QUESTION BANK FOR CLASS TEST 1

BASIC CHEMISTRY

(ALL BRANCHES)

CHAPTER 1 (CHEMICAL BONDING & CATALYSIS)

Q NO	QUESTION	OPTION A	OPTION B	OPTION C	OPTION D	CORRECT OPTION
1	Complete transfer of one or more electrons from one atom to other atom results in _____ bond	ionic bond	covalent bond	co-ordinate bond	dative bond	A
2	When single atom provides both electrons which are needed for completion of covalent bond then it leads to	ionic bond	covalent bond	co-ordinate bond	metallic bond	C
3	When metal atom looses electron it forms	positive ions	negative ions	alkalies	non metals	A
4	During formation of ammonium ion,number of electrons shared between nitrogen atom and hydrogen ion is	1	2	3	4	B
5	Which of the following molecule is an example of co-ordinate bond .	ammonia	ammonium ion	urea	nitrogen	B
6	The pair of outer shell electrons not used in bonding are called as	valence electrons	Free electrons	bonding pair of electrons	lone pair of electrons	D
7	Charge on ion depends upon gain or loss of	electrons	proton	neutrons	none of the above	A
8	The bond formed by sharing of four electrons between two bonding atoms is called as	single covalent bond	triple covalent bond	dative covalent bond	double covalent bond	D
9	The bond formed by sharing two pairs of electrons between two a bonding atoms, is called as	single covalent bond	double covalent bond	triple covalent bond	ionic bond	B
10	The movement of molecules are completelty restricted in following state of matter	solids	liquids	gasese	all of the above	A
11	During chemical bonding ,the electronic configuration of the bonding atoms is similar to	alkali metal atom	noble gas atom	non metallic atom	no change in configuration	B
12	Electrons are usually lost by	metals	non-metals	inert gases	all of the above	A
13	During formation of nitrogen molecule,number of electrons pair shared between the two bonding nitrogen atoms are	1	2	3	4	C
14	Name the type of bonding in MgO molecule	ionic bond	covalent bond	metallic bond	dative bond	A
15	Metals are good conductors due to	inner shell electrons	lone pair of electrons	localized electrons	delocalized electrons	D
16	Conduction of electricity in metallic bonding is due to the presence of	inner shell electrons	lone pair of electrons	delocalized electrons	localized electrons	C
17	In an ionic bond bonding atoms combine together so as	to get rid of excess electrons	to attain configuration of noble gases	to avoid further reaction	all of the above	B
18	Which of the following conditions are true for metal atom	lose their outer electrons	become positively charged	become negatively charged	both (a) and (b)	D
19	Nitrogen molecule is an example of	single covalent bond	double covalent bond	triple covalent bond	co-ordinate bond	C
20	Which of the following bond is represented by single,double or triple line	metallic bond	co-ordinate bond	covalent bond	ionic bond	C
21	covalent compounds are	good conductors of electricity	non-conductors of electricity	poor conductors of electriciy	none of the above	B
22	A metal on loosing electrons turns into	anodes	cathodes	negative ions	posiive ions	D
23	molecules which have permanent dipole are known as	polar	dipolar	non-polar	tripolar	A
24	Electrvalent bond is also called as....	metallic bond	covalent bond	ionic bond	co-ordinate bond	C
25	Chemical bond is formed by participation of	nucleus of atoms	valence electrons of atoms	valence electrons and inner shell electrrons of atoms	none of the above	B
26	Which of the following is not a characterstic of metal	luster	ductiliy	increase in conductance by increase in temperature	malleability	C
27	conduction in metal is due to presence of _____	ions	free elctrons	neutrons	protons	B
28	The difference between the number of atoms in a unit cell of a BCC crystal and an FCC crystal is	1	2	4	6	B
29	an atom to acheive a stable state requires to have _____number of electrons in its valence shell	5	6	7	8	D
30	Anions are formed as a result of	loose of electron	gain of electron	loose of protons	gain of protons	B
31	two identical atoms share electron pairs the resulting bond formed is	non -polar covalent bond	polar covalent bond	double covalent bond	ionic bond	A
32	Three dimensional ordered arrangment of point in space is called as	crystal lattice	lattice points	unit cell	triangles of point	A
33	In crystal lattice,particles are arranged in	two dimensions	four dimension	Three dimension	single dimension	C
34	unit cell is the smallest building unit of	crystal lattice	liquids	gases	none of the above	A
35	which of the following is an amorphous solid	diamond	glass	sodium chloride	none of the above	B
36	The lattice site in a pure crystal cannot be occupied by	molecule	ion	electron	atom	C
37	substances which alter the rate of chemical reaction without undergoing any chemical changes are called as	Reactant	catalysts	products	byproducts	B
38	The process in which catalyst has a different phase to a reaction mixture is known as	homogeneous catalysis	hypergeneous catalysis	heterogenous catalysis	hypogeneous catalyst	C
39	The substances that reduces the effectiveness of a catalyst are called	promoters	autocataysts	inhibitors	none of the above	C
40	when catalyst and reactants are in the same phase then it is called	homogeneous catalysis	heterogenous catalysis	autocataystsis	catalysis	A
41	When a product acts as a catalyst then it is called as	self catalysis	positive catalysis	autocataystsis	negative catalysis	C
42	Which types of bonds are present in a water (H2O)	hydrogen bonds only	ionic and non polar hydrogen bonds	non polar covalent bonds only	both polar covalent and hydrogen bonds	D
43	The bond between two identical non metals atoms has pair of electrons:	unequally shared between the two	transferred fully from one atom to another	equally shared between them	shared by only one atom	C
44	which one is an example of polar molecule	N2	CO2	CH4	H2O	D
45	Electron sea exist in _____	polar bond	ionic bond	covalent bond	metallic bond	D
46	the number of atom nearest neighbour atoms , for an atom in a hexagonal close pack crystal structure is	6	12	18	24	A
47	which one is an example of non polar molecule	CO2	H2O	NH3	HCL	A
48	What is the reason for ionic compounds having high melting point and boiling point	bonds between atoms are strong	small amount of energy is needed to break the bond	ionic compound has giant crystalline structure	the bond between ions are strong	D
49	Number of atom in simple cubic unit cell is	8	9	14	15	A
50	Number of atoms in body centered cubic unit cell (BCC) is	8	9	14	15	B
51	Number of atoms in face centered cubic unit cell (FCC) is	8	9	14	15	C
52	The number of atoms per unit cell of BCC structure is	1	2	4	6	B
53	The number of atoms per unit cell of FCC structure is	1	2	4	6	C
54	The number of atoms per unit cell of SC structure is	1	2	4	6	A
55	Name the metal which increases the activity of iron metal catalyst when added in small quantity in Haber process	Cu	Mo	Al	Mn	B
56	What is the role of Mo in manufacturing of ammonia by habers process where Fe is used as catalyst	catalytic inhibitor	catalytic promoter	catalyst	auto catalyst	B
57	The strength of metallic bond increses with	increase in number of valence electron	decrease in number of valence electron	the decrease in size of atom	increase in size of atom	A
58	crystalline solids are anisotropic because	their physical properties are same in different direction	their physical properties are diferent in different direction	their physical properties are different in same direction	their physical properties are sar	B
59	Amorphous solids are isotropic because	their physical properties are same in different direction	their physical properties are diferent in different direction	their physical properties are different in same direction	their physical properties are sar	A
60	The adsorption theory is applicable to	homogeneous catalysis	heterogenous catalysis	enzymatic catalysis	none of the above	B

