

CHEM 26200 (Thermodynamics) Notes

Steven Labalme

January 24, 2022

Contents

I	Statistical Mechanics	1
1	The Boltzmann Factor and Partition Functions	2
1.1	Overview of Major Results	2
1.2	Boltzmann Factor Examples / Partition Function	5
1.3	Calculating Average Energies	8
1.4	Chapter 17: The Boltzmann Factor and Partition Functions	10
2	Partition Functions and Ideal Gases	19
2.1	System Partition Functions	19
2.2	Molecular Partition Functions	20
2.3	Chapter 18: Partition Functions and Ideal Gases	22
3	Kinetic Theory of Gases	23
3.1	Maxwell-Boltzmann Distribution	23
	References	25

List of Figures

1.1 Sodium lamp energy levels.	5
--	---

Part I

Statistical Mechanics

Week 1

The Boltzmann Factor and Partition Functions

1.1 Overview of Major Results

- 1/10:
- In this course, we will review thermochemistry from intro chem, but go deeper with statistical mechanics.
 - TA: Haozhi.
 - Did his undergrad at Oxford.
 - Has already taught this class in the PME.

- **Boltzmann constant:** The following constant. *Denoted by k_B . Given by*

$$k_B = 1.381 \times 10^{-23} \text{ J/K}$$

- Equal to the quotient of the ideal gas constant and Avogadro's constant.
- **Ideal gas law:** The following relationship between the pressure P , volume V , number of moles n , and temperature T of an ideal gas, and the ideal gas constant R .

$$PV = nRT$$

- Multiplying by the quotient of Avogadro's constant with itself yields

$$PV = nN_A \frac{R}{N_A} T$$

$$PV = Nk_B T$$

where N is the number of molecules in the system.

- The unit for PV is Joules.
 - Thus, the above form states that PV is equal to the number of particles times a tiny unit of energy.
- Relating PV to the kinetic energy of gas molecules/atoms^[1].
 - Pressure originates microscopically from the collisions of particles with the walls of their container.
 - As such, we first seek to derive an expression for the number of collisions per second per area.

¹This derivation differs from that on Labalme (2021a, pp. 3–4) and Labalme (2021b, pp. 18–19), in that its approach is from a flux perspective.

- Consider the number $N(v_x)$ particles with speed v_x in the x -direction.
- The quotient $N(v_x)/V$ is the density in the container of particles with speed v_x .
- Thus, the flux “through”/to/at the wall is this density, times the area of the wall, times the x -velocity of the particles.
- Assume an elastic collision of each particle with the wall. Thus, when each particle of mass m collides with the wall, it transfers $2mv_x$ of momentum.
- Therefore, since $F = dp/dt$, the overall force exerted on the wall by the gas particles moving with speed v_x is $2mv_x \cdot N(v_x)/V \cdot v_x \cdot \text{Area}$ times per second.
- But, of course, we must sum over all possible v_x , so the total force

$$F = \int_{v_x > 0} 2mv_x \cdot \frac{N(v_x)}{V} \cdot v_x \cdot \text{Area} dv_x$$

- It follows that

$$\begin{aligned} P &= \frac{F}{\text{Area}} \\ &= \int_{v_x > 0} 2mv_x^2 \cdot \frac{N(v_x)}{V} dv_x \end{aligned}$$

The factor of 1/2 in the following line comes from the fact that we are only integrating over half of the possible v_x^2 s (i.e., the positive ones).

$$\begin{aligned} &= 2m \cdot \frac{N}{V} \cdot \frac{1}{2} \langle v_x^2 \rangle \\ &= \frac{N}{V} m \langle v_x^2 \rangle \\ PV &= Nm \cdot \langle v_x^2 \rangle \end{aligned}$$

Assuming that the gas is not moving in any one direction means that $\langle v_x^2 \rangle = \langle v_y^2 \rangle = \langle v_z^2 \rangle = \frac{1}{3} \langle v^2 \rangle$. Therefore,

$$\begin{aligned} &= Nm \cdot \frac{1}{3} \langle v^2 \rangle \\ &= \frac{2}{3} N \cdot \frac{1}{2} m \langle v^2 \rangle \\ &= \frac{2}{3} N \cdot \langle E_{KE} \rangle \\ \langle E_{KE} \rangle &= \frac{3}{2} \frac{PV}{N} \\ \langle E_{KE} \rangle &= \frac{3}{2} k_B T \end{aligned}$$

- Note that this applies to all sorts of regimes — we used no properties of the particles (e.g., atom vs. molecule) to derive this relationship.
- Getting the distribution of the gas energies or speed is the next logical step.
- First, though, we consider alternate occurrences of $k_B T$.
 - The activation energy of Arrhenius (1889): “To collide is to react” is inaccurate; it must collide with sufficient energy. The molecule must be “activated.”

$$k = Ae^{-E_a/RT} = Ae^{-E_a/k_B T}$$

- The first E_a is the molar energy of activation; the second is the molecular energy of activation.
 - Yields the probability distribution of a molecule reacting.
- Nernst equation:

$$E_{\text{cell}} = E_{\text{cell}}^0 - \frac{RT}{nF} \ln Q$$

- $\ln Q$ is the ratio inside vs. outside the membrane.
 - $F = N_A e$ where e is the charge of an electron.
 - Thus,
- $$\Delta E = \frac{RT}{nF} = \frac{k_B T}{ne}$$
- If the potential across the membrane is approximately $k_B T$, then $\ln Q \approx 1$, so $Q \approx e$.
 - Thus, at body temperature ($T = 310 \text{ K}$), $k_B T/e = 26 \text{ mV}$.
- The speed of sound: Certainly sound cannot travel faster than the molecules. Therefore, we can derive the following approximation for the speed of sound.

$$\begin{aligned} \frac{1}{2} m \langle v^2 \rangle &= \frac{3}{2} k_B T \\ \sqrt{\langle v^2 \rangle} &= \sqrt{\frac{3 k_B T}{m}} \\ v_{\text{rms}} &= \sqrt{\frac{3 k_B T}{m}} \end{aligned}$$

- This estimate is within 20 – 30 % — take m to be the average mass of air.
- de Broglie wavelength: A molecule has a kinetic energy approximately equal to $k_B T$. Additionally, the quantum mechanical kinetic energy of a molecule aligns with this, as $\hbar^2 k^2 / 2m \approx k_B T$. Furthermore, the particle-wave duality relates the momentum to wavelength by $p = \hbar k = h / \lambda$. Therefore,

$$\lambda \approx \sqrt{\frac{h^2}{2m k_B T}}$$

- Thus, a gas at STP has a very small de Broglie wavelength and behaves classically.
- Only at very low temperatures with very light gasses do quantum considerations come into play.
- A H_2 molecule at 300 K has de Broglie wavelength $\lambda = 1.78 \text{ \AA}$.
- Note that the quantum mechanical kinetic energy of a free particle is derived as follows.

$$\begin{aligned} \hat{H}\psi &= E\psi \\ -\frac{\hbar^2}{2m} \frac{\partial^2}{\partial x^2} (e^{ikx}) &= E e^{ikx} \\ \frac{\hbar^2 k^2}{2m} e^{ikx} &= E e^{ikx} \\ E &= \frac{\hbar^2 k^2}{2m} \end{aligned}$$

- **Boltzmann factor:** Gives the relative probability p_2/p_1 of two states E_1, E_2 , provided their respective energies E_1, E_2 . Given by

$$\frac{p_2}{p_1} = e^{-(E_2 - E_1)/k_B T}$$

- Consider states E_1, E_2, E_3, \dots , denoted by their energies.

