

Chapter 17

Carboxylic Acids and Derivatives

17.1 Carboxylic Acids and Derivatives 1

- 4/7:
- We now consider compounds that have heteroatoms where the α carbon of the carbonyl used to be.
 - The heteroatoms can be oxygen (esters), nitrogen, etc.
 - Today, we will do oxygen and nitrogen nucleophiles but in this context.
 - Next Tuesday, we will do carbon and hydrogen nucleophiles in this context.
 - Carboxylic acid derivatives.

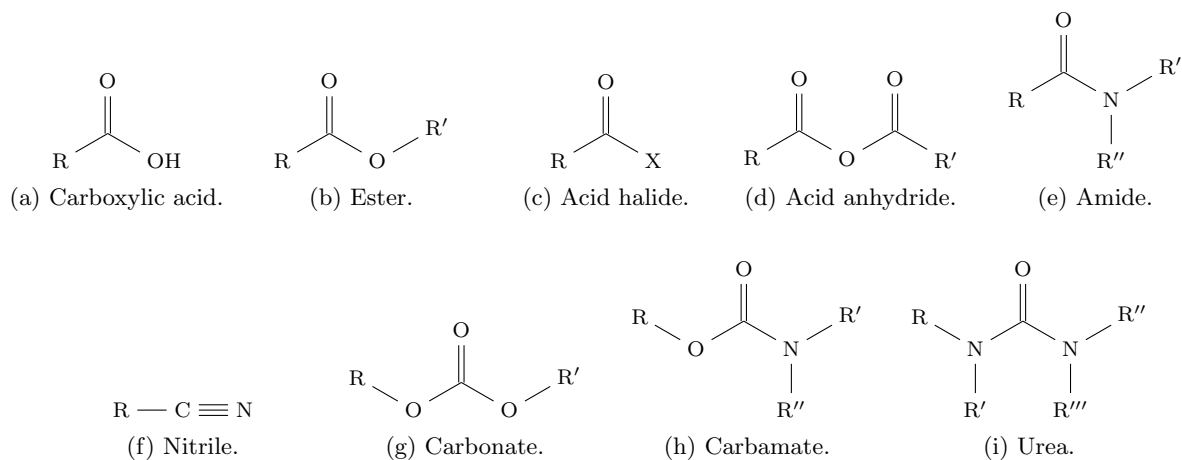


Figure 17.1: Carboxylic acid derivatives.

- Once again, we will not be tested on nomenclature, but it's good to know.
- Acid anhydrides are so named because it is two carboxylic acids, minus a water molecule.
- Nitriles are still a carbon bonded to three heteroatoms; it's just the same heteroatom.
- A key property of carboxylic acids is that they're... acidic.
- Acidity.
 - Gives the $\text{p}K_{\text{a}}$'s of benzoic acid, benzyl alcohol, and phenol to demonstrate that resonance is king in determining acidity.
 - Benzoic acid is more acidic than phenol, which is more acidic than benzyl alcohol.

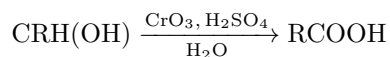
- Inductive effects (changes to the α carbon) play a smaller role.
- EWGs on arene rings when present play an even smaller role.
- These latter two effects allow us to fine-tune acidity.

- Methods of carboxylic acid synthesis.

1. Overoxidation.
2. Carboxylation of Grignards or lithiates.
3. Nitrile hydrolysis.

- Overoxidation.

- General form.



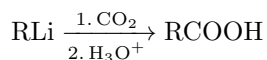
- Note that the reagents constitute Jones reagent.

- Mechanism.

- Virtually identical to that from Labalme (2022).

- Carboxylation of Grignards and lithiates.

- General form.



- Note that we may use either lithiates (RLi) or Grignards (RMgBr), even though only an organo-lithium compound is shown above.

- Mechanism.

