

CHEM 30200 (Synthesis and Physical Methods in Inorganic Chemistry) Notes

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Week 1

Nanomaterials and Intro to XRD

1.1 Nanomaterials

- 1/3:
- Contact Dr. Shevchenko at eshevchenko@anl.gov or eshevchenko@uchicago.edu.
 - How to make nano.
 - Top-down approach: Start with large, end with nano. Includes nanofabrication.
 - Bottom-up approach: Solution-based approach.
 - Scalable and cheap.
 - Use an inorganic core with a coating.
 - What nanoparticles look like.
 - Differences in size, size distribution, shape, chemical composition, and structures.
 - Different sizes (like atoms) and different shapes (like bacteria and viri).
 - Ancient nanoscience.
 - The Lycurgus cup.
 - A 4th-century Roman glass cage cup.
 - Currently housed at the British Museum.
 - Contains ~ 70 nm gold/silver nanoparticles.
 - When front-lit, it appears green (light is scattered by larger NPs).
 - When back-lit, it appears red (light is absorbed by NPs).
 - Hair dye.
 - 2000 years ago from Greco-Roman times.
 - Made of lead oxide (PbO), slaked lime ($Ca(OH)_2$), and water (H_2O).
 - The lead oxide combines with sulfur-rich peptides in the hair to make ~ 5 nm PbS NPs.
 - Applications of nanoparticles: Catalysis.
 - Refining of petroleum (transformation of crude oil into gasoline, jet fuel, diesel oil, and fuel oils).
 - Converter of automobile exhaust (reduction of nitrogen oxides [NO_x] to N_2 and O_2 ; oxidation of CO to CO_2 ; oxidation of unburnt hydrocarbons to $CO_2 + H_2O$).
 - Hydrogenation of CO (synthesis of fuels such as methane or methanol).
 - Selective oxidation of hydrocarbons (synthetic fibers, plastics, and fine chemicals).
 - Methods of NP analysis: XRD, TEM, XANES, and XPS.

- Applications of nanoparticles: Displays.
 - Semiconductor nanoparticles (e.g., solutions of CdSe or InP nanoparticles) emit different colors.
 - Sony has announced that it will embed quantum dots in its latest flat-screen TV.
 - QLEDs can be made out of CdSe, CdS, InP coatings with silica, perovskite (CsPbX where $\text{X} = \text{Cl}, \text{Br}, \text{I}$), and cesium lead halide salts.
- Milestones in the synthesis of nanomaterials (a subjective and incomplete POV).
 - Alexei Ekimov (late 1970s-1981, USSR): CuCl_x and CdSe in molten glass matrix (fluorescence, gradient colors).
 - Alexander Afros (1982): Theoretical description of size effect.
 - Louis Brus (1983, Bell Labs, US): CdS in solution.
 - Paul Alivisatos (UChicago) and Moungi Bawendi (MIT).
 - Moungi Bawendi et al. (1993): Synthesis of monodisperse CdSe nanoparticles — a big one!
 - Philippe Guyot-Sionnest (1996): Synthesis of core/shell nanoparticles.
 - Paul Alivisatos (1997 and 2003): Synthesis of nanorods and tetrapods.
 - Chris Murray and Shouheng Sun (2000): Synthesis of magnetic FePt nanoparticles.
 - Benoit Dubert (2007, France): Synthesis of CdSe nanoplates (more stable, emission is polarized and directional).
 - Maksym Kovalenko (2015): Synthesis of perovskites.
 - Mostafe El-Sayed, Catherine Murphy, Peidong Yang, and Yunan Xia: Synthesis of Au and Ag nanoparticles.
- Synthesis of nanoparticles.
 - The 1993 Bawendi paper.
 - The innovation was the synthesis of NPs in organic solvents, still widely used today.
- LaMer model.
 - Precursors undergo nucleation, focusing, and “nano”-Ostwald ripening.
 - Key idea: Separation of nucleation and growth in time.
- Nanomaterials: State-of-the-art.
 - The chemistry behind QD synthesis is rather simple compared to what is used by organic or coordination chemists, but the field sometimes lacks depth and chemical understanding.
 - Indeed, only a fraction of reported results have been reproduced, and only a fraction of those have been understood and optimized.
 - This is a big problem for AI/ML approaches.
 - During the next 5-10 years, nanomaterials synthesis will progress mostly through systematic mechanistic studies.
- Synthesis of nanocrystals without Ostwald Ripening.
 - The nanocrystals form and grow during 0.1-1 minute after the start of the reaction.
 - Annealing at high temperatures (250-280 °C) is required to improve crystallinity.
 - No change in particle size takes place upon the annealing.
 - Tune particle size with nucleation, since growth proceeds until all monomer is consumed — fast nucleation leads to many particles, which can only grow so large; slow nucleation leads to a few particles which can grow very large (conservation of end volume).

- Synthesis with “artificial molecules.”
 - Rearrangement, addition, substitution, and elimination.
- Hollow nanocrystals: Kirkendal Effect at Nanoscale.
 - Uniform spherical cobalt nanocrystals can be synthesized by rapid pyrolysis of cobalt carbonyl in hot solvent.
 - Hollow nanocrystals form after sulfidation reaction.
- Kirkendal effect.

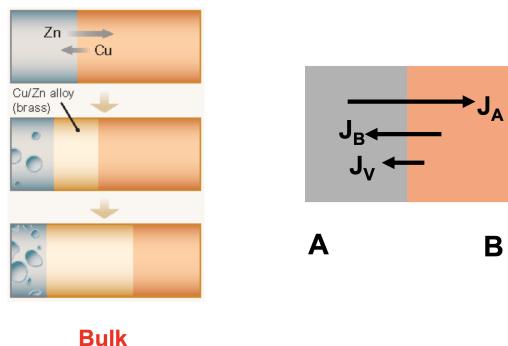
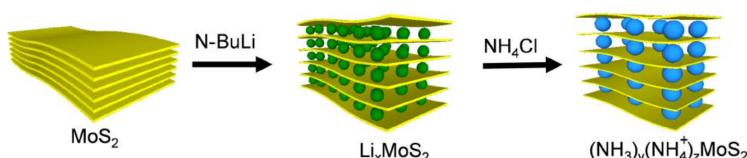


Figure 1.1: Kirkendal effect.

- Occurs when the diffusion rates of two species are different.
- When vacancies become supersaturated, they condense into voids in the fast diffusion species side.
- The Kirkendall voids result in weak bonding and lead to brittle fracture at the bonding interface.
- Top-down approaches for the synthesis of MoS₂.

Figure 1.2: Top-down synthesis of MoS₂.

- This is the synthesis of interlayer expanded (IE) MoS₂ through chemical intercalation of Li and the following exchange with NH₃ and NH₄.
- Each MoS₂ layer is composed of an atomic layer of Mo sandwiched between atomic layers of S through strong ionic/covalent bonds. Weak van der Waals forces link individual MoS₂ layers with an interlayer spacing of 0.615 nm.
- Procedure.
 - Electron-donating species, e.g., alkali metals, Lewis bases, and organolithium compounds can intercalate between layers.
 - Alkali metals can after that be evaporated or react with water.
 - **Exfoliation** with ultrasound (mechanical exfoliation).
- Useful characterization methods: TEM, XRD, XPS.
- **Exfoliation:** The complete separation of the layers of a material.

- Bottom-up approaches for the synthesis of MoS₂.
 - Chemical synthesis of interlayer expanded (IE) MoS₂.
 - Chemical Vapor Deposition (CVD).
- MXene = 2D metal and surface chemistry.
 - Applications to supercapacitors, batteries, conductors, catalysts, and composites.
 - Most made out of Ti₃AlC₂.
- MXenes: Solution processed 2D transition metal carbides and nitrides.
 - Scanning electron microscopy (SEM) and HRTEM images shown.
 - Etching and delamination phases.
- MXenes: Variable termination groups.
 - Ask in OH??
- MXenes are solution processable 2D transition metal carbides and nitrides.
 - Lists various experimentally synthesized structures.
- What can and cannot be synthesized in solution?
 - Current solution synthesis methodology can be applied to materials that crystallize below 400 °C.
Many materials that require higher temperatures to form, e.g., nitrides, carbides, GaAs.
 - Higher temperatures: Gas phase and solid state synthesis, as well as synthesis in molten salts.
- Nanocrystals in molten salts.
 - QDs are synthesized at high T in molten salts.
 - There is typically a postpreparative treatment phase still in molten salts.
- Nanoparticles as building blocks to make new materials.
 - The idea behind nanoparticles is encapsulated by the synthesis of NaCl from Na° and Cl₂°: Two substances with certain properties combine to form a new material with very different properties.
 - Assembly of atoms leads to new materials and new properties!
- Self-assembly of nanoparticles.
 - Can be multilayered (up to five).
- Crystals of nanocrystals.
 - Example: 3D crystals (~ 30 µm) have been assembled from 3.3 nm CdSe nanocrystals.
 - Example: 3D crystals (~ 10 µm) have been assembled from 6 nm CoPt₃ nanocrystals.
 - More conventional example: Crystals of quartz (made by atoms).
- Binary nanoparticle superlattices.
 - Natural opal vs. synthetic opals (very similar appearance and properties).
 - Formation of binary superlattices depends on the ratio of nanoparticle radii (γ), the concentration of nanoparticles, the size distribution of nanoparticles, the nature of the capping ligand, and the substrate.
 - Evaporating colloidal solutions of binary nanoparticle mixtures at 45 °C leads to the formation of binary nanoparticle superlattices.