- Consistency check: Given

$$\frac{p_2}{p_1} = e^{\frac{-(E_2-E_1)}{k_B T}} \qquad \frac{p_3}{p_2} = e^{\frac{-(E_3-E_2)}{k_B T}}$$

we do indeed have

$$\frac{p_3}{p_1} = \frac{p_3}{p_2} \cdot \frac{p_2}{p_1} = e^{\frac{-(E_3-E_2)}{k_B T} + \frac{-(E_2-E_1)}{k_B T}} = e^{\frac{-(E_3-E_1)}{k_B T}}$$

- We'll take this as God-given for now. Boltzmann derived it with a very good knowledge of the thermodynamics of freshman chemistry.
- We're starting with the above exciting result, and then going back and building up to it over the next three weeks.
- We write the Boltzmann factor for degenerate states as follows.
 - Consider four states at E_2 and one state at E_1 .
 - The Boltzmann factor still tells us that $p_2/p_1 = e^{-(E_2-E_1)/k_B T}$, but we have to make the following adjustment. Indeed, the total probability of being in one of the four states at energy E_2 is $p(E_2) = 4p_2$, while the total probability of being in the one state at energy E_1 is still just $p(E_1) = 1p_1$.
 - In each state E_2 ,

$$\frac{p(E_2)}{p(E_1)} = \frac{N_2}{N_1} e^{-(E_2-E_1)/k_B T}$$

- The weekly quiz.
 - The first quiz will be next week.
 - A Canvas quiz – we'll have 24 hours to take it, but only 1 hour to take it.

1.2 Boltzmann Factor Examples / Partition Function

1/12:

- We will apply the Boltzmann factor to electronic, magnetic, translational, rotational, and vibrational molecular states.
- Example: Sodium lamp – two lines at 589.6 nm and 589.0 nm with intensity ratio 1:2.

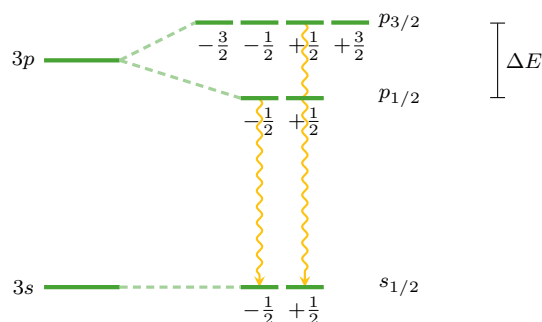


Figure 1.1: Sodium lamp energy levels.

- Street lamps use this (very efficient).
- Also used in astronomy.
- In the sodium atom, there are two energy levels (3s and 3p).
- The states have a spin-orbit coupling effect.

- $3s$ (with $S = 1/2$) splits into two degenerate states $s_{\pm 1/2}$ based on spin.
- $3p$ (with $L = 1$ and $S = 1/2$) splits into two nondegenerate states ($l = \pm 1$ [called $p_{3/2}$] and $l = 0$ [called $p_{1/2}$]), which further subdivide into four (resp. two) degenerate states ($-3/2, -1/2, 1/2, 3/2$ and $-1/2, 1/2$).
- Let ΔE be the difference in energy between the $p_{3/2}$ and $p_{1/2}$. Then

$$\frac{\Delta E}{k_B} = \frac{1}{k_B} \left(\frac{hc}{\lambda_1} - \frac{hc}{\lambda_2} \right) = 25 \text{ K}$$

where $\lambda_1 = 589.6 \text{ nm}$ and $\lambda_2 = 589.0 \text{ nm}$.

- Thus, $e^{-\Delta E/k_B T} \approx 1$ for $T = 300 \text{ K}$ (the temperature in the sodium vapor lamp).
- Therefore,

$$\begin{aligned} \frac{p(E_2)}{p(E_1)} &= \frac{4}{2} \cdot 1 \\ p(E_2) &= 2p(E_1) \end{aligned}$$

- Example: MRI.

- The magnetic field polarizes the spins of the hydrogen protons in our body with $\Delta E = \mu_B B$.
- If we also take $B = 6 \text{ T}$ and $T = 310 \text{ K}$ (body temperature), then

$$\frac{\mu_B B}{k_B T} = 2 \times 10^{-5}$$

- Thus, very few protons actually flip, but with modern technology we can still measure this.

- **Proton magnetic moment:** The magnetic moment of a proton. *Denoted by μ_B . Given by*

$$\mu_B = 1.4 \times 10^{-26} \text{ J/T}$$

- Example: Rotational.

- The rotational energy E_J of a molecule depends on the angular momentum quantum number J and the moment of inertia of the molecule $I = \mu R^2$ via the following relation.

$$E_J = \frac{\hbar^2}{2I} J(J+1)$$

- Microwave spectroscopy can be used to find molecules out in the universe.
- At 300 K,

$$\frac{p(J=1)}{p(J=0)} = \frac{3}{1} e^{\frac{-(E_1-E_0)}{k_B T}} = 2.95$$

■ As before $J = 1$ corresponds to states $j = -1, 0, 1$.

- See Figure 18.5 in the textbook.
- There is a range of angular momenta due to the temperature that for $T = 300 \text{ K}$ peaks around $J = 5$.

- Example: Vibrational.

- Here, $\Delta E = E_n - E_{n-1} = h\nu$ for every energy level since $E_n = h\nu(n + 1/2)$.
- It follows that

$$\frac{h\nu}{k_B} = 2800 \text{ K}$$

for CO, meaning that at 300 K, CO will be largely in its ground state.

- The partition function tells us everything we would want to know about a system.

$$Q = \sum_i e^{-E_i/k_B T}$$

- All we need to know is the energy of every state in the system.
- This is impossible for an infinite system, but the Schrödinger equation gives us the energy of a system, so its a great place to start.

- Calculating the total energy from the partition function.

- To construct it, start with

$$Q = \frac{p_1}{p_1} + \frac{p_2}{p_1} + \frac{p_3}{p_1} + \dots = 1 + e^{\frac{-(E_2-E_1)}{k_B T}} + e^{\frac{-(E_3-E_1)}{k_B T}} + \dots$$

- The total energy is equal to

$$\langle E \rangle = E_1 p_1 + E_2 p_2 + E_3 p_3 + \dots$$

- Taking $E_1 = 0$ gives

$$\langle E \rangle = p_1 \left[E_2 \frac{p_2}{p_1} + E_3 \frac{p_3}{p_1} + \dots \right]$$

- Note that

$$\frac{\partial}{\partial T} \left(e^{-E_2/k_B T} \right) = \frac{E_2}{k_B T^2} e^{-E_2/k_B T} = \frac{1}{k_B T^2} \left(E_2 \frac{p_2}{p_1} \right)$$

- Additionally,

$$\begin{aligned} p_1 &= 1 - (p_2 + p_3 + \dots) \\ &= 1 - p_1 \left(\frac{p_2}{p_1} + \frac{p_3}{p_1} + \dots \right) \\ &= 1 - p_1 (Q - 1) \\ p_1 &= \frac{1}{Q} \end{aligned}$$

- Therefore,

$$\begin{aligned} \langle E \rangle &= p_1 k_B T^2 \frac{\partial}{\partial T} \left(\frac{p_1}{p_1} + \frac{p_2}{p_1} + \dots \right) \\ &= p_1 k_B T^2 \frac{\partial Q}{\partial T} \\ &= \frac{1}{Q} k_B T^2 \frac{\partial Q}{\partial T} \\ \langle E \rangle &= k_B T^2 \frac{\partial}{\partial T} (\ln Q) \end{aligned}$$

- The above is an important result.

