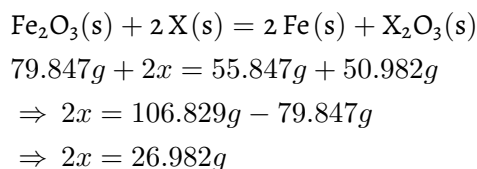


Ap Chem Summer Assignment #3

Shaurya Singh

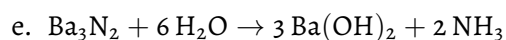
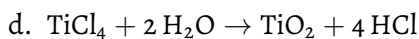
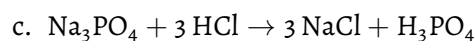
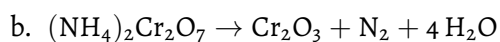
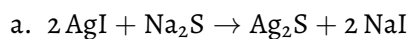
September 6, 2021

1 The following reaction was performed, Identify element X.



Since the atomic weight of Fe is proportional to the given weight (55.847g), the atomic weight of X is equally proportional to the derived weight (26.982g). **Therefore, the element is Aluminium (Al).**

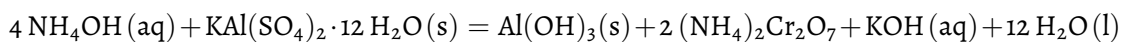
2 Fill in the blanks to balance the following chemical equations:



3 Balance the following equation:

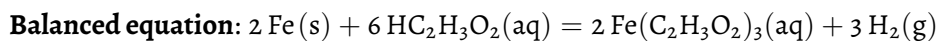
We can multiple NH_4OH by 4, and increase NH_4 and H_2O on the product side to compensate

Balanced equation:

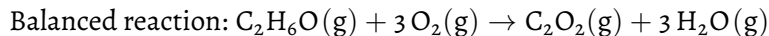


4 Balance the following equation

We need to increase the number of acetate ions on the right side to match the ones on the left. We also need to increase the right hand side to match the number of hydrogen atoms on the left.



5 How many grams of water vapor can be generated from the combustion of 18.74 g of ethanol (C 2 H 6 O)?



Given: ethanol ($\text{C}_2\text{H}_6\text{O}$) = 18.74 g

To be calculated : gms of water vapour (H_2O) generated after combustion

Calculating how many mol of ethanol we have:

Molar mass of ethanol = 46.07 g/mol

Mass(g) of ethanol = 18.74g

Mol = Mass in grams(g) / molar mass (g/mol)

$$\begin{aligned} \text{mol} &= 18.74/46.07 \\ &= .4068 \end{aligned}$$

Multiply this by the ratio of H₂O produced from the combustion of ethanol 3 H₂O molecules were produced from 1 C₂H₆O molecule

$$\begin{aligned}\text{mol of H}_2\text{O} &= \text{mol of C}_2\text{H}_6\text{O} * (\text{molecules produced of H}_2\text{O} / \text{molecules of reactant of C}_2\text{H}_6\text{O}) \\ &= 0.4068 \text{ mol} \times (3/1) \\ &= 1.22 \text{ mol}\end{aligned}$$

Convert mol back to grams

$$\text{Mass(g)} / \text{molar mass(g/mol)} = \# \text{ of mol}$$

$$\text{Mass(g)} = \# \text{ of mol} * \text{molar mass(g/mol)}$$

$$\begin{aligned}\text{Mass of H}_2\text{O(g)} &= 1.22 \text{ mol} * 18.015 (\text{g/mol}) \\ &= 21.99 \text{ g}\end{aligned}$$

Therefore, we will get 21.99g of water vapor

6 How many grams of potassium iodide are necessary to completely react with 20.61g of Mercury (II) chloride

Balance the equation: $\text{HgCl}_2 + 2 \text{KI (aq)} \rightarrow \text{HgI}_2(\text{s}) + 2 \text{KCl(aq)}$

