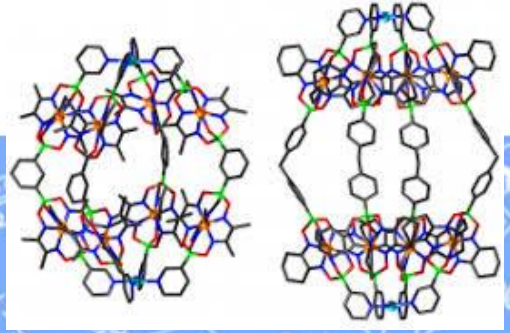
P BLOCK ELEMENTS

#BOUNCEBACK

GROUP 17 & 18



Sakshi Vora

IIT Roorkee

- 7+ years Teaching experience
- 10th, 12th CBSE State Topper
- KVPY fellow

BounceBack

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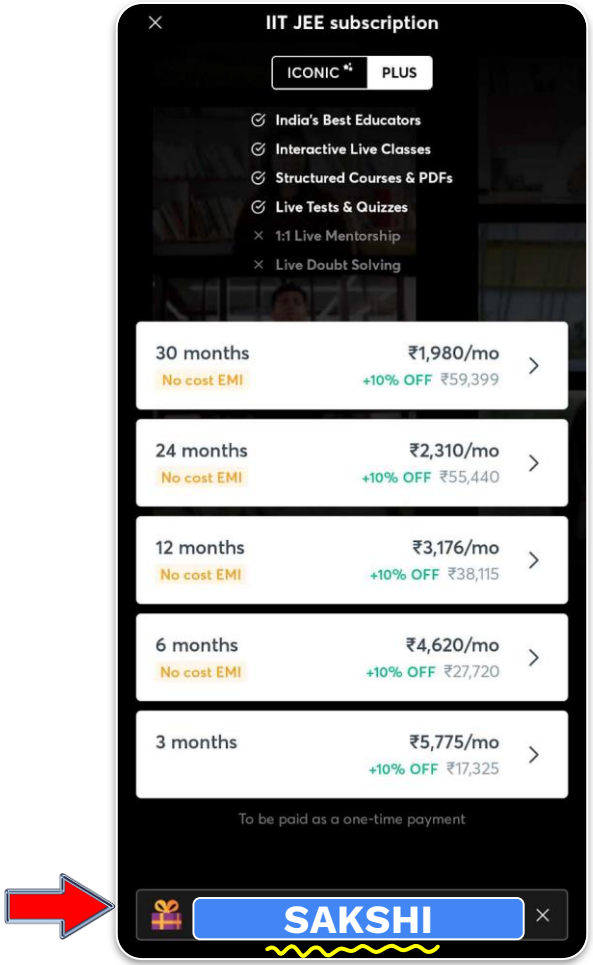
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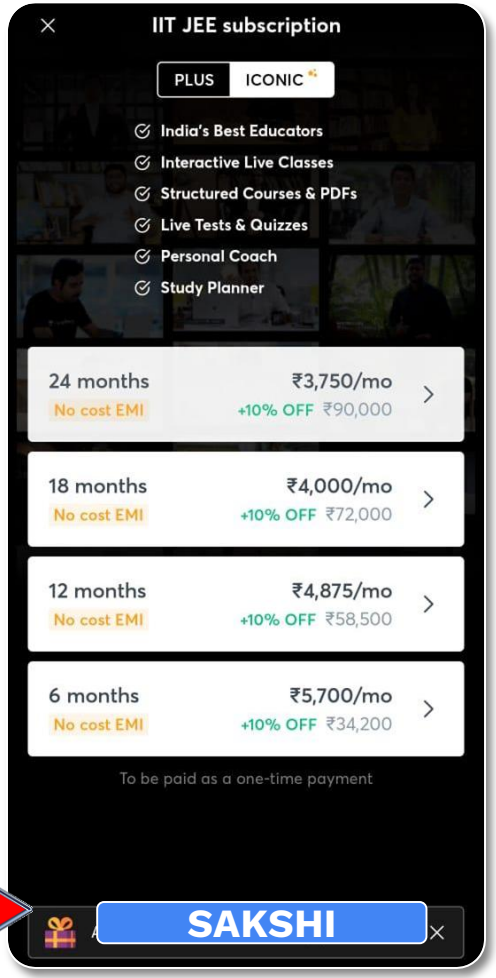
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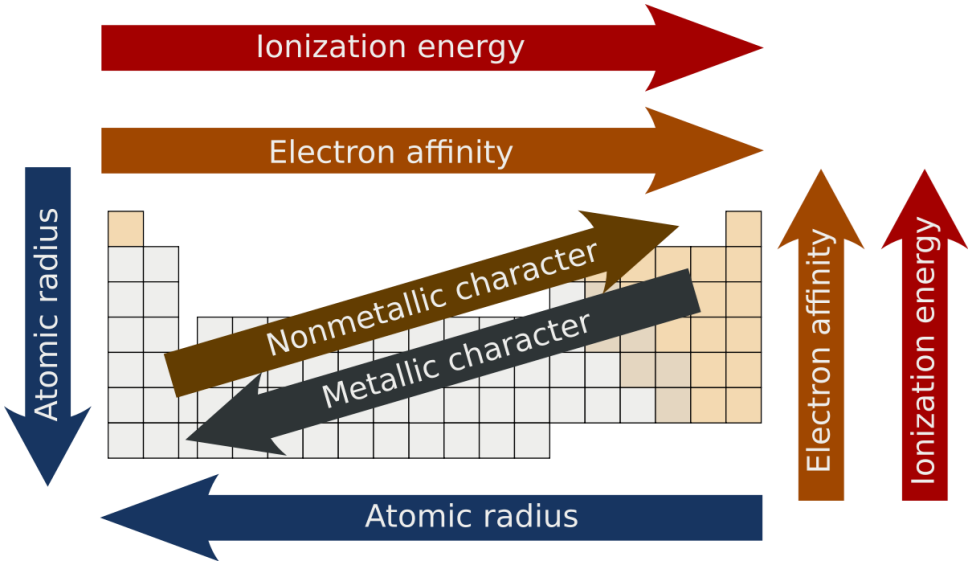
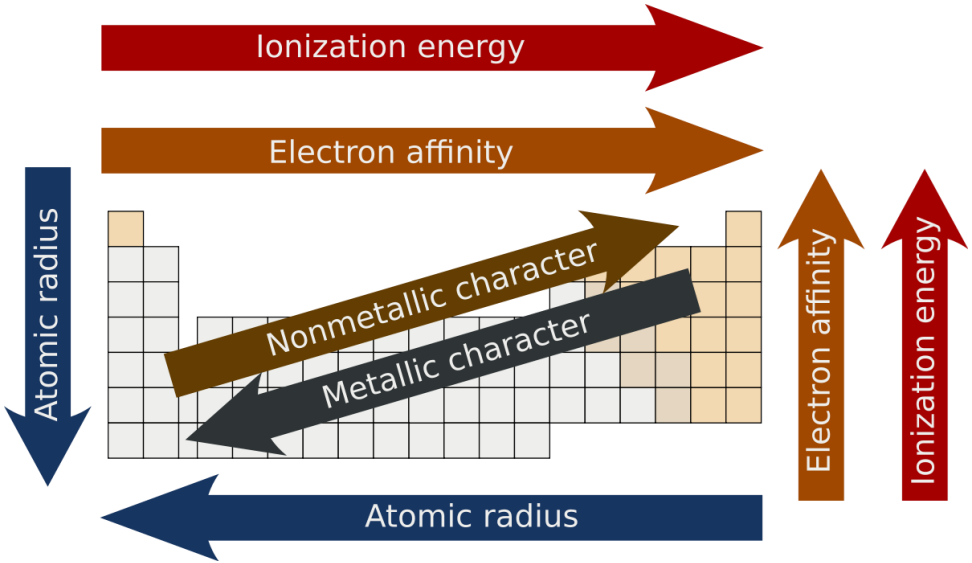
p block elements : group 17 & 18



Group 17


Ionization energy
Electron affinity
Atomic radius
Nonmetallic character
Metallic character
Electron affinity
Ionization energy
Atomic radius

Group - 17 Halogens

F Cl Br I At

All non-metals

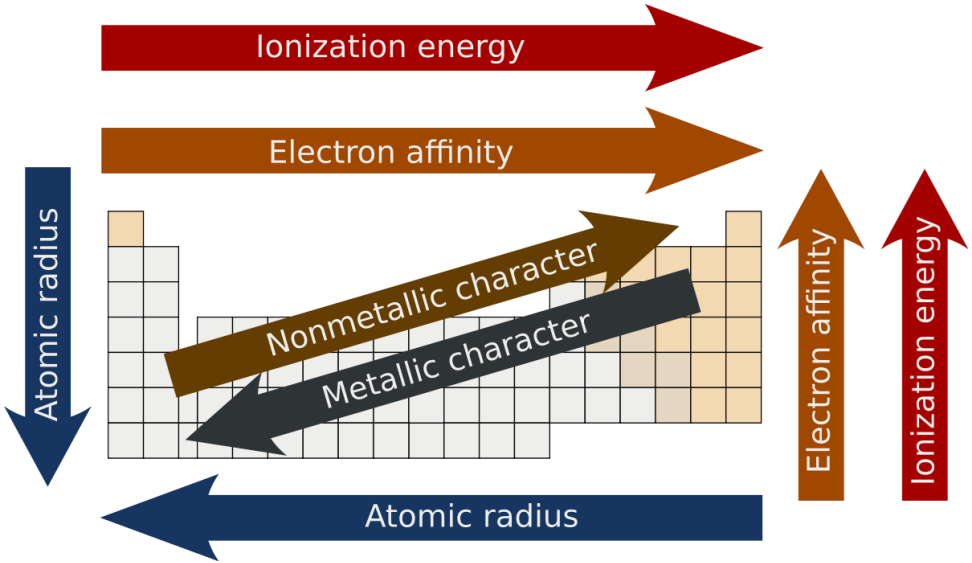
$1e^-$ deficiency

General electronic configuration : $ns^2\ np^5 + 1e^- \longrightarrow ns^2np^6$

(-1)

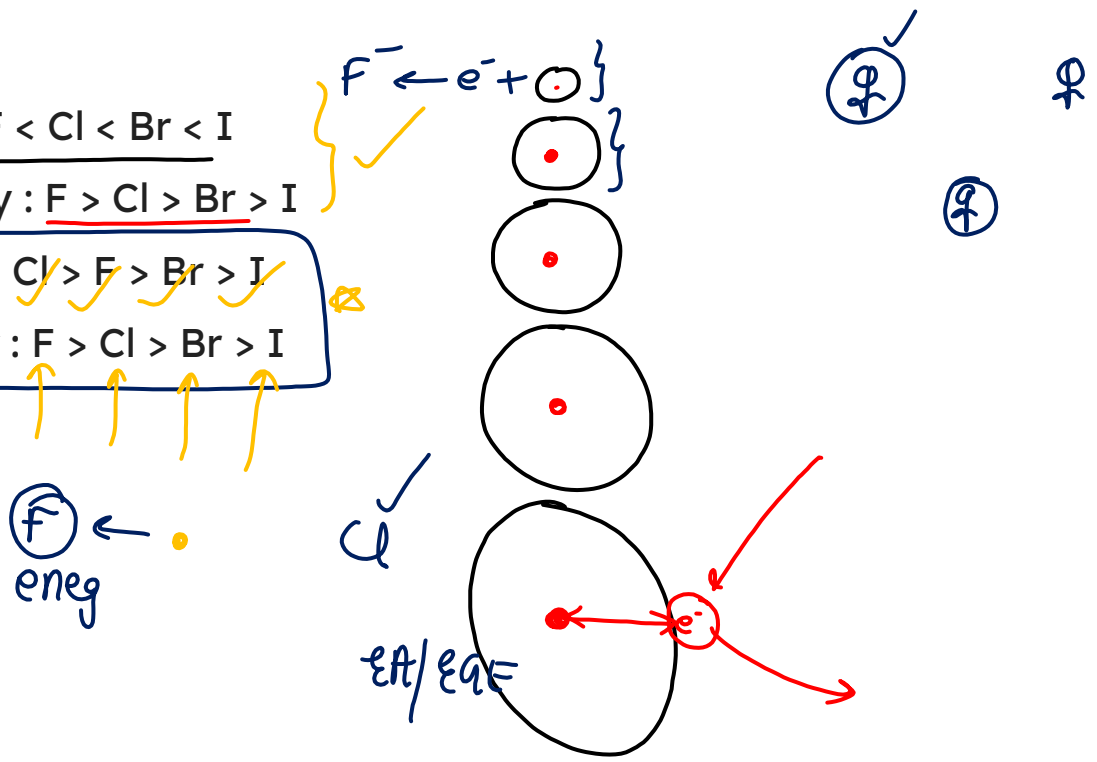
F
Cl
Br
I
At

Variations in Group 17



Variations in group - 17

- Atomic Radius : F < Cl < Br < I
- Ionization Energy : F > Cl > Br > I
- Electron Affinity : Cl > F > Br > I
- Electronegativity : F > Cl > Br > I