- Changing the origin of energy.

- We know that

$$\begin{aligned} Q(E_0) &= Q(E'_0) e^{-(E'_0 - E_0)/k_B T} \\ \ln Q(E_0) &= \ln Q(E'_0) - \frac{E'_0 - E_0}{k_B T} \end{aligned}$$

– Thus,

$$\begin{aligned}\langle E \rangle_{E_0} &= k_B T^2 \frac{\partial}{\partial T} (\ln Q(E_0)) \\ &= k_B T^2 \left(\frac{\partial}{\partial T} (\ln Q(E'_0)) - \frac{\partial}{\partial T} \left(\frac{E'_0 - E_0}{k_B T} \right) \right) \\ &= \langle E \rangle_{E'_0} + (E'_0 - E_0) \\ \langle E \rangle_{E_0} + E_0 &= \langle E \rangle_{E'_0} + E'_0\end{aligned}$$

– So the change of the energy origin does indeed change the total energy by the same amount.

1.3 Calculating Average Energies

- 1/14:
- We derived that for an ideal gas, $\langle E \rangle = 3k_B T/2$. But this may change at higher pressures.
 - Calculating the average kinetic energy at higher temperatures.
 - Use the main result from last time, which gives us the energy in terms of the partition function.
 - We have different degrees of freedom since KE and PE are on different coordinates (KE is on speed and PE is on position).
 - When we write the Boltzmann factor, we'll have an exponential with the sum of the kinetic and potential energy.

$$Q = \sum_{ij} e^{-(E_{KE_i} - E_{PE_j})/k_B T} = \sum_{ij} e^{-E_{KE_i}/k_B T} e^{-E_{PE_j}/k_B T} = Q_{KE} Q_{PE}$$

- The second equality holds because KE depends on the velocity coordinates and PE depends on position coordinates; thus, they are independent.
- Kinetic energy partition function.

$$E_{KE} = \frac{1}{2} m v_x^2$$

■ Thus,

$$Q_{KE_{v_x}} = \int_{-\infty}^{\infty} e^{-\frac{1}{2} m v_x^2 / k_B T} dv_x = \sqrt{\frac{2\pi k_B T}{m}}$$

- This function doesn't depend on anything of significant import.
- It follows that

$$\langle KE_x \rangle = k_B T^2 \frac{\partial}{\partial T} (\ln Q_{KE_{v_x}}) = k_B T^2 \frac{\partial}{\partial T} \left(\ln \sqrt{\frac{2\pi k_B T}{m}} + \frac{1}{2} \ln T \right) = \frac{k_B T}{2}$$

and

$$\langle KE \rangle = \langle KE_x \rangle + \langle KE_y \rangle + \langle KE_z \rangle = \frac{3}{2} k_B T$$

- Therefore, this result holds beyond the specific case of an ideal gas!
- Now for the potential energy of a harmonic oscillator.
 - $PE = \frac{1}{2} k x^2$; calculate the partition function for the coordinate x .

$$Q_x = \int_{-\infty}^{\infty} e^{-\frac{1}{2} k x^2 / k_B T} dx = \sqrt{\frac{2\pi k_B T}{k}}$$

- Thus,

$$\langle PE_x \rangle = \frac{k_B T}{2}$$

- For a 3D harmonic oscillator,

$$\langle PE \rangle = \frac{3}{2} k_B T$$

- Average potential energy of a gravitational potential.

- Apply the virial theorem (relates the average kinetic energy of a system in a conservative potential to the potential energy).
- Since we've shown that for any system, the average kinetic energy in one dimension is $k_B T/2$, the potential in any system will be related (i.e., have a factor of $k_B T$).

- What it means to cool something down, if KE always follows the same formula.

- Although the formula does not change, $\langle KE \rangle \propto T$, so decreasing the temperature decreases the kinetic energy.
- Similarly, as things change phase, more and more potentials take hold (e.g., in the gas phase, there is no potential energy, but there is significant potential energy in the solid and liquid phases).

- Rotational kinetic energy.

- Consider N_2 , with its two rotational degrees of freedom.
- Classically,

$$E_{\text{rot}} = \frac{1}{2} I \omega^2$$

- Thus, once again,

$$Q_\omega = \int_{-\infty}^{\infty} e^{-\frac{1}{2} I \omega^2 / k_B T} d\omega = \sqrt{\frac{2\pi k_B T}{I}}$$

making

$$\langle E_{\text{rot}} \rangle = \frac{k_B T}{2}$$

for one degree of freedom.

- **Law of Dulong and Petit:** The heat capacity of elemental solids is about $3nR$.

- Observed in 1819.
- A major result in an era where atomic structure was just emerging.
- Imagine an atom bound in a three-dimensional (octahedral) potential. It's energy is thus

$$\frac{1}{2} m v^2 + \frac{1}{2} k r^2$$

- Thus,

$$\begin{aligned} \langle E_{\text{atom}} \rangle &= \frac{3}{2} k_B T + \frac{3}{2} k_B T = 3k_B T \\ \langle E_{\text{solid}} \rangle &= 3N k_B T = 3n N_A k_B T = 3n R T \end{aligned}$$

- Some heat capacities are lower than $3nR$ (solids of rare gases that are heavier and need more heat to behave ideally), and some are higher (the potential is not a harmonic potential).
- As experiments got better, people realized that heat capacity, as a function of temperature, decreases as $T \rightarrow 0$ K, and was only asymptotic at $3nR$ at temperatures sufficiently close to room temperature.

- Quantum mechanics, especially the work of Einstein, solved this mystery.
- Atomic motion is quantized in units of energy.
 - If the temperature is much higher than the quantized energies, the system behaves classically.
 - If the temperature drops below the quantization energies of the vibration, we will not have equal population of energy levels (most will be in the ground state, making the energy 0; thus, there is no derivative of it and no heat capacity).
- Partition function of a quantum harmonic oscillator and the energy of the oscillator.
 - Recall that the energies are given by $(n + 1/2)h\nu$.
 - The partition function of the vibration of the quantum harmonic oscillator is

$$\begin{aligned}
 Q &= 1 + e^{-h\nu/k_B T} + e^{-2h\nu/k_B T} + \dots \\
 Q &= (e^{-h\nu/k_B T})^0 + (e^{-h\nu/k_B T})^1 + (e^{-h\nu/k_B T})^2 + \dots \\
 Q - Qe^{-h\nu/k_B T} &= 1 \\
 Q &= \frac{1}{1 - e^{-h\nu/k_B T}}
 \end{aligned}$$

when we take the zero point energy as our zero of energy.

