$

$$\begin{aligned}
a_{B_{2u}} &= \frac{1}{8} \sum_{R_c} g_c \chi_{\Gamma_{3N}}(R_c) \chi_{B_{2u}}(R_c) \\
&= \frac{1}{8} [(1)(45)(1) + (1)(-3)(-1) + (1)(-1)(1) + (1)(-1)(-1) + (1)(-3)(-1) + (1)(1)(1) + (1)(7)(-1) + (1)(3)(1)] \\
&= 6
\end{aligned}$$

$$\begin{aligned}
a_{B_{3u}} &= \frac{1}{8} \sum_{R_c} g_c \chi_{\Gamma_{3N}}(R_c) \chi_{B_{3u}}(R_c) \\
&= \frac{1}{8} [(1)(45)(1) + (1)(-3)(-1) + (1)(-1)(-1) + (1)(-1)(1) + (1)(-3)(-1) + (1)(1)(1) + (1)(7)(1) + (1)(3)(-1)] \\
&= 7
\end{aligned}$$

Therefore, we know that  $\Gamma_{3N} = 6A_g + 4B_{1g} + 6B_{2g} + 5B_{3g} + 4A_u + 7B_{1u} + 6B_{2u} + 7B_{3u}$ . Since  $\Gamma_{\text{trans}} = B_{1u} + B_{2u} + B_{3u}$  and  $\Gamma_{\text{rot}} = B_{1g} + B_{2g} + B_{3g}$ , we have by subtraction that  $\Gamma_{\text{vibs}} = 6A_g + 3B_{1g} + 5B_{2g} + 4B_{3g} + 4A_u + 6B_{1u} + 5B_{2u} + 6B_{3u}$ . Thus, of the 39 normal vibrations, 6 have symmetry  $A_g$ , 3 have symmetry  $B_{1g}$ , 5 have symmetry  $B_{2g}$ , 4 have symmetry  $B_{3g}$ , 4 have symmetry  $A_u$ , 6 have symmetry  $B_{1u}$ , 5 have symmetry  $B_{2u}$ , and 6 have symmetry  $B_{3u}$ .  $\square$

VI) Benzene ( $C_6H_6$ ) is a planar molecule.

a) Assign the symmetry group.

*Answer.* Not high or low symmetry. Has a  $C_6$  axis. Has 6 perpendicular  $C_2$  axes. Has a  $\sigma_h$  plane. Therefore,  $C_6H_6$  is of the  $D_{6h}$  point group.  $\square$

b) Determine the number and symmetries of the C–H stretching modes in benzene.

*Answer.* We can determine that  $\Gamma_\nu = (6, 0, 0, 0, 2, 0, 0, 0, 0, 6, 0, 2)$  by counting how many  $\overrightarrow{C-H}$  vectors stay the same under each symmetry operation. We can reduce this as follows.

$$\begin{aligned} a_{A_{1g}} &= \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma_\nu}(R_c) \chi_{A_{1g}}(R_c) \\ &= \frac{1}{24} [(1)(6)(1) + (2)(0)(1) + (2)(0)(1) + (1)(0)(1) + (3)(2)(1) + (3)(0)(1) \\ &\quad + (1)(0)(1) + (2)(0)(1) + (2)(0)(1) + (1)(6)(1) + (3)(0)(1) + (3)(2)(1)] \\ &= 1 \end{aligned}$$

$$\begin{aligned} a_{A_{2g}} &= \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma_\nu}(R_c) \chi_{A_{2g}}(R_c) \\ &= \frac{1}{24} [(1)(6)(1) + (2)(0)(1) + (2)(0)(1) + (1)(0)(1) + (3)(2)(-1) + (3)(0)(-1) \\ &\quad + (1)(0)(1) + (2)(0)(1) + (2)(0)(1) + (1)(6)(1) + (3)(0)(-1) + (3)(2)(-1)] \\ &= 0 \end{aligned}$$

$$\begin{aligned} a_{B_{1g}} &= \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma_\nu}(R_c) \chi_{B_{1g}}(R_c) \\ &= \frac{1}{24} [(1)(6)(1) + (2)(0)(-1) + (2)(0)(1) + (1)(0)(-1) + (3)(2)(1) + (3)(0)(-1) \\ &\quad + (1)(0)(1) + (2)(0)(-1) + (2)(0)(1) + (1)(6)(-1) + (3)(0)(1) + (3)(2)(-1)] \\ &= 0 \end{aligned}$$

$$\begin{aligned} a_{B_{2g}} &= \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma_\nu}(R_c) \chi_{B_{2g}}(R_c) \\ &= \frac{1}{24} [(1)(6)(1) + (2)(0)(-1) + (2)(0)(1) + (1)(0)(-1) + (3)(2)(-1) + (3)(0)(1) \\ &\quad + (1)(0)(1) + (2)(0)(-1) + (2)(0)(1) + (1)(6)(-1) + (3)(0)(-1) + (3)(2)(1)] \\ &= 0 \end{aligned}$$

$$\begin{aligned} a_{E_{1g}} &= \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma_\nu}(R_c) \chi_{E_{1g}}(R_c) \\ &= \frac{1}{24} [(1)(6)(2) + (2)(0)(1) + (2)(0)(-1) + (1)(0)(-2) + (3)(2)(0) + (3)(0)(0) \\ &\quad + (1)(0)(2) + (2)(0)(1) + (2)(0)(-1) + (1)(6)(-2) + (3)(0)(0) + (3)(2)(0)] \\ &= 0 \end{aligned}$$

$$\begin{aligned}
a_{E_{2g}} &= \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma_\nu}(R_c) \chi_{E_{2g}}(R_c) \\
&= \frac{1}{24} [(1)(6)(2) + (2)(0)(-1) + (2)(0)(-1) + (1)(0)(2) + (3)(2)(0) + (3)(0)(0) \\
&\quad + (1)(0)(2) + (2)(0)(-1) + (2)(0)(-1) + (1)(6)(2) + (3)(0)(0) + (3)(2)(0)] \\
&= 1
\end{aligned}$$

$$\begin{aligned}
a_{A_{1u}} &= \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma_\nu}(R_c) \chi_{A_{1u}}(R_c) \\
&= \frac{1}{24} [(1)(6)(1) + (2)(0)(1) + (2)(0)(1) + (1)(0)(1) + (3)(2)(1) + (3)(0)(1) \\
&\quad + (1)(0)(-1) + (2)(0)(-1) + (2)(0)(-1) + (1)(6)(-1) + (3)(0)(-1) + (3)(2)(-1)] \\
&= 0
\end{aligned}$$

$$\begin{aligned}
a_{A_{2u}} &= \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma_\nu}(R_c) \chi_{A_{2u}}(R_c) \\
&= \frac{1}{24} [(1)(6)(1) + (2)(0)(1) + (2)(0)(1) + (1)(0)(1) + (3)(2)(-1) + (3)(0)(-1) \\
&\quad + (1)(0)(-1) + (2)(0)(-1) + (2)(0)(-1) + (1)(6)(-1) + (3)(0)(1) + (3)(2)(1)] \\
&= 0
\end{aligned}$$

$$\begin{aligned}
a_{B_{1u}} &= \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma_\nu}(R_c) \chi_{B_{1u}}(R_c) \\
&= \frac{1}{24} [(1)(6)(1) + (2)(0)(-1) + (2)(0)(1) + (1)(0)(-1) + (3)(2)(1) + (3)(0)(-1) \\
&\quad + (1)(0)(-1) + (2)(0)(1) + (2)(0)(-1) + (1)(6)(1) + (3)(0)(-1) + (3)(2)(1)] \\
&= 1
\end{aligned}$$

$$\begin{aligned}
a_{B_{2u}} &= \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma_\nu}(R_c) \chi_{B_{2u}}(R_c) \\
&= \frac{1}{24} [(1)(6)(1) + (2)(0)(-1) + (2)(0)(1) + (1)(0)(-1) + (3)(2)(-1) + (3)(0)(1) \\
&\quad + (1)(0)(-1) + (2)(0)(1) + (2)(0)(-1) + (1)(6)(1) + (3)(0)(1) + (3)(2)(-1)] \\
&= 0
\end{aligned}$$

$$\begin{aligned}
a_{E_{1u}} &= \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma_\nu}(R_c) \chi_{E_{1u}}(R_c) \\
&= \frac{1}{24} [(1)(6)(2) + (2)(0)(1) + (2)(0)(-1) + (1)(0)(-2) + (3)(2)(0) + (3)(0)(0) \\
&\quad + (1)(0)(-2) + (2)(0)(-1) + (2)(0)(1) + (1)(6)(2) + (3)(0)(0) + (3)(2)(0)] \\
&= 1
\end{aligned}$$

$$\begin{aligned}
a_{E_{2u}} &= \frac{1}{24} \sum_{R_c} g_c \chi_{\Gamma_\nu}(R_c) \chi_{E_{2u}}(R_c) \\
&= \frac{1}{24} [(1)(6)(2) + (2)(0)(-1) + (2)(0)(-1) + (1)(0)(2) + (3)(2)(0) + (3)(0)(0) \\
&\quad + (1)(0)(-2) + (2)(0)(1) + (2)(0)(1) + (1)(6)(-2) + (3)(0)(0) + (3)(2)(0)] \\
&= 0
\end{aligned}$$

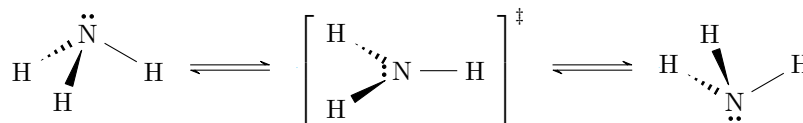
Therefore, we know that  $\Gamma_\nu = A_{1g} + E_{2g} + B_{1u} + E_{1u}$ . Thus, there are 6 C–H stretching modes in benzene: 1 with symmetry  $A_{1g}$ , 2 with symmetry  $E_{2g}$  (note that these 2 modes form a pair vibration modes of the same type), 1 with symmetry  $B_{1u}$ , and 2 with symmetry  $E_{1u}$  (again, these pair up).  $\square$

- c) Determine the Raman and infrared activities for each vibration.

*Answer.* Since the  $A_{1g}$ ,  $E_{1g}$ , and  $E_{2g}$  irreducible representations are Raman active, we know that the  $A_{1g}$  and  $E_{2g}$  C–H stretching modes are Raman active. The others are Raman silent. Additionally, since the  $A_{2u}$  and  $E_{1u}$  irreducible representations are IR active and two of the C–H stretching vibration modes are of the  $E_{1u}$  type, the  $E_{1u}$  vibrations are both IR active. The others are IR silent.  $\square$

### 3 Constructing Molecular Orbitals

- 2/4: I) Ammonia undergoes a facile inversion (“umbrella flip”) as shown below. The activation barrier for inversion is low ( $\Delta G^\ddagger \sim 5$  kcal/mol), and the transition state for this motion is planar  $\text{NH}_3$ . Note that the relevant valence shell IP’s are  $\text{N}_{2s} = -26.0$  eV,  $\text{N}_{2p} = -13.4$  eV, and  $\text{H}_{1s} = -13.6$  eV.