Variations in group - 17

Bond energy: $Cl_2 > Br_2 > F_2 > I_2$

Bond length : $F_2 < Cl_2 < Br_2 < I_2$

M.Pt / BpT : $F_2 < Cl_2 < Br_2 < I_2$

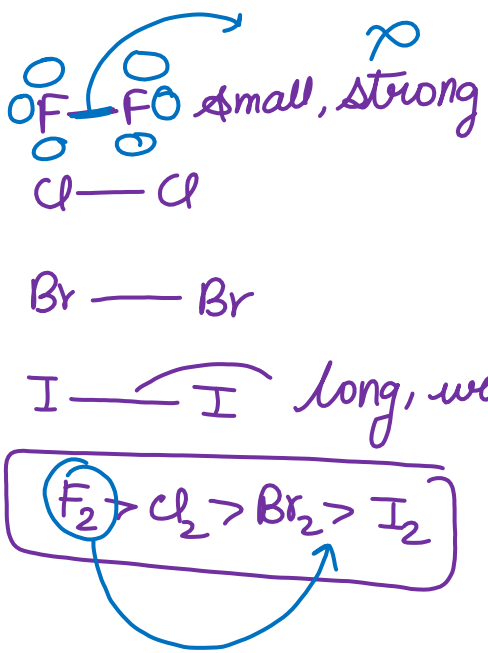
F_2 : Pale yellow gas

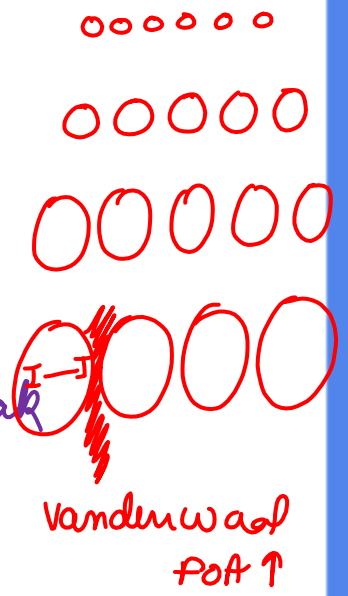
Cl_2 : dark yellow gas

Br_2 : Reddish brown liquid

I_2 : Violet Black solid

colour?





Arrange the following bonds according to their average bond energies in descending order: C – Cl, C – Br, C – F, C – I

A. C – F > C – Cl > C – Br > C – I

B. C – Br > C – I > C – Cl > C – F

C. C – I > C – Br > C – Cl > C – F

D. C – Cl > C – Br > C – I > C – F

long

weak

bond – small → strong

Jan. 08, 2020 (II)

F–F

Cl–Cl

Br–Br

I–I

$Cl_2 > Br_2 > F_2 > I_2$

The colour of the X_2 molecules of group 17 elements changes gradually from yellow to violet down the group. This is due to

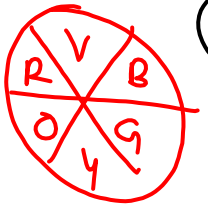
[Adv.2017]

A. The physical state of X_2 at room temperature changes from gas to solid down the group

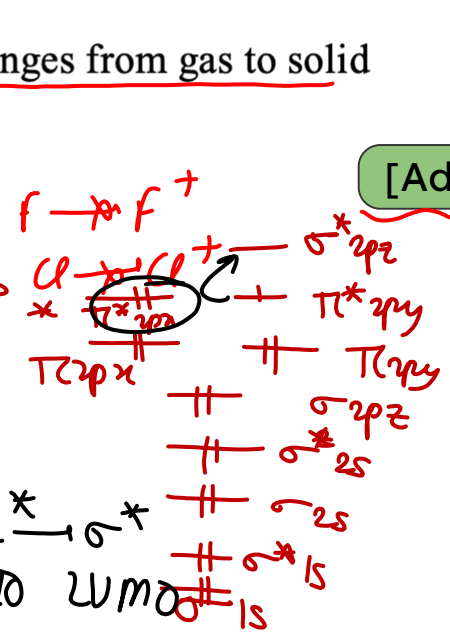
B. Decrease in ionization energy down the group

C. Decrease in π^* - σ^* gap down the group

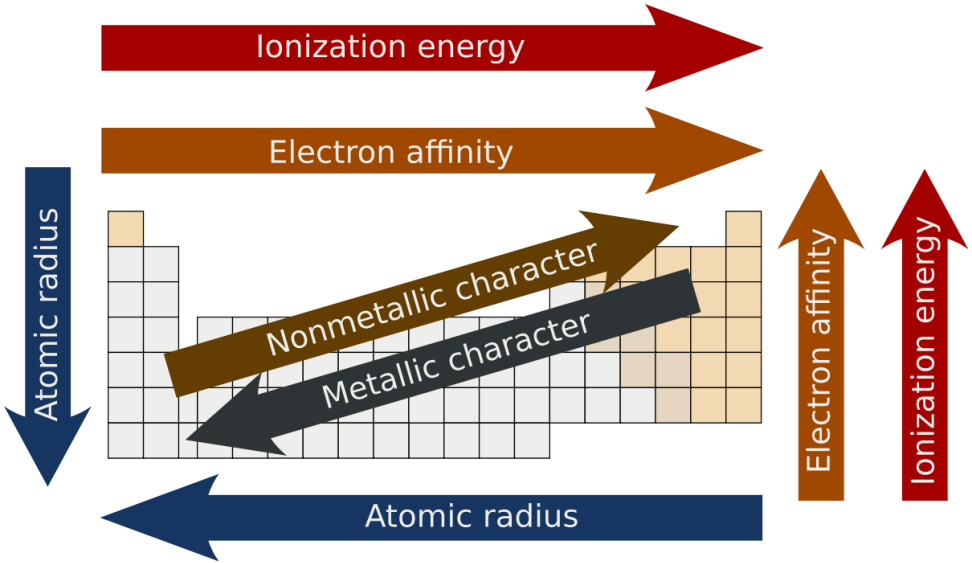
D. Decrease in HOMO-LUMO gap down the group



(C,D)



Anomalous behavior of F



Anomalous Behavior of F

- Small size
- Highest electronegativity
- Low covalency
- Low bond dissociation energy
- Highest oxidising nature
- Absence of d orbitals

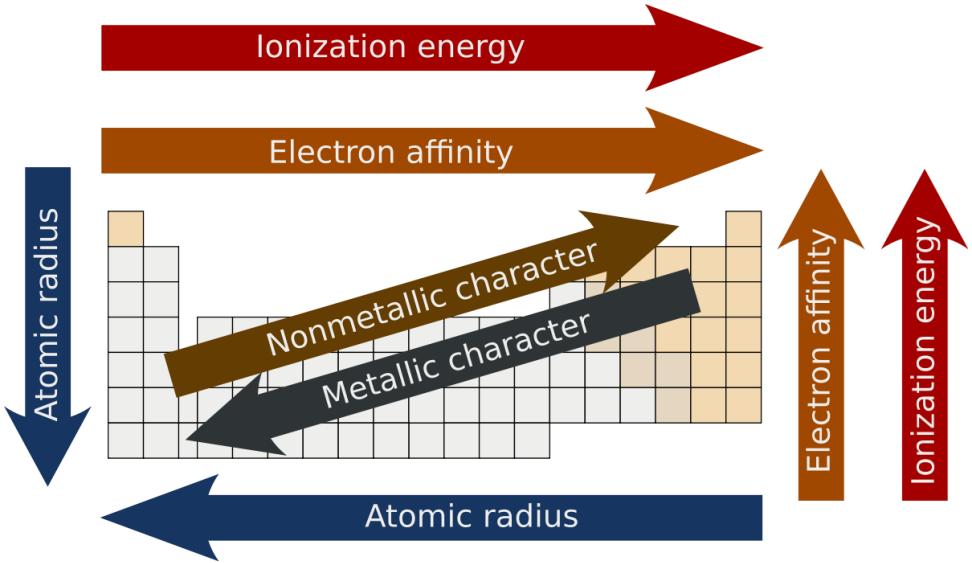
H—F

BDE

F–F

why ?

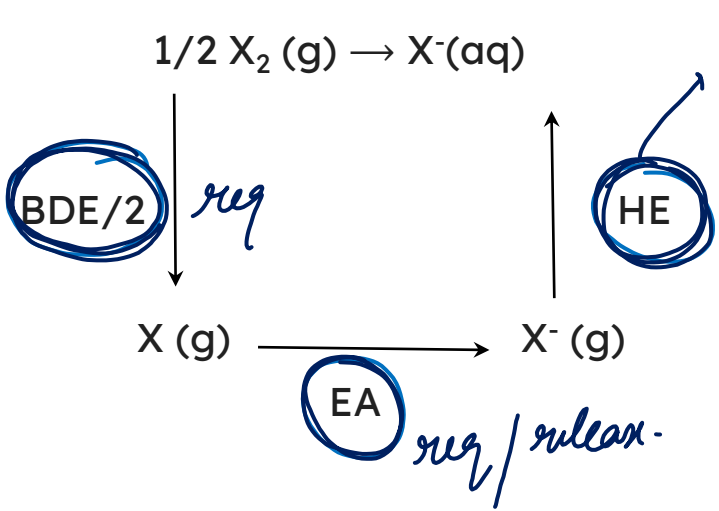
Oxidising nature



Oxidising & reducing nature of halogens

$1/2\ X_2\ (g) \longrightarrow X^-(aq)$

$1/2\ X_2\ (g) \longrightarrow X^-(aq)$



$\frac{1}{2}X_2(g) \longrightarrow X^-(aq)$

BDE/2 ↓ ; ↑ H_{hyd} (HE)