Find the molar mass of HgCl and 2KI (g/mol)

$$\text{HgCl}_2 = 200.59 + 2(35.45)$$

$$= 271.49$$

$$2 \text{ KI} = 2(39.10 + 126.90)$$

$$= 332$$

We calculate the ratio of KI to HgCl₂

$$\text{Ratio} = 2\text{KI}/\text{HgCl}_2$$

$$= 332/271.49$$

$$= 1.22$$

Multiply this by the amount of HgCl₂ used to get the amount of Potassium Iodide needed

$$20.61 \cdot 1.22 = 25.20g$$

Therefore, 25.20g of Potassium Iodide is needed

7 A reaction combines 113.484 g of lead (II) nitrate with 45.010 g of sodium hydroxide (NaOH[aq]).

The equation for the reaction is $\text{Pb}(\text{NO}_3)_2 + 2 \text{NaOH} \rightarrow \text{Pb}(\text{OH})_2 + 2 \text{NaNO}_3$

The molecules molar mass's are

$$\text{Pb}(\text{OH})_2 = 241.196 \text{amu}$$

$$\text{Pb}(\text{NO}_3)_2 = 331.207 \text{amu}$$

$$2 \text{NaOH} = 79.94 \text{amu}$$

$$2 \text{NaNO}_3 = 169.99 \text{amu}$$

- a. We know that 331.207 amu of lead nitrate can react with 79.94 amu of sodium hydroxide. Therefore, 113.484g of lead nitrate will react with $\frac{113.484 \cdot 79.94}{331.207} g$, or 27.41g of sodium hydroxide

Similarly, we know that 331.207 amu of lead nitrate can produce 241.196 amu of lead hydroxide. Therefore, 113.484g of lead nitrate will produce $\frac{113.484 \cdot 241.196}{331.207} g$, or 82.643g of lead hydroxide

- b. From the above we can see the limiting reactant is lead (II) nitrate and the excess reactant left over is sodium hydroxide.
- c. There is $45.010 - 27.408 = 17.602$ grams of the excess reactant left over.
- d. From the above we get an experimental yield of 80.02 percent. We know the limiting reactant gives us a theoretical yield of 82.463 percent. Therefore, the percent yield is $(\frac{80.02}{82.463} \times 100 = 97.04)$, or 97.04%

8 A reaction combines 64.81 grams of silver nitrate with 92.67 grams of potassium bromide

The equation for the reaction is already balanced: $\text{AgNO}_3 + \text{KBr} \rightarrow \text{AgBr} + \text{KNO}_3$

- a. Calculate the atomic weight of the reactants and AgBr:

$$\text{AgNO}_3 = 169.872g$$

$$\text{KBr} = 119.002g$$

$$\text{AgBr} = 187.772g$$

Now we can calculate how much AgBr each reactant made

$$\text{AgNO}_3: \frac{64.81 \times 187.772g}{169.872} = 71.64g$$

$$\text{KBr}: \frac{92.67 \times 187.772g}{119.002} = 146.2g$$

- b. AgNO_3 is the limiting reactant since it produces the least amount of AgBr. Therefore the excessive reactant is KBr since it produced the most amount of AgBr
- c. To calculate how much excessive reactant is left over, we can use the theoretical yield and find the mass of the excessive reactant used:

$$\frac{71.64 \times 119.002g}{187.772} = 45.40, \text{ or } 45.40g \text{ of KBr}$$

- d. In order to find the percent yeild, divide the actual yield by the theoretical yield, then multiply by 100:

$$\frac{14.77g}{71.64g} \times 100 = 20.62, \text{ or } 20.62 \text{ percent.}$$

9 The molecular weight of an insecticide, dibromoethane, is 187.9. Its molecular formula is $C_2H_4Br_2$, What percent by mass of bromine does dibromoethane contain?

We must calculate the atomic weight for each element

$$C = 12.011$$

$$H = 1.008$$

$$Br = 79.90$$

Since the formula is $C_2H_4Br_2$, we can substitute the atomic weights in place of the elements

$$\begin{aligned} &= 2(C) + 4(H) + 2(Br) \\ &= 2(12.011) + 4(1.008) + 2(79.90) \\ &= 24.022 + 4.032 + 159.8 \\ &= 187.9 \end{aligned}$$

Finally, we need to divide the amount of bromine by the total amount in order to find the percent by mass of bromine in $C_2H_4Br_2$

$$\begin{aligned} &= \frac{159.8}{187.9} \\ &= .8505 \end{aligned}$$

Therefore, dibromoethane contains 85.05% by mass of bromine.