- It follows that

$$\begin{aligned}
 \langle E \rangle &= k_B T^2 \frac{\partial}{\partial T} \left[\ln \left(\frac{1}{1 - e^{-h\nu/k_B T}} \right) \right] \\
 &= \frac{h\nu}{e^{h\nu/k_B T} - 1}
 \end{aligned}$$

- As $T \rightarrow \infty$, $h\nu/k_B T$ gets very small. But since $e^x \approx 1 + x$ at small x , as $T \rightarrow \infty$, we have that

$$\langle E \rangle \approx \frac{h\nu}{(1 + h\nu/k_B T) - 1} = k_B T$$

- Therefore, as $T \rightarrow \infty$, we recover the energy of a classical harmonic oscillator.
- On the other hand, as $T \rightarrow 0$, $E \rightarrow 0$.
- Note that heat capacity $C = \partial E / \partial T$.

1.4 Chapter 17: The Boltzmann Factor and Partition Functions

From McQuarrie and Simon (1997).

1/23:

- Up to this point, we have established that all physical systems' energy states are quantized. Now, we address questions such as "what fraction of the molecules are to be found in the ground vibrational state, the first excited vibrational state, and so on" (McQuarrie & Simon, 1997, p. 693).
 - We will see how notions such as 'higher temperature systems should have more populated excited states' translate into precise mathematics.
 - Our two most important tools to address such questions are the **Boltzmann factor** and the **partition function**.
- **Boltzmann factor:** The relation between the probability that a system will be in a given state to the energy of that state. *Given by*

$$p_j \propto e^{-E_j/k_B T}$$

- **Partition function:** A function in terms of which we can express all of the macroscopic properties of a given system, such as energy, heat capacity, and pressure. Denoted by Q , $Q(N, V, \beta)$, $Q(N, V, T)$. Given by

$$Q = \sum_j e^{-E_j/k_B T}$$

- Determining on what quantities the energies of a macroscopic system (such as some volume of fluid, gas, or solid) depend.
 - Consider an ideal gas confined to a cubic box of side length a (we will not generalize this result explicitly even though it can be).
 - In such a system, the constituent particles do not interact, so the energy of the system will be a simple sum

$$E_j(N, V) = \epsilon_1 + \cdots + \epsilon_N$$

of the energies $\epsilon_1, \dots, \epsilon_N$ of the N particles with no higher-degree interaction terms necessary.

- Additionally, the confinement means that quantum mechanically, every particle exists in a potential of zero within the cubic box and is subject to infinite potential outside the box.
- Thus, if we consider only the translational energies of each particle, we may apply the particle in a 3D cubic box model from Chapter 3 of McQuarrie and Simon (1997) to learn that each

$$\epsilon_i = \frac{h^2}{8ma^2}(n_x^2 + n_y^2 + n_z^2)$$

- Notice that E_j depends on N and V in this system via the N terms in the summation and the dependence of each ϵ_i on $a = \sqrt[3]{V}$.
- These are the most important (and general) factors on which E_j depends, and hence we often denote the energy of the j^{th} state of the system by $E_j(N, V)$
- **Heat reservoir:** An essentially infinite heat bath.
- **Ensemble:** A huge collection of systems with identical values of N , V , and T in thermal contact with a heat reservoir at a temperature T .
 - For a given ensemble, we denote the number of systems in state j by a_j and the total number of systems by \mathcal{A} .
- Finding the relative number a_m/a_n of systems in the ensemble in states a_n and a_m .
 - a_m/a_n will depend on the energies E_n and E_m via some function f , i.e.,

$$\frac{a_m}{a_n} = f(E_n, E_m)$$

- Energies are given relative to some zero, but a_m/a_n will not depend on the arbitrary choice of this zero. Thus, the math must make said zero cancel, so we take

$$\frac{a_m}{a_n} = f(E_n - E_m)$$

- The above equation must hold for any two energy states, so we can also write $a_l/a_m = f(E_m - E_l)$ and $a_l/a_n = f(E_n - E_l)$ for instance. But this implies that

$$\frac{a_l}{a_n} = \frac{a_m}{a_n} \cdot \frac{a_l}{a_m}$$

$$f(E_n - E_l) = f(E_n - E_m)f(E_m - E_l)$$

- The above equation uniquely describes an exponential function, so we take $f(E) = e^{\beta E}$ ^[2].

■ To check that our definition of f satisfies the above equation, note that

$$e^{\beta(E_n - E_l)} = e^{\beta(E_n - E_m)} e^{\beta(E_m - E_l)}$$

- Deriving an expression for a_m .

- From the above result, we have that

$$\begin{aligned}\frac{a_m}{a_n} &= e^{\beta(E_n - E_m)} \\ a_m e^{\beta E_m} &= a_n e^{\beta E_n}\end{aligned}$$

i.e., that the value of $a_n e^{\beta E_n}$ is the same (hence constant) for any n since m is arbitrary. Thus, let $C = a_n e^{\beta E_n}$.

- It follows that

$$\begin{aligned}\frac{a_m}{a_n} &= e^{\beta(E_n - E_m)} \\ a_m &= a_n e^{\beta E_n} e^{-\beta E_m} \\ a_m &= C e^{-\beta E_m}\end{aligned}$$

- Determining C .

- We have that

$$\begin{aligned}C \sum_j e^{-\beta E_j} &= \sum_j a_j = \mathcal{A} \\ C &= \frac{\mathcal{A}}{\sum_j e^{-\beta E_j}}\end{aligned}$$

- Thus,

$$\frac{a_j}{\mathcal{A}} = \frac{1}{\sum_j e^{-\beta E_j}} e^{-\beta E_j(N, V)}$$

- Taking the limit as the number of systems in the ensemble goes to infinity makes $a_j/\mathcal{A} \rightarrow p_j$, where p_j is the *probability* that a system will be in state j (see MathChapter B).
- Recognizing that the denominator above is the partition function (and a function of N , V , and β), we have that

$$p_j(N, V, \beta) = \frac{1}{Q(N, V, \beta)} e^{\beta E_j(N, V)}$$

- We will later show that $\beta = 1/k_B T$.

- Note, however, that from a theoretical point of view, β can be just as useful as T .

- Expressions for the average energy $\langle E \rangle$ of a system.

- From the definition of $\langle E \rangle$ in MathChapter B, we have that

$$\langle E \rangle = \sum_j p_j E_j = \sum_j \frac{E_j e^{-\beta E_j}}{Q}$$

²Note that the base need not be e , but we can take it to be e WLOG since $e^{\beta E} = (e^\beta)^E$ and we may take β such that e^β equals any positive real number.

- We can also express $\langle E \rangle$ entirely in terms of Q since

$$\frac{\partial \ln Q}{\partial \beta} = - \sum_j \frac{E_j e^{-\beta E_j}}{Q} = - \langle E \rangle$$

$$\langle E \rangle = - \frac{\partial \ln Q}{\partial \beta}$$

- Substituting $\beta = 1/k_B T$ and applying the chain rule to $\partial \ln Q / \partial T$ yields

$$\frac{\partial f}{\partial T} = \frac{\partial f}{\partial \beta} \cdot \frac{\partial \beta}{\partial T} = \frac{\partial f}{\partial \beta} \cdot \frac{1}{k_B} \frac{\partial}{\partial T} \left(\frac{1}{T} \right) = \frac{\partial f}{\partial \beta} \cdot -\frac{1}{k_B T^2}$$

$$\frac{\partial f}{\partial \beta} = -k_B T^2 \frac{\partial f}{\partial T}$$

so

$$\langle E \rangle = k_B T^2 \frac{\partial \ln Q}{\partial T}$$

- Using results from Chapter 14, McQuarrie and Simon (1997) calculates the average energy of a bare proton in a magnetic field, concluding that at $T = 0$ (i.e., zero thermal energy), the proton orients itself in the direction of the magnetic field with certainty while as $T \rightarrow \infty$, the thermal energy is such that the proton becomes equally likely to be in either state.
- Calculating the average energy of a monatomic ideal gas.