- a) Construct an MO diagram for *planar*  $\text{NH}_3$ .

*Answer.* Point group:  $D_{3h}$

Basis functions: all three H orbitals,  $\text{N}_{2s}$ ,  $\text{N}_{2p_x}$ ,  $\text{N}_{2p_y}$ , and  $\text{N}_{2p_z}$ .

Apply operations, generate reducible representations, and reduce to irreducible representations:

$$\begin{aligned}\Gamma_{\text{H}} &= (3, 0, 1, 3, 0, 1) = A'_1 + E' \\ \Gamma_{\text{N}_{2s}} &= A'_1 \\ \Gamma_{\text{N}_{2p_x}} &= E' \\ \Gamma_{\text{N}_{2p_y}} &= E' \\ \Gamma_{\text{N}_{2p_z}} &= A''_2\end{aligned}$$

Combine central and peripheral orbitals by their symmetry:

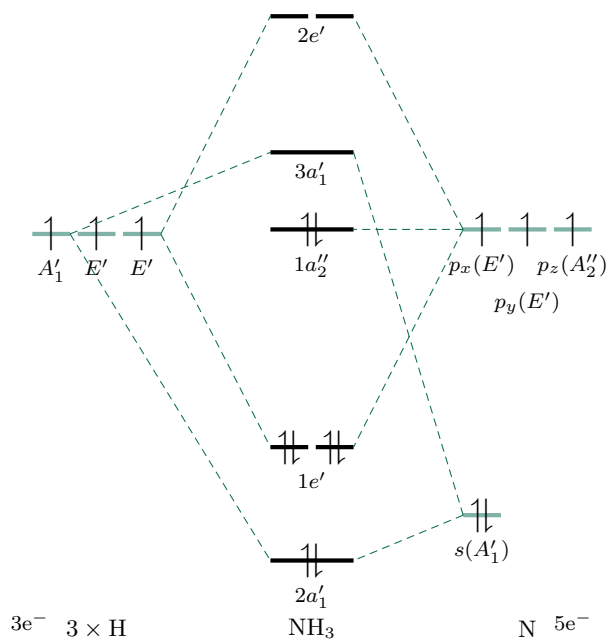


Figure 3: Planar  $\text{NH}_3^\ddagger$  orbital diagram.

□

- b) Label the MOs with the appropriate Mulliken symbols ( $a_{1g}$ ,  $e_g$ , etc.) and add electrons to show the proper orbital occupancies.

*Answer.* See Figure 3.

□



- c) Compare your MO diagram with that for pyramidal  $\text{NH}_3$  (Figure 5.30 in your text), and comment qualitatively on why this process is a low-energy one.

*Answer.* It appears that the only change between the two MO diagrams is that the two  $3a_1$  electrons in the pyramidal  $\text{NH}_3$  diagram must be excited to the  $1a_2''$  orbital in the planar  $\text{NH}_3$  diagram. Since  $1a_2''$  is higher in energy than  $3a_1$ , there will be an increase in energy, but since it is only marginally higher, the increase will be very small.  $\square$

- d) What vibrational mode is responsible for the inversion?

*Answer.* If any vibrational mode is responsible for the inversion, it certainly won't be a stretching mode since these have no effect on molecular geometry about the central atom. On the other hand, a bending mode could well achieve such a transition. Thus, we will find the bending modes in both pyramidal and planar  $\text{NH}_3$  and compare.

For pyramidal  $\text{NH}_3$ , we can determine that  $\Gamma_{x,y,z} = (3, 0, 1)$ . We can also figure out that the number of atoms unmoved after applying each symmetry operation is  $(4, 1, 2)$ . Thus,  $\Gamma_{3N} = (12, 0, 2)$ . We can decompose this by inspection to  $\Gamma_{3N} = 3A_1 + A_2 + 4E$ . Since  $\Gamma_{\text{trans}} = A_1 + E$  and  $\Gamma_{\text{rot}} = A_2 + E$ , we have by subtraction that  $\Gamma_{\text{vibs}} = 2A_1 + 2E$ .

We can determine that  $\Gamma_\nu = (3, 0, 1)$  by counting how many  $\overrightarrow{\text{N}-\text{H}}$  vectors stay the same under each symmetry operation. We can decompose this by inspection to  $\Gamma_\nu = A_1 + E$ . Thus, we have by subtraction that  $\Gamma_\delta = A_1 + E$ .

For planar  $\text{NH}_3$ , we can determine that  $\Gamma_{x,y,z} = (3, 0, -1, 1, -2, 1)$ . We can also figure out that the number of atoms unmoved after applying each symmetry operation is  $(4, 1, 2, 4, 1, 2)$ . Thus,  $\Gamma_{3N} = (12, 0, -2, 4, -2, 2)$ . We can decompose this by repeated applications of the reduction formula to  $\Gamma_{3N} = A'_1 + A'_2 + 3E' + 2A''_2 + E''$ . Since  $\Gamma_{\text{trans}} = E' + A''_2$  and  $\Gamma_{\text{rot}} = A'_2 + E''$ , we have by subtraction that  $\Gamma_{\text{vibs}} = A'_1 + 2E' + A''_2$ .

We can determine that  $\Gamma_\nu = (3, 0, 1, 3, 0, 1)$  by counting how many  $\overrightarrow{\text{N}-\text{H}}$  vectors stay the same under each symmetry operation. We can decompose this by inspection to  $\Gamma_\nu = A'_1 + E'$ . Thus, we have by subtraction that  $\Gamma_\delta = E' + A''_2$ .

Since the  $E$  pyramidal bending modes transform into the analogous  $E'$  planar bending modes, but the  $A_1$  pyramidal bending mode has no planar analogue, it is the  $A_1$  bending mode in pyramidal  $\text{NH}_3$  that causes the inversion.  $\square$

II)

- a) Use group theory to construct an MO diagram for octahedral  $\text{SF}_6$ . Consider only  $\sigma$ -bonding between S and the F's and use only the sulfur 3s and 3p valence orbitals (i.e., ignore the 3d-orbital involvement). For fluorine, just use a " $\sigma$ -type" orbital to determine the  $6 \times \text{F}$  group orbitals.

*Answer.* Point group:  $O_h$

Basis functions: all six F orbitals,  $S_{3s}$ ,  $S_{3p_x}$ ,  $S_{3p_y}$ , and  $S_{3p_z}$ .

Apply operations, generate reducible representations, and reduce to irreducible representations:

$$\begin{aligned}\Gamma_F &= (6, 0, 0, 2, 2, 0, 0, 0, 4, 2) = A_{1g} + E_g + T_{1u} \\ \Gamma_{S_{3s}} &= A_{1g} \\ \Gamma_{S_{3p_x}} &= T_{1u} \\ \Gamma_{S_{3p_y}} &= T_{1u} \\ \Gamma_{S_{3p_z}} &= T_{1u}\end{aligned}$$

Combine central and peripheral orbitals by their symmetry: □

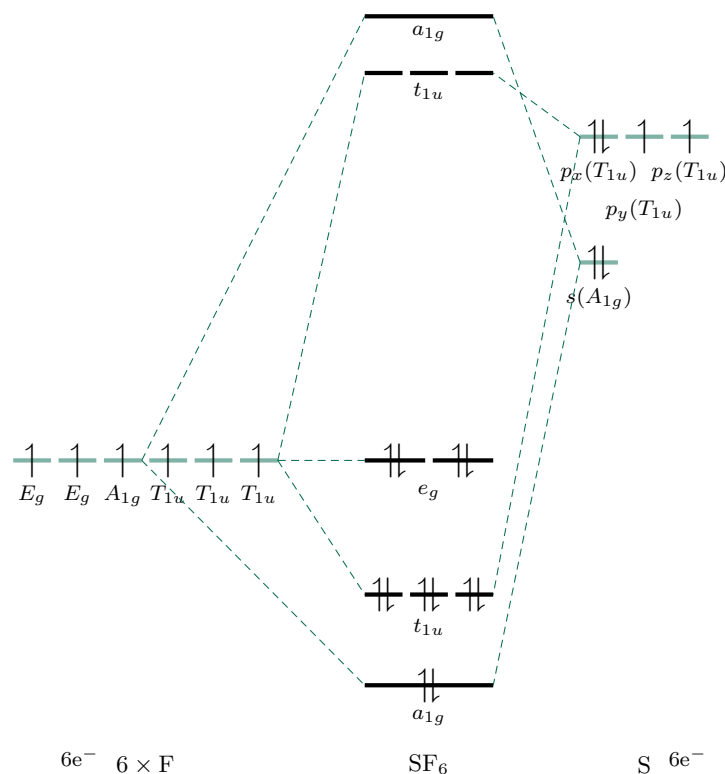


Figure 4:  $\text{SF}_6$  orbital diagram.

- b) Label the MO's with the appropriate Mulliken symbols and show the orbital occupancies (i.e., fill in the MO levels with the proper number of electrons).

*Answer.* See Figure 4. □

- c) Based on the MO diagram, comment on the number of bonding electrons in  $\text{SF}_6$  and the bond-order of each S–F bond.

*Answer.* There are 8 bonding electrons (the two in the  $1a_{1g}$  orbital, and the six in the degenerate  $1t_{1u}$  orbitals; the four in the degenerate  $1e_g$  orbitals are nonbonding and all anti-bonding orbitals are unfilled). Since the bond order is one half the number of bonding electrons divided by the number of bonds, we have  $\text{B.O.} = \frac{2}{3}$ .  $\square$

## 4 Band Theory and Acid-Base Interactions

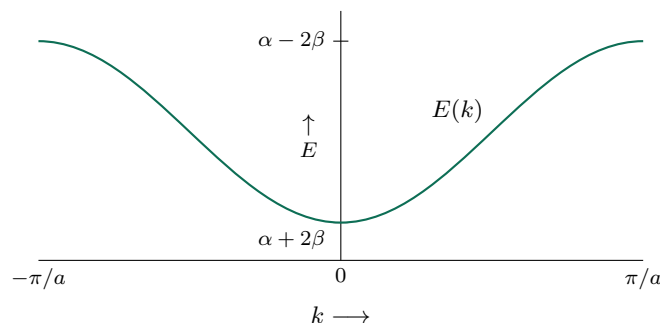
- 2/15: I) In our description for a one-dimensional  $s$ -orbital chain with period  $a$ , the allowed energy states broaden into a band whose dispersion (dependence of the energy  $E$  on the wavevector  $k$ ) is given by

$$E(k) = \alpha + 2\beta \cos ka$$

where  $\beta$  is the nearest-neighbor interaction energy and is less than zero for neighboring  $s$ -type orbitals.

