Q.no	Question	Option A	Option B	Option C	Option D	Correct option
1	The materials which allows heat or electric charge to pass through them.	Insulator	Conductor	substance	none of these	conductor
2	All the materials in solid state which allows electricity to pass through them, ( )	Electrolytic conductor	Thermal conductor	Metallic Conductor	none of these	Metallic Conductor
3	Metals are conductor because they have free electrons.	free electron	Mobile electron	delocalized electron	all of these	all of these
4	The materials which donot allows heat or electric charge to pass through them.	Insulator	Conductor	substance	none of these	Insulator
5	An aqueous or fused solution of Electrovalent salt which allows current to pass through them .	Insulator	Conductor	Thermal conductor	electrolytic conductor	electrolytic conductor
6	An electrolyte conduct due to the presence of	free electron	Mobile electron	delocalized electron	free ion	free ion
7	Free Ions are formed for molecules by the process of	Dissociation	Ionisation	only A	Both A and B	Both A and B
8	..... is a negative charged ion and .....is a positive charged ion.	Anion and Cation	Cation and Anion	electron and proton	none of these	Anion and Cation
9	Those substance which undergoes complete dissociation	Weak electrolyte	Medium electrolyte	Strong Electrolyte	none of these	Strong Electrolyte
10	which of the given is not a strong electrolyte	HCl	HNO3	CH3COOH	H2SO4	CH3COOH
11	Electrolyte: Those substance which undergoes incomplete/partial dissociation.	Weak electrolyte	Medium electrolyte	Strong Electrolyte	none of these	WeakElectrolyte
12	Which of the given is a weak electrolyte	HCl	HNO3	CH3COOH	H2SO4	CH3COOH
13	Ohms Law mathematics expression is..... where I = current in amperes ,E = Electrode potential in volts and R = Resistance in ohms	I=E.R	I=E/R	E=IR	E=R/I	I=E/R
14	the easy of flow of charge its unit is mhos.	Conductance	Resistance	Specific Conductance	Specific resistance	Conductance
15	Obstruction to the flow of Charge,its unit is ohms	Conductance	Resistance	Specific Conductance	Specific resistance	Resistance
16	The electrode made of Silver - Silver chloride is	Hydrogen electrode	calomel electrode	Glass electrode	all of these	glass electrode
17	electrode consist of Platinum (Pt) wire and pure hydrogen gas is passed through	Hydrogen electrode	calomel electrode	Glass electrode	all of these	Hydrogenelectrode
18	The electrode consisting of Mercury-Mercurous Chloride	Hydrogen electrode	calomel electrode	Glass electrode	all of these	Calomel electrode
19	Which of the given electrode is also called SHE is....	Hydrogen electrode	calomel electrode	Glass electrode	all of these	Hydrogen electrode
20	The Electeode consisting of Silver (Ag) electrode coated with Silver chloride	Hydrogen electrode	calomel electrode	Glass electrode	all of these	Glass electrode
21	This electrode is also used as internal reference electrode	Hydrogen electrode	calomel electrode	Glass electrode	all of these	Glass electrode
22	Conductometric titrationanalysis based on the measurements of change in .....of solution	Conductance	Resistance	Heat	Moisture	conductance
23	Conductance of electrolytes depends on .....free ions	Number of free ions	Mobility of free ion	Both A and B	none of these	Both A and B
24	Acid - Base titartion :Storng acid react with Strong base the nature of graph is.....	straight line	Perabola	V shaped	S shaped	V shaped
25	Conductivity of the strong electrolyte in a conductometric titration ..... first and then ..... after reaching equivalance point.	decrease first then increases	increase the. decreases	remains same	none of these	decrease and then increase