- From Chapter 18,

$$Q(N, V, \beta) = \frac{[q(V, \beta)]^N}{N!} \qquad q(V, \beta) = \left(\frac{2\pi m}{h^2 \beta} \right)^{3/2} V$$

- We have that

$$\begin{aligned} \ln Q &= N \ln q - \ln N! \\ &= -\frac{3N}{2} \ln \beta + \frac{3N}{2} \ln \left(\frac{2\pi m}{h^2} \right) + N \ln V - \ln N! \\ &= -\frac{3N}{2} \ln \beta + \text{terms not involving } \beta \end{aligned}$$

- Therefore,

$$\langle E \rangle = - \frac{\partial \ln Q}{\partial \beta} = \frac{3N}{2} \frac{d \ln \beta}{d \beta} = \frac{3N}{2\beta} = \frac{3}{2} N k_B T = \frac{3}{2} n R T$$

- The above result leads us to a fundamental postulate of physical chemistry: “The ensemble average of any quantity, as calculated using the probability distribution $[p_j = \frac{1}{Q} e^{-\beta E_j}]$, is the same as the experimentally observed value of that quantity” (McQuarrie & Simon, 1997, p. 700).
- The experimentally observed energy of a system is denoted by U .
- A molar quantity is denoted by an overbar (e.g., \bar{U} is the experimentally observed energy of one mole of a system).
- Calculating the average energy of a diatomic ideal gas.

- From Chapter 18,

$$Q(N, V, \beta) = \frac{[q(V, \beta)]^N}{N!} \qquad q(V, \beta) = \left(\frac{2\pi m}{h^2 \beta} \right)^{3/2} V \cdot \frac{8\pi^2 I}{h^2 \beta} \cdot \frac{e^{-\beta h \nu / 2}}{1 - e^{-\beta h \nu}}$$

for the rigid rotator-harmonic oscillator model of an ideal diatomic gas.

- The first term in the expression for $q(V, \beta)$ is translational (and identical to that of a monatomic ideal gas), the second term is rotational, and the third term is vibrational.
- Using the same procedure as before, we can calculate that for one mole of a diatomic ideal gas,

$$\bar{U} = \frac{3}{2}RT + RT + \frac{N_A h \nu}{2} + \frac{N_A h \nu e^{-\beta h \nu}}{1 - e^{-\beta h \nu}}$$

- “The first term represents the average translational energy, the second term represents the average rotational energy, the third term represents the zero-point vibrational energy, and the fourth term represents the average vibrational energy” (McQuarrie & Simon, 1997, p. 701).
- Note that the fourth term becomes significant only at higher temperatures.
- **Constant-volume heat capacity:** A measure of how the energy of a system changes with temperature at constant amount and volume. Denoted by C_V . Given by

$$C_V = \frac{\partial \langle E \rangle}{\partial T} = \frac{\partial U}{\partial T}$$

- We can express C_V in terms of Q via our above expression for $\langle E \rangle$ as a function of Q .
- For an ideal monatomic gas,

$$\bar{C}_V = \frac{3}{2}R$$

- For an ideal diatomic gas,

$$\bar{C}_V = \frac{5}{2}R + R \left(\frac{h\nu}{k_B T} \right)^2 \frac{e^{-h\nu/k_B T}}{(1 - e^{-h\nu/k_B T})^2}$$

- Molar heat capacity of a crystal, as per the **Einstein model of atomic crystals**.
 - Because each lattice site is identical, assume further that all atoms vibrate with the same frequency.
 - The associated partition function is thus

$$Q = e^{-\beta U_0} \left(\frac{e^{-\beta h \nu / 2}}{1 - e^{-\beta h \nu}} \right)^{3N}$$

where ν is characteristic of the particular crystal and U_0 is the **sublimation energy** (at 0 K).

- It follows as before that

$$\bar{C}_V = 3R \left(\frac{h\nu}{k_B T} \right)^2 \frac{e^{-h\nu/k_B T}}{(1 - e^{-h\nu/k_B T})^2}$$

- **Einstein model of atomic crystals:** A model of a crystal as N atoms situated at lattice sites, with each atom vibrating as a three-dimensional harmonic oscillator.
- **Sublimation energy** (at T): The energy needed to separate all the atoms from one another at T .
- One important consequence of this result is that by experimentally measuring the heat capacity of a crystal at different temperatures, we can determine its fundamental frequency ν .
- Another is the **law of Dulong and Petit**.
- **Law of Dulong and Petit:** The molar heat capacities of atomic crystals should level off at a value of $3R = 24.9 \text{ J mol}^{-1} \text{ K}^{-1}$ at high temperatures.
- Expressions for the average pressure $\langle P \rangle$ of a system.

- From Chapter 19, the pressure of a macroscopic system in state J is

$$P_j(N, V) = -\frac{\partial E_j}{\partial V}$$

- From the definition of $\langle P \rangle$ in MathChapter B, we have that

$$\langle P \rangle = \sum_j -\frac{\partial E_j}{\partial V} \frac{e^{-\beta E_j}}{Q}$$

- We can make the above more compact since

$$\frac{\partial Q}{\partial V} = -\beta \sum_j \frac{\partial E_j}{\partial V} e^{-\beta E_j} = Q\beta \langle P \rangle$$

so

$$\langle P \rangle = \frac{1}{\beta} \frac{\partial \ln Q}{\partial V} \qquad \langle P \rangle = k_B T \frac{\partial \ln Q}{\partial V}$$

- Just like we did with energy, we equate the ensemble average pressure with the observed pressure via $P = \langle P \rangle$.
- Deriving the ideal-gas equation of state.
 - We restrict ourselves at first to the special case of a monatomic ideal gas.
 - As before, we have that

$$\begin{aligned} \ln Q &= \frac{3N}{2} \ln \left(\frac{2\pi m}{h^2 \beta} \right) + N \ln V - \ln N! \\ &= N \ln V + \text{terms not involving } V \end{aligned}$$

- Therefore,

$$\begin{aligned} \langle P \rangle &= k_B T \frac{\partial \ln Q}{\partial V} = \frac{N k_B T}{V} \\ PV &= N k_B T \end{aligned}$$

- Since only the terms not involving V change for diatomic and polyatomic gases, the above equation of state holds for all ideal gases.
- The partition function associated with the van der Waals equation is

$$Q(N, V, \beta) = \frac{1}{N!} \left(\frac{2\pi m}{h^2 \beta} \right)^{3N/2} (V - Nb)^N e^{\beta a N^2 / V}$$

- Indeed, going through the same process as above with this equation yields

$$\left(P + \frac{aN^2}{V^2} \right) (V - Nb) = N k_B T$$

- Since we lack the computational power to calculate the set $\{E_j\}$ of eigenvalues of the N -body Schrödinger equation, we often approximate each E_j as the sum of the energies of the constituent particles of a system.
- Consider a system of independent, *distinguishable* particles.