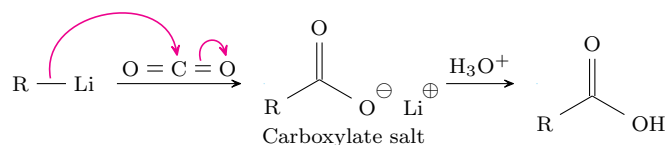


Figure 17.2: Carboxylation of lithiates mechanism.

- Mechanistic interlude: Nucleophilic acyl substitution.

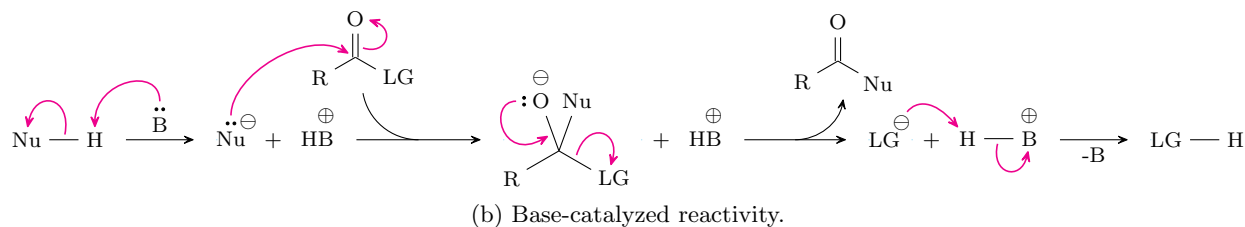
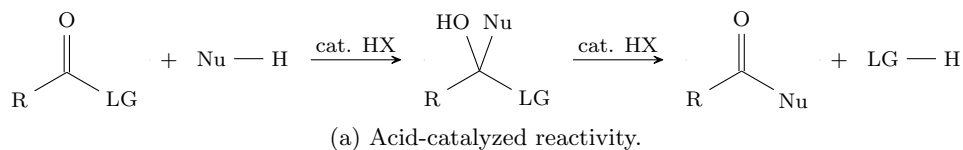


Figure 17.3: The typical reactivity of carboxylic acid derivatives.

- This mode of reactivity is the one that is most typical of carboxylic acid derivatives.
 - It is so-named because the portion of a carboxylic acid derivative that is not the leaving group is called an acyl group, and we are substituting one group on the acyl for another.
- Think of all of the carboxylic acid derivatives (see Figure 17.1) as containing a leaving group on one of their sides.
 - When these compounds react nucleophiles, the nucleophile replaces the leaving group.
- These reactions are either acid- or base-catalyzed.
 - In the acid-catalyzed version (Figure 17.3a), the first step proceeds exactly as in Figure 16.8a, except that $R' = LG$. The second step proceeds exactly as in Figure 16.8b, except that it is the leaving group that is protonated and kicked out instead of the nucleophile we just added in.
 - The basic mechanism is related to Figure 16.9, but rather than being a straight replication, the alkoxide species produced in Figure 16.9a proceeds straight to the reactivity of the alkoxide in Figure 16.9b (see Figure 17.3b).
- **Tetrahedral intermediates:** The nucleophilic acyl substitution intermediates (of both the acidic and basic pathways) that have four groups attached to the central carbon.

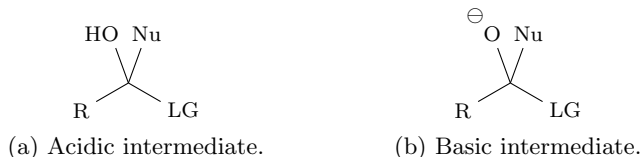


Figure 17.4: The tetrahedral intermediates.