- Periodic table of nanocrystals.
 - Different types of unit cells listed.
 - Examples include AlB₂, MgZn₂, Cu₃Au, Fe₄C, CaCu₅, and CaB₆.
- Directing of self-assembly of nanocrystals.
 - Various additives can be mixed in.
- Metal organic frameworks (MOFs).
 - MOFs (also called porous coordination polymers or PCPs) are two- or three-dimensional porous crystalline materials with infinite lattices synthesized from secondary building units (SBUs) — metal cations, salts, or clusters — and polydentate organic ligands with coordination type connections.
 - Common SBUs pictured.
 - From single-metal nodes to SBUs: More than 20,000 MOF structures have been reported so far.
 - Characterization methods: TEM, SEM, XRD, XANES, XPS, and electron paramagnetic resonance (EPR).
- Summary.
 - Structural information (XRD, ED).
 - Compositional information (XRD, ED, energy dispersive X-ray analysis [EDX], and X-ray fluorescence [XRF]).
 - Size and morphology of materials (TEM, SEM, and XRD).
 - Redox states of bulk and surface (XANES, XPS).
 - Important variables: Quality of synthesized materials, stability of materials during processes, and structure-property correlations.

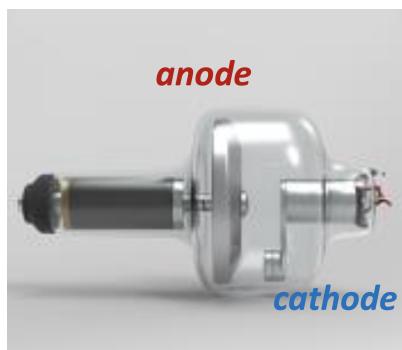
1.2 XRD Analysis 1

1/5:

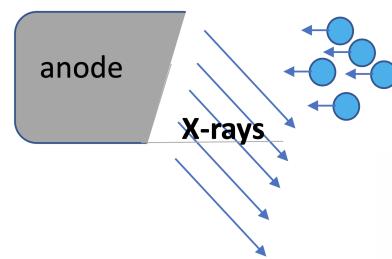
- Applications of X-ray diffraction.
 - Geology, environmental science, material science, engineering, biology.
 - Originally developed for analysis of minerals.
- What you can do with it.
 - Characterization of crystalline materials (structure, unit cell, etc.).
 - Identification of unknown/new crystalline materials (e.g., minerals, meteorites, etc.; rarely something on Earth since all of that has been extensively studied).
 - Evaluation of sample purity.
- XRD sample types: Can be used to characterize powders and thin films.
- Popular because...
 - It's relatively inexpensive.
 - There are benchtop instruments as well as larger ones.
 - It's relatively universal.
 - Standardized procedures.
 - It's nondestructive.

- Building blocks of X-ray diffractometers.
 - **X-ray tube**, sample, **backstop**, **goniometer**, **collimator**, **filter/monochromator**, **detector**, **shutter**, and **safety interlock**.
 - Basic definitions now; more on each of these devices later.
- **X-ray tube**: The shield around the X-ray tube.
- **Backstop**: The device located behind the sample that absorbs non-diffracted X-rays. *Also known as beam trap*.
- **Goniometer**: The mechanism that allows you to adjust the position of the sample relative to the incident X-ray beam.
- **Collimator**: The device that narrows the beam of X-rays.
- **Filter/monochromator**: The instrument that diffracts X-rays from a crystal to produce a beam comprising a narrow range of wavelengths.
- **Detector**: The device that measures diffraction patterns or energy spectra generated as a result of the interaction of X-rays with the sample.
- **Shutter**: The safety device (made of tungsten, tantalum, lead, etc.) inserted into the beam pass to stop all X-rays.
- **Safety interlock**: The device that stops the generation of X-rays when access to the interior part of the diffractometer is possible.
- How are X-rays generated?
 - Originally with a **Crookes tube**.
- **Crookes tube**: A type of discharge tube that facilitated the initial discovery of X-rays.
 - History.
 - Invented between 1869-1875 by William Crookes, who also discovered thallium.
 - Wilhelm Roentgen was responsible for the initial discover of X-rays — using a Crookes tube — in 1895.
 - Design.
 - Cold cathode vacuum tubes (no heated filament; these emit more electrons than can be supplied by thermionic emission alone).
 - Relies on electron emission induced by an electrostatic field or secondary electron emission.
 - Requires a small amount of gas (air) in them to function (10^{-6} - 5×10^{-8} atm).
 - High voltage is applied to the tube.
 - Electric current causes the ionization of gas molecules: Electrons “knock off” electrons from other gas molecules forming positive ions and negative electrons (negative ions are also formed as a result of electron interactions with neutral gas molecules).
 - High velocity electrons hit the anode (metal) that creates the X-rays.
- The first images obtained using X-rays.
 - The first medical X-ray was by Röntgen (alternate spelling of Roentgen) of his wife in 1885.
 - It led Roentgen to win the first ever Nobel in physics in 1901.
 - People started making X-rays of their bones just for fun — like taking selfies — without realizing the harm of X-rays.
 - Some doctors notice improvement in cancer patients after X-ray exposure.

- X-rays were discovered accidentally.
 - A Crookes tube was wrapped in black cardboard to block the visible light.
 - A fluorescent screen painted with phosphor barium platinocyanide ($\text{Ba}[\text{Pt}(\text{CN})_4]$) was $\sim 1\text{ m}$ away.
 - When the tube was activated, the screen glowed green.
- Limitations of Crookes tubes.
 - X-rays originate from a rather large area. Thus, the resulting X-ray images lack contrast.
 - Low intensity of X-rays; necessitates long exposures of the object.
- Thus, Willian Coolidge develops the Coolidge tube in 1913.
 - This is a hot cathode tube.
 - It works with a much higher vacuum ($\sim 10^{-9}\text{ atm}$).
 - Irving Langmuir designs an “extreme vacuum” with a new pump design (the Langmuir pump).
 - This enabled generation of an electron current using a hot thermionic filament as the cathode.
 - Other improvements.
 - Replaced Pt target with W.
 - Mo shield to focus the X-rays.
 - These improvements led to...
 - An X-ray source of remarkable brilliance and reliability;
 - Control over intensity and hence penetrating power.
 - Since then, basically nothing has changed, though we did start using rotating anodes (see below).
- Working principle of the X-ray tube.



(a) Overall design.



(b) Electrons to X-rays.

Figure 1.3: X-ray tube design.

- The **cathode** heats up and induces thermionic emission of a cloud of electrons.
- The **anode** attracts the electrons across the vacuum tube.
- The anode converts the energy of incident electrons into X-rays, dissipating heat as a byproduct.
- You rotate the anode very fast (up to 10 000 rpm) to prevent it from overheating (can get up to 2000°C). This is why you often have a cooling system in high-power X-rays.
- The anode is a beveled disk (shaped to focus the X-rays).
- Most X-ray tubes have an **anode angle** of 12-15°.
 - A smaller angle results in a smaller effective **focal spot**.

- **Cathode:** The negatively charged filament.
- **Anode:** The positively charged disk. *Also known as anticathode.*
- **Anode angle:** The angle between the vertical and the target surface.
- **Focal spot:** The surface area on the anode over which X-rays are produced.

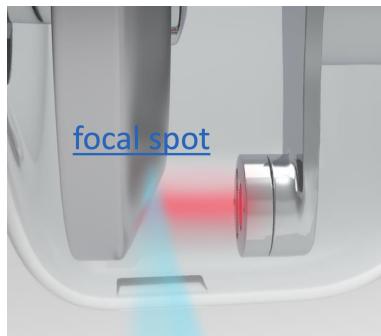


Figure 1.4: Anode focal spot.

- Naturally, increasing the angle will allow the electrons to spread out over more area.
- Design of the X-ray tube.

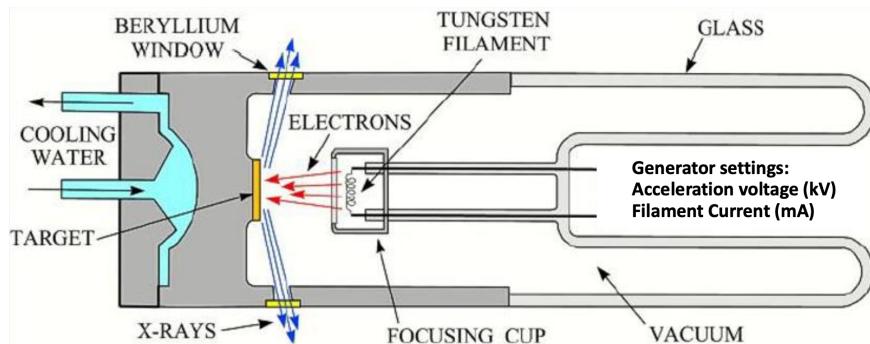


Figure 1.5: Schematic cross section of an X-ray tube.

- Beryllium is used as a window because it doesn't absorb X-rays, even though it's toxic.
- Beryllium has low X-ray absorbance because it is a very small atom. In technical terms, it has low **contact** (see below).
- A few practical comments on the anode (filament).
 - Lifetime of ~ 2000 h (at ~ 50 kV and 200 mA).
 - Reducing the tube current to 100 mA increases the lifetime by 50%.
 - Are there drawbacks to doing this?? Less intense X-rays, perhaps? Ask in OH.
 - The lifetime of the cathode depends on the number of times the X-ray source is turned on and off.
 - Turning the XRD system on and off stresses the filament as well as other components.
 - It is better to keep the XRD system on and in stand-by settings (power of ~ 20 kV and current of ~ 10 mA).

- The spectrum of X-rays generated in an X-ray tube.

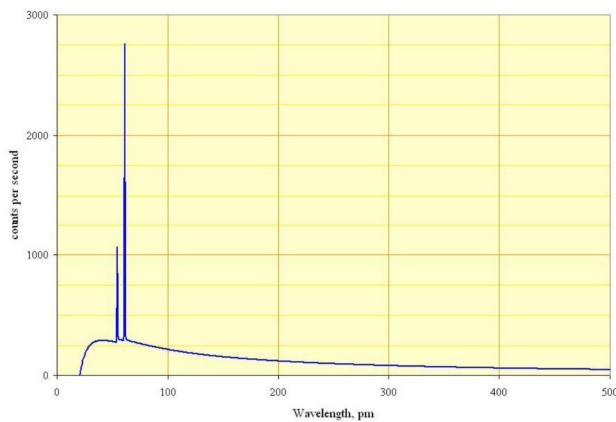


Figure 1.6: The spectrum of X-rays generated in an X-ray tube.