10 A given sample of xenon fluoride contains molecules of a single type of XeF_n , where n is some whole number.

First, we need to calculate how many moles of xenon fluoride there are, and calculate its weight

$$\begin{aligned} \text{moles} &= 9.03 * 10^{20} / 6.022 * 10^{23} \\ &= 1.5 * 10^{-3} \\ &= 0.31g \end{aligned}$$

Now, we can calculate for n

$$\begin{aligned} &= 0.31/131 + 19n \\ &= 186.5 + 23.5n = 310 \\ n &= 4 \end{aligned}$$

Therefore its formula is XeF_4

11 A 6.32 g sample of potassium chlorate was decomposed according to the following equation, how many moles were formed?

We have the following values:

$$\begin{aligned} k &= 39.0983g \\ Cl &= 35.45g \\ O &= 16.00g \end{aligned}$$

From there we can calculate the total molar mass

$$39.0983 + 35.45 + 3 * 16 = 122.55g$$

Now, by performing dimensional analysis we get the following equation to convert grams of potassium chlorate to moles of oxygen

$$\begin{aligned} \text{mol} &= \frac{6.32g}{1} \times \frac{1\text{mol}}{122.548g} \times \frac{3}{2} \\ &= \frac{6.32g * 3}{(122.648 * 2)} \\ &= 7.74 * 10^{-2} \end{aligned}$$

Therefore, $7.74 * 10^{-2}$ moles of O_2 is formed

12 What is the coefficient in front of water, when it is produced from the reaction of hydrochloric acid with calcium hydroxide? Calcium chloride is the other product.

The balanced equation is $\text{Ca}(\text{OH})_2 + 2\text{HCl} = \text{CaCl}_2 + 2\text{H}_2\text{O}$

Therefore the coefficient of water (H_2O) is 2

13 What is the subscript of aluminum in the formula of aluminum phosphate?

Aluminum has a subscript of 1 in AlPO_4

14 The reaction of 11.9 g of CHCl_3 with excess chlorine produced 12.6 g of CCl_4 , carbon tetrachloride, what is the percent yield?

The equation for the reaction is $2\text{CHCl}_3 + 2\text{Cl}_2 = 2\text{CCl}_4 + 2\text{HCl}$

We need to calculate the theoretical yield of this reaction. To do that, we need to calculate the atomic weight of CHCl_3 and CCl_4 .

CHCl_3 : 119.378g

CCl_4 : 153.823g

Now we can find the theoretical yield:

$$\begin{aligned} &= \frac{11.9\text{g}}{1} \times \frac{1\text{mol}}{119.378\text{g}} \times \frac{2}{2} \times \frac{153.823\text{g}}{1} \\ &= \frac{11.9 * 2 * 153.823\text{g}}{(119.378 * 2)} \\ &= 15.3\text{g} \end{aligned}$$

To find the percent yield, divide the experimental yield by the theoretical yield, then multiply by

100

$$\begin{aligned} &= \frac{12.6g}{15.3g} \times 100 \\ &= 82.4 \end{aligned}$$

Therefore the percent yield is 82.4%

15 What mass of CCl₄ is formed by the reaction of 8.00 g of methane with an excess of chlorine? CH₄ is the limiting reactant

The given equation is already balanced. The question is asking us to calculate the theoretical yield

We need to calculate the atomic weight of CCl₄ and CH₄

CCl₄: 153.823g

CH₄: 16.043g

Using dimensional analysis, we can calculate the theoretical yield

$$\begin{aligned} &= \frac{8.00g}{1} \times \frac{1}{16.043g} \times \frac{1}{1} \times \frac{153.823g}{1} \\ &= \frac{8 \times 153.823g}{16.043} \\ &= 76.7g \end{aligned}$$

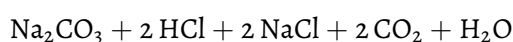
Therefore, 76.72g of CCl₄ is formed by the reaction.