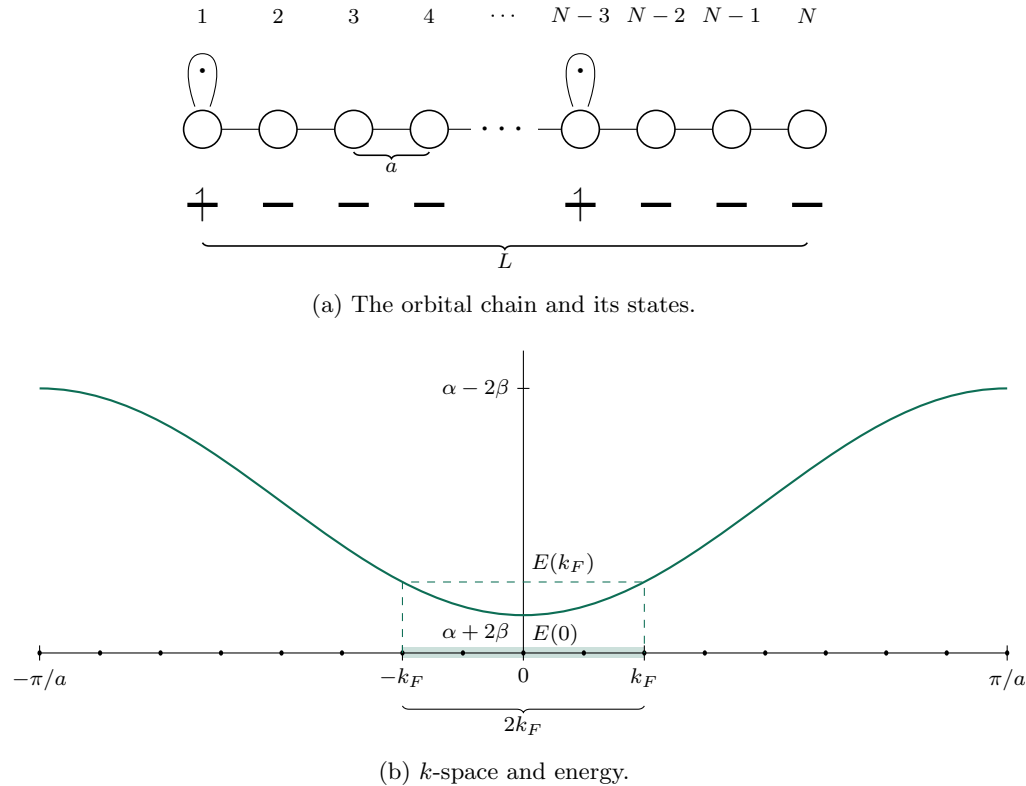
- a. Sketch  $E(k)$  over the range  $-\frac{\pi}{a} < k < \frac{\pi}{a}$ , assuming that  $\beta < 0$ .

*Answer.*



□

- b. Suppose there is one electron for every 4 atoms; that is, if there are  $N$  atoms in the chain, there are  $N/4$  electrons. With  $a = 2.00 \times 10^{-10}$  m (i.e.,  $2.00 \text{ \AA}$ ) and  $\beta = -2.00 \text{ eV}$ , calculate the value of  $k$  corresponding to the highest occupied electronic state (also known as Fermi wavevector  $k_F$ , the wavevector at the Fermi surface) and the height of the Fermi energy  $E_F$  above the bottom of the band. Note that a chain of  $N$  atoms will have a total length  $L = Na$  and that states are uniformly distributed in the  $k$ -space.

Figure 5: Theory describing a one-dimensional  $s$ -orbital chain.

*Answer.* As in Figure 5a, consider a one-dimensional chain of  $N$  atoms with one  $s$ -orbital each. The orbitals are separated by period  $a$ , and the length of the chain is  $L = Na$ . As in Figure 5b, each orbital contributes a state indexed by  $k$ . These states are evenly distributed in one-dimensional  $k$ -space between  $-\frac{\pi}{a}$  and  $\frac{\pi}{a}$ . Now imagine that the one-dimensional  $s$ -orbital chain is devoid of electrons, and we have to fill them in. The first electron will go in the lowest energy state at  $k = 0$  (see Figure 5b). The second will go into either  $k = a$  or  $k = -a$ , and the third will go into the other one. As we add more and more electrons, the states that they go into will expand outwards evenly from the origin, always filling the lowest energy states available first. Eventually, all electrons will be filled in; however, this does not mean that all states will be filled, as there is one electron for every four atoms and, thus, for every eight states (again, see Figure 5a). Indeed, some region evenly surrounding the origin will be completely filled in, and by definition the Fermi surface will be the boundary between occupied and unoccupied states. Since the Fermi wavevector  $k_F$  is the radius of the Fermi surface (i.e., the distance from the origin to the boundary of the Fermi sphere), we can label the boundary points  $k_F$  and  $-k_F$ . Having established the theoretical design, let's start the actual calculations.

Since  $k$ -space within the first Brillouin zone runs from  $-\frac{\pi}{a}$  to  $\frac{\pi}{a}$ , the length of the  $s$ -orbital chain in  $k$ -space is  $L = \frac{\pi}{a} - (-\frac{\pi}{a}) = \frac{2\pi}{a}$ . It follows that each orbital is separated by  $a = \frac{2\pi}{L}$ , and that there are  $\frac{L}{2\pi}$  orbitals per unit length. Now the length of the occupied region is  $k_F - (-k_F) = 2k_F$ , so combining this with the previous result gives us that there are  $2k_F \times \frac{L}{2\pi} = \frac{k_F L}{\pi}$  occupied orbitals. Since each orbital adds 2 quantum states, we have  $\frac{2k_F L}{\pi}$  occupied states. But the number of occupied states is simply equal to the number of electrons  $\frac{N}{4}$ . Since we also know that  $L = Na$ ,

we have that

$$\begin{aligned}\frac{2k_F Na}{\pi} &= \frac{N}{4} \\ \frac{k_F a}{\pi} &= \frac{1}{8} \\ k_F &= \frac{\pi}{8a}\end{aligned}$$

$$k_F = 1.96 \times 10^9 \text{ m}^{-1}$$

As to the other part of the question, refer to Figure 5b once again. It tells us that the height of the Fermi energy above the bottom of the band is given by

$$\begin{aligned}h &= E(k_F) - E(0) \\ &= \alpha + 2\beta \cos\left(\frac{\pi}{8a} \cdot a\right) - (\alpha + 2\beta \cos(0 \cdot a)) \\ &= 2\beta \left(\cos\left(\frac{\pi}{8}\right) - 1\right)\end{aligned}$$

$$h = 0.304 \text{ eV}$$

□

II) Do the following problems from Chapter 6: 13, 25, 30, 37.

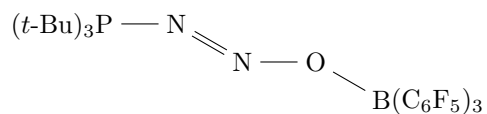
**6.13** If an equimolar mixture of  $P(t\text{-C}_4\text{H}_9)_3$  and  $B(\text{C}_6\text{F}_5)_3$  is mixed with 1 bar of the gas  $\text{N}_2\text{O}$  in bromobenzene solution, a white product is formed in good yield. A variety of NMR evidence has been gathered on the product: there is a single  $^{31}\text{P}$  NMR resonance;  $^{11}\text{B}$  and  $^{19}\text{F}$  NMR are considered consistent with a 4-coordinate boron atom; and  $^{15}\text{N}$  NMR indicates two nonequivalent nitrogen atoms. In addition, no gas is released in the reaction.

a. Suggest the role of  $\text{N}_2\text{O}$  in this reaction.

*Answer.*  $\text{N}_2\text{O}$  reorganizes into a bridge linking the phosphorous and boron central atoms of  $P(t\text{-C}_4\text{H}_9)_3$  and  $B(\text{C}_6\text{F}_5)_3$ , respectively.  $P(t\text{-C}_4\text{H}_9)_3$  contains a frustrated Lewis pair that could not normally react with the boron in  $B(\text{C}_6\text{F}_5)_3$ , but by virtue of this bridge, there is enough separation between all of the substituents (i.e., steric hindrance is sufficiently reduced) to permit the formation of a combined product.  $\square$

b. Propose a structure of the product.

*Answer.*



$\square$

- 6.25** The most common source of mercury is cinnabar ( $\text{HgS}$ ), whereas Zn and Cd in the same group occur as sulfide, carbonate, silicate, and oxide. Why?

*Answer.* Mercury is soft, so it forms a stable product with the similarly soft sulfide ion. However, zinc and cadmium are intermediate-hard, so they can form stable products with harder anions, such as carbonate, silicate, and oxide.  $\square$



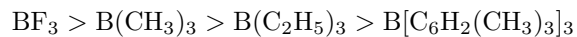
**6.30** CsI is much less soluble in water than CsF, and LiF is much less soluble than LiI. Why?

*Answer.* Cesium iodide is formed from two soft ions, which grants it extra stability in comparison to soft-hard compounds such as cesium fluoride. As such, CsI will dissolve much less in water than CsF since it prefers to be in the molecular state rather than the ionic state.

The same is true for the hard-hard compound lithium fluoride in comparison to the hard-soft compound lithium iodide. □

- 6.37** List the following acids in order of their acid strength when reacting with  $\text{NH}_3$ :  $\text{BF}_3$ ,  $\text{B}(\text{CH}_3)_3$ ,  $\text{B}(\text{C}_2\text{H}_5)_3$ , and  $\text{B}[\text{C}_6\text{H}_2(\text{CH}_3)_3]_3$ .

*Answer.* We have



$\text{BF}_3$  will be the strongest Lewis acid because it is the only one out of the collection that has electron-withdrawing groups as opposed to electron donating groups. Among the three compounds with electron-donating groups, we rank as stronger acids the ones with smaller groups, both because smaller groups donate less and the reduction in steric hindrance.  $\square$

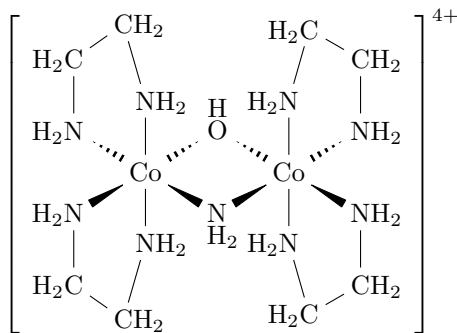
## 5 Coordination Complexes: Name-Structure Conversions and Isomers

2/19: I) Do the following problems from Chapter 9: 7; 9; 12a,b,e; 20; 23.

9.7 Give structures for the following:

a. Bis(en)Co(III)- $\mu$ -amido- $\mu$ -hydroxobis(en)Co(III) ion.

*Answer.*

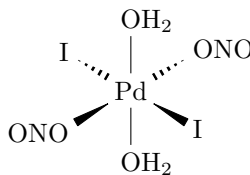


□

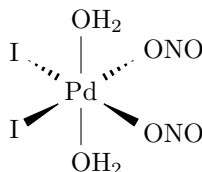
b. DiaquadiiododinitritoPd(IV), all isomers.