- Historically, the name arose when scientists were arguing about whether or not an sp^3 carbon could be in this reaction. Some scientists supported the theory that these tetrahedral intermediates existed, while others disagreed.
- Nitrile hydrolysis.
- General form.

$$RCN + H_3O^+ \longrightarrow RCOOH + NH_4^+$$
 - Note that here we're using a stoichiometric full equivalent of acid, not just catalytic acid, because we are liberating ammonia which mops up our acid, forming NH_4^+ as a byproduct.
 - The existence of this reaction is the reason we consider nitriles to be carboxylic acid derivatives (i.e., because we can interconvert them with carboxylic acids).
- Mechanism.
 - Note that the fourth intermediate is one deprotonation away from being an amide.
 - However, the reaction conditions do not produce an amide but continue as drawn to a carboxylic acid.
 - This is because in general, the amide oxygen is more basic than the nitrile nitrogen, so if the conditions are such that the nitrile will begin the reaction, the amide will certainly finish it.
 - Note that there are some enzymes that can stop at the amide through various mechanisms that recognize one species as substrate but not another.
 - Every once in a while, people will claim that they've isolated the amide in this mechanism, but these results are hard to reproduce because of the above facts.
 - If we do add up all of the equivalents of water and acid added, we can see that only one equivalent of acid is added, overall (and two equivalents of water).

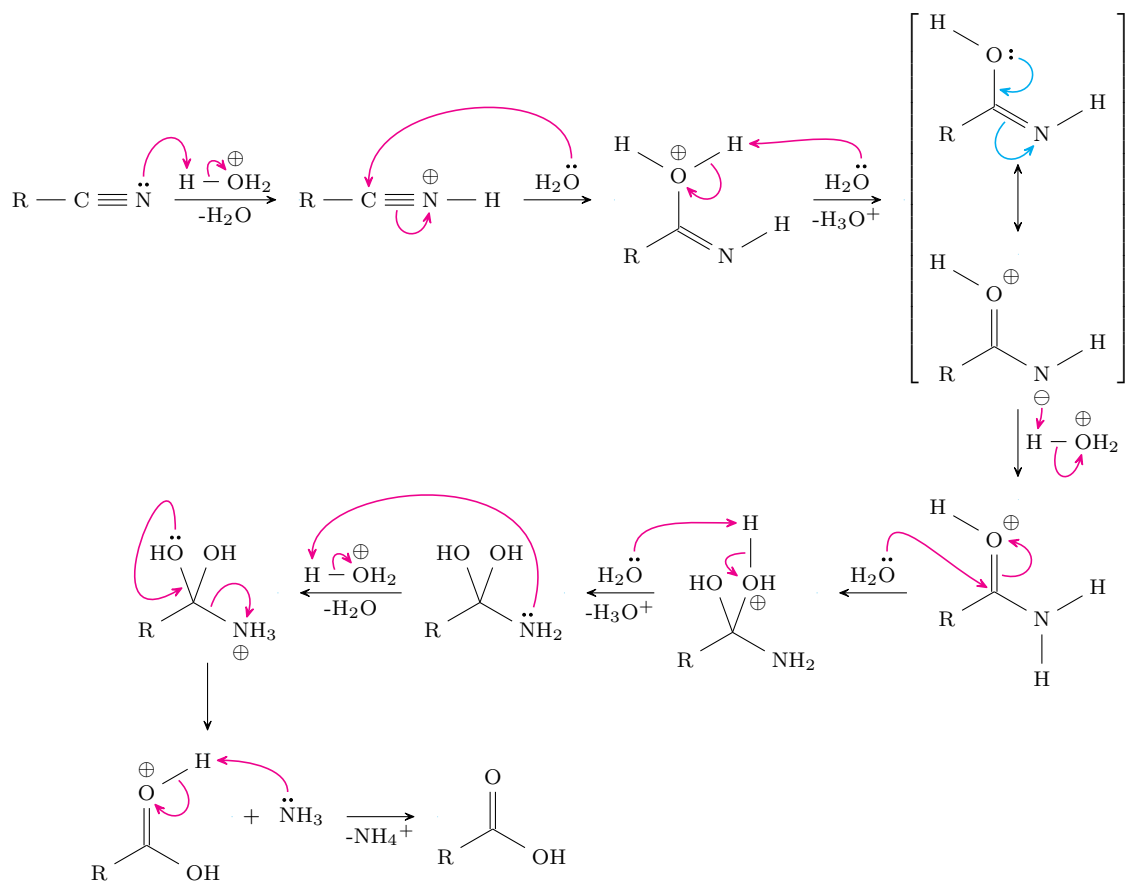
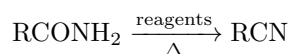


Figure 17.5: Nitrile hydrolysis mechanism.