- A good example of one is the Einstein model of atomic crystals, since each atom is assumed to vibrate independently of the others and each atom is distinguishable by its position in the crystal lattice.
- Applying the summation approximation, we get

$$E_l(N, V) = \underbrace{\varepsilon_i^a(V) + \varepsilon_j^b(V) + \varepsilon_k^c(V) + \cdots}_{N \text{ times}}$$

where each ε_i^a denotes the energy of an individual particle (i being the energy state, and a being the index of the particle [they are distinguishable]).

- Under this approximation, the partition function of the system becomes

$$Q(N, V, T) = \sum_l e^{-\beta E_l} = \sum_{i,j,k,\dots} e^{-\beta(\varepsilon_i^a(V) + \varepsilon_j^b(V) + \varepsilon_k^c(V) + \cdots)}$$

- Since we can sum over the indices separately (i.e., one after another), the above summation can mathematically be rewritten

$$\begin{aligned} Q(N, V, T) &= \sum_i e^{-\beta \varepsilon_i^a} \sum_j e^{-\beta \varepsilon_j^b} \sum_k e^{-\beta \varepsilon_k^c} \cdots \\ &= q_a(V, T) q_b(V, T) q_c(V, T) \cdots \end{aligned}$$

where each $q(V, T)$ is a **molecular partition function**.

- **Molecular partition function:** A partition function pertaining to a particular molecule within a system. Denoted by $q(V, T)$. Given by

$$q(V, T) = \sum_j e^{-\varepsilon_j / k_B T}$$

- Often able to be evaluated since they only depend on the allowed energies of individual atoms or molecules.
- **Boson:** A particle whose wave function must be symmetric under the interchange of two identical particles.
 - Particles of integer spin (such as photons [spin 1] and deuterons [spin 0]) are bosons.
 - Two identical fermions *can* occupy the same single-particle energy state.
- **Fermions:** A particle whose wave function must be antisymmetric under the interchange of two identical particles.
 - Particles of half-integer spin (such as electrons, protons, and neutrons [all with spin 1/2]) are fermions.
 - Two identical fermions *cannot* the same single-particle energy state.
- Consider a system of independent, *indistinguishable* particles.
 - As before, we have that $E_{ijk\dots} = \varepsilon_i + \varepsilon_j + \varepsilon_k + \cdots$ for N terms, but we since we cannot distinguish between particles, we cannot sum over the indices separately.
 - Thus, our partition function is set at

$$Q(N, V, T) = \sum_{i,j,k,\dots} e^{-\beta(\varepsilon_i + \varepsilon_j + \varepsilon_k + \cdots)}$$

- If the particles in question are fermions, the indices are not independent of each other.

- In particular, we cannot have a $\epsilon_1 + \epsilon_1 + \epsilon_2 + \epsilon_3 + \dots$ term in the summation because no two fermions can occupy the same single-particle energy state.
- This restriction omits all terms with more than one particle in the same energy state from the summation.
- If the particles in question are bosons, then we must avoid summing identical terms.
 - In particular, terms such as $\epsilon_1 + \epsilon_2 + \epsilon_2 + \epsilon_2 + \dots$ and $\epsilon_2 + \epsilon_1 + \epsilon_2 + \epsilon_2 + \dots$ represent the same state, and thus should be included only once in the summation. However, an unrestricted summation would include N such terms.
 - On the other end of the spectrum, there are $N!$ terms that include $\epsilon_1 + \epsilon_2 + \epsilon_3 + \dots + \epsilon_N$ in some order.
- In either case, the terms causing the problems are those with two or more identical indices.
- If it were not for such terms, we could carry out the summation in an unrestricted manner, obtaining $[q(V, T)]^N$ as with distinguishable particles (forcing each particle to have a different state is mathematically equivalent to making them distinguishable), and then divide by $N!$, obtaining $[q(V, T)]^N / N!$, to account for the over-counting.
- But such terms do exist in the summation. However, there are times when their presence is negligible.
 - If, for example, the number of quantum states available to any particle is significantly greater than the number of particles, it is very unlikely that two particles will ever be in the same state.
 - Indeed, most quantum-mechanical systems have an infinite number of energy states. However, at any given temperature, many of these states will be energetically inaccessible.
 - Thus, we can only invoke the $[q(V, T)]^N / N!$ approximation if the number of quantum states with energies less than $k_B T$ (which is roughly the average energy of a molecule) is much larger than the number of particles.
- In particular, if

$$\frac{N}{V} \left(\frac{h^2}{8mk_B T} \right)^{3/2} \ll 1$$

then we may approximate

$$Q(N, V, T) = \frac{[q(V, T)]^N}{N!}$$

- This criterion favors large particle mass, high temperature, and low density.
- It is easily satisfied in most conventional cases.
- Quantum systems that do not satisfy this criterion must be treated by special methods beyond the scope of McQuarrie and Simon (1997).
- **Boltzmann statistics:** Statistics in which the number of available molecular states is much greater than the number of particles.
 - Favored by high temperatures.
- Relating the average energy $\langle E \rangle$ of a system to the average energy $\langle \epsilon \rangle$ of a constituent molecule.

$$\begin{aligned} \langle E \rangle &= k_B T^2 \frac{\partial \ln Q}{\partial T} \\ &= k_B T^2 \frac{\partial}{\partial T} \left(\ln \left(\frac{q^N}{N!} \right) \right) \\ &= N k_B T^2 \frac{\partial \ln q}{\partial T} \\ \langle E \rangle &= N \langle \epsilon \rangle \end{aligned}$$

- The probability that a molecule is in its j^{th} molecular energy state is denoted by π_j .
- Assuming $\varepsilon = \varepsilon_i^{\text{trans}} + \varepsilon_i^{\text{rot}} + \varepsilon_i^{\text{vib}} + \varepsilon_i^{\text{elec}}$ allows us to write $q = q_{\text{trans}}q_{\text{rot}}q_{\text{vib}}q_{\text{elec}}$ since the various energy terms are distinguishable here (hence we can sum over the indices separately).
- The probability that a molecule is in its i^{th} translational, j^{th} rotational, k^{th} vibrational, and l^{th} electronic state is given by

$$\pi_{ijkl} = \frac{e^{-\varepsilon_i^{\text{trans}}/k_B T} e^{-\varepsilon_j^{\text{rot}}/k_B T} e^{-\varepsilon_k^{\text{vib}}/k_B T} e^{-\varepsilon_l^{\text{elec}}/k_B T}}{q_{\text{trans}}q_{\text{rot}}q_{\text{vib}}q_{\text{elec}}}$$

- Since the total probability π_k^{vib} that a molecule is in its k^{th} vibrational state (for example) encompasses all probabilities of it being in any translational, rotational, or vibrational state, we have by summation that

$$\pi_k^{\text{vib}} = \sum_{i,j,l} \pi_{ijkl} = \frac{\sum_i \left(e^{-\varepsilon_i^{\text{trans}}/k_B T} \right) \sum_j \left(e^{-\varepsilon_j^{\text{rot}}/k_B T} \right) \sum_l \left(e^{-\varepsilon_l^{\text{elec}}/k_B T} \right) e^{-\varepsilon_k^{\text{vib}}/k_B T}}{q_{\text{trans}}q_{\text{rot}}q_{\text{vib}}q_{\text{elec}}} = \frac{e^{-\varepsilon_k^{\text{vib}}/k_B T}}{q_{\text{vib}}}$$

- It follows that

$$\langle \varepsilon^{\text{vib}} \rangle = \sum_k \varepsilon_k^{\text{vib}} \frac{e^{-\varepsilon_k^{\text{vib}}/k_B T}}{q_{\text{vib}}} = -\frac{\partial \ln q_{\text{vib}}}{\partial \beta} = k_B T^2 \frac{\partial \ln q_{\text{vib}}}{\partial T}$$

– Analogous results hold for $\langle \varepsilon^{\text{trans}} \rangle$, $\langle \varepsilon^{\text{rot}} \rangle$, and $\langle \varepsilon^{\text{elec}} \rangle$.