$X(g) \xrightarrow{e^-} X^-(g)$; EA

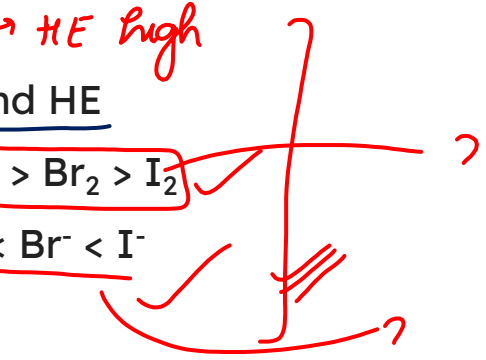
F_2, Cl_2, Br_2, I_2

Oxidising & Reducing Nature of Halogens

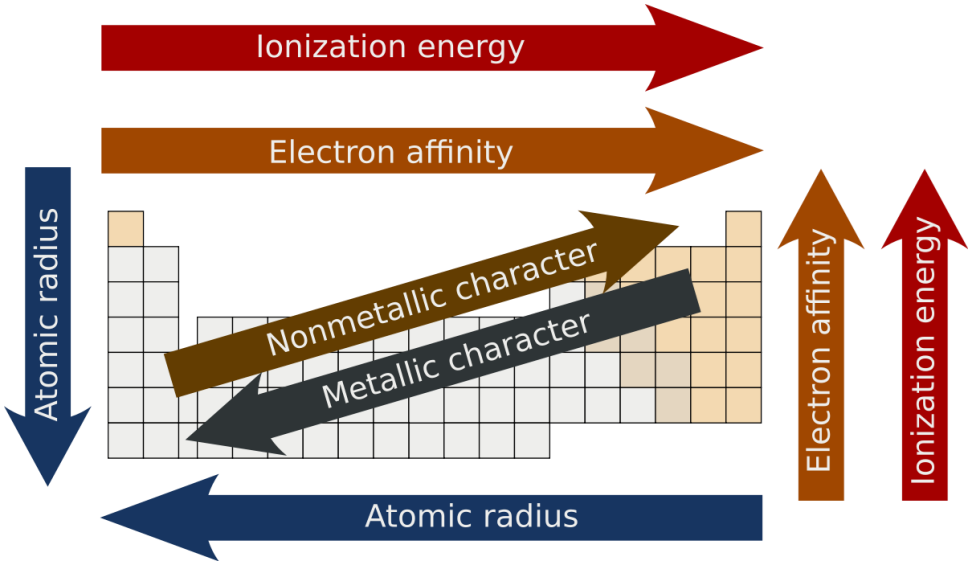
On comparing BDE, EA and HE

Oxidising nature - $F_2 > Cl_2 > Br_2 > I_2$

Reducing nature : $F^- < Cl^- < Br^- < I^-$



Reactivity towards hydrogen



Reactivity towards Hydrogen -

$X_2 + H_2 \rightarrow 2HX$

Reactivity : $F_2 > Cl_2 > Br_2 > I_2$

Reaction of fluorine can occur in dark

Reaction of iodine requires hv/catalyst

Reactivity towards Hydrogen -

$-\triangle H_f$ HF > HCl > HBr > HI (all exothermic)

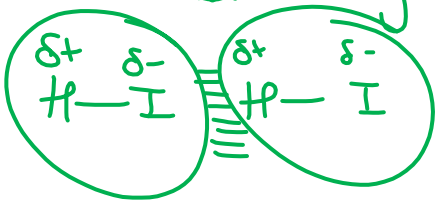
B length : HF < HCl < HBr < HI

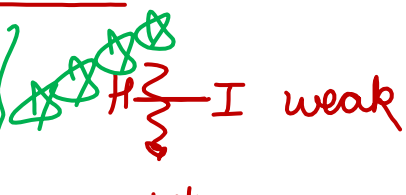
Bond energy : HF > HCl > HBr > HI

Acidic nature : HF < HCl < HBr < HI

M.pt : HI > HF > HBr > HCl

B.pT. : HF > HI > HBr > HCl





H^+, I^-

$H_2 + F_2 \rightarrow HF$

$H_2 + Cl_2 \rightarrow HCl$

HBr

HI

HF has highest boiling point among hydrogen halides, because it has:

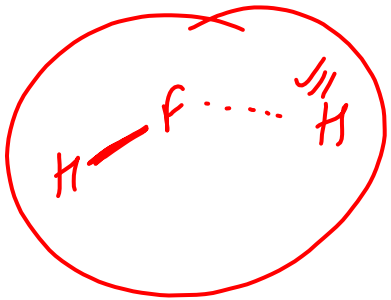
A. strongest van der Waals' interactions

B. lowest ionic character

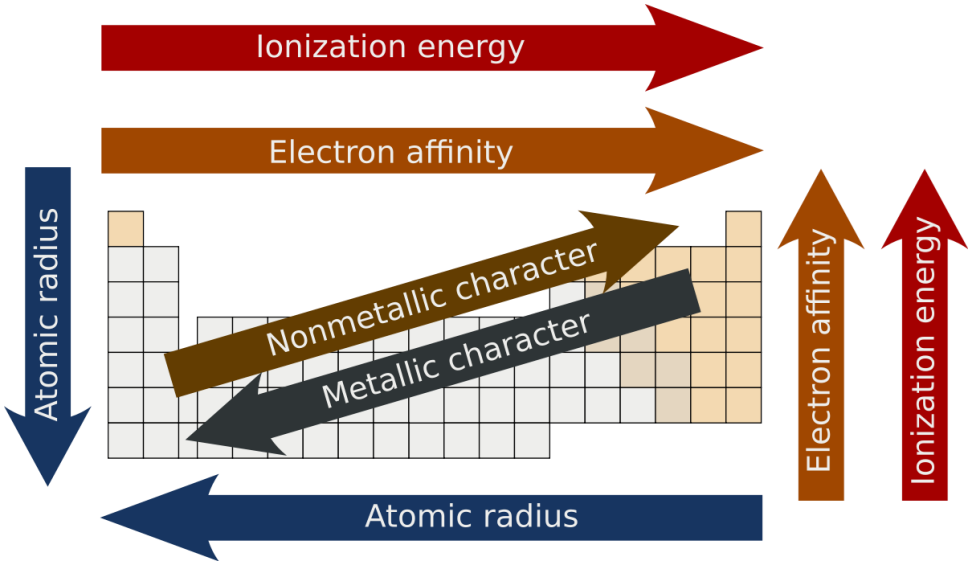
C. strongest hydrogen bonding

D. lowest dissociation enthalpy

[April 9, 2019 (II)]



Reactivity towards Oxygen



Reactivity towards Oxygen -

Most of the compounds are OA (fluorides of oxygen)

$O_2F_2 + X_eF_4 \longrightarrow XeF_6 + O_2$

(+1, +4, +6, 0)

O_2F_2 (+1), OF_2 (+2)

Cl, Br, I form oxides from +1 to +7

As +ve charge inc, stability of oxides inc

For a particular OS,

Stability I > Cl > Br

imp.

O_2F_2 (+1)

$2x + 2(-1) = 0$

$2x = 2$

$x = +1$

F.O

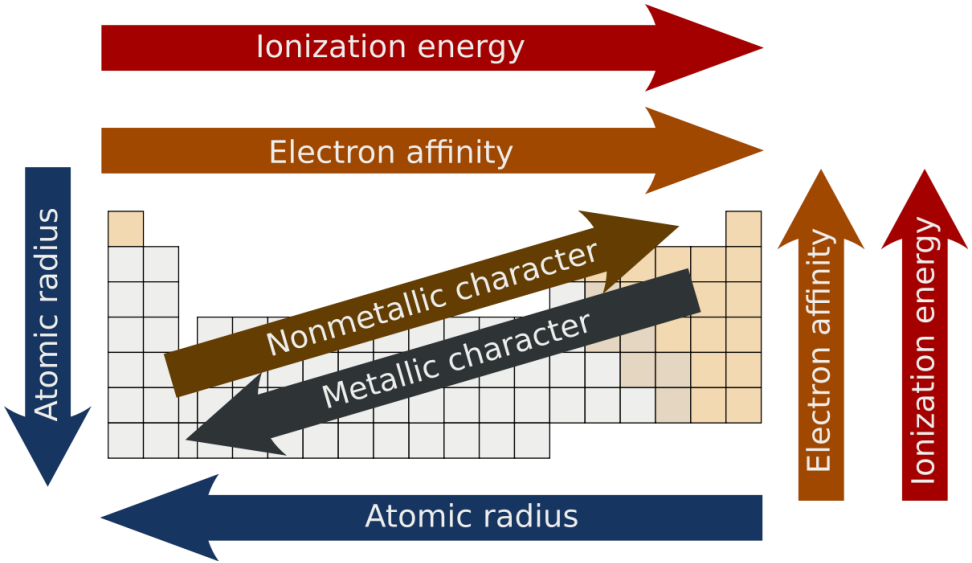
F^- O^+

H_2O

oxide of hydrogen

O^{2-} H^+

Interhalogen compounds



Reactivity towards Halogen -

$X_2 + X_2' \longrightarrow XX'_n$

Interhalogen compounds

n = 1, 3, 5, 7

eneg : X < X'

$X_2 + X'_2 \longrightarrow XX'_n$

$F_2 + Cl_2 \longrightarrow ClF_3$

$X'_2 + X_2 \longrightarrow XX'_n$

central atom: big in size, covalency more, less eneg.

side atom: small, eneg. high, covalency (less)

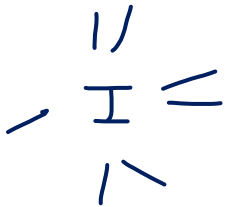
Interhalogen compounds -

XX'	XX'_3	XX'_5	XX'_7
IF	IF_3	IF_5	IF_7
BrF	ClF_3	BrF_5	ICl_7, IBr_7, ClF_7, $BrCl_7$
ClF	BrF_3	ClF_5	
~~BrCl~~	ICl_3		
ICl			
IBr			
linear	T-shaped	Squre pyramidal	Pentagonal Bipyramidal
sp^3d	sp^3d	sp^3d^2	sp^3d^3

$5\sigma + 1lp$

sp^3d^2

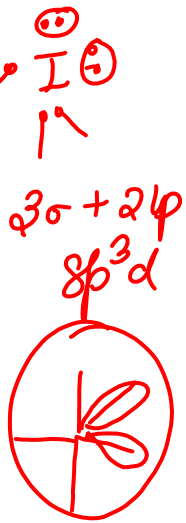
ClF_3 total lp.