*Answer.* There are six stereoisomers of diaquadiiododinitritoPd(IV). We will classify them into five groups by the relative positions of like ligands as follows. One group will contain two enantiomers.

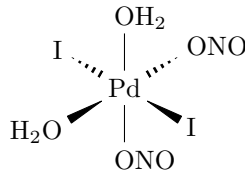
1. All like ligands *trans*.



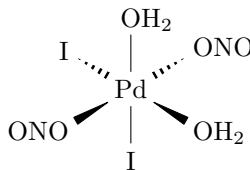
2. *trans*-aqua, *cis*-everything else.



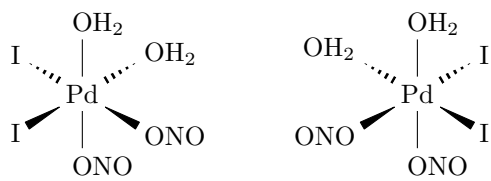
3. *trans*-iodo, *cis*-everything else.



4. *trans*-nitrito, *cis*-everything else.

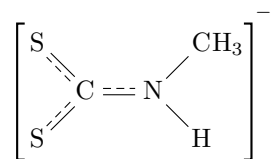


5. All like ligands *cis*.



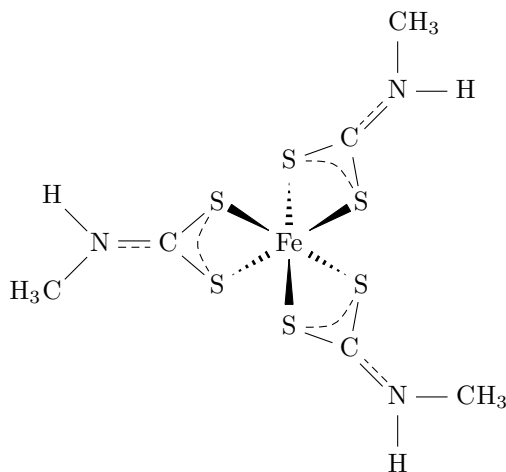
□

c.  $\text{Fe}(\text{dtc})_3$ , all isomers, where dtc is

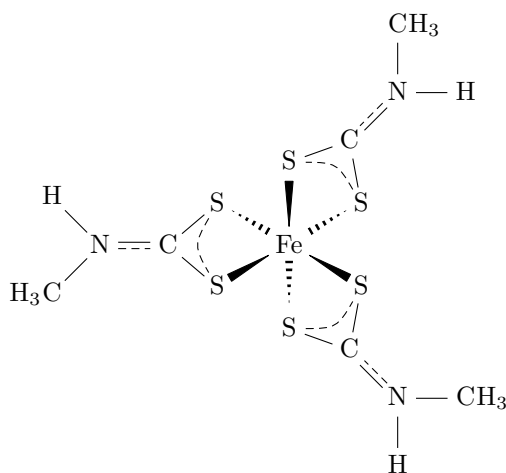


*Answer.* There are two stereoisomers of  $\text{Fe}(\text{dtc})_3$  if we consider the bidentate ligand to be a simple bridge: one  $\Lambda$  form and one  $\Delta$  form.

1.  $\Lambda$  form.



2.  $\Delta$  form.

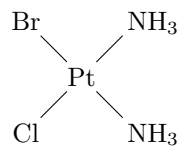


□

**9.9** Show structures for the following:

**a.** *cis*-Diamminebromochloroplatinum(II).

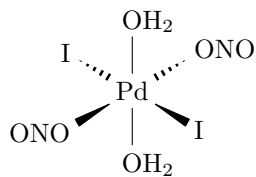
*Answer.*



□

**b.** Diaquadiiododinitritopalladium(IV), all ligands *trans*.

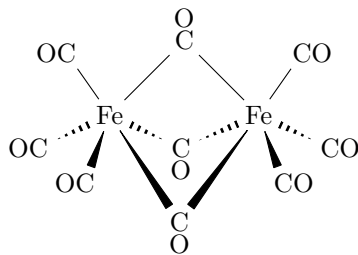
*Answer.*



□

**c.** Tri- $\mu$ -carbonylbis(tricarbonyliron(0)).

*Answer.*



□

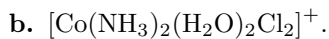
**9.12** Sketch all isomers of the following. Indicate clearly each pair of enantiomers.



*Answer.* There is one facial and one meridional isomer of  $[\text{Pt}(\text{NH}_3)_3\text{Cl}_3]^+$ . There are no enantiomers.

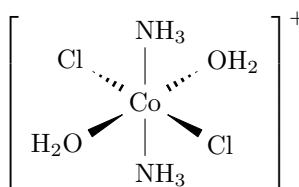


□

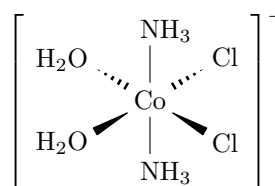


*Answer.*

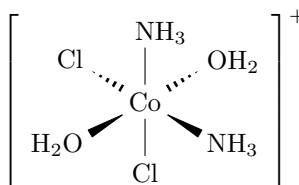
1. All like ligands *trans*.



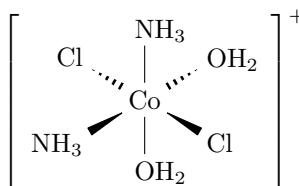
2. *trans*-ammine, *cis*-everything else.



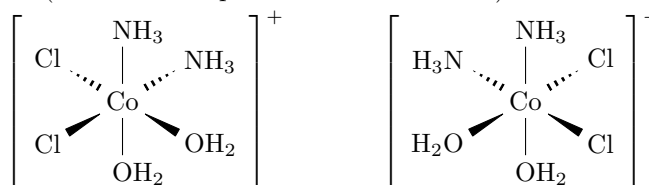
3. *trans*-aqua, *cis*-everything else.



4. *trans*-chloro, *cis*-everything else.



5. All like ligands *cis* (these two compounds are enantiomers).

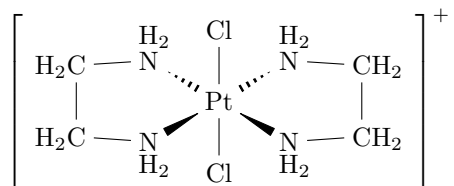


□

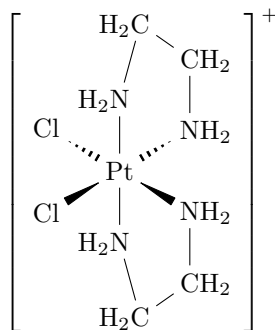
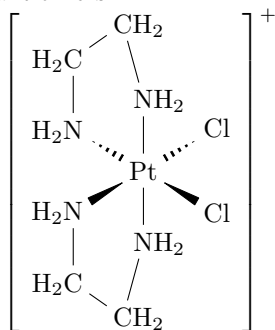
e.  $[\text{Pt}(\text{en})_2\text{Cl}_2]^{2+}$ .

*Answer.* There is one achiral isomer and one pair of chiral enantiomers.

1. Achiral isomer.



2. Chiral enantiomers.



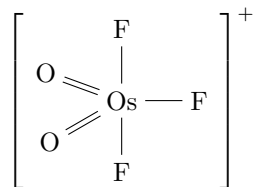
□





- 9.23** When *cis*-OsO<sub>2</sub>F<sub>4</sub> is dissolved in SbF<sub>5</sub>, the cation OsO<sub>2</sub>F<sub>3</sub><sup>+</sup> is formed. The <sup>19</sup>F NMR spectrum of this cation shows two resonances, a doublet and a triplet having relative intensities of 2 : 1. What is the most likely structure of this ion? What is its point group?

*Answer.* Since there are two resonances, there are two types of bonds. Additionally, the 2 : 1 ratio tells us that the doublet bonds are twice as strong as the triplet bonds. Thus, the most likely structure is



Its point group is  $C_{2v}$ .

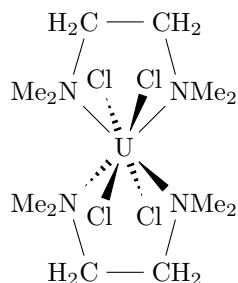
□

- II) The enthalpies of reaction of trimethylboron with  $\text{NH}_3$ ,  $\text{CH}_3\text{NH}_2$ ,  $(\text{CH}_3)_2\text{NH}$ , and  $(\text{CH}_3)_3\text{N}$  are  $-58$ ,  $-74$ ,  $-81$ , and  $-74$  kJ/mol, respectively. Why is trimethylamine out of sequence?

*Answer.* Trimethylboron is not a very small lewis acid like  $\text{BH}_3$  or  $\text{H}^+$ . Thus, its own steric bulk clashes with the considerable bulk of  $(\text{CH}_3)_3\text{N}$ , hindering the reaction in a non-negligible fashion. This hindrance lowers the magnitude of the reaction enthalpy more than the additional electron-donating methyl group can raise it.  $\square$

- III) The complex  $\text{UCl}_4(\text{tmeda})_2$  (where tmeda is an abbreviation for  $\text{Me}_2\text{N}-\text{CH}_2\text{CH}_2-\text{NMe}_2$ , which binds by chelation to metals through both N-atoms) has been shown by X-ray crystallography to have almost perfect  $D_{2d}$  symmetry. Carefully draw a stereochemically accurate picture of this 8-coordinate uranium complex based on what you know about geometrical preferences in 8-coordinate complexes.

*Answer.*



□

## 6 Coordination Complexes: Electron Configurations

- 3/1: I) Derive the proper MO diagram for tetrahedral molecules  $ML_4$  using the  $n$   $d$ -orbitals, the  $(n + l)$   $s$ -orbital, and the  $(n + l)$   $p$ -orbitals on the central atom M for

(a)  $\sigma$ -only ligands;

*Answer.* Point group:  $T_d$ .

Basis functions: All four L  $\sigma$ -orbitals,  $M(ns)$ ,  $M((n + l)p)$ , and  $M((n + l)d)$ .