26	In weak acid and strong base titration the graph the conductance variation will give a graph of shape.....	straight line	Perabola	V shaped	S shaped	S shaped
29	Primary Battery is the one which is	Irreversible	rechargeable	only A	Both A and B	only A
30	Battery which can be recharge is a	Primary	Secondary	Fuel cell	reserve Cell	Secondary
31	Which one is incorrect statement for secondary batteries	Rechargable	reversible	Single use	none of these	single use
32	A device in which electrolyte decompose by passing electricity.	Electrochemical cell	Fuel Cell	Electrolytic Cell	none of these	Electrolytic Cell
33	A device in which electricity is generated by spontaneous redox reaction.	Electrochemical cell	Fuel Cell	Electrolytic Cell	none of these	Electrochemical cell
34	when Current from external circuit reverse the chemical reaction and restore the battery	Discharge	free Charge	Null Discharge	Charge	charge
35	when the battery provided energy to the external source	Discharge	free Charge	Null Discharge	Charge	discharge
36	Dry Cell is an example of	Primary	Secondary	Fuel cell	reserve Cell	Primary
37	Dry Cell is an example of	Electrochemical cell	Fuel Cell	Electrolytic Cell	none of these	Electrochemical cell
38	In an electrochemical cell Cathode is the electrode where .....reaction occurs	Redox reaction	Oxidation reaction	Reduction reaction	Decomposition reaction	Reduction reaction

		39	In an electrochemical cell Anode is the electrode where .....reaction occurs	Redox reaction	Oxidation reaction	Reduction reaction	Decomposition reaction	Oxidation reaction
		40	A battry made of Anode of Zinc vessel and Cathode made of graphite rod surrounded by paste of Carbon and Magnesedioxide	Nickel - Cadmium	lead Acid	Lithium ion	Dry cell	Dry Cell
		41	Electrolyte used in a Dry Cell is	20% KOH	25 % KCl	paste of ZnCl2 + NH4Cl	20 % H2SO4	paste of ZnCl2 + NH4Cl2
		42	The two electrode anode and cathode in an electrochemical cell is charged respectively as	Positive and Negative	Negative and positive	Both electode are positive	both electrode are negative	Positive and Negative
		43	Match : A- Primary cell ,B- Secondary cell , C- Fuel Cell WITH a- H2-O2 cell ,b- Dry Cell,c- Ni-Cd cell	A-a,B-b,C-c	A-b,B-a,C-c	A-c,B-b,C-a	A-b,B-c,C-a	A-b,B-c,C-a
		44	Emf of dry cell is ..... volts	1	1.5	2	2.5	1.5
		45	The battery generally used in torches, doorbell, transistors etc	primary	secondary	Fuel cell	reserve Cell	Primary
		46	Electrolyte used in a Ni-Cd Cell is	20% KOH	25 % KCl	lithium salt in organic solvent	20 % H2SO4	20% KOH
		47	Electrolyte used in a lead acid Cell is	20% KOH	25 % KCl	paste of ZnCl2 + NH4Cl2	20 % H2SO4	20 % H2SO4
		48	Electrolyte used in a Lithium ion cell Cell is	20% KOH	25 % KCl	paste of ZnCl2 + NH4Cl2	none of these	none of these
		49	which of the following Cell is not a secondary cell.	Nickel - Cadmium	lead Acid	Lithium ion	Dry cell	Dry Cell
		50	In an Nickel Cadmium c the Anode electrode is made of	Nickel Oxide	Cadmium Oxide	Lithium salt	Zinc Chloride	Cadmium Oxide
		51	In a Ni - Cd the Cathode electrode is made of ..... with Nickel	Nickel Oxide	Cadmium Oxide	Lithium salt	Zinc Chloride	Nickel Oxide
		52	The cell which is generally not used in power tools, computer power supply and also space Application	Nickel - Cadmium	lead Acid	Lithium ion	Dry cell	Dry cell
		53	Lead Acid Cell is an example of.....	primary	secondary	fuel cell	Reserve cell	secondary
		55	Emf of lead acid cell is .....volts	1	1.5	2	2.5	2
		56	The cell which converts the chemical energy of fuel directly in to electrical energy.	primary	secondary	fuel cell	Reserve cell	Fuel cell
		57	Match:A -Hydrogen Electrode ,B - Calomel Electrode ,C- Glass Elecrrode WITH a- Pt/H2 ,b- Ag/AgCl2 ,c-Hg/HgCl2	A-a ,B-c,C-b	A-a,B-b,C-c	A-c ,B-b ,C-a	none of these	A-a,B-b,C-c