- The spectrum is a smooth, continuous curve with spikes.
- Implies two types of radiation: **Bremsstrahlung radiation** and **Characteristic radiation**.
- You want to get rid of the Bremsstrahlung radiation.
 - The ultimate goal is to get your X-ray tube to emit only a single wavelength at high intensity, and the natural choice is the highest characteristic peak. Thus, everything else must go.
- Bremsstrahlung radiation:** Electromagnetic radiation generated via deceleration of a charged particle when deflected by another charged particles (e.g., electrons by protons). *Also known as braking radiation, continuous X-radiation.* *Etymology* **bremsstrahlung** from German *bremsen* “to brake” and *strahlung* “radiation.”

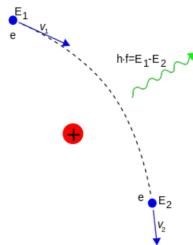


Figure 1.7: Generation mechanism for bremsstrahlung radiation.

- The charged particle loses kinetic energy via generation of photons (law of conservation of energy).
 - How is this “inelastic scattering??” What does that even mean? Ask in OH.
- If you work with high-Z (proton number) elements, these nuclei generate a rather strong electric fields.
- Strong electric fields scatter the electrons.
- The spectrum of these X-rays is continuous (hence the moniker “continuous X-rays”).
- The frequency of bremsstrahlung radiation is limited by the energy of incident electrons.
- Peak intensity shifts toward higher frequencies when the energy of the decelerated particle (particle to be decelerated) increases.
- Peak intensity increases when the energy of the decelerated particle increases.
 - Rationalizing these last two statements?? Ask in OH. Related to current and voltage below??

- X-ray notation for atomic energy levels.

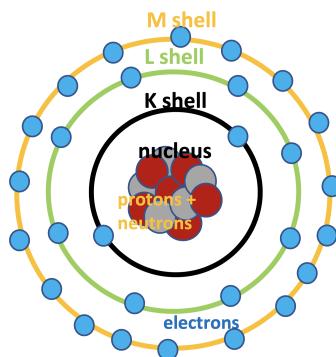


Figure 1.8: Energy levels in the Bohr model.

- The K, M, and L shells correspond to energy levels $n = 1, 2, 3$, respectively.
- Somewhat antiquated notation from the early days of X-ray invention when the Bohr model was still popular.
- Contribution of Bremsstrahlung.

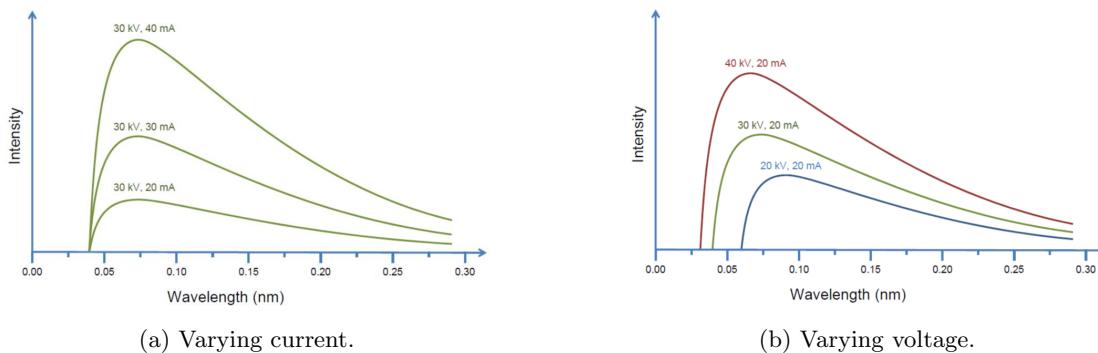


Figure 1.9: X-ray intensity vs. wavelength at different voltages and currents.

- The distribution of intensity (photon counts) against the wavelengths of the emitted radiation.
- Two settings you can change during the measurement: Current and voltage.
 - When you increase the current, the intensity increases (more electrons means more counts per second).
 - When you increase the voltage, the intensity increases and the peak shifts to higher frequencies (higher energy electrons means higher energy [resp. frequency] radiation).
- See **Kramer's law** and its consequence the **Duane-Hunt law** for more on the functional form of these curves. Note that these curves are different from the Planck blackbody distribution law, even though they look somewhat similar.
- **Kramer's law:** The law describing the spectral distribution of X-rays produced by electrons hitting a solid target. *Given by*

$$I(\lambda)d\lambda = K \left(\frac{\lambda}{\lambda_{\min}} - 1 \right) \cdot \frac{1}{\lambda^2} d\lambda$$

- The proportionality constant K is **contact**.
- λ_{\min} is the x -intercept in Figure 1.9.

- **Contact:** A quantity proportional to the atomic number of the target element. *Denoted by K .*
- **Duane-Hunt law:** The relationship between the voltage V applied to an X-ray tube and the maximum frequency f of the x -rays emitted from the target. *Given by*

$$f = \frac{Ve}{h}$$

– Note: $e = 1.602 \times 10^{-19} \text{ C}$ is the electron charge and $h = 6.626 \times 10^{-34} \text{ Js}$ is Planck's constant.

- **Characteristic emission:** The emission of quantized photons with energy corresponding to the energy difference between higher and lower states in the target atoms.
- Mechanism of characteristic emission.

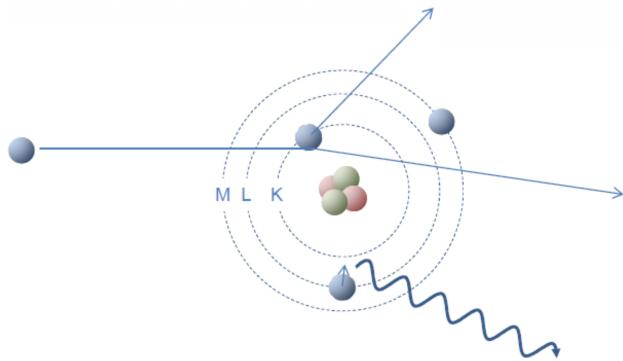


Figure 1.10: Generation mechanism for characteristic radiation.

- The target element is bombarded with high energy electrons.
- Incident electrons can knock orbital electrons out of the inner shell of the target atom.
- When this happens, the atom is left with a **core hole**.
- Outer shell electrons fill the vacancy, losing energy in the form of X-radiation to do so.
- Results in emissions characteristic of the target element (each element has unique energy levels).
- The maximum energy of the generated X-ray photon is limited by the energy of the incident electron, which is equal to the voltage on the tube times the electron charge.
- **Core hole:** A vacant energy level in the core electron shell.
- We often use copper as the target element.

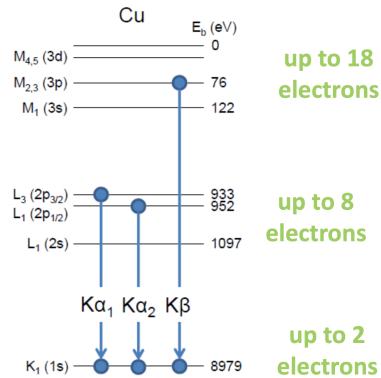


Figure 1.11: Copper spectral lines.

- The copper K_α line has greater intensity than the K_β one and thus is more desirable in diffraction experiments.
- The same holds true in many elements, including the occasional alternative molybdenum.
- Achieving monochromatic emission.
 - The X-rays are released from the anode with a range of energies.
 - There are continuous X-rays produced by the X-ray tube, K_β rays, and fluorescent X-rays from the sample.
 - Applications require monochromatic X-rays, so all of these (except the desired K_α ray) need to be filtered out.
 - X-rays of undesirable energies can be filtered using **filters** and **monochromators**.
 - You need filters to make the spectrum of X-rays really narrow??
- X-ray filters.

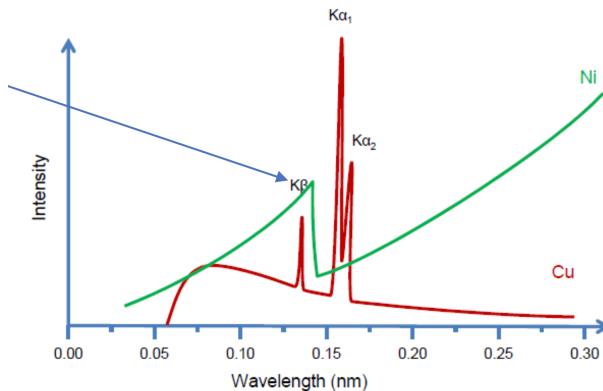


Figure 1.12: Nickel filter for a copper anode.

- The filter is usually made of a metal that has one proton less than the anode material (e.g., Ni filter for Cu anode and Nb filter for Mo anode).
- Main purpose: To filter the K_β line.
- The filtering of K_β line depends on the thickness of the filter material.
- Example: The absorption edge of nickel metal is at 1.488 Å. This is between the K_{α_1} ($\lambda = 1.542 \text{ \AA}$) and K_β ($\lambda = 1.392 \text{ \AA}$) X-ray spectral lines of copper. Hence, nickel foil of an appropriate thickness can be used to reduce the intensity of the Cu K_β line.
 - In particular, nickel foil of thickness 20 μm is typically used. It reduces the K_β intensity by 99% and the K_α intensity by only 58%.
- X-ray targets.
 - Most popular anode is Cu, but can be Co as well.
 - Fe, Cr, and Mn fluoresce under the incident Cu K_α beam, resulting in polychromatic radiation and alteration of the XRD results (strange shapes and elevated background).
 - Implication: Be careful if you see something unusual in your spectra. We may not have won a Nobel prize for something novel; it could just be Fe, Cr, or Mn contamination of the sample.
 - In this case, seek out an instrument that doesn't use Cu or Co as the target element.
 - Or is it that we want to switch from Cu to Co in this case??