- Although we have written partition functions as sums over energy *states* up to this point, we can also sum over energy *levels* by including the degeneracy g_J of the level.

– For example, since the energy and degeneracy of a rigid rotator are, respectively,

$$\varepsilon_J = \frac{\hbar^2}{2I} J(J+1) \qquad g_J = 2J+1$$

we can write

$$q_{\text{rot}}(T) = \sum_{J=0}^{\infty} (2J+1) e^{-\hbar^2 J(J+1)/2Ik_B T}$$

Week 2

Partition Functions and Ideal Gases

2.1 System Partition Functions

1/19:

- Decomposing the partition function of a molecule into the product of separate sums as partitioned by degrees of freedom (e.g., translation, rotation, vibration, and electronic).
- The partition functions of **independent**, distinguishable/indistinguishable molecules.
 - We should not double count the same states.
 - The $N!$ in $Q = q^N/N!$ is not important when calculating energy (because of the properties of the \ln function), but it is very important when calculating quantities such as entropy.
- **Independent** (particles): A set of particles that do not interact with one another.
- Discusses bosons and fermions.
 - We can have a two fermions in the state $|1, 1\rangle$ because it is a symmetric state.
- Recall the Fermi level, the boundary between the filled and unfilled electronic states in a solid.
 - If T is small, this level is a hard boundary.
 - If T is large, electrons can easily be excited and the Fermi level is a soft boundary.
- Does the 3D particle in a box derivation for the translation molecular partition function.
 - Note that since the de Broglie wavelength $\lambda_{\text{DB}} = \sqrt{h^2/2mk_B T}$, we may write

$$q_x = \sum_{n_x} e^{-h^2/8mk_B T L_x^2} = \sum_{n_x} e^{-\lambda_{\text{DB}}^2 n_x^2 / 4L_x^2}$$

- The number of states are occupied/have energy within $k_B T$ of the ground state.
 - $-\lambda_{\text{DB}}^2 n_x^2 / 4L_x^2$ is on the order of 1, implying that n_x is on the order of $2L/\lambda_{\text{DB}}$.
 - It follows if L is on a macroscopic scale (e.g., $L \approx 1$ m) and λ_{DB} is on a sub-angstrom scale that n_x is on the order of 10^{10} . When n_x is at such a scale, $e^{-\lambda_{\text{DB}}^2 n_x^2 / 4L_x^2} \approx 1/e$.
 - It follows that in a 1 m^3 box, we will have about 10^{30} states, so we really are in a regime where the number of states is larger than the number of molecules.
- More precisely, we want

$$N \ll n_x n_y n_z = \left(\frac{8mk_B T}{h^2} \right)^{3/2} L_x L_y L_z$$

where the middle term approximates the number of states so that

$$\frac{N}{V} \ll \left(\frac{8mk_B T}{h^2} \right)^{3/2}$$

- Approximating the translational energy with an integral.
 - Concludes with the translational partition function.
 - Since we can approach this problem from a classical perspective (as we did last Friday) or quantum mechanically (as we did today) to achieve the same result, this system again demonstrates the relation between quantum and classical mechanics.

2.2 Molecular Partition Functions

- 1/21: • We approximate the total molecular energy as

$$q = q_{\text{elec}} q_{\text{trans}} q_{\text{vib}} q_{\text{rot}}$$

- The heat capacity in the very high temperature limit where translations, rotations, and vibrations are classical.
 - Translational: $\frac{3}{2}k_B$.
 - Vibrational: Each degree of freedom ($3N - 5$ for a linear molecule and $3N - 6$ for a nonlinear molecule) contributes k_B .
 - Rotational: Each degree of freedom (2 for a linear molecule and 3 for a nonlinear molecule) contributes $\frac{1}{2}k_B$.
- We can use the above to calculate the heat capacity of various molecules at very high temperatures (note, however, that at such temperatures, molecules would likely dissociate; we're simply theoretically considering the classical limit here).
 - Ne: $\frac{3}{2}k_B$.
 - H₂O: $\frac{3}{2}k_B + 3 \cdot k_B + 3 \cdot \frac{1}{2}k_B = 6k_B$.
 - O₂: $\frac{3}{2} + 1 \cdot k_B + 2 \cdot \frac{1}{2}k_B = \frac{7}{2}k_B$.
 - CO₂: $\frac{3}{2}k_B + 4 \cdot k_B + 2 \cdot \frac{1}{2}k_B = \frac{13}{2}k_B$.
 - CHCl₃: $\frac{3}{2}k_B + 9 \cdot k_B + 3 \cdot \frac{1}{2}k_B = 12k_B$.
- Electronic partition function.
 - Consider the bottom D_e of the potential well of a diatomic.
 - D_0 is the ionization energy from the bottom state ($D_e \neq D_0$, but relations can be obtained via spectroscopy).
 - It follows that

$$q_{\text{elec}} = g_1 e^{-(D_e/k_B T)} + g_2 e^{-E_2/k_B T}$$
 - If $dT \ll (E_2 + D_e)$, then $q_{\text{elec}} = g_1 e^{D_e/k_B T}$.
- Vibrational partition function.
 - As before with the law of Dulong and Petit.
 - It's a special point where $T = h\nu/k_B$.
- Rotational partition function.
 - Almost always classical.
 - The rotational energy of a polyatomic molecule will almost always be $\frac{3}{2}k_B$.
 - Let's look at a heteronuclear diatomic, such as CO. Derives

$$q_{\text{rot}} = \sum_{J=0}^{\infty} (2J+1) e^{-\hbar^2 J(J+1)/2Ik_B T}$$

- The **rotational temperature** leads to

$$q_{\text{rot}} = \sum_{J=0}^{\infty} (2J+1)e^{\Theta_{\text{rot}}/T} = \frac{T}{\Theta_{\text{rot}}}$$

- Thus, at the temperature at which we exist, rotation is equivalent classically to quantum mechanically.

- **Rotational temperature:** The following quantity. Denoted by Θ_{rot} . Given by

$$\Theta_{\text{rot}} = \hbar^2/2Ik_B$$

- PGS will not specify whether we need a quantum vs. classical model.

- Homonuclear diatomic (e.g., H_2).

- The vibrational differences in energy become visible with spectroscopy.

- $q_{\text{rot}} = T/2\Theta_{\text{rot}}$.

- Partition functions:

- If the molecule is linear, it's of the form T/Θ_{rot} .
- If the molecule is nonlinear, it's of the form $T/2\Theta_{\text{rot}}$.
- Spherical top (e.g., CH_4):

$$\frac{\sqrt{\pi}}{\sigma} \left(\frac{T}{\Theta_{\text{rot}}} \right)^{3/2}$$

- Symmetric top (e.g., NH_3):

$$\frac{\sqrt{\pi}}{\sigma} \sqrt{\frac{T^3}{\Theta_{\text{rot},a}^2 \Theta_{\text{rot},b}}}$$

- a and b are the two different symmetry axes.