61	Name the catalyst which is used for the manufacture of glucose from cane sugar	Maltase	Zymase	CuCL2	CuCL	A
62	Which of the following process is used for the prepartion of chlorine gas	Deacon's process	Bergius process	Ostwalds process	Habers process	A
63	Which of the following process is used for the prepartion of Sulphuric acid	Deacon's process	Chamber process	Ostwalds process	Habers process	B
64	Name the catalyst in Haber Process	Mo	Fe	NO	Pt/V2O5	B
65	Name the catalyst used in manufacturing of sulphuric acid by contact process	pt/V2O5	NO	Fe	Mo	A
66	Select the catalyst which is used for manufacture of ethanol from glucose	Maltase	Pt/V2O5	Zymase	Fe	C
67	2% Ethanol added to chloroform to prevent its oxidation into phosphene gas , which act as	negative catalyst	positive catalyst	catalytic promoter	catalytic inhibitor	A
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	CLASS TEST 1 (BASIC CHEMISTRY) ALL BRANCHES 2.CORROSION					
	QUESTIONS	OPTION A	OPTION B	OPTION C	OPTION D	CORRECT ANSWER
1	Dry corrosion is also called as	atmospheric corrosion	wet corrosion	electrchemical corrosion	Immersed corrosion	A
2	Wet crrosion is aslo called as	electrchemical corrosion	atmospheric corrosion	dry corrosion	direct corrosion	A
3	Which of the following is least protective oxide film	stable porous	stable non porous	unstable	volatile	D
4	In which of the following, rate of corrosion slows or stops after formation of oxide film	stable porous	stable non porous	unstable	volatile	B
5	In which of the following, rate of corrosion is continous and rapid after formation of oxide film	stable porous	stable non porous	unstable	volatile	D
6	A stable porous oxide film is formed on	Fe	AL	Cu	Au	A
7	The corrosion of metal due to oxygen in air is called as	wet corrosion	Oxidation corrosion	electrchemical corrosion	Immersed corrosion	B
8	Gold metal do not corrode as it forms	stable porous oxide film	stable non porous oxide film	unstable oxide film	volatile oxide film	C
9	Metallic Corrosion always occures at the metal	Surface	edges	inner body	corners	A
10	The Process in which an atom looses one or more electrons is called as	reduction	Oxidation	redox	catalysis	B
11	The Process in which an atom gain one or more electrons is called as	reduction	Oxidation	redox	catalysis	A