- Monochromators.
 - Main purpose: To remove all X-rays but the K_{α} ones.
 - The monochromator is a crystal.
 - The monochromator works by reflecting wavelengths that obey Bragg's Law (see next lecture and Labalme (2022)) for the particular d spacings of the monochromator.
 - Desired characteristics.
 - The crystals should have suitable interplanar distances d so that the desired wavelength λ can be obtained.
 - The crystals must be mechanically strong and stable in the beam. This is why graphite and silicon are so popular.
 - Low absorbance of X-rays (we are interested in reflection).
 - Easy to grow single crystals at a reasonable cost.
 - Examples: pyrolytic graphite crystals, Si, Ge, LiF, and LiCl.
 - PG crystals and Si are particularly popular choices.
 - Graphite is a broad band monochromator: The variance around the allowed λ is relatively large due to its **mosaicity**.
 - In contrast, silicon is a narrow band monochromator.
 - With copper, the Bragg equation gives a very small difference in Bragg angle between K_{α_1} and K_{α_2} . Thus, we need a narrow band monochromator such as Si.
- **Mosaicity:** The spread of crystal plane orientations.
 - Essentially, less perfect crystals have planes that aren't perfectly aligned, creating spread.
- Effect of monochromators.

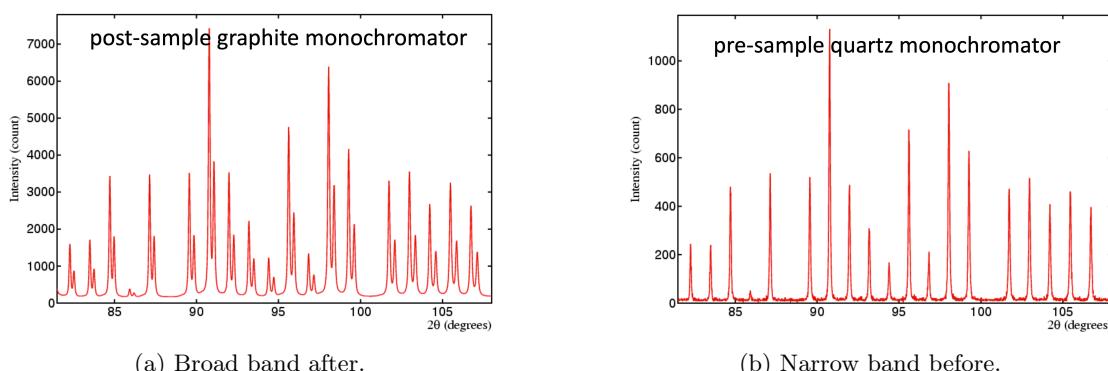


Figure 1.13: Effect of monochromator type and placement.

- Monochromators can be placed before or after the sample.
- Most instruments have both.
- The spectrum in Figure 1.13a was collected using a broad band monochromator after the sample.
 - Splitting of the diffraction peaks is due to the presence of both K_{α_1} and K_{α_2} radiation.
- The spectrum in Figure 1.13b was collected using a narrow band monochromator before the sample.
 - Benefit: The peaks due to K_{α_2} radiation are removed.
 - Drawback: Diffraction intensity is decreased.

- Working principle of the monochromator.
 - Monochromators work based on Bragg's law ($\lambda = 2d \sin \theta$) for the particular d spacings of the monochromator.
 - Let's investigate how a silicon monochromator can select for the K_{α_1} line of copper.
 - Silicon has a cubic crystalline lattice with side length 0.543 nm.
 - Thus, the spacing of the 111 planes is

$$d_{111} = \sqrt{\frac{0.543^2}{1^2 + 1^2 + 1^2}} = 0.3135 \text{ nm}$$

- Using this value and the known wavelengths of the copper K_{α_1} and K_{α_2} lines (i.e., $\lambda = 154.051 \text{ pm}$, 154.433 pm , respectively), we can use Bragg's law to calculate the angle 2θ at which we should orient the silicon 111 planes relative to the incident X-rays in order to select for one line or the other. In particular, if we want to select for the copper K_{α_1} line, we should use the angle of incidence

$$2\theta = 2 \sin^{-1} \left(\frac{\lambda}{2d} \right) = 2 \sin^{-1} \left(\frac{154.051 \text{ pm}}{2 \cdot 313.5 \text{ pm}} \right) \approx 28.442^\circ$$

- Owing to its high crystalline purity (low mosaicity), silicon at this angle can select for the desired wavelength.
- Graphite monochromators will pass both K_α wavelengths, but not K_β for which the Bragg angle is considerably different.
- Silicon has a peak position at 28.46. This number is right between the copper peaks; thus, it can separate the K_{α_1} and K_{α_2} wavelenghts from a laboratory X-ray source??
- Is this all correct??

- Filters vs. monochromators.

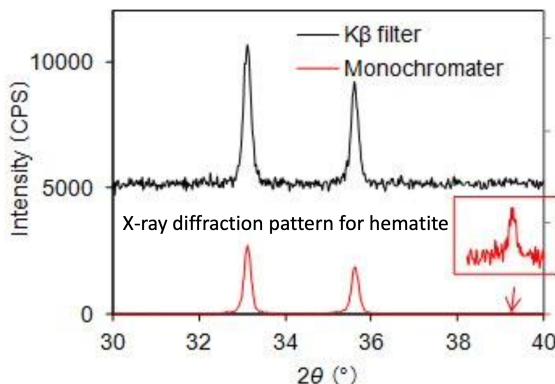


Figure 1.14: Filters vs. monochromators.

- Metal filter vs. a crystal monochromator.
- By using a monochromator, it is possible to obtain X-ray diffraction patterns with high signal-to-background ratios since the monochromator can remove interfering components such as continuous X-rays produced by the X-ray tube, K_β rays, and fluorescent X-rays from the sample.
- Notice the extra noise in the spectrum taken with a K_β filter in Figure 1.14.
- One nice thing about using a filter, though, is that you preserve more intensity (monochromators rely on relatively low-probability reflection events as opposed to high probability transmission events).

- Beam divergence.

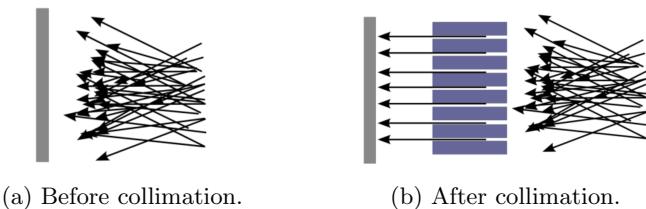


Figure 1.15: Collimator function.

- A generated X-ray beam is far from perfect; the X-rays go in many directions.
 - Solution: Use a collimator.
 - Typically lead, but can be tungsten, molybdenum, tin, bismuth, high density plastics, etc.
 - Lead is preferred because of its high density and low cost.
 - Mechanism: Most X-rays get absorbed; the right ones (straight forward ones) pass through.
 - How a diffractometer is made.

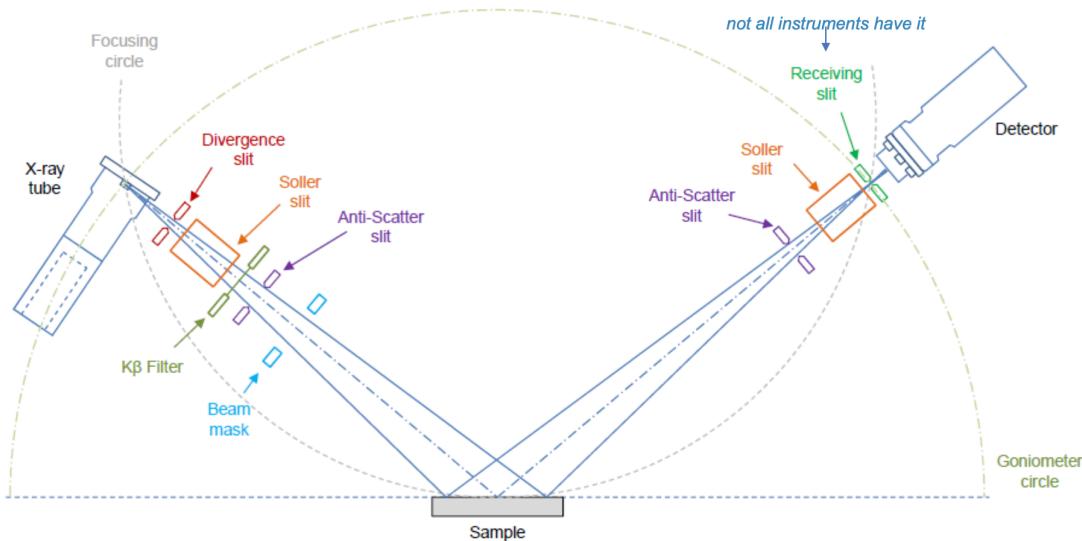


Figure 1.16: Typical diffractometer configuration.

- X-rays are generated and emitted by an X-ray source.
 - Next, they pass through a **divergence slit**, which limits the total irradiation area of the sample.
 - At higher 2θ angles, less area is irradiated, which decreases diffraction intensity. The depth of penetration of the beam becomes commensurably deeper with higher angles.
 - Then its into the sample chamber.
 - After that, they pass through **scatter slits**, which address the scattering because of “too thick” samples, rough samples, scattering from the substrate or material matrix, etc.
 - Another round of slits is the **Soller slits**.
 - The final “filter” is a monochromator, which removes 75% of the unwanted wavelengths.
 - Finally, the X-rays arrive at the detector.

- There is another diffractometer schematic and a labeled picture of an actual one in the slides.
- Examples of slits and masks.
 - **Divergence slits, beam masks, programmable divergence slits**, and Soller slits.
- **Divergence slit:** A piece of metal with an opening of a certain width.
- Fixed vs. variable divergence slits.