$\sigma = 5$

sp^3d^3

IF_3



$3\sigma + 2lp$

sp^3d

Hydrolysis of interhalogen compounds -

$XX'_n + H_2O \longrightarrow$ oxyacid + hydracid

$XX'_7 + H_2O \longrightarrow HXO_4 + HX'$

$XX'_5 + H_2O \longrightarrow HXO_3 + HX'$

$XX'_3 + H_2O \longrightarrow HXO_2 + HX'$

$XX' + H_2O \longrightarrow HXO + HX'$

$\overset{+7}{IF_7} \longrightarrow \overset{+7}{HIO_4}$

$1 + x - 8 = 0$

$IF_7 + H_2O \rightarrow HIO_4 + HF$

$ClF_3 + H_2O \rightarrow HClO_2 + HF$

$IF_5 + H_2O \rightarrow HIO_3 + HF$

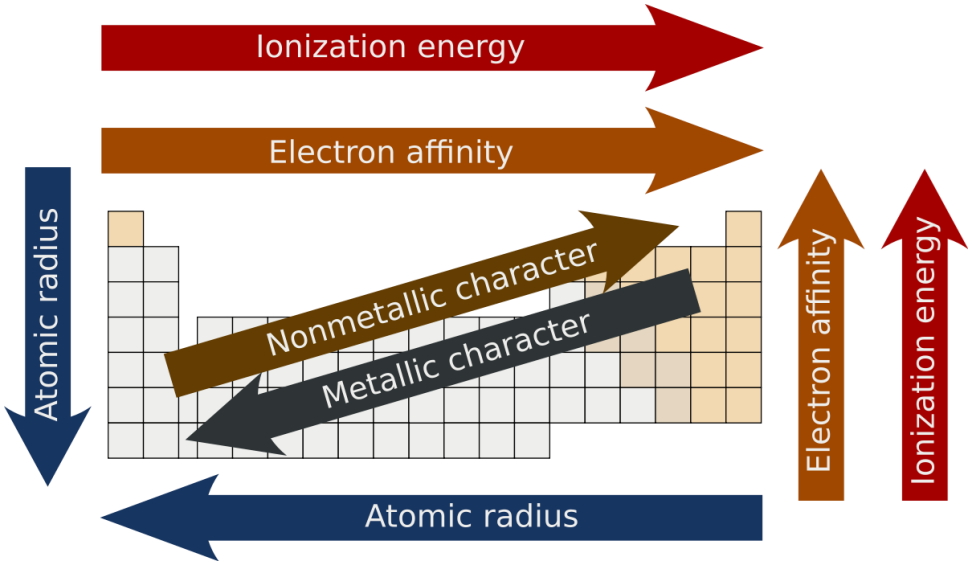
$XX'_7 \rightarrow HXO_4 + HX'$

$XX'_5 \rightarrow HXO_3 + HX'$

$XX'_3 \rightarrow HXO_2 + HX'$

$XX' \rightarrow HXO + HX'$

Reactivity order

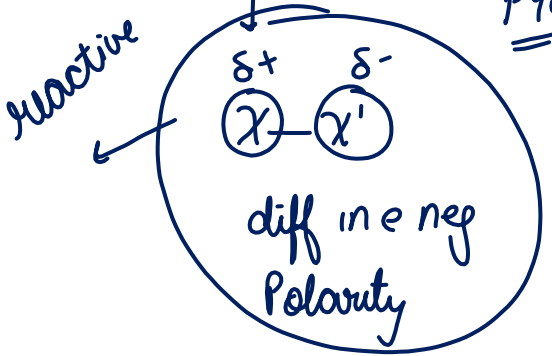


Reactivity order-

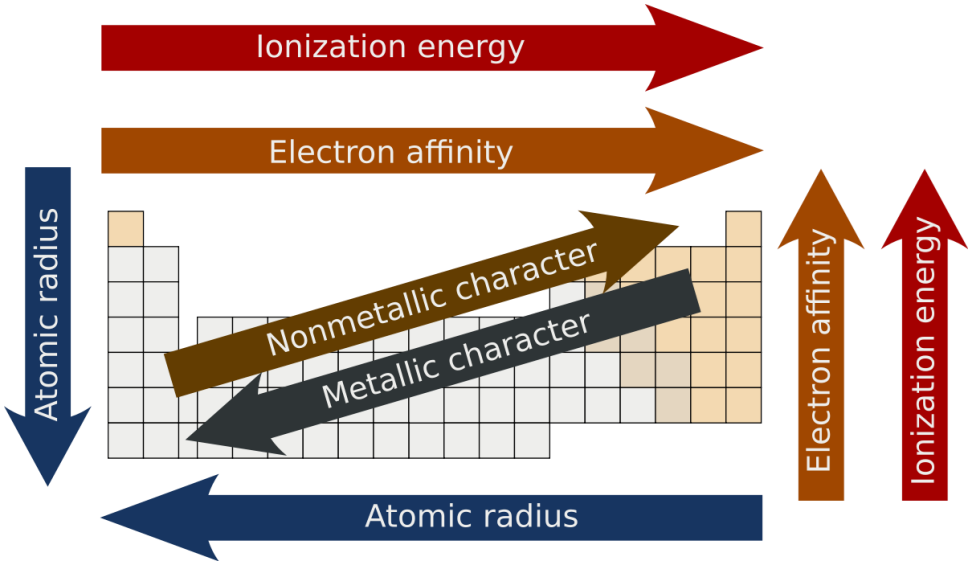
F_2 > Interhalogen compounds > Cl_2 > Br_2 > I_2

chemical Bonding

PYQ.



Reactivity towards metals



Reactivity towards metals -

$M + X_2 \longrightarrow MX_n$

As +ve charge ↑

Covalent nature ↑

Ionic nature ↓

$Na + Cl_2 \longrightarrow \overset{+1}{NaCl}$ — ionic + cov

$Mg + Cl_2 \longrightarrow \overset{+2}{MgCl_2}$ — ionic + cov

$Al + Cl_2 \longrightarrow \overset{+3}{AlCl_3}$ — cov + ion

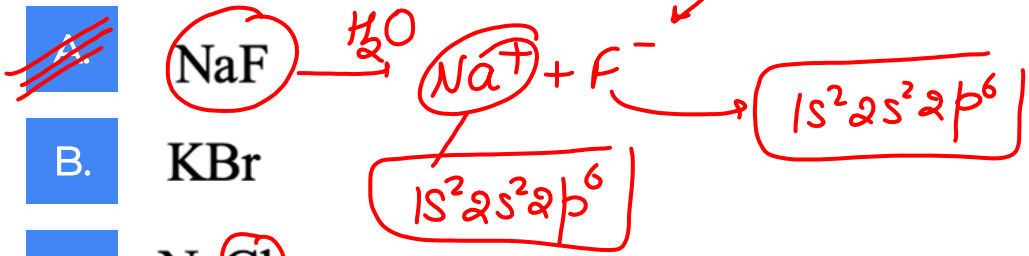
Aqueous solution of which salt will not contain ions with the electronic configuration $1s^2 2s^2 2p^6 3s^2 3p^6$?

A. NaF

B. KBr

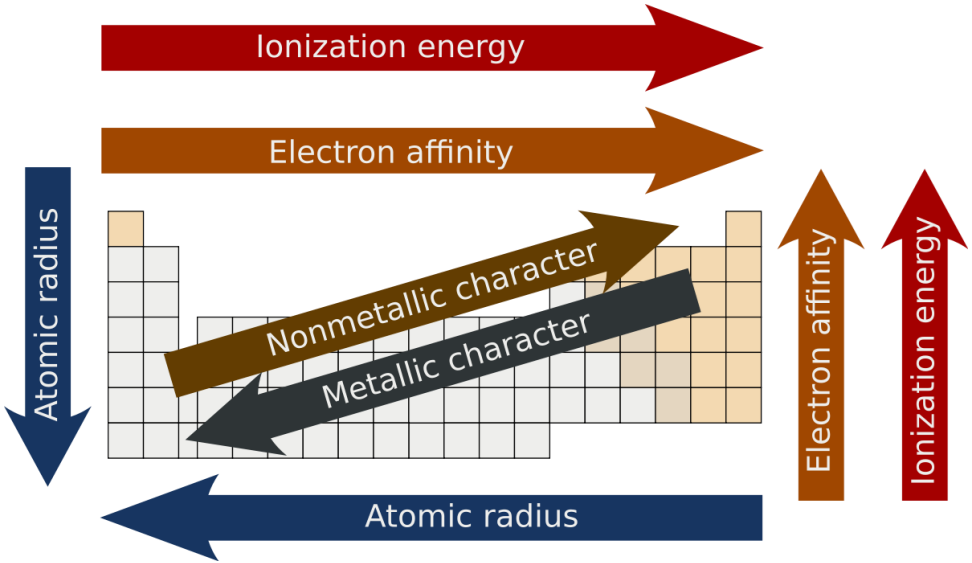
C. NaCl

D. CaI_2



[April 10, 2016]

Preparation of Halogens



Preparation of halogens -

a. F_2

Electrolysis of KHF_2 : Moisson process

$K^+ HF_2^-$

at anode : $2F^- \longrightarrow 2e^- + F_2$

at cathode : $2H^+ + 2e^- \longrightarrow H_2$

$K^+ \; HF_2^-$

b. Cl_2

Deacon's process

$HCl + O_2 \longrightarrow Cl_2 (\uparrow) + H_2O$

$CuCl_2$

Preparation of halogens -

Preparation of Br_2

$Br^- + O_3 / KMnO_4 / H^+ / K_2Cr_2O_7 / H^+ \longrightarrow Br_2 + H_2O / Mn^{2+} / Cr^{3+}$

$Br^- + MnO_2 / Conc.\ HNO_3 / Conc\ H_2SO_4 \longrightarrow Br_2 + / Mn^{2+} / NO_2 / SO_2$

① $Br^- + O_3 \longrightarrow Br_2 + H_2O$

$Br^- + KMnO_4/H^+ \longrightarrow Br_2 + Mn^{2+}$

$Br^- + K_2Cr_2O_7/H^+ \longrightarrow Br_2 + Cr^{3+}$

$+ MnO_2 \longrightarrow$

Preparation of halogens -

Preparation of I_2

$I^- + MnO_2 / K_2Cr_2O_7 / H^+ / KMnO_4 / H^+ \longrightarrow I_2 + Mn^{2+} / Cr^{3+} / Mn^{2+}$

$I^- + Cl_2 / Br_2 / \text{Conc } HNO_3 / \text{Conc } H_2SO_4 \longrightarrow I_2 + Cl^- / Br^- / NO_2 / SO_2$

$I^- + MnO_2 \longrightarrow I_2 + Mn^{2+}$

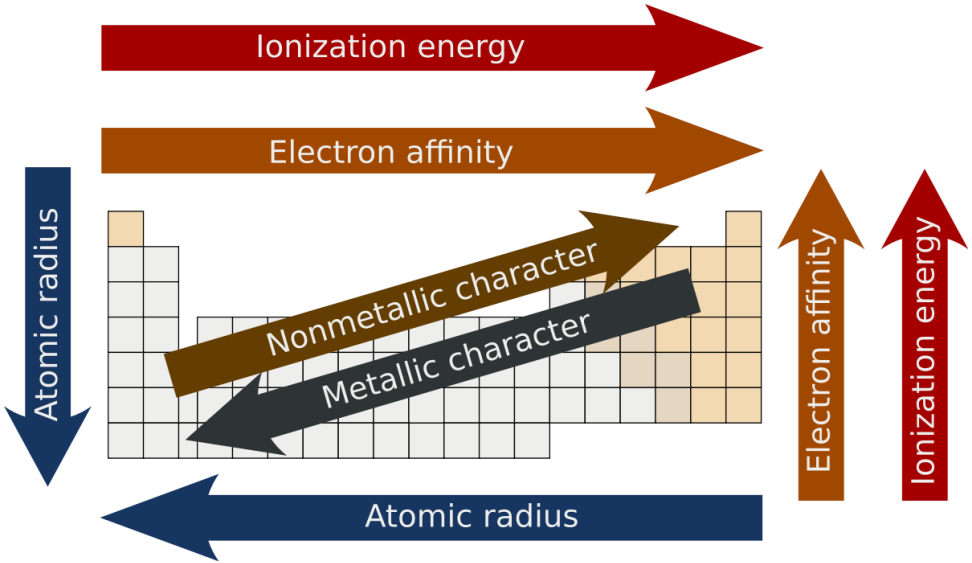
$K_2Cr_2O_7/H^+ \longrightarrow I_2 + Cr^{3+}$