- Asymmetric top (e.g., H_2O):

$$\frac{\sqrt{\pi}}{\sigma} \sqrt{\frac{T^3}{\Theta_{\text{rot},a} \Theta_{\text{rot},b} \Theta_{\text{rot},c}}}$$

- Application to total energy and heat capacity of a molecule.

- We have that

$$q = \left(\frac{2\pi mk_B T}{h^2} \right)^{3/2} \cdot V \cdot \sqrt{\frac{T^2}{\sigma \Theta_{\text{rot},a} \Theta_{\text{rot},b} \Theta_{\text{rot},c}}} \cdot \sum_1^{3N-6} \frac{e^{-\Theta_{\text{vib}}/2T}}{1 - e^{-\Theta_{\text{vib}}/T}} \cdot g_1 e^{D_e/k_B T}$$

- Thus,

$$\langle E \rangle = k_B T^2 \frac{\partial \ln q}{\partial T} = k_B T^2 \frac{\partial}{\partial T} \left(\frac{3}{2} \ln T + \text{constant} + \frac{3}{2} \ln T + \text{vibration} - D_e \right)$$

- The energy of the vibration is $E = k_B \Theta_{\text{vib}} / (e^{\Theta_{\text{vib}}/k_B T} - 1) + k_B \Theta_{\text{vib}}/2$. It follows that

$$C = \frac{\partial E}{\partial T} = k_B \frac{\Theta_{\text{vib}}^2}{T^2} \frac{e^{-\Theta_{\text{vib}}/T}}{(1 - e^{-\Theta_{\text{vib}}/T})}$$

2.3 Chapter 18: Partition Functions and Ideal Gases

From McQuarrie and Simon (1997).

- Herein, we will calculate the partition functions and heat capacities of ideal gases.
 - We heavily rely on the expression of the partition function for a system of independent, indistinguishable particles, which ideal gases are likely to satisfy because of their low density.
- Deriving the translational molecular partition function of an atom in a monatomic ideal gas.
 - As mentioned in Chapter 17, if we let the container be cubic, then

$$\varepsilon(n_x, n_y, n_z) = \frac{h^2}{8ma^2}(n_x^2 + n_y^2 + n_z^2)$$

- It follows that

$$\begin{aligned} q_{\text{trans}} &= \sum_{n_x, n_y, n_z=1}^{\infty} e^{-\beta \varepsilon(n_x, n_y, n_z)} \\ &= \sum_{n_x=1}^{\infty} \exp\left(-\frac{\beta h^2 n_x^2}{8ma^2}\right) \sum_{n_y=1}^{\infty} \exp\left(-\frac{\beta h^2 n_y^2}{8ma^2}\right) \sum_{n_z=1}^{\infty} \exp\left(-\frac{\beta h^2 n_z^2}{8ma^2}\right) \\ &= \left[\sum_{n=1}^{\infty} \exp\left(-\frac{\beta h^2 n^2}{8ma^2}\right) \right]^3 \end{aligned}$$

- The above sum cannot be evaluated in closed form. However, since later terms in the summation get very small, it is an excellent approximation to replace the summation with an integral, i.e.,

$$\begin{aligned} q_{\text{trans}} &= \left(\int_0^{\infty} e^{-\beta h^2 n^2 / 8ma^2} dn \right)^3 \\ &= \left(\sqrt{\frac{\pi}{4\beta h^2 / 8ma^2}} \right)^3 \\ &= \left(\sqrt{\frac{2\pi m}{\beta h^2}} \right)^3 a^3 \\ &= \left(\frac{2\pi m k_B T}{h^2} \right)^{3/2} V \end{aligned}$$

Week 3

Kinetic Theory of Gases

3.1 Maxwell-Boltzmann Distribution

- 1/24:
- Applying the molecular partition function to the heat capacity of a water molecule.
 - A water molecule has three vibrational modes, which we will denote by ν_1, ν_2, ν_3 (corresponding to symmetric stretch, antisymmetric stretch, and bend).
 - Main takeaway: Heat capacity can change with temperature.
 - After a while (at several thousand kelvin), it will level off (see Figure 18.7).
 - Considers CO₂'s vibrational modes, too.
 - The infrared absorption of the bending mode is what's associated with the Greenhouse Effect.
 - The symmetric stretch is IR inactive due to its lack of change of dipole moment.
 - Raman active: Change in the polarizability of the molecule.
 - The Maxwell-Boltzmann distribution.
 - Maxwell derived it long before Boltzmann, but Boltzmann's thermodynamic derivation is much easier.
 - We know from the boltzmann factor that $p(E) \propto e^{-E/k_B T}$.
 - Thus, to get the probability $p(v)$ of some speed v , we should have $p(v) \propto e^{-mv^2/2k_B T}$ times a constant giving the number of molecules of each speed? This yields

$$p(v) = A4\pi v^2 e^{-mv^2/2k_B T}$$

where A is a normalization constant.

- The Maxwell-Boltzmann distribution is such that

$$\begin{aligned} 1 &= \int_0^\infty p(v) dv \\ &= A \int_0^\infty 4\pi v^2 e^{-mv^2/2k_B T} dv \\ &= A \int_0^\infty 4\pi \left(\frac{2k_B T}{m} \right)^{3/2} u^2 e^{-u^2} du \\ &= A4\pi \left(\frac{2k_B T}{m} \right)^{3/2} \int_0^\infty u^2 e^{-u^2} du \\ &= A4\pi \left(\frac{2k_B T}{m} \right)^{3/2} \frac{\sqrt{\pi}}{4} \end{aligned}$$

$$A = \left(\frac{m}{2\pi k_B T} \right)^{3/2}$$

– Therefore,

$$p(v) = 4\pi \left(\frac{m}{2\pi k_B T} \right)^{3/2} v^2 e^{-mv^2/2k_B T}$$

– Any distribution that doesn't look like this isn't in thermal equilibrium.

- A system with all particles having $v = 0$ is at thermal equilibrium with $T = 0$ K.
- A system with all particles having constant velocity in the same direction is at thermal equilibrium with $T = 0$ K.
 - Think relativity; if you're moving with them, it looks like they're not moving and thus this case is the same as the last one because your movement doesn't affect the thermodynamics of that system.
- A system with all particles having constant velocity in different directions is not at thermal equilibrium since it does not fit the bell curve but is rather a spike.

References

- Arrhenius, S. A. (1889). Über die dissociationswärme und den einfluß der temperatur auf den dissociationsgrad der elektrolyte. *Zeitschrift für Physikalische Chemie*, 4, 96–116. <https://doi.org/10.1515/zpch-1889-0408>
- Labalme, S. (2021a). *AP Chemistry notes*. Retrieved January 11, 2022, from <https://github.com/shadypuck/APChemNotes/blob/master/main.pdf>
- Labalme, S. (2021b). *PHYS 13300 (Waves, Optics, and Heat) notes*. Retrieved January 11, 2022, from <https://github.com/shadypuck/PHYS13300Notes/blob/master/Notes/notes.pdf>
- McQuarrie, D. A., & Simon, J. D. (1997). *Physical chemistry: A molecular approach*. University Science Books.