16 A reaction occurs between sodium carbonate and hydrochloric acid producing sodium chloride, carbon dioxide, and water. Write the balanced chemical equation for the reaction.

The equation will be sodium carbonate + hydrochloric acid = sodium chloride + carbon dioxide + water. In correct notation this is written as:



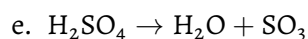
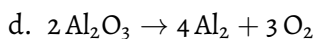
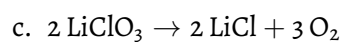
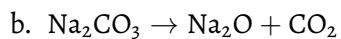
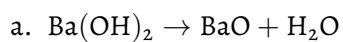
Balanced, this equation is



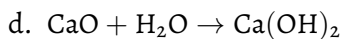
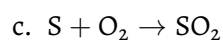
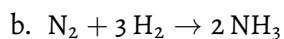
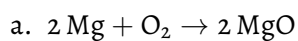
17 Classify the type of reaction from the five major type of reactions you learned in your first year chemistry course and write word equations. If necessary, balance.

- This equation is already balanced, and is a double replacement reaction since the two positive ions switched the negative ions they are bonded with.
- $\text{CH}_4 + 2 \text{O}_2 = \text{CO}_2 + 2 \text{H}_2\text{O}$ is a combustion reaction. since CH_4 reacted with oxygen gas.
- $\text{Fe} + 3 \text{NaBr} = \text{FeBr}_2 + 3 \text{Na}$ is a single replacement reaction since the negative ion in the reactants switches the positive ion its bonded with.
- This equation is already balanced, and is a double replacement reaction since the two positive ions switched the negative ions they are bonded with. The equation was already balanced out
- This equation is already balanced, and is a double replacement reaction since the two positive ions switched the negative ions they are bonded with. The equation was already balanced out
- This equation is already balanced, and is a synthesis reaction since the two reactants combined to form one product.
- This equation is already balanced, and is a decomposition reaction since a compound (in this case the reactant) breaks down

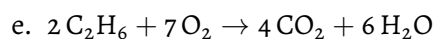
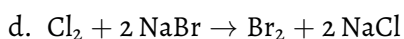
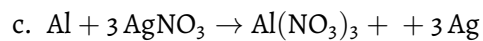
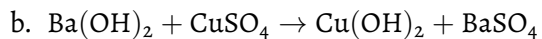
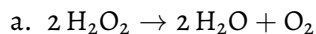
18 Now try these reaction types, Rewrite as a balanced equation with the products predicted



19 Now try these reaction types, Rewrite as a balanced equation with the products predicted



20 Attempt to write and predict products the following chemical reactions:



21 Using the solubility rules table, classify each of the substances as being soluble or insoluble in water. Then, Identify the two new compounds that form if the solutions, as suggested by the following table, were mixed via a double displacement reaction.

21.1 Part A

a. KBr = Soluble

Based on the soluble salt rules, bromine anions are soluble when bonded with a cation that isn't Pb, Ag, or Hg_2^+ . This means that KBr is soluble.

b. PbCO_3 = Insoluble

Since Pb isn't from group 1 and there aren't any rules for CO_3 , PbCO_3 is insoluble.

c. BaSO_4 = Insoluble

Based on the soluble salt rules, when SO_4 is bonded with Ba, it makes it insoluble

d. zinc hydroxide = Insoluble

$\text{Zn}(\text{OH})_2$ is amphoteric but it isn't a strong acid or a strong base, so its insoluble

e. sodium acetate = Soluble

NaCH_3COO is soluble since the cation is from the first group of the periodic table and the anion is the polyatomic ion acetate, both of which are always soluble in a compound

f. silver iodide = Insoluble

AgI is insoluble since the rules of soluble salts state that when the anion I is in a compound with Ag, then it is insoluble

g. cadmium (II) sulfide = Insoluble

CdS is insoluble since there is no rule written about sulfur anions nor cadmium. Cadmium also isn't a group 1 element, making the compound insoluble