$KMnO_4/H^+ \longrightarrow I_2 + Mn^{2+}$

$I^- + Cl_2 \longrightarrow Cl^- + I_2$

$Br_2 \longrightarrow Br^- + I_2$

Reaction with water



Reaction of Halogens with water -

$F_2 + H_2O \longrightarrow HF + O_2$

$Cl_2 + H_2O \longrightarrow HCl + HOCl$

$Br_2 + H_2O \longrightarrow HBr + HOBr$

$I_2 + H_2O \longrightarrow$ no reaction

$Cl_2 + H_2O \longrightarrow HClO + HCl$

$Br_2 + H_2O \longrightarrow HBrO + HBr$

$HF + NaOH \longrightarrow NaF + H_2O + O_2$

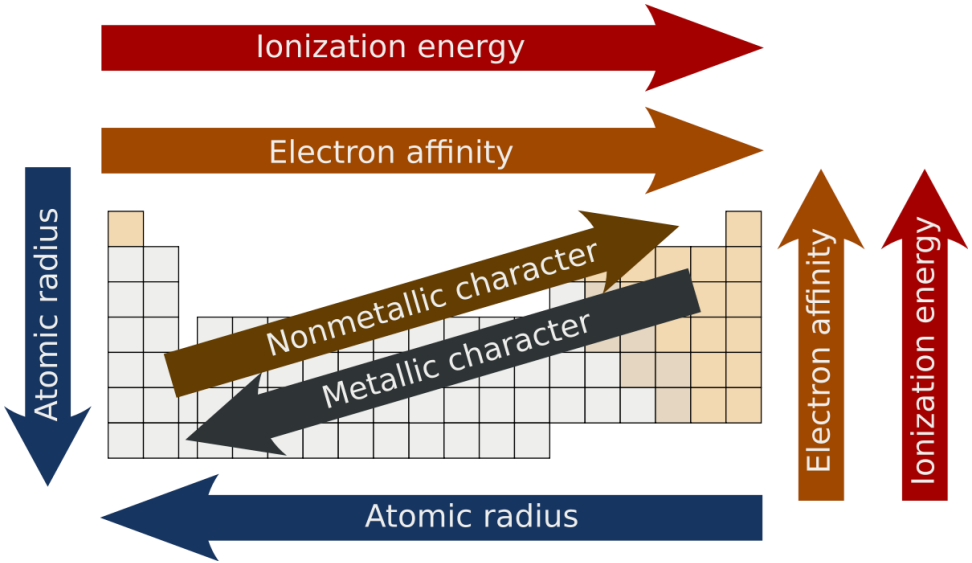
$\overset{0}{F_2} + \overset{+1\ -2}{H_2O} \longrightarrow \overset{0}{O_2} + HF$ (−1)

$F_2 + H_2O \longrightarrow HF + O_2$

$Cl_2/Br_2 + H_2O \longrightarrow HCl + HClO$
$HBr + HBrO$

$I_2 + H_2O \longrightarrow \times$

Reaction with NaOH



Reaction with NaOH -

F_2 + cold NaOH $\longrightarrow$ NaF + H_2O + O_2

F_2 + hot conc NaOH $\longrightarrow$ NaF + H_2O + O_2

Cl_2 + cold dil NaOH $\longrightarrow$ NaCl + NaClO

Cl_2 + hot conc NaOH $\longrightarrow$ NaCl + $NaClO_3$

Br_2 + cold dil NaOH $\longrightarrow$ NaBr + NaBrO

Br_2 + hot conc NaOH $\longrightarrow$ NaBr + $NaBrO_3$

I_2 + cold dil NaOH $\longrightarrow$ NaI + NaIO

I_2 + hot conc NaOH $\longrightarrow$ NaI + $NaIO_3$

$1 + x - 2 = 0$

$x - 1 = 0$

$x = +1$

NaCl + NaClO

NaCl + $NaClO_3$

$1 + x - 6 = 0$

NaF, H_2O, O_2

In the following reactions, products (A) and (B), respectively, are:

$NaOH + Cl_2 \rightarrow (A) + \text{side products (hot and conc.)}$

$Ca(OH)_2 + Cl_2 \rightarrow (B) + \text{side products (dry)}$

Jan. 07, 2020 (II)

A. $NaClO_3$, and $Ca(OCl)_2$

B. $NaClO_3$ and $Ca(ClO_3)_2$

C. $NaOCl$ and $Ca(OCl)_2$

D. $NaOCl$ and $Ca(ClO_3)_2$

$Ca(OH)_2 + Cl_2 \rightarrow Ca(OCl)_2 + CaCl_2 + H_2O$

Chlorine on reaction with hot and concentrated sodium hydroxide gives :

A. Cl^- and ClO_3^-

B. Cl^- and ClO^-

C. ClO_3^- and ClO_2^-

D. Cl^- and ClO_2^-

$Cl_2 + \text{hot conc } NaOH \rightarrow NaCl + NaClO_3$

Cl^-, ClO_3^-

[Jan. 12, 2019 (II)]

The products obtained when chlorine gas reacts with cold and dilute aqueous NaOH are :

A. ClO^- and ClO_3^-

B. ClO_2 and ClO_3^-

C. Cl^- and ClO^-

D. Cl^- and ClO_2^-

NaCl, NaClO

Cl^-, ClO^-

[Main 2017]

Reaction of Br_2 with Na_2CO_3 in aqueous solution gives sodium bromide and sodium bromate with evolution of CO_2 gas. The number of sodium bromide molecules involved in the balanced chemical equation is

[Sep. 06,2020 (1)]

P-Block

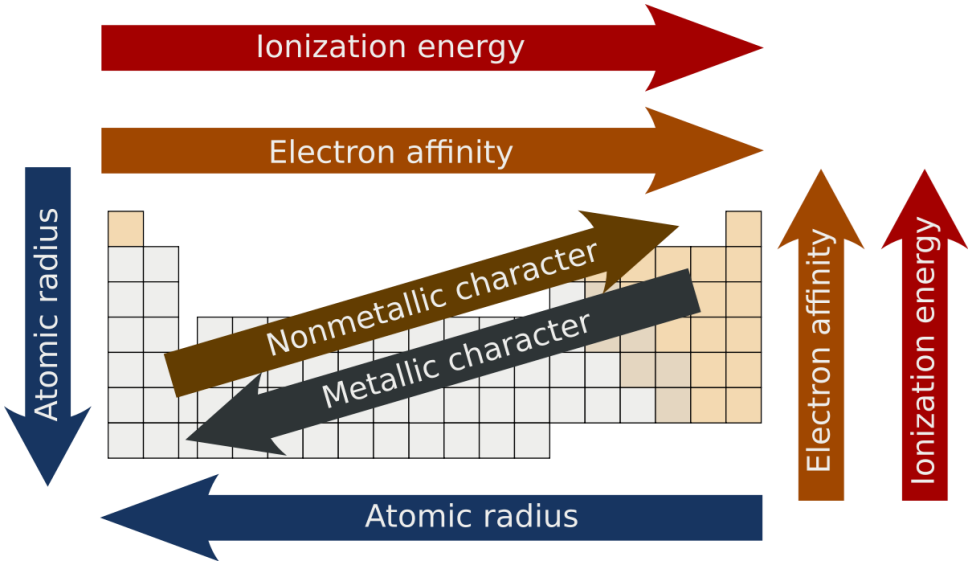
Redox

$$\overset{0}{Br_2} + Na_2CO_3\,(aq) \longrightarrow \overset{-1}{5NaBr} + \overset{+5}{NaBrO_3} + CO_2$$

(∵ $NaOH + H_2CO_3$)

$\hookrightarrow H_2O + CO_2$

Reaction with Ammonia



Reaction with NH_3 -

$Cl_2 + NH_3$ (excess) $\longrightarrow NH_4Cl + N_2$

Cl_2(excess) $+ NH_3 \longrightarrow NCl_3 + HCl$

already gpt Ques.

basic

NH_3 (excess) + Cl_2 $\longrightarrow$ HCl + N_2

Cl ♡ H

acidic

$\downarrow$

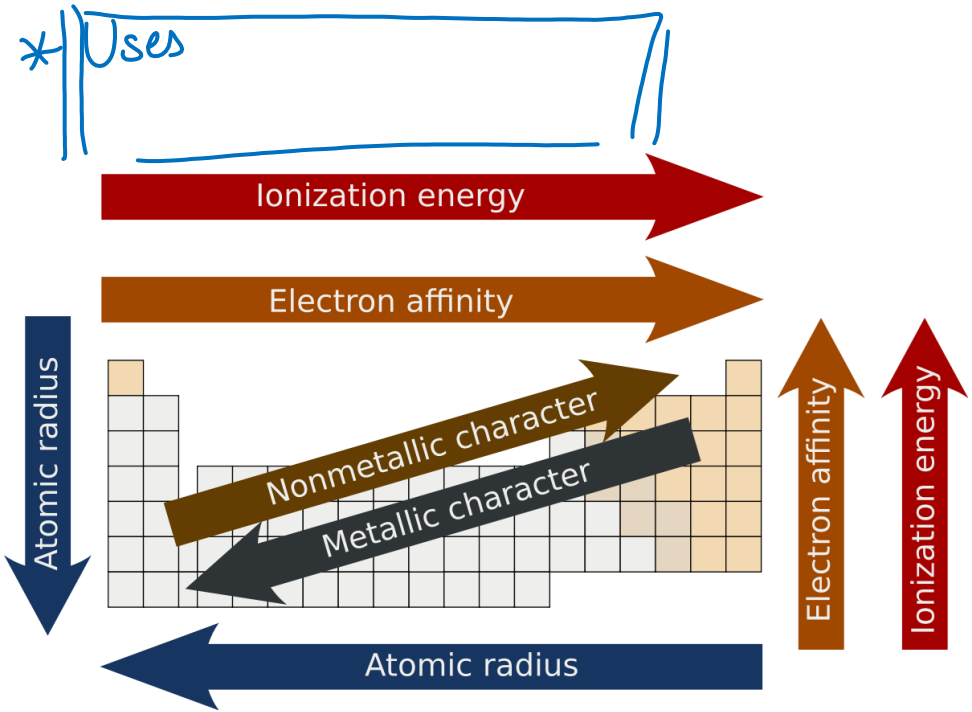
NH_4Cl

$NH_3 + Cl_2 \longrightarrow HCl + N_2$

excess

$\hookrightarrow NCl_3$

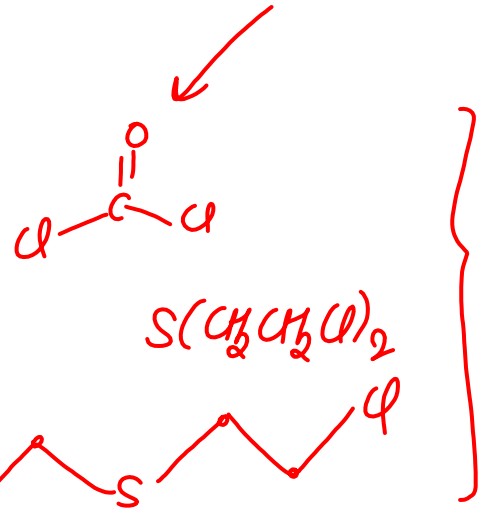
Important points



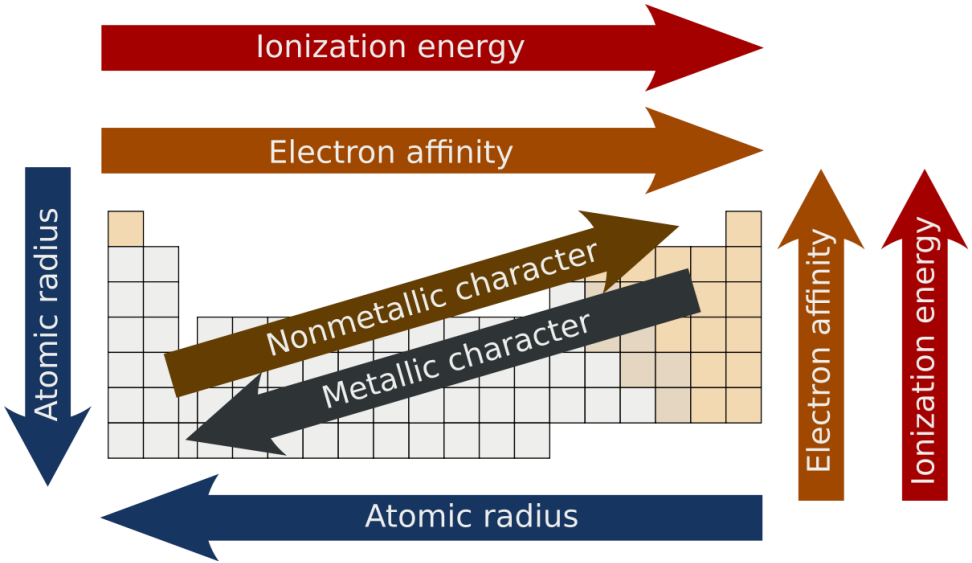
Important Note

Poisonous Gases

Phosgene	$COCl_2$
Tear gas	CCl_3NO_2
Mustard gas	$S(CH_2CH_2Cl)_2$



Hydrochloric acid



HCl

Preparation : $NaCl + H_2SO_4 \longrightarrow NaHSO_4 + HCl\ (\uparrow)$

$NaHSO_4 + NaCl \longrightarrow Na_2SO_4 + HCl\ (\uparrow)$

- HCl can be dried by conc. H_2SO_4
- HCl is colourless and has pungent smell and acidic