Apply operations, generate reducible representations, and reduce to irreducible representations:

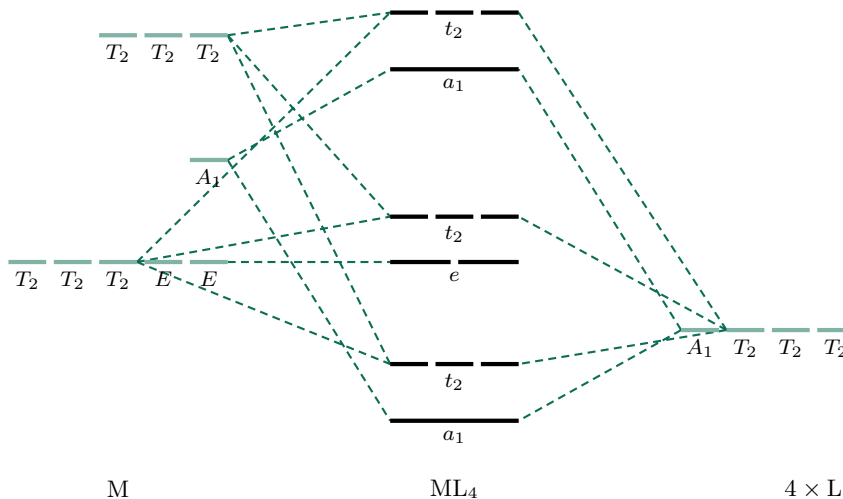
$$\Gamma_{L\sigma} = (4, 1, 0, 0, 2) = A_1 + T_2$$

$$\Gamma_{M_{ns}} = A_1$$

$$\Gamma_{M_{(n+l)p}} = T_2$$

$$\Gamma_{M_{(n+l)d}} = E + T_2$$

Combine orbitals by their symmetry.



□

(b)  $\pi$ -donor ligands (having low-lying, filled  $\pi$  orbitals);

*Answer.* Point group:  $T_d$ .

Basis functions: All four L  $\sigma$ -orbitals, all eight L  $\pi_{x,z}$  orbitals,  $M(ns)$ ,  $M((n + l)p)$ , and  $M((n + l)d)$ .

Apply operations, generate reducible representations, and reduce to irreducible representations:

$$\Gamma_{L\sigma} = (4, 1, 0, 0, 2) = A_1 + T_2$$

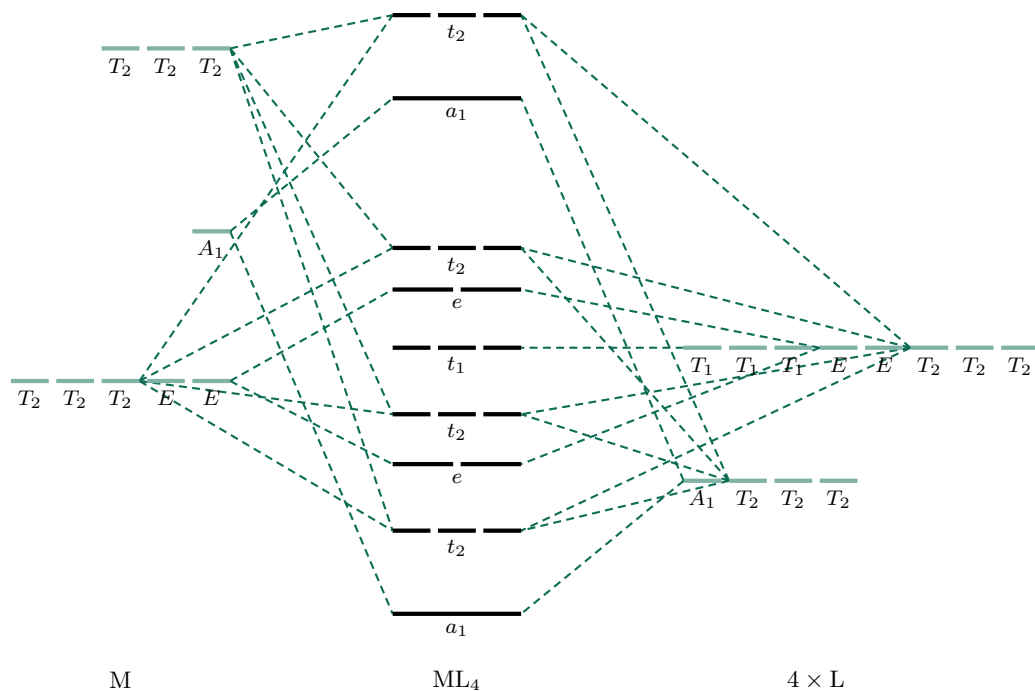
$$\Gamma_{L\pi_{x,z}} = (8, -1, 0, 0, 0) = E + T_1 + T_2$$

$$\Gamma_{M_{ns}} = A_1$$

$$\Gamma_{M_{(n+l)p}} = T_2$$

$$\Gamma_{M_{(n+l)d}} = E + T_2$$

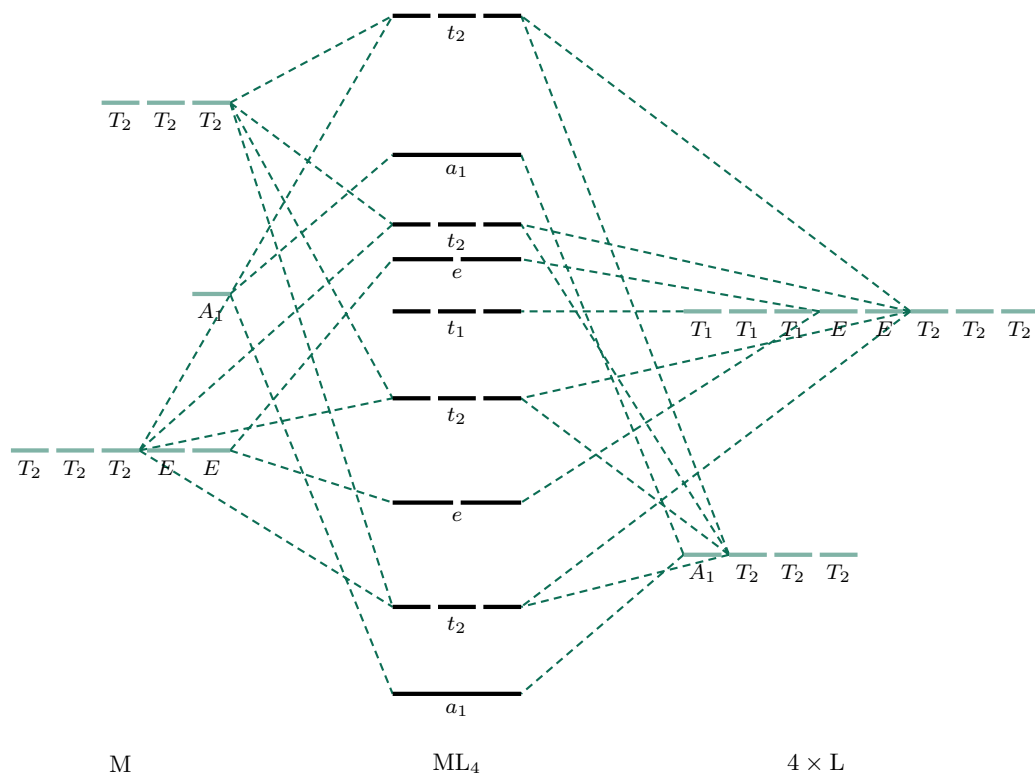
Combine orbitals by their symmetry.



□

(c)  $\pi$ -acceptor ligands (having high-lying, empty  $\pi$  orbitals).

*Answer.* Same intro analysis as part (b).



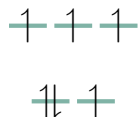
□

II) Do the following problems from Chapter 10: 1, 6, 7, 8, 19, 22.

**10.1** Predict the number of unpaired electrons for each of the following:

a. A tetrahedral  $d^6$  ion.

*Answer.* Tetrahedral means high spin and  $t_2$  orbitals above  $e$ . Thus, there will be 4 unpaired electrons.



□

b.  $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$ .

*Answer.* Octahedral cobalt with a 2+ oxidation state means high spin. Thus, there will be 3 unpaired electrons.



□

c.  $[\text{Cr}(\text{H}_2\text{O})_6]^{3+}$ .

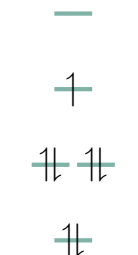
*Answer.* High spin/low spin doesn't matter here; there will be 3 unpaired electrons, regardless.



□

d. A square-planar  $d^7$  ion.

*Answer.* Square planar compounds have an orbital arrangement determined by the angular overlap model, and are low spin. Thus, there will be 1 unpaired electron.



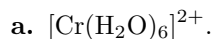
□

e. A coordination compound with a magnetic moment of 5.1 Bohr magnetons.

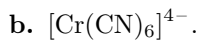
*Answer.* Using the spin-only magnetic moment formula  $\mu_S = \sqrt{n(n+2)}$ , we can solve for  $n$  with the quadratic formula, take the positive answer, and round.

$$\begin{aligned} 5.1 &= \sqrt{n(n+2)} \\ 0 &= n^2 + 2n - 5.1^2 \\ n &\approx 4 \end{aligned}$$

□

**10.6** Predict the magnetic moments (spin-only) of the following species.

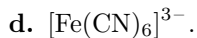
*Answer.* We know that  $[\text{Cr}(\text{H}_2\text{O})_6]^{2+}$  has four unpaired electrons. Thus,  $\mu = \sqrt{4(4+2)} = \sqrt{24} \approx 4.9$ .  $\square$



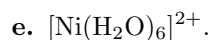
*Answer.* We know that  $[\text{Cr}(\text{CN})_6]^{4-}$  has two unpaired electrons. Thus,  $\mu = \sqrt{2(2+2)} = \sqrt{8} \approx 2.82$ .  $\square$



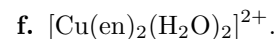
*Answer.* We know that  $[\text{FeCl}_4]^-$  has five unpaired electrons. Thus,  $\mu = \sqrt{5(5+2)} = \sqrt{15} \approx 5.92$ .  $\square$



*Answer.* We know that  $[\text{Fe}(\text{CN})_6]^{3-}$  has one unpaired electron. Thus,  $\mu = \sqrt{1(1+2)} = \sqrt{3} \approx 1.73$ .  $\square$



*Answer.* We know that  $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$  has two unpaired electrons. Thus,  $\mu = \sqrt{2(2+2)} = \sqrt{8} \approx 2.82$ .  $\square$



*Answer.* We know that  $[\text{Cu}(\text{en})_2(\text{H}_2\text{O})_2]^{2+}$  has one unpaired electron. Thus,  $\mu = \sqrt{1(1+2)} = \sqrt{3} \approx 1.73$ .  $\square$



- 10.7** A compound with the empirical formula  $\text{Fe}(\text{H}_2\text{O})_4(\text{CN})_2$  has a magnetic moment corresponding to  $2\frac{2}{3}$  unpaired electrons per iron. How is this possible? (Hint: Two octahedral  $\text{Fe}(\text{II})$  species are involved, each containing a single type of ligand.)

*Answer.*  $\text{Fe}(\text{H}_2\text{O})_6$  is a low spin complex with 4 unpaired electrons.  $\text{Fe}(\text{CN})_6$  is a high spin complex with 0 unpaired electrons. If we take a weighted average of the spins, we find that the magnetic moment of  $\text{Fe}(\text{H}_2\text{O})_4(\text{CN})_2$  is  $\mu = 4 \cdot \frac{4}{6} + 0 \cdot \frac{2}{6} = 2\frac{2}{3}$ .  $\square$

- 10.8** What are the possible magnetic moments of Co(II) in tetrahedral, octahedral, and square-planar complexes?