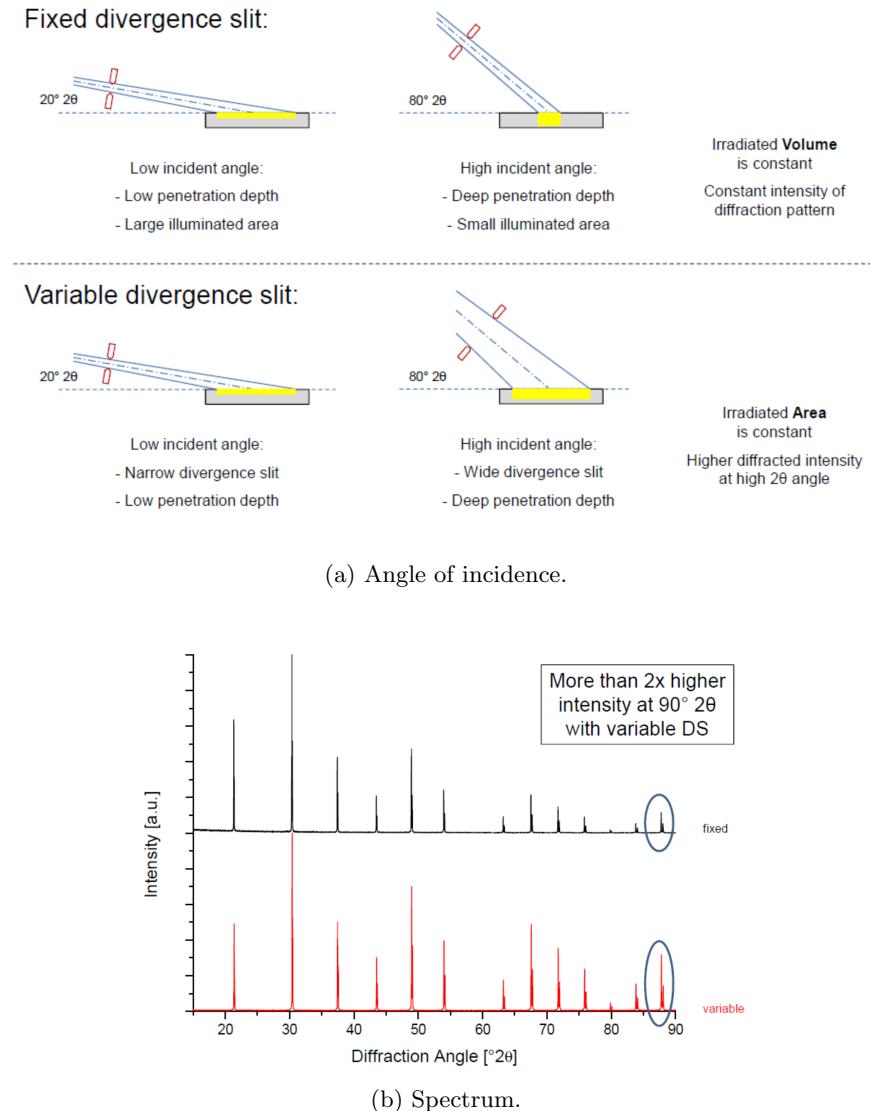


Figure 1.17: Fixed vs. variable divergence slits.

- Fixed divergence slit: As the angle changes, the amount of surface area illuminated decreases.
- With a variable divergence slit, we can solve this issue.
- Signals are the same at low 2θ ; signals differ in intensity at high 2θ .
- **Beam mask:** Similar to a divergence slit.
- **Programmable divergence slit:** A divergence slit with slots for a beam mask, Soller slits, and attenuation foil or a β -filter.

- **Soller slits:** A stack of metal films that enable removal of $\sim 99\%$ of the unwanted wavelengths of radiation from the beam. *Also known as receiving slits, RSm slits.*
 - These are a part of the instrument and you cannot touch them.
 - Not all instruments have these.
- Most common targets for X-ray tubes.
 - Usually copper, especially for crystallography.
 - Medical X-ray tubes use tungsten.
 - Softer X-rays are needed for mammography, so molybdenum is used.
- Detectors.
 - For lab experiments, the characteristic emission lines of metal targets such as Cr, Cu, Co, and Mo are used in X-ray tubes.
 - Therefore, the detector should detect X-ray photons with energies in the range of 5-20 keV. The background bremsstrahlung photons (an unwanted byproduct) have energies up to 55 keV.
 - There are two types of detectors.
 - (i) **Photon counting** detectors.
 - (ii) **Integrating** detectors.
- **Photon counting** (detector): A detector that, during a measurement, converts the energy of each individual photon into a charge and registers the charge package. *Also known as digital.*
- **Integrating** (detector): A detector that integrates (or “adds up”) the charge that is generated due to the conversion of the photon’s energy into electric charge. *Also known as analog.*
 - The energy information of the photons that are detected is lost and cannot be recovered.
- Types of the detectors.
 - **Point, linear, and area** detectors.
 - These are 0D, 1D, and 2D, respectively.
 - What are these??
 - Typical brands listed.
- **Point** (detector): A detector in which the receiving slit determines the active height.
- **Linear** (detector): A detector comprised of a linear array of solid state detectors.
- **Area** (detector): A detector comprised of a 2D array of solid state detectors.
- What regimes are used in different instruments?
 - Powder sample?
 - The scanner moves along a line.
 - Other important takeaways here??
- Examples of detectors.
 - Bruker uses a 1D detector and a Ni filter.
 - Certain Panalytical machines also use a 1D detector and a Ni filter.
 - Another type of Panalytical machine uses a 0D detector and a graphite filter.
 - We can’t change the detectors.

- Bragg-Brentano geometry.

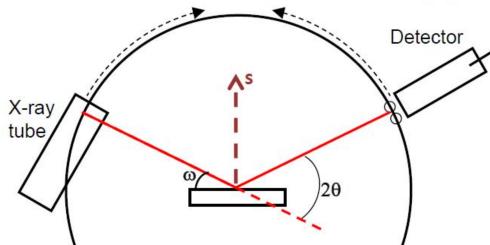


Figure 1.18: Bragg-Brentano geometry.

- Many diffractometers use this parafocusing geometry.
 - Most common; can be used in all types of analysis.
 - Efficiency is not that great for thin films, though.
- The incident- and diffracted-beam slits move on a circle that is centered at the sample. X-rays from the source hit the sample at different points on its surface. During the diffraction process, the X-rays are refocused on the detector slit.
- The incident angle ω between X-ray source and the sample is always half of the detector angle 2θ .
- Two types of setups.
 1. Fixed X-ray tube: The sample rotates at θ/min and the detector is always at $2\theta/\text{min}$.
 2. Fixed sample: The tube rotates at the same rate as the detector at θ/min .
- Sample surface is kept on the tangent plane of the **focusing circle**.
- **Focusing circle:** The circle lying in the unique plane containing the sample, X-ray source, and receiving slit.
- Grazing incidence XRD.
 - Helps with thin films.
 - Other takeaways??
- The X-ray spectrum.

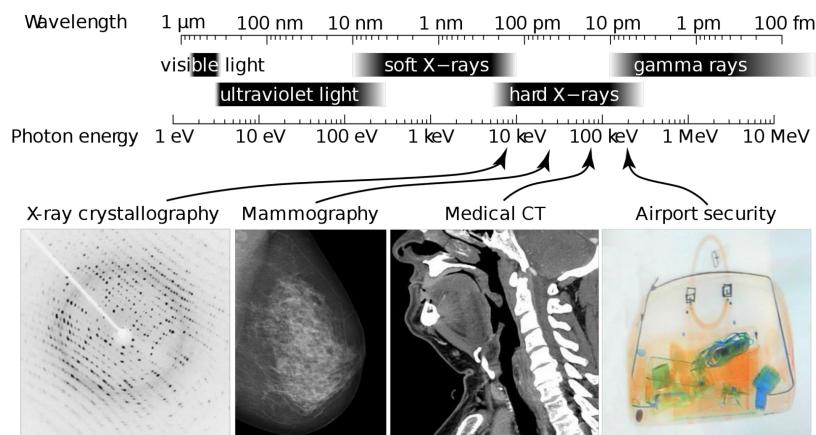


Figure 1.19: X-ray spectrum.

- Airport (1000 keV), medical (100 keV), mammography (10 keV), X-ray crystallography (8 keV).

- Applications of X-ray diffraction.
 - Same slide as at the beginning.
 - Shevchenko is really good at circling back to things over and over again so they stick!
- Diffraction basics.
 - X-rays interact with matter, resulting in **absorption**, **elastic scattering**, and **inelastic scattering**.
- **Absorption:** The energy of the incident photon has to be larger than the **ionization energy threshold** of an atom or molecule, or the **work function** of a metal.
 - Results in the photoelectric effect, which in turn leads to **fluorescence**.
 - Extremely beneficial for elemental analysis.
 - Absorption is not something we want to see in XRD.
 - See earlier discussion of fluorescence.
- **Fluorescence:** The two step process consisting of (1) absorption of radiation by an atom and ionization and (2) relaxation and emission of characteristic radiation.

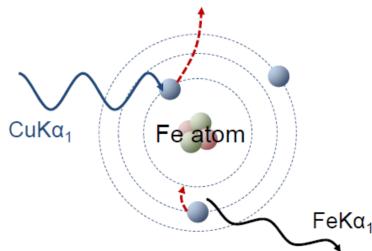


Figure 1.20: Absorption and fluorescence.

- **Elastic scattering:** Only the *direction* of the scattered photon changes, without changing its initial energy. *Also known as Thomson scattering.*
 - The scattering of optical light from electrons was first observed in 1906 by J. J. Thomson.
 - X-rays are scattered at the electrons of the atomic shell.
 - X-ray diffraction is a special case of elastic scattering.
- **Inelastic scattering:** The interaction of X-ray photons with a charged particle (electron) resulting in a decrease in energy of the photon. *Also known as Compton scattering.*
 - First observed by Arthur Compton in 1923.
 - The frequency shift has been explained by the momentum transfer of light quanta to the electrons of the material. This provided direct evidence for the quantum nature of light.
 - Initial energy of the photon is *changing*.
- Why X-rays?
 - X-ray scattering experiments have played a central role in nearly all areas of physics, chemistry, and material science to study microscopic structure and the state of matter.
 - We use X-rays because the interatomic distances in crystals are between 0.15-0.4 nm and X-ray wavelengths span this range. Thus, the match is perfect for structure determination.
 - Indeed, interference phenomena are possible only for features of about λ .

