Calculation: volume of $KMnO_4$ from plot-2 = 3.8 mL

(N x V) of steel sample solution = (N x V) of $KMnO_4$

N of steel sample solution = 0.05 N x Volume of $KMnO_4$ from Plot-2
_____
20 mL of steel sample

$= 9.5 \times 10^{-3}$ N

Amount of Fe present in 1 L of sample solution = Normality of steel sample x At. wt. of Fe (55.85)

Amount of Fe present in given (100 ml) sample solution = Normality of steel sample x 55.85 x 100
_____
1000

= 0.053 grams in 100 mL

Result: The amount of Iron present in given steel sample is found to be = 0.053 grams.

Evaluation of Result:

Sample number	Experimental value	Actual Value	Percentage of error	Marks awarded