h. zinc carbonate = Insoluble

ZnCO_3 is insoluble since there is no rule written about carbonate polyatomic anions. Zinc also isn't a group 1 element

i. silver acetate = Soluble

$\text{AgC}_2\text{H}_3\text{O}_2$ is soluble since the compound contains an acetate polyatomic anion. The soluble salt rule states that any compound with acetate should be soluble

j. copper (II) sulfide = Insoluble

CuS is insoluble since neither copper nor sulfur is included in the soluble salt rules. Copper also isn't a group 1 element, so the compound is insoluble

k. $\text{Mg}_3(\text{PO}_4)_2$ = Insoluble

Phosphate is not mentioned in the soluble salt rules, and magnesium isn't a group one element, so this compound is insoluble

l. KOH = Soluble

It is in the list of the 8 strong bases

m. NiCl_2 = Soluble

The soluble salt rule states that when a compound contains chlorine as anion, as long as the cation it is bonded to isn't Pb, Ag, or Hg_2^{2+} it will be a soluble compound.

n. NH_4OH = Soluble

NH_4OH contains NH_4 , and based on the soluble salt rules, the compound containing it is soluble.

o. Hg_2SO_4 = Insoluble

The soluble salt rules state when sulfate is bonded with Hg_2 in a compound, it renders it insoluble

p. PbI_2 = Insoluble

The soluble salt rule states that when the anion iodine is bonded to the cation lead, it will be insoluble

21.2 Part B

Underlined compounds are precipitates of the reaction

<u>AgBr(s)</u> , KNO ₃ (aq)	<u>Ag₂CO₃(s)</u> , NaNO ₃ (aq)	<u>Ag₂S(s)</u> , Ca(NO ₃) ₂ (aq)	<u>AgOH(s)</u> , NH ₄ NO ₃ (aq)
BaBr ₂ (aq), KCl(aq)	NaCl(aq), <u>BaCO₃(s)</u>	CaCl(aq), BaS(aq)	Ba(OH) ₂ (aq), NH ₄ Cl(aq)
AlBr ₃ (aq), KNO ₃ (aq)	<u>Al₂(CO₃)₃(s)</u> , NaNO ₃ (aq)	AlBr ₃ (aq), <u>Al₂S₃(s)</u>	<u>Al(OH)₃(aq)</u> , NH ₄ NO ₃ (aq)
K ₂ SO ₄ (aq), CuBr ₂ (aq)	<u>CuCO₃(s)</u> , NaSO ₄ (aq)	K ₂ SO ₄ (aq), <u>CuS(s)</u>	NH ₄ (SO ₄) ₂ (aq), <u>Cu(OH)₂(s)</u>

22 Name the following, then draw the Lewis Structure for the following hydrocarbons from their full names.

a. CH₄ - methane

Since the compound contains one carbon atom (meth) and there are sigma bonds (ane), the name of this compound is methane. From the diagram we can see the molecular formula is CH₄

b. C₃H₈ - propane

Since the compound has three carbon atoms (prop) and there are sigma bonds (ane), the name of the compound is propane. With the help of the diagram, we can see the molecular formula is C₃H₈

c. C₄H₈ - 1-butene

Since the compound contains four carbon atoms (but) and central double bond (ene), the name of the compound is 1-butene. With the help of the lewis diagram the molecular formula is C₄H₈

d. C₄H₈ - 2-butene

Since the compound contains four carbon atoms (but) and a central double bond (ene), the name of this compound is 2-butene. The arrangement is different than 1-butene, hence the 2- at the start of the name

22.1 Draw Lewis Structures for the following

- a. Ethane C_2H_6

Based on the name, there are two carbon atoms and sigma bonds are present. The lewis diagram would be

- b. Methane CH_4

Based on the name, there should be one carbon atom and sigma bonds are present. The lewis diagram would be

- c. Propyne C_3H_4

Based on the name, there should be three carbon atoms and a triple bond. The lewis diagram would be

- d. 2 Butene C_4H_8

Based on the name, there should be four carbon atoms and a double bond. The 2 in front signifies this is an isomer of butene. Its lewis diagram would be