- Dehydration of amides.
- General form.



- This is the reverse reaction to nitrile hydrolysis.
- Reagents is either SOCl_2 or POCl_3 .
- SOCl_2 and POCl_3 are **dehydrating agents**.

- **Dehydrating agent:** A chemical that drives conversions in which water is lost from a molecule.
 - Notice how the amide overall loses two hydrogens and an oxygen (i.e., a water molecule overall) in Figure 17.6.
- Mechanism.
 - Part of the reason the amide oxygen is such a good nucleophile is because the nitrogen can participate, as in step 1 above.
 - Driving force: Kicking out a gas (SO_2) and chloride.
 - Note that the mechanism implies that we must have an amide with two H's (esp., we cannot have one or two R groups in their place).
 - Although only the mechanism for SOCl_2 is illustrated, the mechanism is virtually identical for POCl_3 .

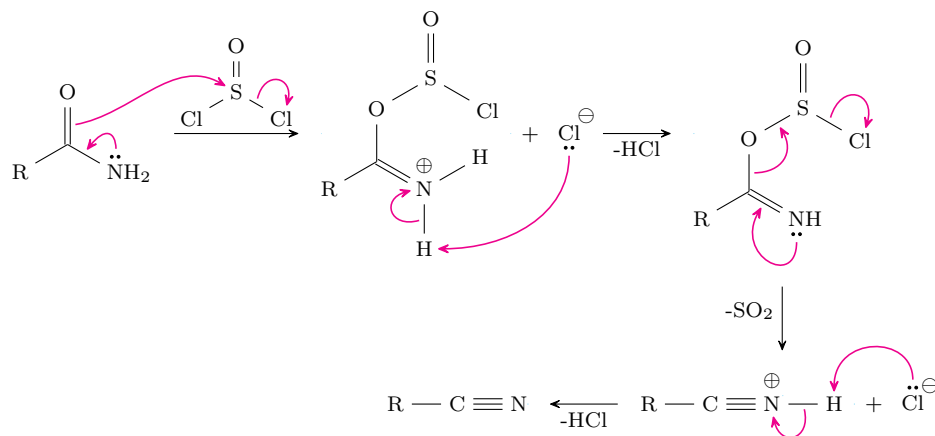


Figure 17.6: Dehydration of amides mechanism.

- Comparing methods 2 and 3 of synthesizing carboxylic acids.

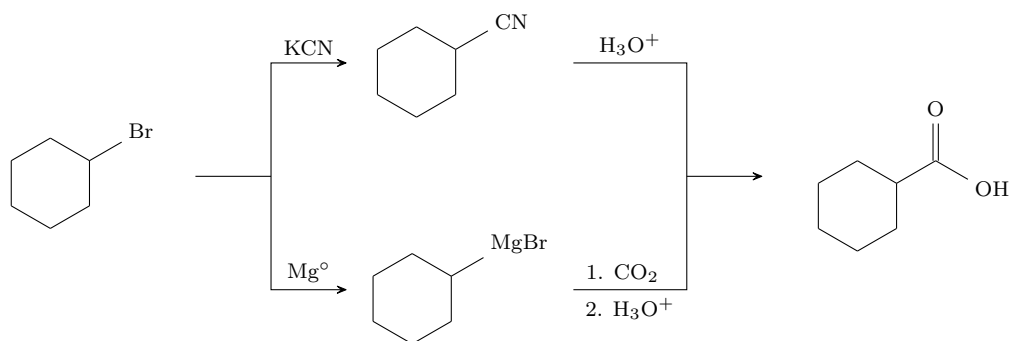


Figure 17.7: Two ways to synthesize a carboxylic acid from an alkyl halide.