- The first use of X-rays to study crystal structure.
 - Reported by Max von Laue (Nobel Prize 1914).
 - Laue considered crystals in terms of a 3D network of rows of atoms. His analysis is based on the notion that crystals are like 3D diffraction gratings.
 - Laue utilized white X-radiation (broad spectrum, unfiltered).
 - The diffraction spots that surrounded the central spot of the primary beam could be explained as an interference pattern due to the crystal's space lattice. Each spot was caused by X-rays that corresponded to a certain lattice constant and wavelength.
- Bragg's law.
 - Obtained as a result of experiments by William Lawrence Bragg and his father, Sir William Henry Bragg in 1912.
 - Nobel Prize (1915) for their work determining the crystal structures of NaCl, ZnS, and diamond, and for "their services in the analysis of crystal structure by means of X-rays."
 - The Bragg spectrometer had an X-ray tube, filters, collimators, a crystal sample, and a detector.
- The principle of Bragg's law.
 - We'll start here next class.
- Write out ppt. notes ahead of time next time!

Week 2

???

2.1 XRD Analysis 2

1/10: • Dealing with broadening of the XRD beam.

- Synchrotron radiation gets you better resolution.

• How do you get a cleaner spectra, given this one?

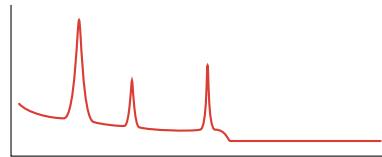


Figure 2.1: Enhancing XRD results.

– In general:

- You can change divergence slits, filter, and masks.
- You can't change the monochromator, however, because it's built into the machine.

– In this particular case:

- The issue is the high background which may be covering up smaller peaks.
- Solution: Use a smaller divergence slit and more masks.
- Another problem that could be causing high background is X-ray fluorescence, so check to make sure that your sample doesn't have iron or other similar things contaminating it.

• The principle of Bragg's law.

- The Braggs proposed that crystals can be described in terms of layers or planes of atoms.
- Their theoretical planes behave like **reflecting planes**.
- Strong “reflected” beams are produced when the path differences between reflections from successive planes in a family is equal to a whole number of wavelengths.
 - If we want to see something, neighboring planes' waves must be in phase, interfering constructively to amplify their intensity rather than dampening it with destructive interference.
- This approach is not correct in a physical sense — planes do not reflect X-rays. However, it is correct in a geometrical sense and provides us with a very simple expression for the analysis of crystal structure.

• **Reflecting plane:** A plane for which the angle of incidence equals the angle of reflection.

- Conditions that are necessary to make the phases of the beams coincide.
 - Refer to Figure 2.2 throughout the following.
 - The angles of the incident and “reflected” photons are equal.
 - The rays of the incident are always in phase and are parallel up to the point at which the top beam reaches the top layer at atom O .
 - The second beam continues to the next layer where it is scattered at atom B . If the two beams travel in adjacent and parallel fashion, the beam scattered at atom B travels an extra distance $AB + BC$. This extra distance should be equal to a whole number of wavelengths.
 - Again, a diffracted beam *looks* reflected, but what it really is is scattered radiation. Drill home that planes are not physically accurate!
- How to derive the Bragg's Law formula.

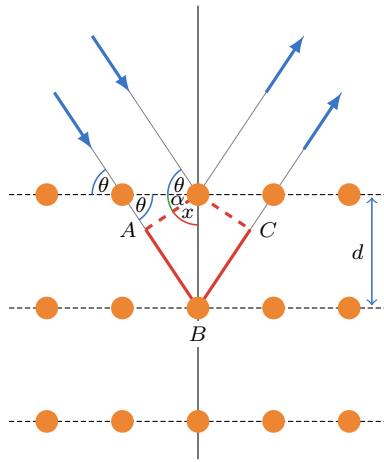


Figure 2.2: Bragg's Law derivation.

- Since $\theta + \alpha = 90^\circ$ and $\theta + x = 90^\circ$, we know that

$$x = \theta$$

- This combined with the observation that $\sin(x) = AB/d$ implies that

$$\begin{aligned} \sin \theta &= \frac{AB}{d} \\ AB &= d \sin \theta \end{aligned}$$

- Lastly, we may observe that $AB = BC$. Therefore, the total phase shift is

$$2AB = 2d \sin \theta$$

- Since we require that this is a whole number of wavelengths, our final condition is

$$n\lambda = 2d \sin \theta$$

where n is an integer determined by the order given, λ is the wavelength of the X-rays, d is the spacing between the planes in the atomic lattice, and θ is the angle between the incident ray and the scattering planes.

- This condition is called **Bragg's Law**.
- Note that copper's $\lambda = 1.54 \text{ \AA}$ is the most common wavelength with which to work.

- Mineralogy — Inspiration for crystallography.
 - Happened way before all of this math, when scientists had far fewer tools.
 - Researchers could only observe a crystal's **habit** and cleavage planes, and measure interfacial angles with a **goniometer**.
 - René Just Haüy postulates in 1801: Crystal structures are made up of orderly arrangements of integrant molecules in successive layers, according to geometrical laws of crystallization.
 - Haüy formulated the **theory of the rational indices** of the faces of a crystal, which is important for crystallographic calculations.
 - By 1792, he had identified several parallelepipeds to explain shapes of a few crystals.
 - This work, which may now seem elementary, is extremely impressive since he had so few tools.
- **Habit:** The tendency for specimens of a mineral to repeatedly grow into characteristic shapes.
- **Goniometer:** An instrument that either measures an angle or allows an object to be rotated to a precise angular position.
- **Theory of rational indices:** The theory that the intercepts of a crystal face with the crystallographic axes can be expressed as a/h , b/k , and c/l , where $1/h$, $1/k$, and $1/l$ are three simple rational numbers.
- Type of Lattice Systems.
 - Haüy (1784): The periodicity of crystalline materials involves the basic repetition of a basic unit called the **unit cell**.
 - Crystalline materials are formed by the repetition in (2D, 3D, etc.) space of cells (or **crystallites**).
 - In 3D space, cells are defined by three non-coplanar vectors (called **fundamental translations**).
 - There are 7 types of cells that together cover all possible point lattices.
 - Important: Crystal structures are defined by a **basis** and a **lattice**.
- **Basis:** *What* gets repeated in the crystal structure.
- **Lattice:** *How* it gets repeated.
- Bravais lattices.

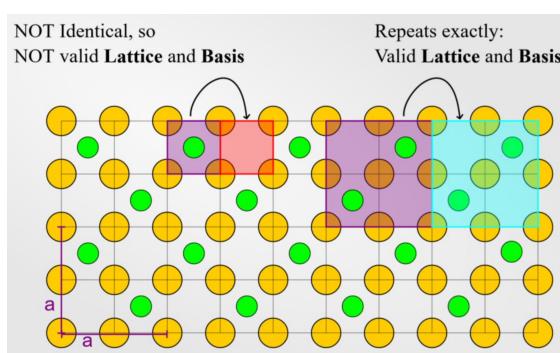


Figure 2.3: Decomposition of a crystal into its Bravais lattice.

- Auguste Bravais (1848) found mathematically that the number of crystalline lattices is finite.
- It is a fully geometrical concept and has nothing to do with atoms or crystalline planes.
- Bravais lattices are the basic lattice arrangements. All other lattices can simplify into one of the Bravais lattices. Bravais lattices move a *specific* basis by a translation.

- There are 14 3D Bravais lattices.
 - Bravais lattices only take into account *translational* symmetry (this is important!).
 - If you can exactly repeat the entire structure by a set of translations, that is the Bravais lattice.
 - Other symmetries, like reflection or inversion, are captured in point groups and space groups, not by Bravais lattices.
- 1D Bravais lattice.

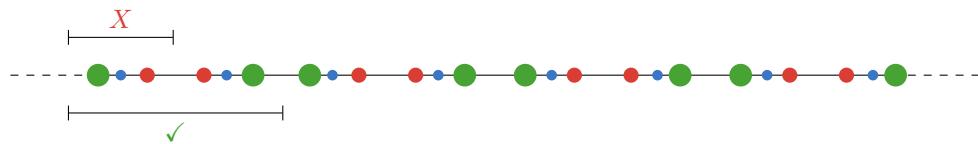


Figure 2.4: 1D Bravais lattices.

- Only one vector, hence only one possible Bravais lattice.
 - Bravais lattices do not allow mirror symmetry, only translation. Thus, we must choose as our basis the smallest structure that repeats *translationally*.
- 2D Bravais lattice.

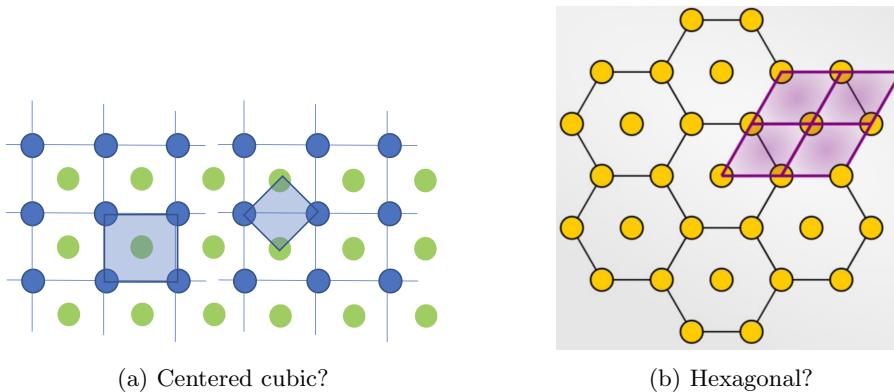


Figure 2.5: 2D Bravais lattices.

- There are five 2D Bravais lattices.
 1. Square ($a = b, \theta = 90^\circ$).
 2. Hexagonal ($a = b, \theta = 120^\circ$).
 3. Rectangular ($a \neq b, \theta = 90^\circ$).
 4. Centered rectangular (see below).
 5. Rhomboidal ($a \neq b, \theta \neq 90^\circ$).
- There does exist centered rectangular (a rectangular lattice with an additional vertex in the center of each rectangle), but there does not exist “centered cubic” because a smaller, rotated square can represent the entire lattice, so “centered cubic” is really just square. See Figure 2.5a.
- The “hexagonal” Bravais lattice can be simplified into a rhombus, but hexagon shows “true” symmetry (i.e., rotation, inversion, etc.). See Figure 2.5b.
- The honeycomb is not a Bravais lattice. Why??

