

Week 8

Electronic Structure of Atoms and Molecules

8.1 Hybridization and Time-Dependent Quantum Mechanics

11/15:

- Consider BeH₂.
 - Then we can form the *sp* molecular orbital

$$\psi_{sp} = \frac{1}{\sqrt{2}}(2s \pm 2p_z)$$

- Draws an orbital diagram of BeH₂.
- BH₃.
 - Lists *sp*² hybrid orbital wave functions.
- CH₄.
 - Lists *sp*³ hybrid orbital wave functions.
- The probability of finding a ψ_1 electron in the $2p_z$ orbital.
 - The absolute value of the coefficient of $2p_z$ in the wave function, squared.
 - Alternatively,

$$\text{Prob} = \left| \int 2p_z^* \psi_1 \, d\tau \right|^2$$

- Time-dependent Schrödinger equation:

$$\hat{H}\psi = i\hbar \frac{d\psi}{dt}$$

- Let $\psi_n(r, t) = \phi_n(r)e^{-iE_n t/\hbar}$. Plugging this into the time-dependent Schrödinger equation yields

$$\begin{aligned}\hat{H}\psi &= i\hbar \frac{d}{dt} \left(\phi_n(r) e^{-iE_n t/\hbar} \right) \\ &= i\hbar \cdot -\frac{iE_n}{\hbar} \cdot \phi_n(r) e^{-iE_n t/\hbar} \\ &= E_n \psi_n\end{aligned}$$

- Thus, to this point, we've been solving the spatial part of the time-dependent Schrödinger equation.

- The probability density of the time-dependent solution is equal to the probability density of the time-independent solution.

$$\begin{aligned}
 |\psi(r, t)|^2 &= \left| \phi(r) e^{-iEt/\hbar} \right|^2 \\
 &= \psi^*(r) \psi(r) e^{iEt/\hbar} e^{-iEt/\hbar} \\
 &= \psi^*(r) \psi(r) \\
 &= |\psi(r)|^2
 \end{aligned}$$

- The time-dependent Schrödinger equation allows us to make predictions for how long it takes an electron to relax back to a lower energy state after excitation, for instance.
- Using the time-dependent Schrödinger equation to describe the interaction of the system with electromagnetic radiation.

1. Solution: $\vec{E} = \vec{E}_0 \cos(\omega t)$ ^[1].

2. Dipole moment of the molecule (μ).

- We let $\hat{V} = -\vec{\mu} \cdot \vec{E}$. This is how radiation interacts with matter.
- This leads to the overall Hamiltonian for the molecule interacting with the electric field as $\hat{H} = \hat{H}_0 + \hat{V}$, where \hat{H}_0 is the Hamiltonian of the molecule and \hat{V} is the dipole-electric field perturbation defined above.
- If the electron is transitioning from the ground to the first-excited state, we let

$$\psi(t) = a_1(t)\phi_1 e^{-iE_1 t/\hbar} + a_2(t)\phi_2 e^{-iE_2 t/\hbar}$$

where the first part refers to the ground state, the second part refers to the excited state, and a_1, a_2 are time-dependent expansion coefficients that we are solving for.

- Substitution into the time-dependent Schrödinger equation yields

$$\begin{aligned}
 i\hbar \left(\psi_1 \frac{da_1}{dt} + \psi_2 \frac{da_2}{dt} \right) &= a_1(t) \hat{V} \psi_1 + a_2(t) \hat{V} \psi_2 \\
 i\hbar e^{-iE_2 t/\hbar} \frac{da_2}{dt} &= a_1(t) \int \phi_2^* \hat{V} \psi_1 d\vec{r} + a_2(t) \int \phi_2^* \hat{V} \psi_2 d\vec{r}
 \end{aligned}$$

8.2 Chapter 10: Bonding in Polyatomic Molecules

From McQuarrie and Simon (1997).

- Consider BeH₂.
 - The two Be–H bonds are equivalent with an angle of 180° between them
 - MOs are formed from orbitals that are similar in energy, so in addition to the H(1s) and Be(2s), throw in the energetically close Be(2p).
 - We also add in a 2p orbital because we need the resultant molecular orbitals to point in opposite directions so as to explain the linear structure of the molecule, and directionality is not something the spherically symmetric 2s orbital can provide.