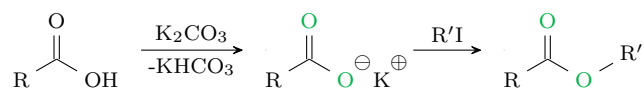
- Both carboxylation and nitrile hydrolysis achieve the same end result from the same starting material, begging the question of why both are necessary.
- The answer lies in the fact that both suit different types of reaction conditions.
- Carboxylation is strongly basic, so we can't use molecules with free H's.
- Nitrile hydrolysis proceeds through $\text{S}_{\text{N}}2$ to start, so we can't use tertiary bromides.

■ This is important on part of PSet 1!

- Methods of ester synthesis.

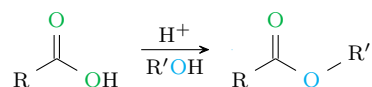
1. Nucleophilic.
2. Fischer esterification.

- Nucleophilic.
- General form.

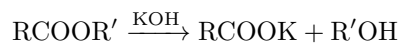


- We deprotonate the carboxylic acid using a relatively weak base.

- K_2CO_3 is often the weak base of choice because it's insoluble in most solvents but will react in a biphasic mixture.
 - Additionally, since $KHCO_3$ is usually insoluble and the carboxylate is typically soluble in the organic solvent in which the reaction is being carried out, it's really easy to separate the two.
- The second step proceeds via an S_N2 mechanism, so methyl or primary alkyl halides are best.
- Note that the two initial oxygens (green) proceed through the whole of the process and end up in the product.
- Fischer esterification.
- General form.



- The acid is a catalyst, and we need an excess of the alcohol, which we typically just use as our solvent.
- Reasons we need an excess of the alcohol.
 - This is essentially a thermoneutral reaction; there's not a great thermodynamic driving force between the carboxylic acid and ester.
 - Thus, the only way to get the reaction to go forward is to overwhelm it with an excess of the alcohol so that Le Châtelier's principle comes into play.
- Removing water can also help drive the reaction.
- H_3O^+ (i.e., excess water) reverses the reaction.
- Note that the mechanism here is a nucleophilic attack, and it is the *methanol* oxygen (blue) that gets incorporated into the final ester (whose initial oxygens are colored green).
- **Saponification:** Subjecting an ester to a single equivalent of KOH (or any other hydroxide base) to form the carboxylate and the alcohol.
 - This is very old chemistry.
 - Sapon- is the Latin prefix for soap.
 - Ancient peoples discovered that combining and heating animal fat, wood ash, and a bit of water creates soap.
 - Combining triglycerides with pot ash yields glycerol soap and long-chain fatty acid carboxylates.
 - Pot ash is where we get the name for potassium, because the ashes from a wood stove are rich in potassium hydroxide.
 - Fatty acid carboxylates serve to solublize grease in water because the lipid end interacts with the grease and the carboxylate end interacts with the water. This is how all soaps work!
- General form.



- The carboxylate is an end-stage product. Resonance delocalizes the negative charge over the carbon atom, significantly decreasing its electrophilicity and hence its capacity to participate in future reactions.
- The presence of basic conditions make it so that this reaction is not reversible. Indeed, if we mix a base with $RCOOH$, we will just deprotonate the acid and return to the carboxylate form.
- Mechanism.

- Hydroxide attacks the ester as a nucleophile, and OR^- leaves to form a carboxylic acid. But OR^- (a strong base) will then deprotonate RCOOH (a strong acid) to form the carboxylate and alcohol.

- Acid chloride synthesis.
- General form.



- Pyridine is not strictly necessary, but it greatly increases the reaction rate.
- Driven in a similar way to the dehydration of amides; we release SO_2 gas, expel a water molecule, and mop up the extra Cl^- with pyridine.

- Mechanism.