$NH_3 + HCl \longrightarrow NH_4Cl$

basic acidic

A colorless aqueous solution contains nitrates of two metals, **X** and **Y**. When it was added to an aqueous solution of NaCl, a white precipitate was formed. This precipitate was found to be partly soluble in hot water to give a residue **P** and a solution **Q**. The residue **P** was soluble in aq. NH_3 and also in excess sodium thiosulfate. The hot solution **Q** gave a yellow precipitate with KI. The metals **X** and **Y**, respectively, are

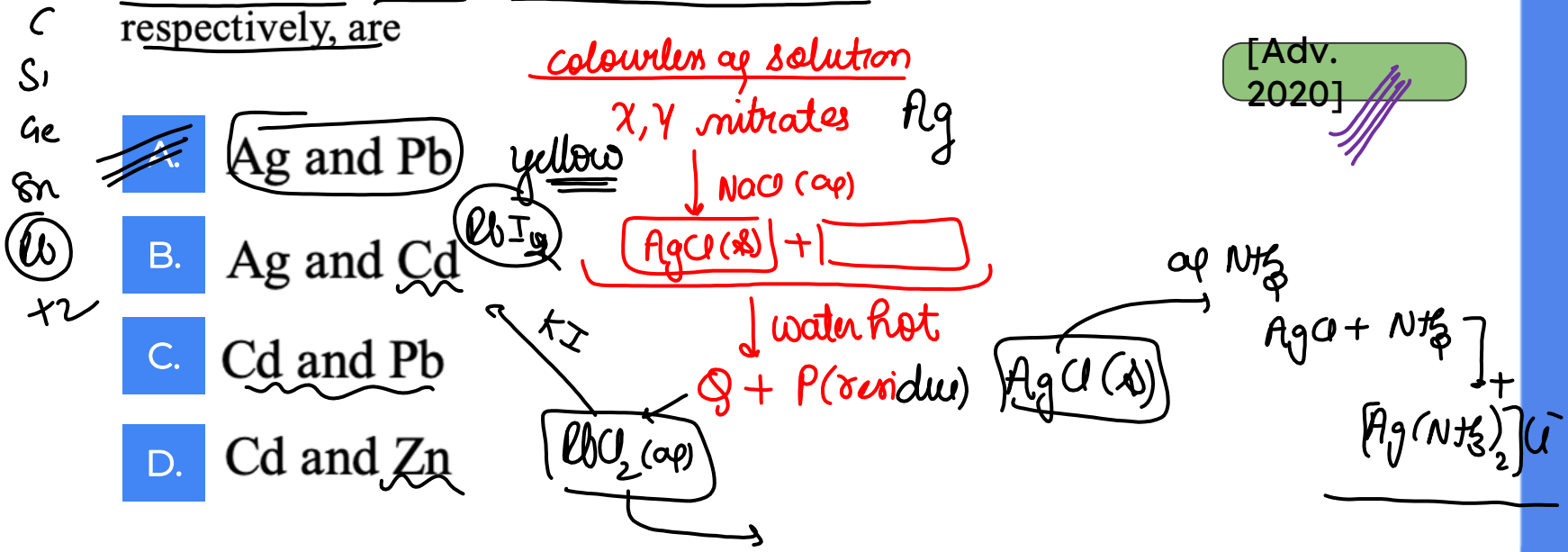
[Adv. 2020]

A. Ag and Pb

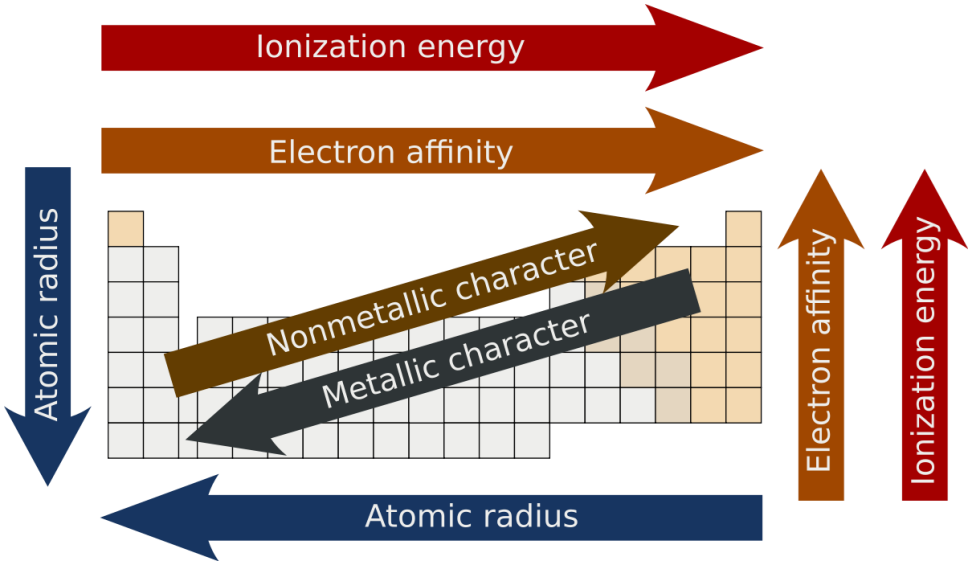
B. Ag and Cd

C. Cd and Pb

D. Cd and Zn



Aqua Regia


Ionization energy
Electron affinity
Atomic radius
Nonmetallic character
Metallic character
Electron affinity
Ionization energy
Atomic radius

Aqua Regia

- Conc. HCl : Conc. HNO_3

  3:1

- Very strong oxidising agent
- $Au + HCl + HNO_3 \longrightarrow H[AuCl_4] + NO + H_2O$
- $Pt + HCl + HNO_3 \longrightarrow H_2[PtCl_6] + NO + H_2O$

$HCl + HNO_3 \rightarrow [Cl]$

**

group 15

Oxyacids of Halogens

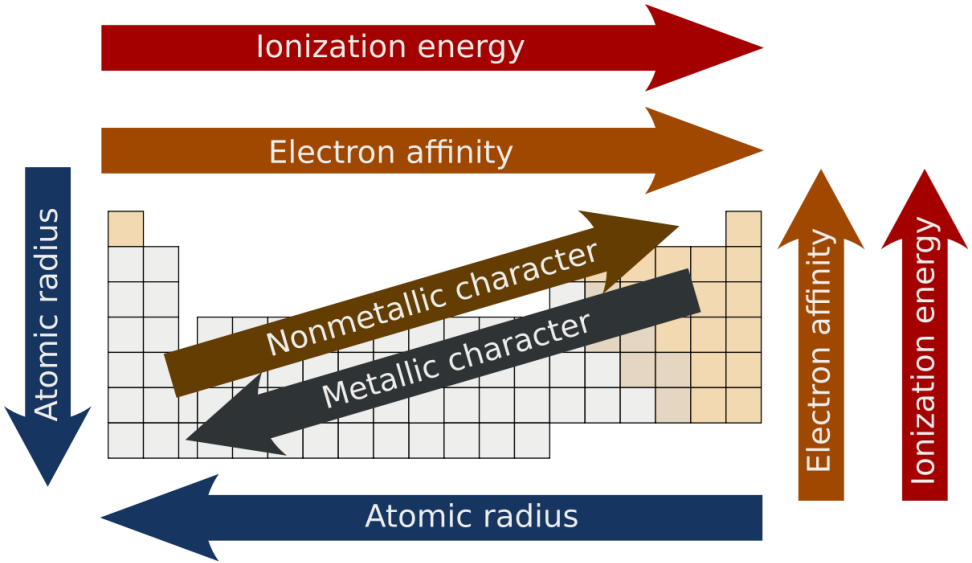
F

Halic (I) acid (Hypohalous acid)	HOF (Hypofluorous acid)	HOCl (Hypochlorous acid)	HOBr (Hypobromous acid)	HOI (Hypoiodous acid)
Halic (III) acid (Halous acid)	HFO_2	HOClO (chlorous acid)		
Halic (V) acid (Halic acid)	HFO_3	$HOClO_2$ (chloric acid)	$HOBrO_2$ (bromic acid)	$HOIO_2$ (iodic acid)
Halic (VII) acid (Perhalic acid)	HFO_4	$HOClO_3$ (perchloric acid)	$HOBrO_3$ (perbromic acid)	HIO_4 (periodic acid)

HIO_3

F–OH

Oxyacids of Halogens



Oxyacids of Halogens

→ $\overset{+7}{H}XO_4$ → Peric acid — $HClO_4$ Perchloric acid

→ $\overset{+5}{H}XO_3$ → ic acid — $HClO_3$ chloric acid

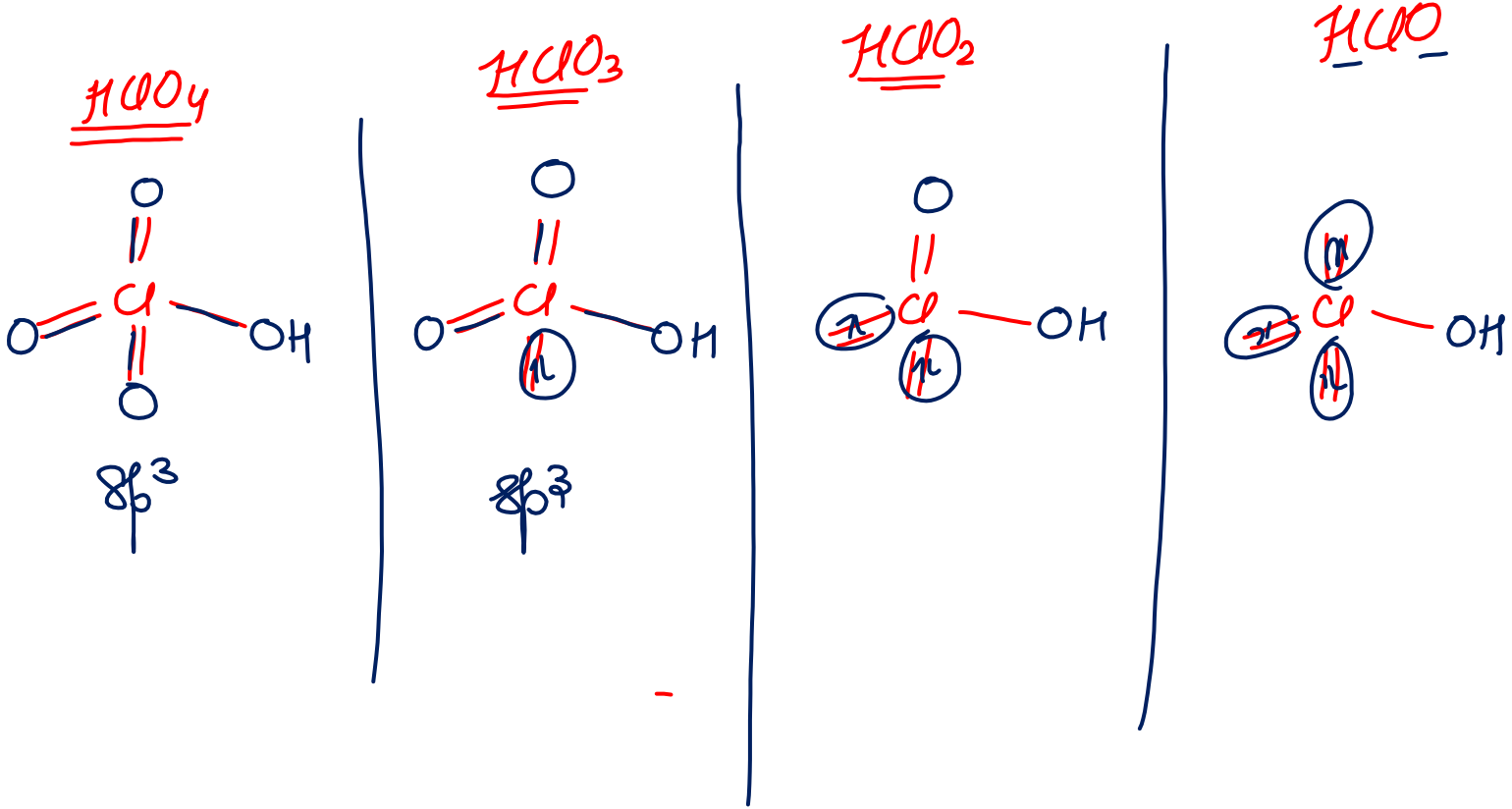
→ $\overset{+3}{H}XO_2$ → ous acid — $HClO_2$ chlorous acid