*Answer.* For tetrahedral and octahedral complexes, we have  $\mu = \sqrt{3(3+2)} = \sqrt{15} \approx 3.87$ . For square planar, we have  $\mu = \sqrt{1(1+2)} = \sqrt{3} \approx 1.73$ .  $\square$

- 10.19** Explain the order of the magnitudes of the following  $\Delta_o$  values for Cr(III) complexes in terms of the  $\sigma$  and  $\pi$  donor and acceptor properties of the ligands.

| Ligand                         | F <sup>-</sup> | Cl <sup>-</sup> | H <sub>2</sub> O | NH <sub>3</sub> | en     | CN <sup>-</sup> |
|--------------------------------|----------------|-----------------|------------------|-----------------|--------|-----------------|
| $\Delta_o$ (cm <sup>-1</sup> ) | 15 200         | 13 200          | 17 400           | 21 600          | 21 900 | 33 500          |

*Answer.* Cyanide has by far the greatest magnitude  $\Delta_o$  because it is the only  $\pi$ -accepting ligand. Ethylenediamine and ammonia form the next group down because they are pure  $\pi$ -donating ligands. Lastly, we have water, chloride, and fluoride because they have  $\pi$ -donating character.  $\square$

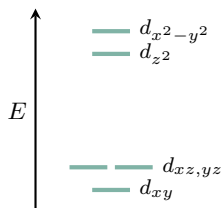
- 10.22** Solid  $\text{CrF}_3$  contains a  $\text{Cr(III)}$  ion surrounded by six  $\text{F}^-$  ions in an octahedral geometry, all at distances of 190 pm. However,  $\text{MnF}_3$  is in a distorted geometry, with  $\text{Mn}-\text{F}$  distances of 179, 191, and 209 pm (two of each). Explain.

*Answer.* The manganese ion has unequally occupied  $e_g$  orbitals whereas the chromium one does not, so the former is subject to Jahn-Teller distortion while the latter is not.  $\square$

III) Use the Angular Overlap Model to derive the  $d$ -orbital splitting diagrams for  $M(\text{CO})_5$  complexes having

(a) Square pyramidal geometry;

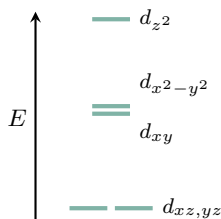
*Answer.* From the charts in Module 36, we have  $e_\sigma = (2, 3, 0, 0, 0)$  and  $e_\pi = (0, 0, 4, 3, 3)$ . Thus we know that the  $d$ -orbital splitting diagram is



□

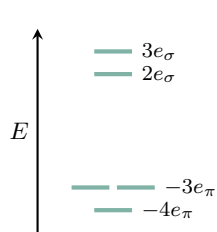
(b) Trigonal bipyramidal geometry.

*Answer.* From the charts in Module 36, we have  $e_\sigma = (\frac{11}{4}, \frac{11}{8}, \frac{9}{8}, 0, 0)$  and  $e_\pi = (0, \frac{3}{2}, \frac{3}{2}, \frac{7}{2}, \frac{7}{2})$ . Thus we know that the  $d$ -orbital splitting diagram is

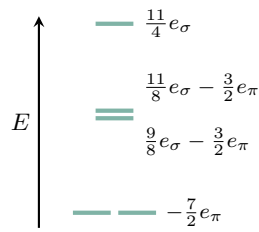


□

(c) Label the orbitals with the appropriate energies in units of  $e_\sigma$  and  $e_\pi$ . Note that degenerate orbitals are those with the same energies — you don't have to use group theory to get this information!



(a) Square pyramidal.



(b) Trigonal bipyramidal.

(d) Determine the relative energies of  $\text{Cr}(\text{CO})_5$  (based on orbital population) in these two geometries, and use this to predict the structure of  $\text{Cr}(\text{CO})_5$ .

*Answer.* Assuming a high spin configuration since chromium is neutral and noting that chromium has a  $d^4$  configuration, we have

$$E_{\text{square pyramidal}} = 1(-4e_\pi) + 2(-3e_\pi) + 2e_\sigma = 2e_\sigma - 10e_\pi$$

$$E_{\text{trigonal bipyramidal}} = 2\left(-\frac{7}{2}e_\pi\right) + 1\left(\frac{9}{8}e_\sigma - \frac{3}{2}e_\pi\right) + 1\left(\frac{11}{8}e_\sigma - \frac{3}{2}e_\pi\right) = \frac{5}{2}e_\sigma - 10e_\pi$$

Since  $E_{\text{trigonal bipyramidal}} > E_{\text{square pyramidal}}$ , we know that  $\text{Cr}(\text{CO})_5$  is square pyramidal.

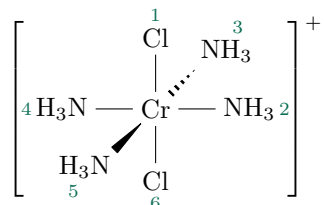
□

## 7 Coordination Complexes: Bonding Models

3/5: I) Do the following problems from Chapter 10: 13, 14, 27, 30, 33.

**10.13** Use the angular overlap method to calculate the energies of both ligand and metal orbitals for  $\text{trans}[\text{Cr}(\text{NH}_3)_4\text{Cl}_2]^+$ , taking into account that ammonia is a stronger  $\sigma$ -donor ligand than chloride, but chloride is a stronger  $\pi$  donor. Use the 1 and 6 positions for the chloride ions.

Answer.



We have from the question that  $\text{NH}_3$  is a pure  $\sigma$  donor whereas  $\text{Cl}^-$  is a  $\sigma$  and  $\pi$  donor. Thus, we can add up

| $\sigma(\text{NH}_3)$ | $z^2$         | $x^2 - y^2$   | $xy$ | $xz$ | $yz$ |
|-----------------------|---------------|---------------|------|------|------|
| 2                     | $\frac{1}{4}$ | $\frac{3}{4}$ | 0    | 0    | 0    |
| 3                     | $\frac{1}{4}$ | $\frac{3}{4}$ | 0    | 0    | 0    |
| 4                     | $\frac{1}{4}$ | $\frac{3}{4}$ | 0    | 0    | 0    |
| 5                     | $\frac{1}{4}$ | $\frac{3}{4}$ | 0    | 0    | 0    |
| Sum:                  | 1             | 3             | 0    | 0    | 0    |

| $\sigma(\text{Cl})$ | $z^2$ | $x^2 - y^2$ | $xy$ | $xz$ | $yz$ |
|---------------------|-------|-------------|------|------|------|
| 1                   | 1     | 0           | 0    | 0    | 0    |
| 6                   | 1     | 0           | 0    | 0    | 0    |
| Sum:                | 2     | 0           | 0    | 0    | 0    |

| $\pi(\text{Cl})$ | $z^2$ | $x^2 - y^2$ | $xy$ | $xz$ | $yz$ |
|------------------|-------|-------------|------|------|------|
| 1                | 0     | 0           | 0    | 1    | 1    |
| 6                | 0     | 0           | 0    | 1    | 1    |
| Sum:             | 0     | 0           | 0    | 2    | 2    |

Adding up the sums for each  $d$  orbital reveals that the metal orbitals are destabilized by<sup>[2]</sup>:

$$\begin{aligned}
 E(d_{z^2}) &= 1e_{\sigma}(\text{NH}_3) + 2e_{\sigma}(\text{Cl}) \\
 E(d_{x^2-y^2}) &= 3e_{\sigma}(\text{NH}_3) \\
 E(d_{xy}) &= 0 \\
 E(d_{xz}) &= 2e_{\pi}(\text{Cl}) \\
 E(d_{yz}) &= 2e_{\pi}(\text{Cl})
 \end{aligned}$$

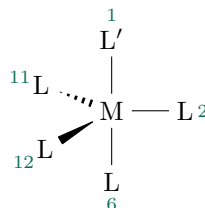
If we now add up the values in the sum rows and divide by the corresponding number of orbitals, we'll get the stabilization energy contribution of each table to its orbitals. For example, the sum in the  $\sigma(\text{NH}_3)$  table is 4, and there are 4  $\text{NH}_3(\sigma)$  orbitals; thus, each  $\text{NH}_3$  ligand orbital is stabilized, at least in part, by  $\frac{4}{4} = 1e_{\sigma}(\text{NH}_3)$ . If we continue in this fashion for the other two, we'll eventually end up with four ligand orbitals stabilized by  $1e_{\sigma}(\text{NH}_3)$  and two ligand orbitals stabilized by  $1e_{\sigma}(\text{Cl}) + 2e_{\pi}(\text{Cl})$ .  $\square$

<sup>2</sup>Note that by  $e_{\sigma}(\text{NH}_3)$ , for example, we mean the angular overlap parameter for a  $\sigma$  bond formed in an  $\text{MA}_2\text{B}_4$ -type complex containing amine groups as the B ligands.

**10.14** Consider a transition metal complex of formula  $ML_4L'$ . Using the angular overlap model and assuming trigonal-bipyramidal geometry, determine the energies of the  $d$  orbitals...

a. considering sigma interactions only (assume L and  $L'$  are similar in donor ability);

*Answer.*



Although  $L'$  is pictured as axial above, it does not matter where it is because we're considering it to be equivalent to L in  $\sigma$ -donor ability. Having established that, we can add up

| $\sigma$ | $z^2$          | $x^2 - y^2$    | $xy$           | $xz$ | $yz$ |
|----------|----------------|----------------|----------------|------|------|
| 1        | 1              | 0              | 0              | 0    | 0    |
| 2        | $\frac{1}{4}$  | $\frac{3}{4}$  | 0              | 0    | 0    |
| 6        | 1              | 0              | 0              | 0    | 0    |
| 11       | $\frac{1}{4}$  | $\frac{3}{16}$ | $\frac{9}{16}$ | 0    | 0    |
| 12       | $\frac{1}{4}$  | $\frac{3}{16}$ | $\frac{9}{16}$ | 0    | 0    |
| Sum:     | $\frac{11}{4}$ | $\frac{9}{8}$  | $\frac{9}{8}$  | 0    | 0    |

With the sums for each  $d$  orbital, we can determine that the metal orbitals are destabilized by<sup>[3]</sup>:

$$E(d_{z^2}) = \frac{11}{4}e_\sigma$$

$$E(d_{x^2-y^2}) = \frac{9}{8}e_\sigma$$

$$E(d_{xy}) = \frac{9}{8}e_\sigma$$

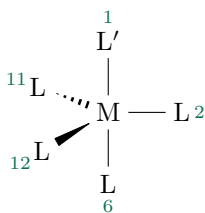
$$E(d_{xz}) = 0$$

$$E(d_{yz}) = 0$$

□

b. and considering  $L'$  as a  $\pi$  acceptor as well. Consider  $L'$  in both  
(1) an axial position;

*Answer.*



The  $\sigma$  contributions to the metal orbitals will be identical to the above. However, we now have to account for the  $\pi$ -acceptor contribution of the axial  $L'$  ligand. Since the rows in the table for positions 1 and 6 are equal, arbitrarily consider row 1:

| $\pi(L')$ | $z^2$ | $x^2 - y^2$ | $xy$ | $xz$ | $yz$ |
|-----------|-------|-------------|------|------|------|
| 1         | 0     | 0           | 0    | 1    | 1    |
| Sum:      | 0     | 0           | 0    | 1    | 1    |

<sup>3</sup>We don't distinguish between  $e_\sigma(L)$  and  $e_\sigma(L')$  here because we are told that both ligands have comparable  $\sigma$ -donating abilities.