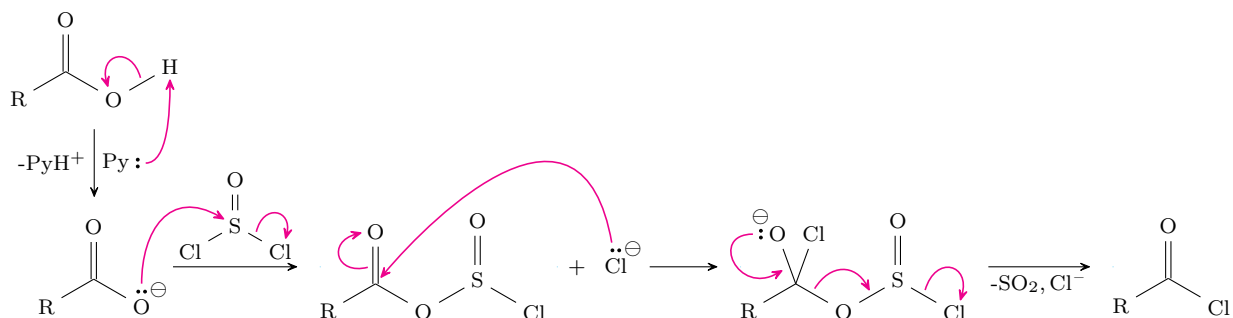
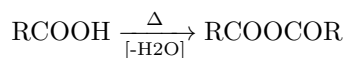


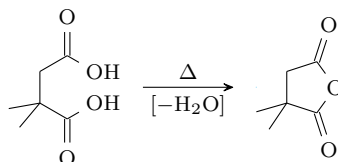
Figure 17.8: Acid chloride synthesis mechanism.

- Since chloride is a fairly weak nucleophile, its addition in step 3 takes a while and is reversible.
 - However, this step is driven in the forward direction by releasing SO_2 gas from the resulting tetrahedral intermediate (Le Châtelier's principle).

- Anhydride synthesis.
- General form (standard).

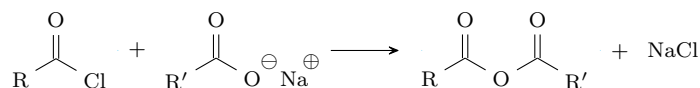


- High heat is required.
- If you use two different carboxylic acids, you will get a statistical mixture (no real selectivity).
- You can selectively create 5-6 membered rings containing anhydrides because this reaction proceeds intramolecularly as well as intermolecularly.
- General form (intramolecular).



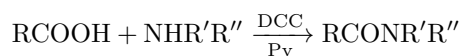
- In particular, if you have a single molecule with two different carboxylic acid groups 2-3 carbons apart, then heating a sample of said molecule while removing water will result in a ring-closing anhydridization.
- If we want to make a ring with another number of carbons, we should go through acid chlorides (see below).

- A way to selectively create anhydrides is via acid chlorides and sodium carboxylates.
- Mixed anhydride synthesis.
- General form.



– This reaction proceeds via nucleophilic substitution.

- Amide synthesis.
- General form.



- Mechanism.