→ $\overset{+1}{H}XO$ → hypous acid — $HClO$ hypochlorous acid

max+2 Peric
max ic
max-2 ous
ous-2 hypous

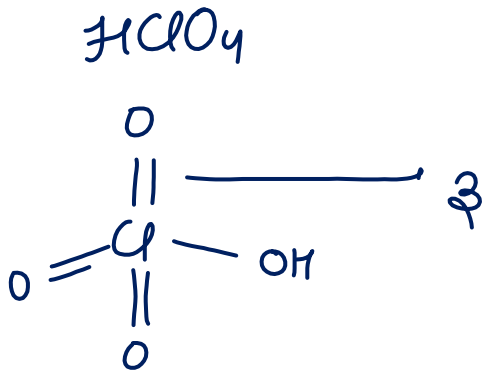
15/16

Oxyacids of Halogens



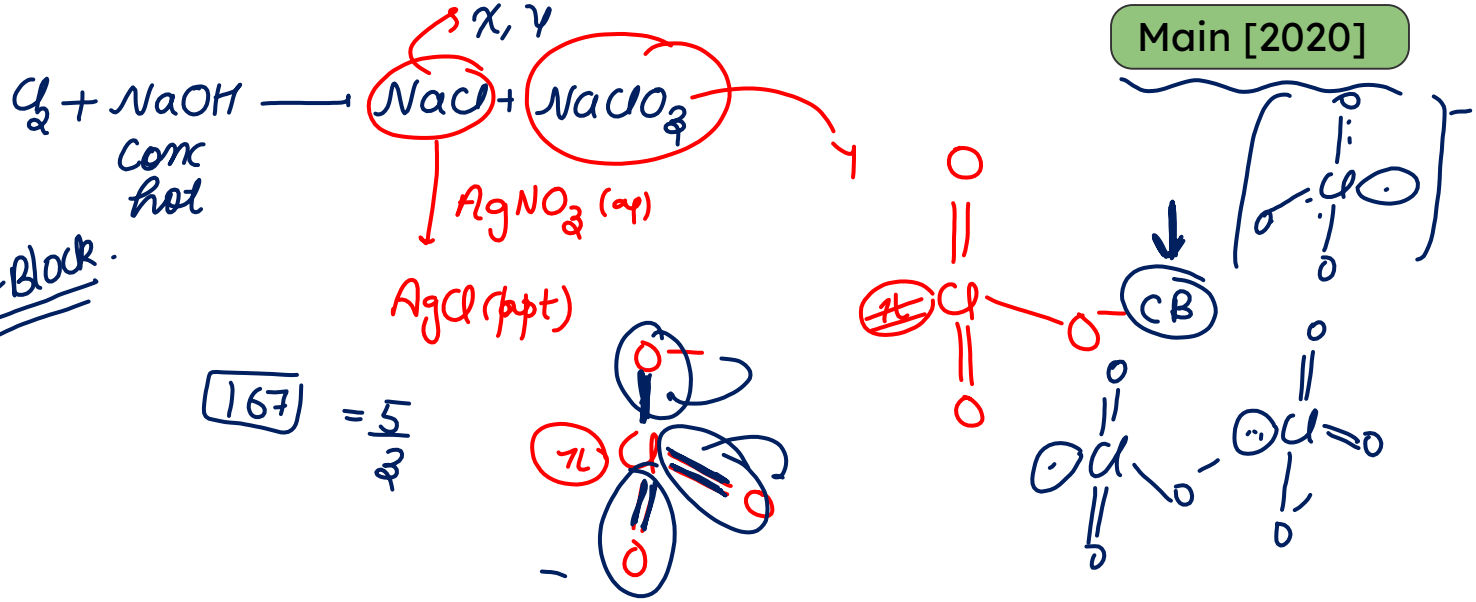
The number of Cl = O bonds in perchloric acid is,

[Main. 2020]



Chlorine reacts with hot and concentrated NaOH and produces compounds (*X*) and (*Y*). Compound (*X*) gives white precipitate with silver nitrate solution. The average bond order between Cl and O atoms in (*Y*) is _____.

Main [2020]



The correct statement(s) about the oxoacids, $HClO_4$ and $HClO$, is (are)

abd

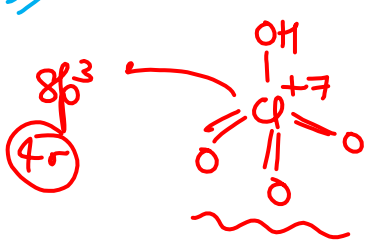
[Adv.2017]

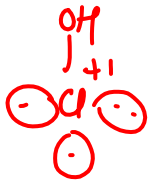
A. The central atom in both $HClO_4$ and $HClO$ is sp^3 hybridized

B. $HClO_4$ is more acidic than $HClO$ because of the resonance stabilization of its anion

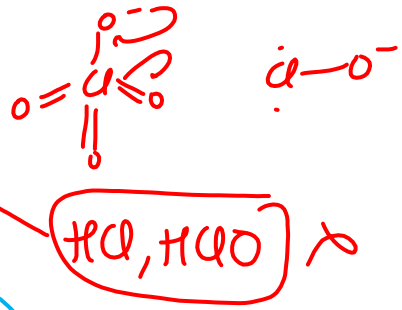
C. $HClO_4$ is formed in the reaction between Cl_2 and H_2O

D. The conjugate base of $HClO_4$ is weaker base than H_2O

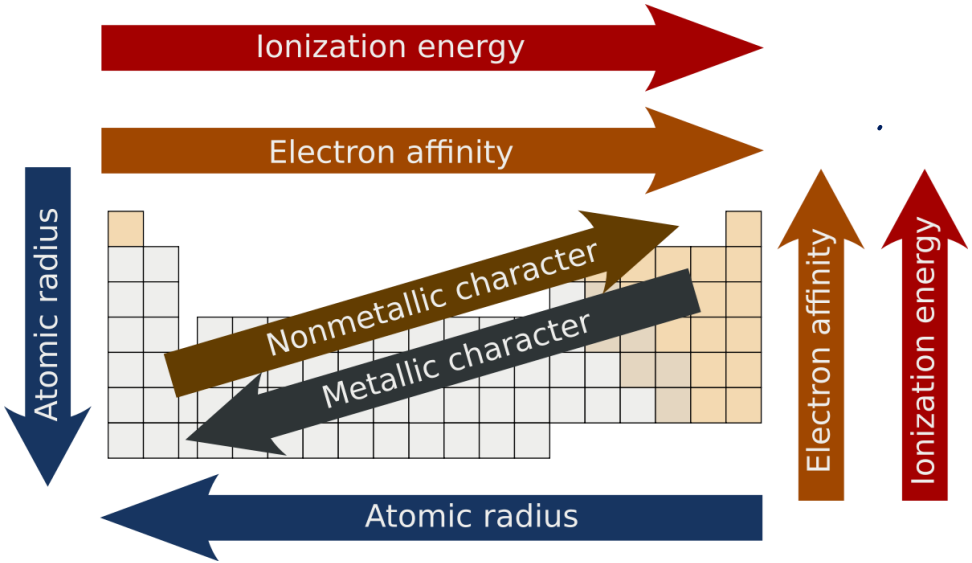




$1s + 3p = sp^3$



Group 18


Ionization energy
Electron affinity
Atomic radius
Nonmetallic character
Metallic character
Electron affinity
Ionization energy
Atomic radius

Group 18

He Ne Ar Kr Xe Rn

Ns^2 ns^2np^6

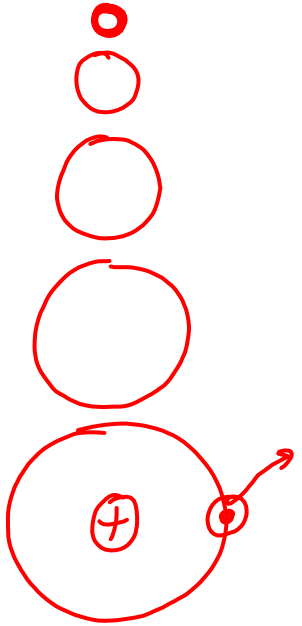
$1s^2$ ns^2

- **All are non metals**
- Rn is radioactive
- **Size** : He < Ne < Ar < Kr < Xe
- **Ionization energy** : He > Ne > Ar > Kr > Xe
- **Electron affinity** : He < Ne < Ar < Kr < Xe

Radon Rn ✓

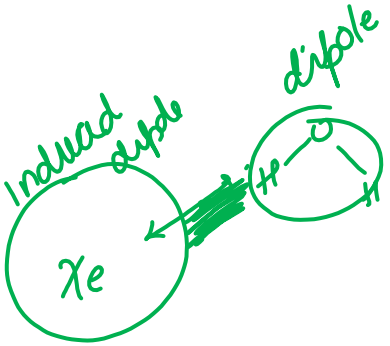
→ Radium Ra

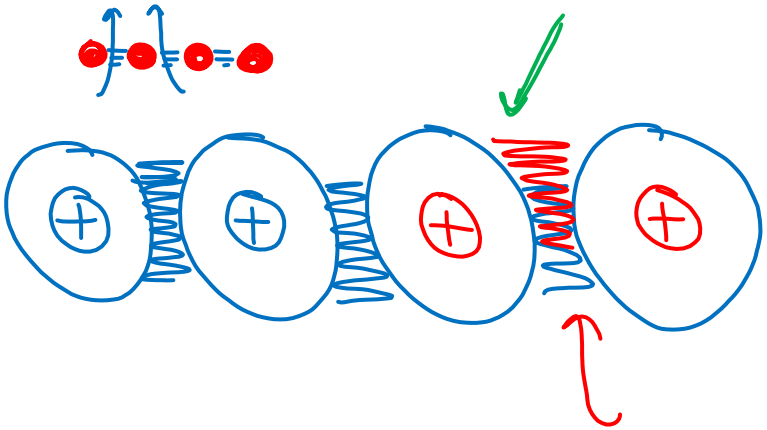
Vanderwaal Radius.