■ Ask in OH.

- 3D Bravais lattice.

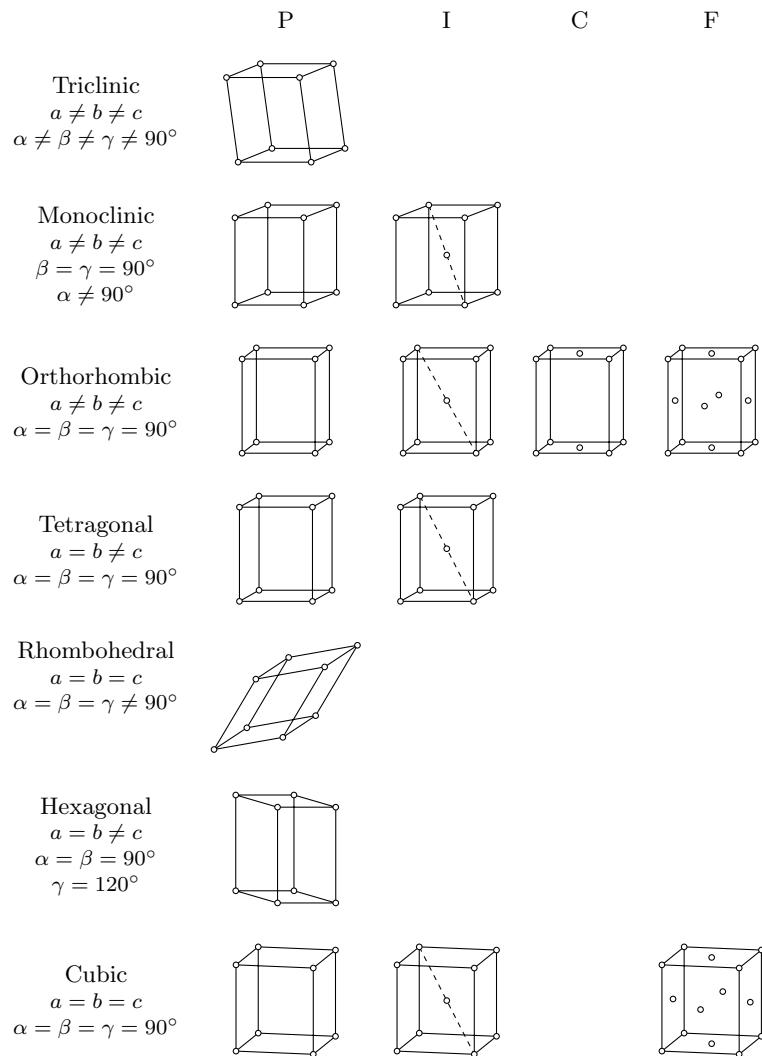


Table 2.1: Bravais lattices.

- Each lattice is a polyhedron.
 - The polyhedrons can be described using three different vectors.
- Some of the Bravais lattices can be expressed by other simple lattices: In 3D, the FCC lattice is also described by a rhombohedral lattice.
- There is no base-centered cubic Bravais lattice because what might be that is actually simple tetragonal.
- There are 4 types of Bravais lattices.
 - P - Primitive.
 - I - Body centered.
 - C - Base-centered.
 - F - Face centered.
 - More on these and how they correspond to space groups later.

- **Unit cell:** The smallest group of atoms which has the overall symmetry of a crystal.
 - Can be used to build the entire lattice by repetition in three dimensions.
 - A 3D structure.
- **Primitive cell:** The smallest possible element of a lattice.
 - May or may not include all symmetry elements.
 - Can have 2D or 3D structure.
- Conventional and primitive cells.

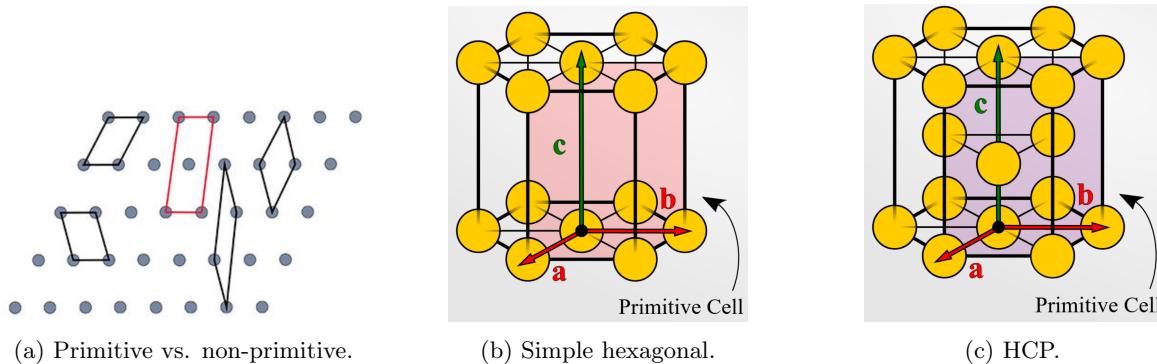


Figure 2.6: Conventional vs. primitive cells.

- The hexagonal 2D Bravais lattice (the conventional cell) might also be described as rhombic (the primitive cell). See Figure 2.5b.
- The hexagonal 3D Bravais lattice is a hexagonal prism that can also be constructed from a primitive cell which is a parallelepiped.
- Know simple hexagonal vs. hexagonal close-packed (HCP). See Figures 2.6b-2.6c.
- Clarification on primitive vs. non-primitive?? See Figure 2.6a.
- What is a lattice point??
- Miller indices.
 - In 1839, the British mineralogist William H. Miller used *reciprocal quantities* — namely, the integer numbers h, k, l — to describe crystal faces.
 - Miller indices are used to specify directions and planes.
 - These directions and planes could be in lattices or in crystals.
 - Notation (important information!):
 - (h, k, l) represents a **point** (commas are used).
 - $[hkl]$ represents a **direction**.
 - $< hkl >$ represents a **family of directions**.
 - (hkl) represents a **plane**.
 - $\{hkl\}$ represents a **family of planes**.
 - Be careful when writing/reading research literature to use/interpret the write notation.
 - Negative numbers and directions are depicted with a bar on top of the number.
- Miller indices for directions.
 - Let's consider a 2D lattice with Miller indices $(4, -2)$.

- Defines a vector pointing in the direction $4\vec{a} - 2\vec{b}$. It is parallel to many other vectors.
- The index $(4, -2)$ [notational issue here??] represents the set of all such parallel vectors.
- The number of indices matches the dimension of lattice (e.g., 1D lattice has 1 Miller index, 2D lattice has 2 Miller indices, etc.).
- Fractions in $(r_1 r_2 r_3)$ are eliminated by multiplying all components by their common denominator. Example: $(1, 3/4, 1/2)$ will be expressed as $(4, 3, 2)$.
- Miller indices: hkl review.
 - See Labalme (2022) for more.
 - You just have to remember that Miller indices represent the reciprocals of the fractional intercepts which the plane makes with crystallographic axes.
 - Notice how (421) [i.e., parentheses] is used to denote a plane!
- How to find Miller indices for planes.
 - Great slide in the slideshow; one stop shop for Miller indices.
 - The planes we will most commonly study are (100) , (001) , and (010) planes.
 - Algorithm.
 - Identity the plane intercepts on the x -, y -, and z -axes.
 - Define intercepts in fractional coordinates.
 - Take the reciprocals of the fractional intercepts.
- Miller indices.
 - Continuation of the previous slide but for slanted planes.
 - Keep in mind that different planes have different chemical distributions. One plane in an oxide may be mostly oxygen; another may be mostly copper.
- Crystal structure, lattice, etc.
 - Crystal structure combines **lattice** with the **basis** again.
 - A lattice is not a crystal. However, if the basis consists of one atom, crystal structures look exactly like the Bravais lattice.
 - Common metallic crystal structures: BCC, FCC, hexagonal close-packed (HCP).
- Example 1: Diamond.

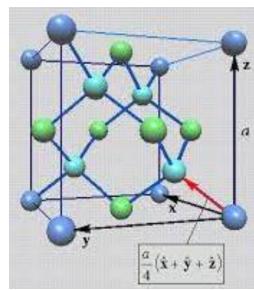


Figure 2.7: Diamond crystal structure.

- Bravais lattice: FCC with a two-atom basis.
- Crystal structure: Cubic diamond.
- Two atom basis at $(0, 0, 0)$ and $(1/4, 1/4, 1/4)$.

- Example 2: NaCl.
 - Bravais lattice: FCC.
 - Crystal structure: FCC.
 - Both atoms make FCC lattices and you get the overall structure by inserting one lattice into the other.
- Example 3: Primitive cubic substances.
 - Examples: Fe, CsCl, and NiAl.
 - Bravais lattice: Primitive cubic with a two-atom basis (Cs at $(0, 0, 0)$ and Cl at $(1/2, 1/2, 1/2)$).
 - Crystal structure: Primitive.
- From Bravais lattices to a full description of crystalline structure.
 - Bravais lattices: There are 14 and they account for translational symmetry. But there are also additional symmetry elements (rotation, inversion, reflection).
 - We discount additional potential translation symmetry operations for now.
 - If you apply all of these other operations to the Bravais lattices, you get 32 crystal classes/point groups.
 - If you add in **screw** and **glide** operations, you get 230 total space groups. This number does depend on the dimension of the space, though, i.e., fewer space groups exist in 2D (to a significant extent).
- **Screw:** Rotation followed by a translation.
- **Glide:** Reflection followed by a translation.
- Screw and glide operations.

