

## Week 5

# Stereochemistry / Reactions of Alkenes

## 5.1 Stereochemistry Nomenclature and Intro to Alkenes

10/28:

- Last time:
  - Chapter 4 (di-substituted cyclohexanes, bicyclic/polycyclic rings).
  - Chapter 5 (stereochemistry, stereoisomers/chirality center, chirality tests [plane/center symmetry], R/S system [nomenclature; a very important survival skill for this class], physical properties of enantiomers, achiral environment: same; chiral environment: different, rotate plane polarized light [left (-) levorotatory, right (+) dextrorotatory, racemic ( $\pm$ ) mixture],  $[\alpha]_D^{25}$  specific rotation, enantiometric excess, Fischer projections [less important]).
- Intro to Guangbin Dong.
- The textbook is only a guide — follow the lecture. Also, if you see discrepancies either with the book or with Piccirilli, bring them up.
- Capital *E* is energy and lowercase *e* is electrons in this class.
- Reading assignment: Chapter 5.12, 5.14, 4.5 (review), 4.17, 7.1-7.4.
- Nongraded homework: 6.39, 5.40, 5.46, 5.48, 7.1, 7.17, 7.18, 30a-d.
- Multiple stereocenters.
- Worked example: Naming (2S,3R)-2-chloro-3-iodobutane.
  - $n = 2$  stereocenters yields at most  $2^n = 2^2 = 4$  stereoisomers.
  - Draws all enantiomers.
- **Diastereomers:** Have at least 2 stereocenters, same formula, same connectivity, but different orientation in space and are not mirror images of each other.
  - Special case: cis/trans isomerism.
- Compounds with 2 stereocenters don't always have 4 stereoisomers.
- Example: Tartaric acid, for which the (2S,3R) compound is superimposable on the (2R,3S) compound. This stereoisomer is a **meso** isomer and not a chiral molecule (there exists a plane of symmetry).
  - Thus, a molecule can be achiral even if it has a chiral center!

- **Meso** (compound): A compound with chiral stereocenters and an internal plane of symmetry.
  - The following compound (two conformers shown) is also meso because the Conformer 2 has the desired plane of symmetry.

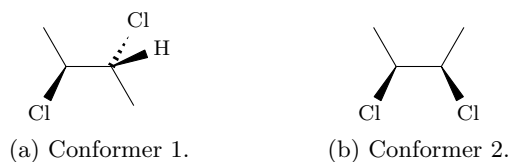


Figure 5.1: Meso compounds.

- **Alkenes:**
  - One of the most important functional groups.
  - Basic industry materials: polyethylene and polypropylene.
  - Biological systems: Fatty acids, vitamins (Vitamin A), drugs/natural products, important building block in other FGs.
- Alkenes have a  $\pi$  bond —  $2s$ ,  $2p_x$ , and  $2p_y$  get hybridized into  $sp^2$  orbitals, and  $2p_z$  forms the  $\pi$  bond.
  - The  $\pi$  bond leads to the enforced coplanar geometry of the alkenes.
  - The bond energy of a C=C bond is significantly greater than that of a C–C bond.
- Alkene cis/trans isomerism.
  - Trans is more stable b/c of steric hindrance.
  - E/Z system.