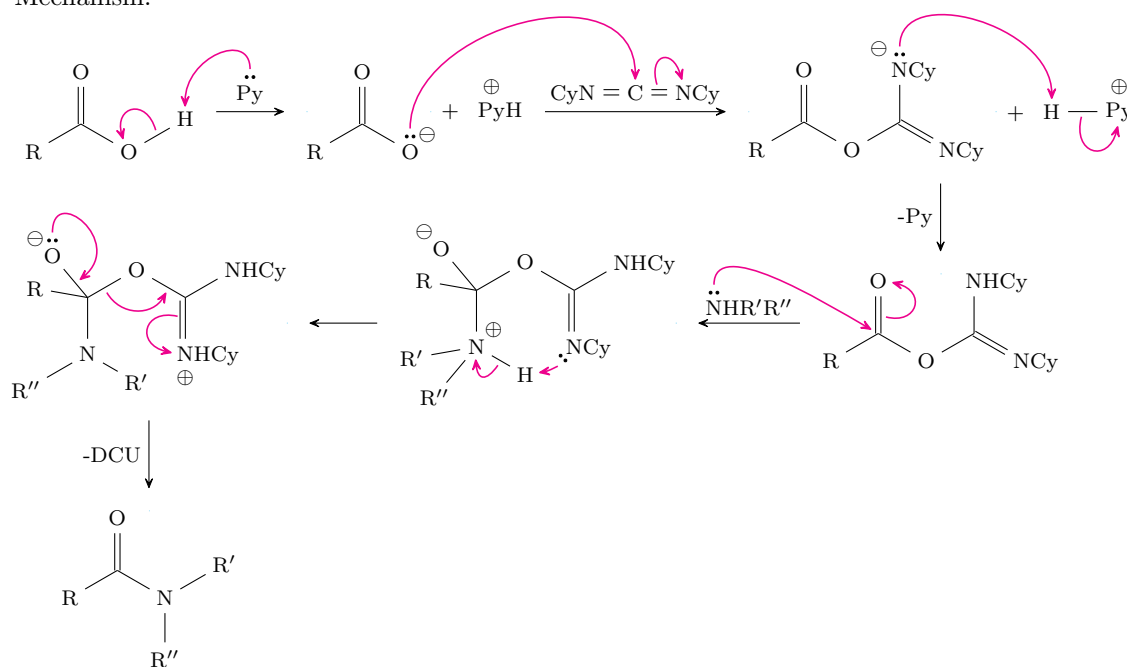


Figure 17.9: Amide synthesis mechanism.

- Note that as in other mechanisms, DCC eventually transforms into a type of leaving group.
- Normally, we use external reagents for proton transfers because doing an internal one would in most cases involve a transition state with a 4-membered ring, which is highly strained.
 - However, in step 5 here, we can do an internal proton transfer because the transition state's conformation is that of a 6-membered ring.
- **DCC:** Dicyclohexylcarbodiimide, a dehydrating reagent key to amide synthesis. *Structure*

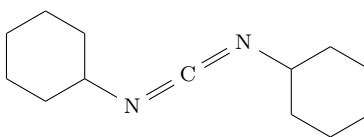


Figure 17.10: Dicyclohexylcarbodiimide (DCC).

- DCC reacts with water as follows.

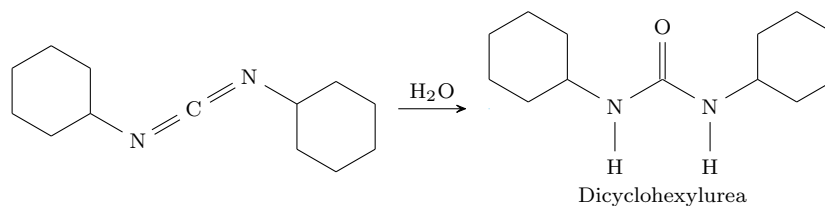


Figure 17.11: DCC and water.

- **DCU**: Dicyclohexylurea, the product of the reaction of DCC and water.
- Reactivity scale.
 - acid chloride > anhydride > ester > amide > carboxylate
 - It should make intuitive sense that acid chlorides are the most reactive carboxylic acid derivatives and carboxylates are the least.
 - Acid chlorides have an electronegative group on the already electrophilic carbon, exacerbating the molecular dipole.
 - Carboxylates delocalize their negative charge over the carbon (as discussed earlier), greatly reducing or eliminating the molecular dipole.
 - A good rule of thumb is that the compound with the best leaving group and worst nucleophile (an acid chloride) is the most reactive, and vice versa in that the compound with the worst leaving group and the best nucleophile (a carboxylate) is the most reactive.
 - What we mean by “reactivity” is that compounds higher on the reactive scale can react with an appropriate nucleophile to become compounds lower on the scale.
 - For instance, we can take an acid chloride to an anhydride, ester, amide, or carboxylate (and we have reactions to do that), but we cannot take all (or any) of these molecules back to an acid chloride without forcing conditions.
 - Some things that qualify as forcing conditions are the use of acidic conditions and dehydrating reagents.
 - In other words, this reactivity scale is for the compounds in basic media with no dehydrating reagents present.
- MCAT comments.
- Trialkyl amines and pyridines.
 - According to our reactivity scale, we should be able to react NEt_3 with RCOCl to yield an amine, for example.
 - However, this leads to a positively charged nitrogen in the amine that cannot be quenched (e.g., by deprotonation). Thus, this is a highly reversible reaction that favors the reactants.
 - Similarly, we should be able to react an anhydride with pyridine.
 - But since pyridine cannot be deprotonated either, the reactants are favored in this reversible reaction once again.
- However, this implies that pyridines can be used to catalyze nucleophilic acyl substitutions.
- **DMAP**: Dimethylaminopyridine, which is one of the best catalysts for nucleophilic acyl substitutions.
Structure
 - Levin gives an example synthesis using DMAP, namely nucleophilic addition to an anhydride.