	BASIC CHEMISTRY CLASS TEST 01 2018-2019 ALL BRANCHES	
Q NO	QUESTION	ANSWERS
1	Complete transfer of one or more electrons from one atom to other atom results in _____ bond	Ionic bond
2	When single atom provides both electrons which are needed for completion of covalent bond then it leads to	CO-ordinate
3	When metal atom looses electron it forms	positive ions
4	During formation of ammonium ion,number of electrons shared between nitrogen atom and hydrogen ion is	2
5	Which of the following molecule is an example of co-ordinate bond .	ammonium ion
6	The pair of outer shell electrons not used in bonding are called as	lone pair
7	Charge on ion depends upon gain or loss of	Electron
8	The bond formed by sharing of four electrons between two bonding atoms is called as	Double covalent bond
9	The bond formed by sharing two pairs of electrons between two a bonding atoms, is called as	Double covalent bond
10	The movement of molecules are completelty restricted in following state of matter	solids
11	During chemical bonding ,the electronic configuration of the bonding atoms is similar to	noble gas
12	Electrons are usually lost by	Metals
13	During formation of nitrogen molecule,number of electron pair shared between the two bonding nitrogen atoms are	3
14	Name the type of bonding in MgO molecule	ionic
15	Metals are good conductors due to	delocalised electon
16	Conduction of electricity in metallic bonding is due to the presence of	delocalised electon
17	In an ionic bond bonding atoms combine together so as	to attain noble
18	Which of the following conditions are true for metal atom	both a n b
19	Nitrogen molecule is an example of	Triple covalent
20	Which of the following bond is represented by single,double or triple line	covalent
21	covalent compounds are	non conductors
22	A metal on loosing electrons turns into	positive ions
23	molecules which have permanent dipole are known as	polar
24	Electrvalent bond is also called as....	ionic bond
25	Chemical bond is formed by participation of	valence electron
26	Which of the following is not a characterstic of metal	increase in conductance
27	conduction in metal is due to presence of _____	free electrons
28	The difference between the number of atoms in a unit cell of a BCC crystal and an FCC crystal is	2
29	an atom to acheive a stable state requires to have _____number of electrons in its valence shell	8
30	Anions are formed as a result of	gain of electrons
31	two identical atoms share electron pairs the resulting bond formed is	non polar covalent
32	Three dimensional ordered arrangment of point in space is called as	crystal lattice
33	In crystal lattice,particles are arranged in	3d
34	unit cell is the smallest building unit of	crystal lattice
35	which of the following is an amorphous solid	GLASS
36	The lattice site in a pure crystal cannot be occupied by	ELECTRONS
37	substances which alter the rate of chemical reaction without undergoing any chemical changes are called as	CATALYST
38	The process in which catalyst has a different phase to a reaction mixture is known as	Hetergenous catalysis
39	The substances that reduces the effectiveness of a catalyst are called	inhhibitors
40	when catalyst and reactants are in the same phase then it is called	Homgenous
41	When a product acts as a catalyst then it is called as	Autocatalyst
42	Which types of bonds are present in a water (H2O)	polar and hydrogen bond
43	The bond between two identical non metals atoms has pair of electrons:	equally shared
44	which one is an example of polar molecule	h2o
45	Electron sea exist in _____	metallic bond
46	the number of atom nearest neighbour atoms , for an atom in a hexagonal close pack crystal structure is	6
47	which one is an example of non polar molecule	CO2
48	What is the reason for ionic compounds having high melting point and boiling point	strong bond
49	Number of atom in simple cubic unit cell is	8
50	Number of atoms in body centered cubic unit cell (BCC) is	9
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52	The number of atoms per unit cell of BCC structure is	2
53	The number of atoms per unit cell of FCC structure is	4
54	The number of atoms per unit cell of SC structure is	1
55	Name the metal which increases the activity of iron metal catalyst when added in small quantity in Haber process	Mo
56	What is the role of Mo in manufacturing of ammonia by habers process where Fe is used as catalyst	promoter
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67	2% Ethanol added to chloroform to prevent its oxidation into phosphene gas , which act as	negative catalyst
1	Dry corrosion is also called as	atmospheric
2	Wet crrosion is aslo called as	electrochemical
3	Which of the following is least protective oxide film	voltalie
4	In which of the following, rate of corrosion slows or stops after formation of oxide film	non porous
5	In which of the following, rate of corrosion is continous and rapid after formation of oxide film	volatile
6	A stable porous oxide film is formed on	Fe
7	The corrosion of metal due to oxygen in air is called as	oxidation
8	Gold metal do not corrode as it forms	unstBLE
9	Metallic Corrosion always occures at the metal	SURFACE
10	The Process in which an atom looses one or more electrons is called as	oxidation
11	The Process in which an atom gain one or more electrons is called as	REDUCTION