- Thus, choose

$$\psi_{\text{Be-H}} = c_1 \psi_{\text{Be}(2s)} + c_2 \psi_{\text{Be}(2p)} + c_3 \psi_{\text{H}(1s)}$$

- The first two terms in the above linear combination can be thought of as representing a new **hybrid** “orbital” on beryllium.

¹This is the definition of a classical electric radiation field. In Advanced Quantum Mechanics, we can turn this into a photon field for quantum electrodynamics.

- The normalized sp hybrid orbitals are given by

$$\psi_{sp} = \frac{1}{\sqrt{2}}(2s \pm 2p_z)$$

where the z -axis is defined to be the internuclear H–Be–H axis.

- Note that sp hybrid orbitals concentrate electron density in one specific direction since the sign of the $2p$ wave function is different in the two direction ($\pm z$) but the sign of the $2s$ wave function is everywhere positive.

- **Hybrid orbital:** A linear combination of atomic orbitals on the same atom.
- The normalized sp^2 hybrid orbitals are

$$\begin{aligned}\psi_1 &= \frac{1}{\sqrt{3}}2s + \sqrt{\frac{2}{3}}2p_z & \psi_2 &= \frac{1}{\sqrt{3}}2s - \frac{1}{\sqrt{6}}2p_z + \frac{1}{\sqrt{2}}2p_x \\ \psi_3 &= \frac{1}{\sqrt{3}}2s - \frac{1}{\sqrt{6}}2p_z - \frac{1}{\sqrt{2}}2p_x\end{aligned}$$

- The normalized sp^3 hybrid orbitals are

$$\begin{aligned}\psi_1 &= \frac{1}{2}(2s + 2p_x + 2p_y + 2p_z) & \psi_2 &= \frac{1}{2}(2s - 2p_x - 2p_y + 2p_z) \\ \psi_3 &= \frac{1}{2}(2s + 2p_x - 2p_y - 2p_z) & \psi_4 &= \frac{1}{2}(2s - 2p_x + 2p_y - 2p_z)\end{aligned}$$

- Bonding vs. lone pair electrons: H_2O .

- We could form bond orbitals out of the singly occupied $p_{y,z}$ orbitals of O, but this would predict a bond angle of 90° . Thus, we need hybrid orbitals.
- The bond angle in water is between that predicted using sp^2 hybrid orbitals and $2p$ orbitals.
- Since oxygen has two bonding and four lone pair electrons, we let

$$\psi = c_1 2s + c_2 2p_y + c_3 2p_z$$

- Determining the coefficients such that two orthogonal orbitals directed at an angle of 104.5° are generated yields

$$\psi_1 = 0.45 \cdot 2s + 0.71 \cdot 2p_y + 0.55 \cdot 2p_z \quad \psi_2 = 0.45 \cdot 2s - 0.71 \cdot 2p_y + 0.55 \cdot 2p_z$$

- Physical interpretation of the normalization constants: c_i^2 is the fractional character of the i^{th} orbital^[2].
- ψ_1, ψ_2 accommodate the bonding electrons. Our unused hybrid orbital and unused $2p_x$ orbital accommodate the lone pair electrons as a linear combination (since symmetry necessitates that the two lone pair orbitals be equivalent).
- Why BeH_2 is linear but H_2O is bent.
 - Using a more general LCAO-MO with all relevant orbitals and solving yields a set of molecular orbitals.
 - Upon bending, the degeneracy of the π orbitals is lifted as the hydrogen orbitals move to intersect more with one specific side of the molecule.

²Think s -character, p -character.

- Such changes are summarized by a **Walsh correlation diagram** (see Labalme (2021, p. 65)).
- We want the minimal energy for however many electrons we have, and with the additional electrons of H_2O , it is more energetically favorable to bend then to remain linear.
- **Walsh correlation diagram:** A plot of the energy of a molecular orbital as a function of a systematic change in molecular geometry.
- **σ -bond framework:** The collection of all σ bonds in a molecule.