With the sums for each  $d$  orbital, we can determine that three of the metal orbitals are destabilized by:

$$E(d_{z^2}) = \frac{11}{4}e_\sigma$$

$$E(d_{x^2-y^2}) = \frac{9}{8}e_\sigma$$

$$E(d_{xy}) = \frac{9}{8}e_\sigma$$

and the other two are stabilized by

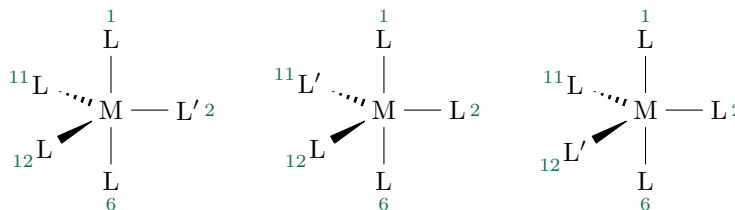
$$E(d_{xz}) = e_\pi(L')$$

$$E(d_{yz}) = e_\pi(L')$$

□

(2) and an equatorial position.

Answer.



The  $\sigma$  contributions to the metal orbitals will be identical to the above. However, we now have to account for the  $\pi$ -acceptor contribution of the equatorial  $L'$  ligand. Since the rows in the table for positions 1, 11, and 12 vary, we will average the three:

| $\pi(L')$ | $z^2$ | $x^2 - y^2$   | $xy$          | $xz$          | $yz$          |
|-----------|-------|---------------|---------------|---------------|---------------|
| 2         | 0     | 0             | 1             | 1             | 0             |
| 11        | 0     | $\frac{3}{4}$ | $\frac{1}{4}$ | $\frac{1}{4}$ | $\frac{3}{4}$ |
| 12        | 0     | $\frac{3}{4}$ | $\frac{1}{4}$ | $\frac{1}{4}$ | $\frac{3}{4}$ |
| Avg:      | 0     | $\frac{1}{2}$ | $\frac{1}{2}$ | $\frac{1}{2}$ | $\frac{1}{2}$ |

With the average contribution for each  $d$  orbital, we can determine that the metal orbitals are destabilized by:

$$E(d_{z^2}) = \frac{11}{4}e_\sigma$$

$$E(d_{x^2-y^2}) = \frac{9}{8}e_\sigma - \frac{1}{2}e_\pi(L')$$

$$E(d_{xy}) = \frac{9}{8}e_\sigma - \frac{1}{2}e_\pi(L')$$

$$E(d_{xz}) = -\frac{1}{2}e_\pi(L')$$

$$E(d_{yz}) = -\frac{1}{2}e_\pi(L')$$

Note that we subtract the  $e_\pi(L')$  contributions because for  $\pi$ -acceptor ligands, these contributions are stabilizing. Thus, they reduce the destabilization energy (they don't increase it). □



- c. Based on the preceding answers, would you expect  $\pi$ -acceptor ligands to preferentially occupy axial or equatorial positions in five-coordinate complexes? What other factors should be considered in addition to angular overlap?

*Answer.* We can construct orbital diagrams for both the axial  $L'$  case and the equatorial  $L'$  case.

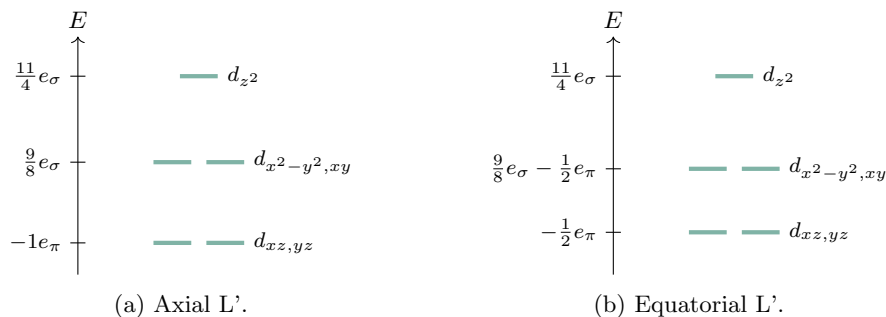


Figure 7: Axial and equatorial  $ML_4L'$   $d$ -orbital diagrams.

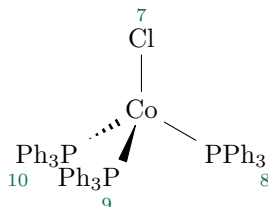
The important factor to be considered in addition to angular overlap is electron configuration — for  $d^1$  to  $d^7$  configurations, the axial complex is preferred; however, for  $d^8$  to  $d^{10}$  configurations, neither is preferred. We can determine this by calculating the energies of each configuration from Figure 7. In the following equations, the axial energy is on the left and the equatorial on the right.

$$\begin{aligned}
 -1e_\pi &< -\frac{1}{2}e_\pi & (d^1) \\
 -2e_\pi &< -1e_\pi & (d^2) \\
 -3e_\pi &< -\frac{3}{2}e_\pi & (d^3) \\
 -4e_\pi &< -2e_\pi & (d^4) \\
 \frac{9}{8}e_\sigma - 4e_\pi &< \frac{9}{8}e_\sigma - \frac{5}{2}e_\pi & (d^5) \\
 \frac{9}{4}e_\sigma - 4e_\pi &< \frac{9}{4}e_\sigma - 3e_\pi & (d^6) \\
 \frac{27}{8}e_\sigma - 4e_\pi &< \frac{27}{8}e_\sigma - \frac{7}{2}e_\pi & (d^7) \\
 \frac{9}{2}e_\sigma - 4e_\pi &= \frac{9}{2}e_\sigma - 4e_\pi & (d^8) \\
 \frac{29}{4}e_\sigma - 4e_\pi &= \frac{29}{4}e_\sigma - 4e_\pi & (d^9) \\
 10e_\sigma - 4e_\pi &= 10e_\sigma - 4e_\pi & (d^{10})
 \end{aligned}$$

□

**10.27** Cobalt(I) complexes are relatively rare compared to Co(II) and Co(0), but the complexes  $\text{CoX}(\text{PPh}_3)_3$  where  $\text{X} = \text{Cl}, \text{Br}, \text{I}$  are known with approximate tetrahedral coordination geometry about the high spin  $d^8$  metal center. The angular overlap model was used to analyze the electronic structure of  $\text{CoCl}(\text{PPh}_3)_3$ , where three independent molecules (with very similar yet statistically different bond lengths and angles) were observed in the unit cell (Krzystek et al., 2012). Using the angular overlap parameters for molecule 1 in Table 3 of this reference, generate an energy-level diagram for  $\text{CoCl}(\text{PPh}_3)_3$ . Does the electronic structure predicted by this method surprise you? Explain. On the basis of Table 3, the chloride ligands are better  $\pi$  donors, and the triphenylphosphine ligands better  $\sigma$  donors in  $\text{NiCl}_2(\text{PPh}_3)_2$  relative to  $\text{CoCl}(\text{PPh}_3)_3$ . What is probably the most important factor that causes these differences?

*Answer.*



From Krzystek et al. (2012), we know that

$$\varepsilon_{\sigma}(\text{Cl}) = 5430 \text{ cm}^{-1} \quad \varepsilon_{\pi}(\text{Cl}) = +1380 \text{ cm}^{-1} \quad \varepsilon_{\sigma}(\text{P}) = 3340 \text{ cm}^{-1} \quad \varepsilon_{\pi}(\text{P}) = -310 \text{ cm}^{-1}$$

We now add up

| $\sigma(\text{Cl})$ | $z^2$ | $x^2 - y^2$ | $xy$          | $xz$          | $yz$          |
|---------------------|-------|-------------|---------------|---------------|---------------|
| 7                   | 0     | 0           | $\frac{1}{3}$ | $\frac{1}{3}$ | $\frac{1}{3}$ |
| Sum:                | 0     | 0           | $\frac{1}{3}$ | $\frac{1}{3}$ | $\frac{1}{3}$ |

| $\sigma(\text{P})$ | $z^2$ | $x^2 - y^2$ | $xy$          | $xz$          | $yz$          |
|--------------------|-------|-------------|---------------|---------------|---------------|
| 8                  | 0     | 0           | $\frac{1}{3}$ | $\frac{1}{3}$ | $\frac{1}{3}$ |
| 9                  | 0     | 0           | $\frac{1}{3}$ | $\frac{1}{3}$ | $\frac{1}{3}$ |
| 10                 | 0     | 0           | $\frac{1}{3}$ | $\frac{1}{3}$ | $\frac{1}{3}$ |
| Sum:               | 0     | 0           | 1             | 1             | 1             |

| $\pi(\text{Cl})$ | $z^2$         | $x^2 - y^2$   | $xy$          | $xz$          | $yz$          |
|------------------|---------------|---------------|---------------|---------------|---------------|
| 7                | $\frac{2}{3}$ | $\frac{2}{3}$ | $\frac{2}{9}$ | $\frac{2}{9}$ | $\frac{2}{9}$ |
| Sum:             | $\frac{2}{3}$ | $\frac{2}{3}$ | $\frac{2}{9}$ | $\frac{2}{9}$ | $\frac{2}{9}$ |

| $\pi(\text{P})$ | $z^2$         | $x^2 - y^2$   | $xy$          | $xz$          | $yz$          |
|-----------------|---------------|---------------|---------------|---------------|---------------|
| 8               | $\frac{2}{3}$ | $\frac{2}{3}$ | $\frac{2}{9}$ | $\frac{2}{9}$ | $\frac{2}{9}$ |
| 9               | $\frac{2}{3}$ | $\frac{2}{3}$ | $\frac{2}{9}$ | $\frac{2}{9}$ | $\frac{2}{9}$ |
| 10              | $\frac{2}{3}$ | $\frac{2}{3}$ | $\frac{2}{9}$ | $\frac{2}{9}$ | $\frac{2}{9}$ |
| Sum:            | 2             | 2             | $\frac{2}{3}$ | $\frac{2}{3}$ | $\frac{2}{3}$ |