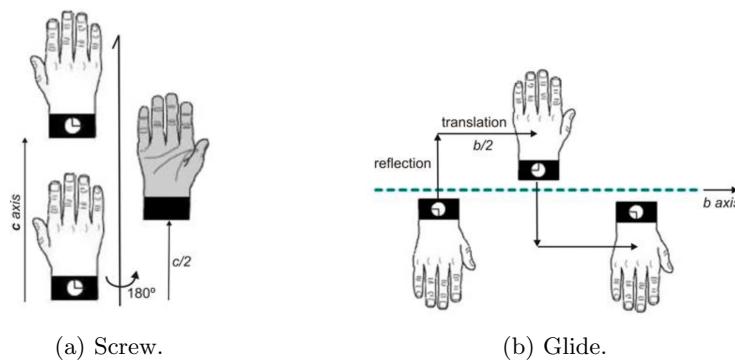


Figure 2.8: Screw and glide operations.

- These are essentially just combinations of the rotation axes and the mirror planes with the characteristic translations of the crystals.
- Discovery of the space groups.
 - Retrat de Arthur Schoenflies (Germany) and Evgraf Fedorov (Russia) proposed space groups while in correspondence via mail.
 - The triumph of their studies was only after the discovery of the utilization of X-rays in structural studies of minerals.
 - Developed 1890-92.

- Symmetry operator notation (Hermann-Mauguin).

Type of symmetry element	Written symbol	Graphical symbol	
Centre of symmetry	$\bar{1}$		
Mirror plane	m	Perpendicular to paper	In plane of paper
Glide planes	$a\ b\ c$	glide in plane of paper	arrow shows glide direction
		glide out of plane of paper	
	n	
Rotation	2 3 4 6	● ▲ ◆ ◆	→
Screw axes	2_1 $3_1, 3_2$ $4_1, 4_2, 4_3$ $6_1, 6_2, 6_3, 6_4, 6_5$	● ▲▲ ◆◆◆ ◆◆◆◆◆◆	→
Inversion axes	$\bar{3}$ $\bar{4}$ $\bar{6}$	△ ◊ ◊	

Table 2.2: Hermann-Mauguin notation.

- Shevchenko won't go into too much depth; she doesn't even remember that much herself, but it's good to know how to read one.
- How to "read" space groups.
 - In the notation for a space group, the first letter is the Bravais lattice and then there are three symmetry elements with respect to 3 viewing directions.
 - Example: NiAsS is orthorhombic with space group $Pca2_1$.
 - P refers to the Bravais lattice.
 - c refers to a glide plane $c \perp a$.
 - a refers to a glide plane $a \perp b$. What do these mean??
 - 2_1 refers to a screw axis parallel to c .
 - In screw axis notation, "the big number is how many stops you make???"
 - Can 3 symmetry elements describe the full symmetry?
 - There are structures with more symmetry elements (e.g., 8, 16, etc.).
 - Their symmetries can be derived from generators, though. So yes??
- **Space group:** The symmetry group of an object in space.
 - Alternate definition: A set of symmetry elements and respective operations that completely describes the spatial arrangements of a given 3D periodic system.
 - In crystals, space is three dimensional.

- Viewing directions.
 - The position of the symmetry element depends on the type of lattice.
 - If it is triclinic, the symmetry elements are always “around” the center of inversion.
 - If it’s monoclinic, there is one viewing direction besides the inversion center (mathematicians have arbitrarily chosen b to be said direction).
 - Shevchenko doesn’t think many people memorize this unless they’re really into it, but it’s worth understanding once. How much do we need to know??
 - Consider building models.
- XRD databases.
 - This is probably the most important/relevant information in this lecture for our research.
 - In the 1940s, the best crystallographers in the world analyzed a bunch of materials and started to build a database.
 - Started by Hanawalt and associates while he was at Dow chemicals. They built a database and used it for chemical analogies.
 - The principal of the analysis is based on the d spacings of the strongest reflections.
 - \$50 per set was very expensive at the time, but worth it because it saved so much work.
 - In 1941, the JCPDS (Joint Committee on Powder Diffraction Standards) was founded.
 - 1978: Became the ICDD (International Center for Diffraction Data).
 - Still have a ton of scientists working on diffraction analysis (around 300 in 1978).
- Data analysis.
 - The slides list databases that contain powder diffraction data (line positions and their intensities).
 - We have access as UChicago students.
 - Different databases have different specialties.
 - Programs for search and match are available, too.
 - Website for visualization, coordinates, and finding primitive and basis vectors ([link](#))
- XRD analysis can help you with...
 - Phase identification;
 - Crystallite size measurements.
 - Texture analysis.
 - Etc.
 - We can also study processes (this is very useful).
 - Examples:
 - Phase transitions.
 - Crystallite growth.
 - Thermal expansion.
 - Ion intercalation.
 - Decomposition.
 - Oxidation.
 - We can observe a flattening/broadening of curves. The oxidized material may be crystalline, but it may well be amorphous, too, leading to said broadening.
 - Stability of the catalysts.

- Diffractograms.

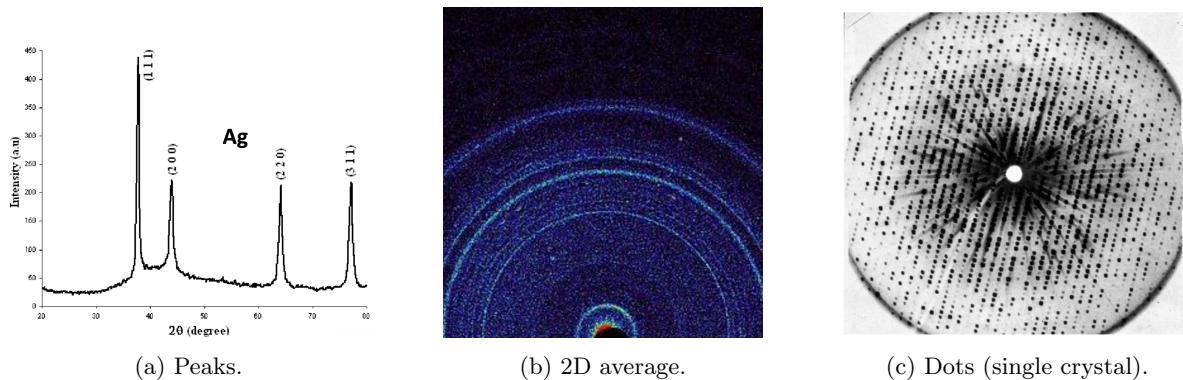


Figure 2.9: Diffractogram types.

- Our diffractograms will largely look like peaks (see Figure 2.9a).
- 2D detectors (as with synchrotrons) yield circular patterns or circular dot patterns.
- Dots (see Figure 2.9c) are generated by single crystals.
 - Myoglobin is a protein in muscles. It stores oxygen to bind and release oxygen depending on the oxygen concentrations in the cell, has functions in the hemostasis of nitric oxide and in the detoxification of reactive oxygen species. Myoglobin is the reason for the red color of the muscle of most vertebrates.
- Averaged rings (see Figure 2.9b) are generated by polycrystalline materials.
 - What is pictured is actually the first X-ray diffraction pattern of Martian soil (from the Curiosity rover at “Rocknest” on October 17, 2012).
 - **Feldspar, pyroxenes, olivine**, etc. were identified (this kind of identification and confirmation can take experts years since the data is so cluttered).
 - We read these with software in general. Calculate distances, work with the parameters of our instrument, etc.
 - 2D is better because gaps in the circles provide information, too??
 - Decades of experience and the help of your peers helps you decipher these images.
- Most of us will work with the APS at Argonne (shut down in April 2023 and will be back online a year later).
- **Feldspar:** Aluminum tectosilicate minerals, e.g., $KAlSi_3O_8$, $NaAlSi_3O_8$, and $CaAl_2Si_2O_8$.
- **Pyroxenes:** General form $XY(Si, Al)_2O_6$, where X can be Ca, Na, Fe^II, Mg, etc. and Y can be Cr, Al, Mg, Co, Mn, Sc, V, etc.
- **Olivine:** General form $(Mg^{2+}, Fe^{2+})_2SiO_4$.
- Sample preparation.
 - Methods:
 - Drop and dry: Make a suspension of nanoparticles or a colloidal solution, drop it onto the plate, and wait for it to dry.
 - Grinding: Powder gets grinded and then compacted.
 - Crystallization: Proteins get crystallized, for example.
 - Diffractogram depends on:
 - Graininess.

- Micro-absorption.
- Texture.
- Sample height displacement/adjustments.
- Surface roughness.
- Sample transparency.
- Diffractograms.
 - Peak positions are determined by the size and shape of the unit cell.
 - Peak intensities are determined by the atomic number and position of the various atoms within the unit cell.
 - Peak widths determined by instrument parameters (and other factors, discussed later).
 - Temperature, crystal size, strain, and other imperfections in the material.
- Sample's graininess.
 - Single crystals should generate "spotty diffracted rays." Powder samples should generate continuous rings.
 - Grainy samples lead to variation in the intensity of the peaks, missing peaks, etc.
 - If you work with grainy samples, grind first and adjust the divergence slits second. You can also spin the sample and/or open the divergence slits to increase the probability of something getting hit.
- Micro-absorption.
 - There are materials and elements with high and low absorption of X-rays. Your spectrum will depend on the high absorption ones??
 - In XRD, we don't care about elements, but we do care about their Z-number.
 - CsCl and CsI are interesting examples. They have the same structure but different numbers of lines in the X-ray pattern.
 - Cs and Cl are not isoelectronic but Cs and I are. X-rays are scattered by electrons, so to X-rays, these atoms look the same. Leads to systematic peak absences.
- Size-effect.
 - Thicker films have more peaks in general.
- Texture/preferred orientation.
 - Nanoparticles vs. short rods vs. long rods.
 - It looks a lot better when you have more nanoparticles aligned in particular directions.
 - At 75%, we're pretty good.
 - Same story with the types of rods.
 - There are various creative solutions found on the internet (including Vaseline, hair spray, etc.) to get the particles to line up the way you want.
 - The underlying tactic is always providing some medium in which the particles can interact/rearrange.

References

Labalme, S. (2022). *CHEM 26300 (Chemical Kinetics and Dynamics) notes*. Retrieved January 7, 2023, from
<https://github.com/shadypuck/CHEM26300Notes/blob/master/Notes/notes.pdf>