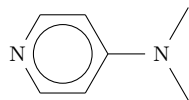


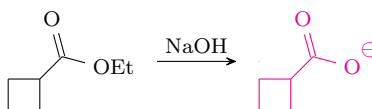
Figure 17.12: Dimethylaminopyridine (DMAP).

- In essence, DMAP adds to the carbonyl, kicks out the leaving group, and then the nucleophile adds to the carbonyl and kicks out DMAP.
- Adding DMAP can accelerate a reaction that would take overnight to taking only a few minutes.
- Acid chlorides, anhydrides, and esters all create the same product (an amide) when reacting with an amine.
- But, you need only one equivalent of the amine for esters while you need two equivalents for the first two.
- This is because of the pK_a 's. The first two byproducts (HCl and RCOOH) protonate amines in solution, whereas ROH does not (as much).

17.2 Discussion Section

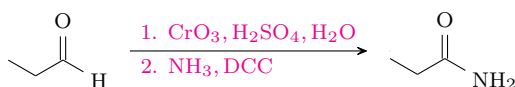
- 4/8:
- We will be working with hot sand baths in the next lab, so just leave them to cool and do not dispose of the contents unless you're sure they're cool.
 - Practice problems.

1.



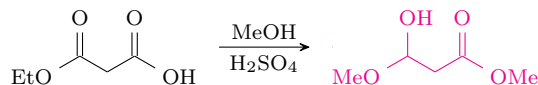
- We form a COO^- ion instead of the carboxylic acid because we are in basic solution.
- The mechanism is a nucleophilic attack on the carbonyl, the oxygen electrons swinging back down and kicking out EtO^- , and then deprotonation of the acid.

2.



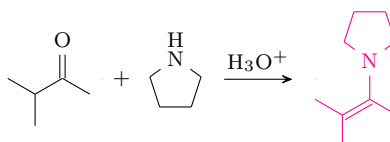
- The intermediate after step 1 is the carboxylic acid, as we have used aqueous Jones reagent.

3.



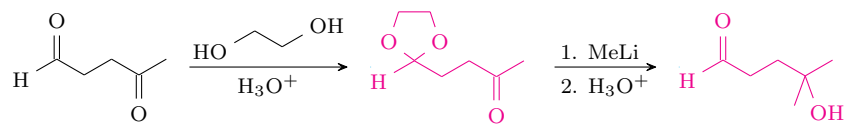
- The reaction of the ester (left) is called **transesterification**; the reaction of the carboxylic acid (right) is called ether formation.
- It's important to know that you can get ester formation in both of these cases.
- This is a common problematic side reaction in synthetic chemistry.
- Mechanism: Methanol attacks each carbonyl, the other group leaves, and then deprotonation.

4.

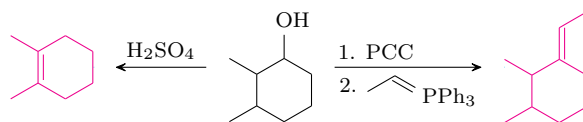


– We choose this enamine as the major product by Zaitsev's rule.

5.



6.



7.