Adding up the sums for each  $d$  orbital (and accounting for the fact that  $\text{Cl}^-$  is a  $\pi$  donor whereas  $\text{PPh}_3$  is a  $\pi$  acceptor) reveals that the metal orbitals are destabilized by

$$E(d_{z^2}) = \frac{2}{3}\varepsilon_{\pi}(\text{Cl}) + 2\varepsilon_{\pi}(\text{P}) = 300 \text{ cm}^{-1}$$

$$E(d_{x^2-y^2}) = \frac{2}{3}\varepsilon_{\pi}(\text{Cl}) + 2\varepsilon_{\pi}(\text{P}) = 300 \text{ cm}^{-1}$$

$$E(d_{xy}) = \frac{1}{3}\varepsilon_{\sigma}(\text{Cl}) + 1\varepsilon_{\sigma}(\text{P}) + \frac{2}{9}\varepsilon_{\pi}(\text{Cl}) + \frac{2}{3}e_{\pi}(\text{P}) = 5250 \text{ cm}^{-1}$$

$$E(d_{xz}) = \frac{1}{3}\varepsilon_{\sigma}(\text{Cl}) + 1\varepsilon_{\sigma}(\text{P}) + \frac{2}{9}\varepsilon_{\pi}(\text{Cl}) + \frac{2}{3}e_{\pi}(\text{P}) = 5250 \text{ cm}^{-1}$$

$$E(d_{yz}) = \frac{1}{3}\varepsilon_{\sigma}(\text{Cl}) + 1\varepsilon_{\sigma}(\text{P}) + \frac{2}{9}\varepsilon_{\pi}(\text{Cl}) + \frac{2}{3}e_{\pi}(\text{P}) = 5250 \text{ cm}^{-1}$$

As to the ligand pairs, we must add up the values in the sum rows and divide by the corresponding number of orbitals (note that  $L_n$  corresponds to the ligand numbered  $n$ ), as follows. For the chloride ligand, there is a simple stabilization of the combined  $\sigma, \pi$ -ligand orbital corresponding to a destabilization of every metal  $d$  orbital. For the triphenylphosphine ligands, there is a stabilization of the  $\sigma$ -ligand orbitals corresponding to a destabilization of the metal  $d_{xy, xz, yz}$  orbitals, and there is a destabilization of the  $\pi$ -ligand orbitals corresponding to a stabilization of the metal  $d_{xy, xz, yz}$  orbitals.

$$E(L_7) = 1\varepsilon_{\sigma}(\text{Cl}) + 2\varepsilon_{\pi}(\text{Cl}) = 8190 \text{ cm}^{-1}$$

$$E(L_{8\sigma}) = 1\varepsilon_{\sigma}(\text{P}) = 3340 \text{ cm}^{-1}$$

$$E(L_{9\sigma}) = 1\varepsilon_{\sigma}(\text{P}) = 3340 \text{ cm}^{-1}$$

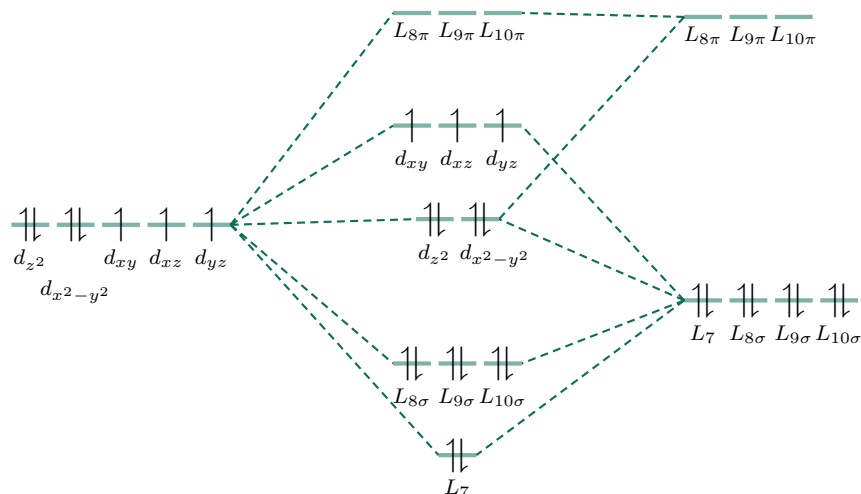
$$E(L_{10\sigma}) = 1\varepsilon_{\sigma}(\text{P}) = 3340 \text{ cm}^{-1}$$

$$E(L_{8\pi}) = \frac{2}{3}\varepsilon_{\pi}(\text{P}) = 207 \text{ cm}^{-1}$$

$$E(L_{9\pi}) = \frac{2}{3}\varepsilon_{\pi}(\text{P}) = 207 \text{ cm}^{-1}$$

$$E(L_{10\pi}) = \frac{2}{3}\varepsilon_{\pi}(\text{P}) = 207 \text{ cm}^{-1}$$

Putting it all together, we have:

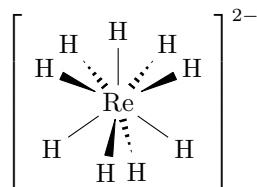


What's surprising: In the angular overlap model, chloride is clearly a stabilizing ligand overall. However, in ligand field theory,  $\pi$ -donor ligands are destabilizing.

Chloride is probably a better  $\pi$  donor in  $\text{NiCl}_2(\text{PPh}_3)_2$  than  $\text{CoCl}(\text{PPh}_3)_3$  because the increased oxidation state of Ni(II) means it has a greater tendency to attract electrons. On the other hand, triphenylphosphine is likely a better  $\sigma$  donor in  $\text{NiCl}_2(\text{PPh}_3)_2$  than  $\text{CoCl}(\text{PPh}_3)_3$  because the reduction in steric bulk allows for better head-on overlap with the orbitals of the metal center.  $\square$

- 10.30** One of the more striking hydride complexes is enneahydridorhenate ( $[\text{ReH}_9]^{2-}$ ), which has tricapped trigonal-pyramidal geometry. What is the point group of this ion? Construct a representation using the hydrogen orbitals as a basis. Reduce this to its component irreducible representations, and determine which orbitals of Re are of suitable symmetry to interact with the hydrogen group orbitals.

*Answer.*



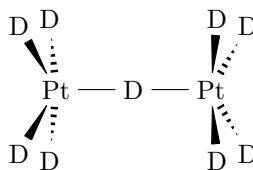
Point group:  $D_{3h}$ .

Construct a reducible representation and reduce it to its component irreducible representations:

$$\Gamma_{\text{H}} = (9, 0, 1, 3, 0, 3) = 2A'_1 + 2E' + A''_2 + E''$$

Rhenium's  $6s$  orbital interacts with the  $A'_1$  hydrogen group orbitals, its  $5p$  orbitals interact with the  $E'$  ( $p_{x,y}$ ) and  $A''_2$  ( $p_z$ ) hydrogen group orbitals, and its  $5d$  orbitals interact with the  $A'_1$  ( $d_{z^2}$ ),  $E'$  ( $d_{x^2-y^2, xy}$ ), and  $E''$  ( $d_{xz, yz}$ ) hydrogen group orbitals.  $\square$

**10.33** The ion  $[\text{Pt}_2\text{D}_9]^{5-}$ , shown below, has eclipsed geometry.



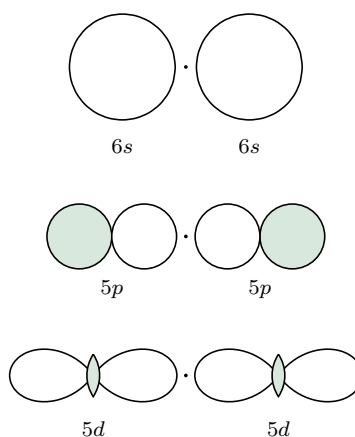
- a. What is point group of this ion?

*Answer.*  $D_{4h}$ . □

- b. Assume that the platinum atoms can potentially use  $s$ ,  $p$ , and  $d$  orbitals to interact with the central deuterium. If the  $z$ -axis is chosen to be collinear with the principal axis of rotation:

1. Sketch the group orbitals of the platinum atoms that potentially could interact with the central D. Be sure to label all orbitals.

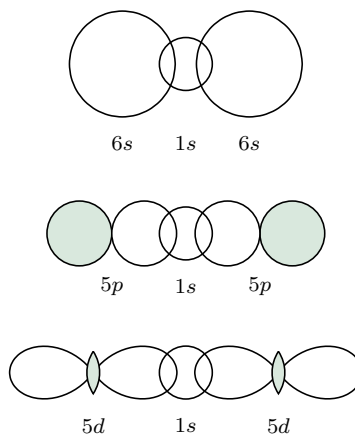
*Answer.*



□

2. Show how the group orbitals and the central atom would interact.

*Answer.*



□

3. Which interaction would you expect to be the strongest, and why?

*Answer.* The  $d_{z^2}$  interaction should be the strongest. In the MO energy diagram for  $\text{MH}_6$  bonding, the hydrogen orbitals are closest in energy to the metal  $d$  orbitals. It stands to reason that it would be similar for square pyramidal bonding and deuterium. □

- II) We derived in class the MO diagram for  $\text{Re}_2\text{Cl}_8^{2-}$ , a complex with a Re–Re quadruple bond. Two chloride ligands on each Re can be replaced with two (neutral)  $\text{PMe}_2\text{Ph}$  ligands to give the cationic complex  $\text{Re}_2(\text{PMe}_2\text{Ph})_4\text{Cl}_4^{2+}$ , which has a similar  $\sigma^2\pi^4\delta^2$  electronic configuration and a Re–Re quadruple bond of 2.215 Å in length. The 1-electron and 2-electron reduced species  $\text{Re}_2(\text{PMe}_2\text{Ph})_4\text{Cl}_4^+$  and  $\text{Re}_2(\text{PMe}_2\text{Ph})_4\text{Cl}_4$  can be prepared, and their structures show Re–Re bond lengths of 2.218 Å and 2.241 Å, respectively. Give the electronic configurations for the 1- and 2-electron reduced species (as given above for the unreduced species), and suggest how this could account for the different Re–Re bond lengths in the three compounds.

*Answer.* From the MO diagram discussed in lecture, we know that adding more electrons to a complex isoelectronic to  $\text{Re}_2\text{Cl}_8^{2-}$  will begin to occupy the  $\delta^*$  molecular orbitals. Thus, the electronic configuration of  $\text{Re}_2(\text{PMe}_2\text{Ph})_4\text{Cl}_4^+$  is  $\sigma^2\pi^4\delta^2\delta^{*1}$ , and the electronic configuration of  $\text{Re}_2(\text{PMe}_2\text{Ph})_4\text{Cl}_4$  is  $\sigma^2\pi^4\delta^2\delta^{*2}$ .

Because the increasingly reduced species have increasing antibonding character, it stands to reason that the bond will weaken and lengthen.  $\square$

## References

Krzystek, J., Ozarowski, A., Zvyagin, S. A., & Telser, J. (2012). High spin Co(I): High-frequency and -field EPR spectroscopy of CoX(PPh<sub>3</sub>)<sub>3</sub> (X = Cl, Br). *Inorganic Chemistry*, 51(9), 4954–4964. <https://doi.org/10.1021/ic202185x>