Group 18

- **Volatile nature** : He > Ne > Ar > Kr > Xe
- **Bpt / Mpt** : He < Ne < Ar < Kr < Xe
- **Ease of liquefaction** : He < Ne < Ar < Kr < Xe
- **Solubility in water** : He < Ne < Ar < Kr < Xe





The noble gas that does NOT occur in the atmosphere is :

A. He

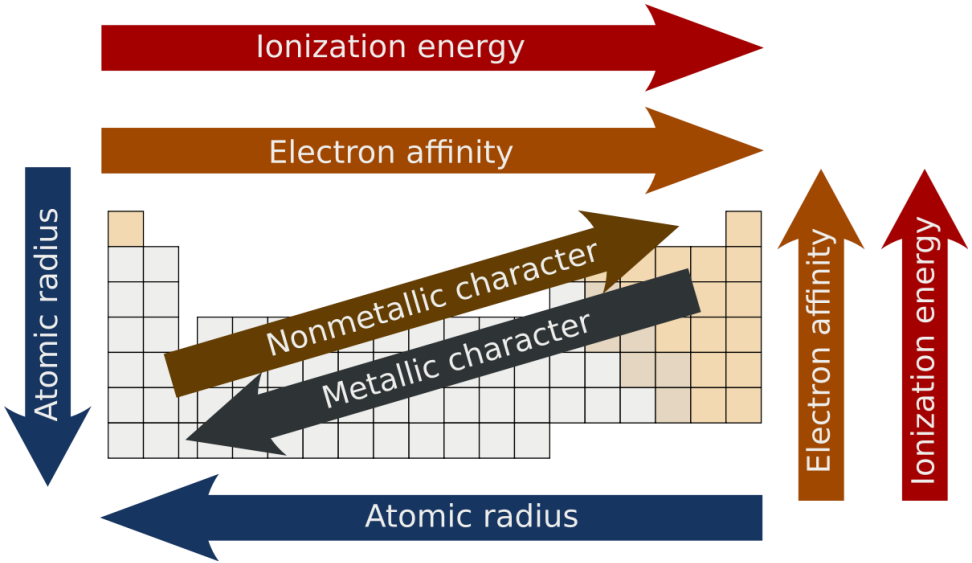
B. Kr

C. Ne

D. Ra — "Radium"

[April 10, 2019 (II)]

Chemical Reactions of Group 18



Chemical properties

Neil Bartlett

$O_2 + PtF_6 \longrightarrow O_2^+ [PtF_6]$

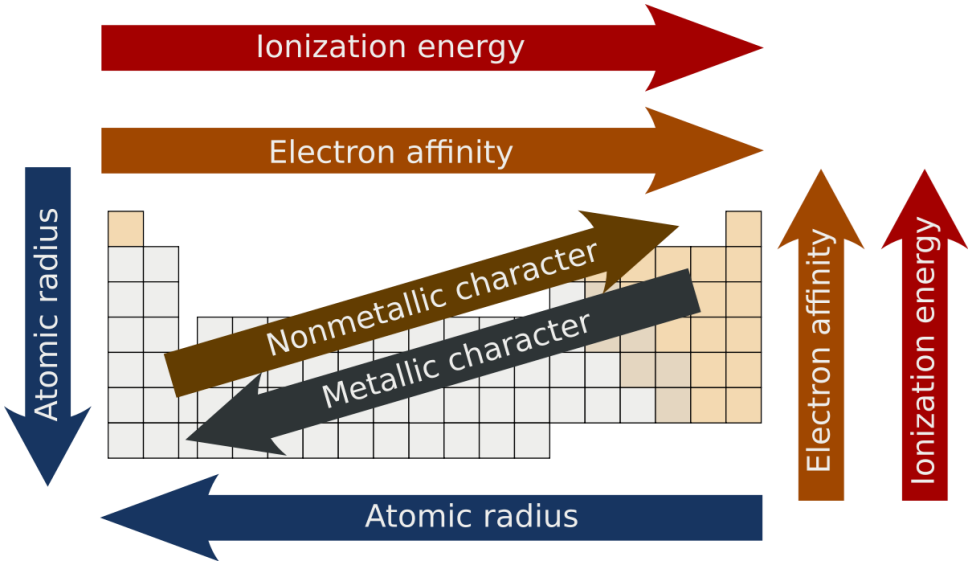
$Xe + PtF_6 \longrightarrow Xe^+ [PtF_6]$ red colour

He
Ne
Ar
Kr
Xe

$O_2 + PtF_6 \longrightarrow O_2^+ [PtF_6]^-$

$Xe + PtF_6 \longrightarrow [Xe]^+ PtF_6$

Fluorides of Xenon



Xenon Fluorides

excess

$Xe + F_2 \rightarrow XeF_2$

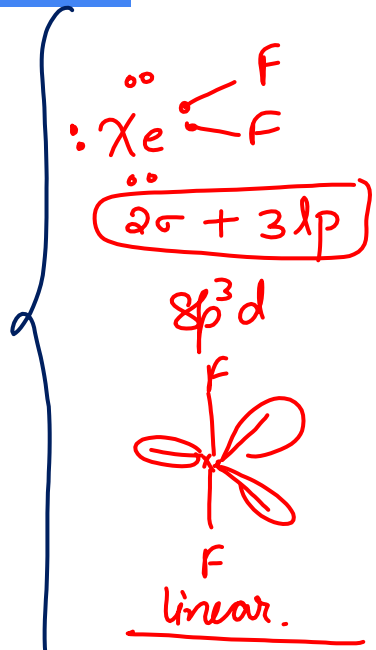
2 : 1

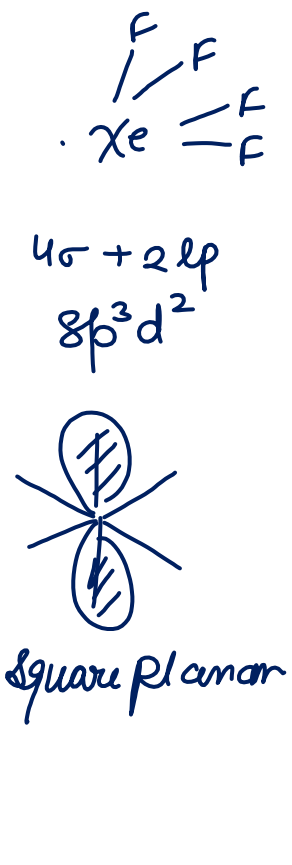
$Xe + F_2 \rightarrow XeF_4$

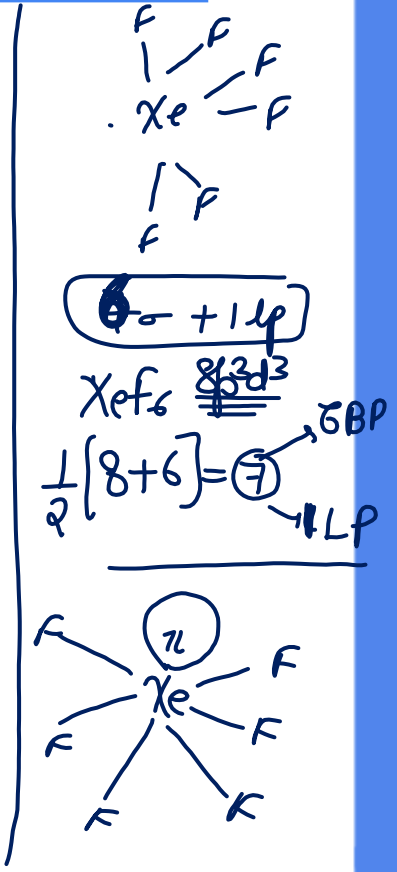
1 : 5

$Xe + F_2 \rightarrow XeF_6$

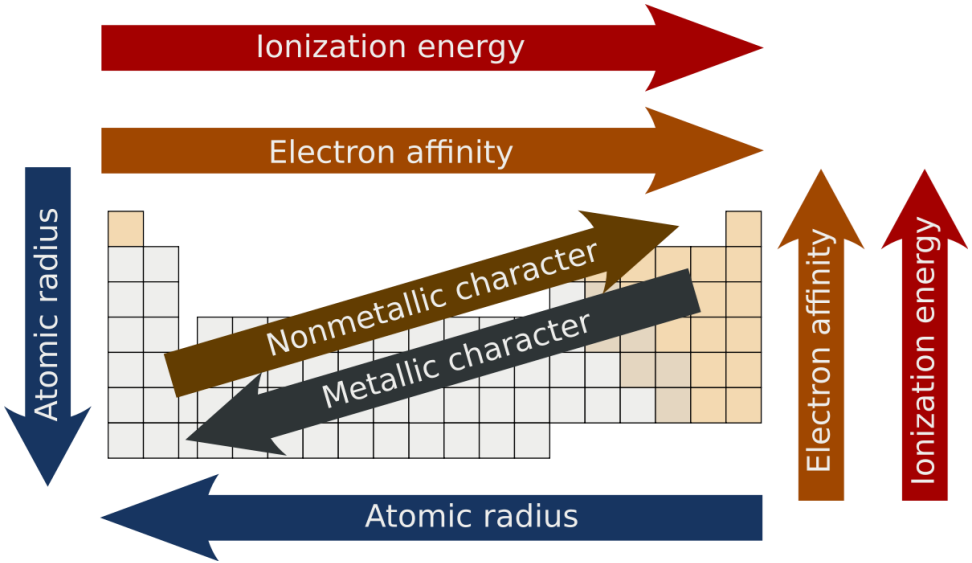
1 : 20



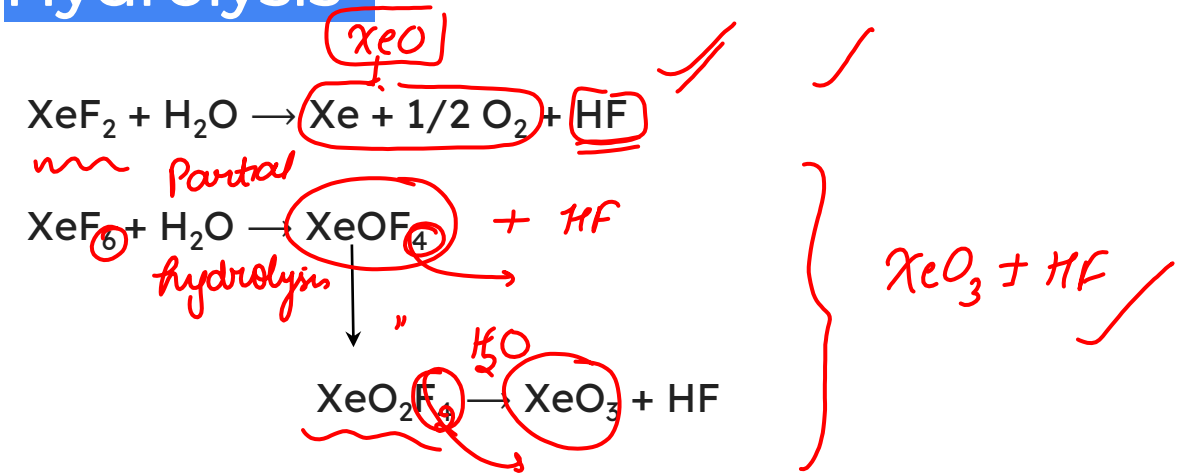




Hydrolysis of Xe Fluorides



Hydrolysis



$XeF_4 + H_2O \longrightarrow Xe + O_2 + HF + XeO_3$

The type of hybridisation and number of lone pair(s) of electrons of Xe in $XeOF_4$, respectively, are:

A. sp^3d^2 and 1

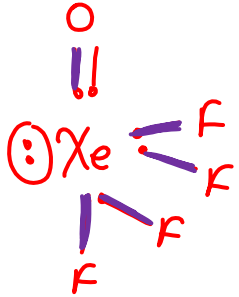
B. sp^3d and 2

C. sp^3d^2 and 2

D. sp^3d and 1

[Jan 10, 2019 (I)]

$XeOF_4$



$5\sigma + 1lp = 6$

sp^3d^2

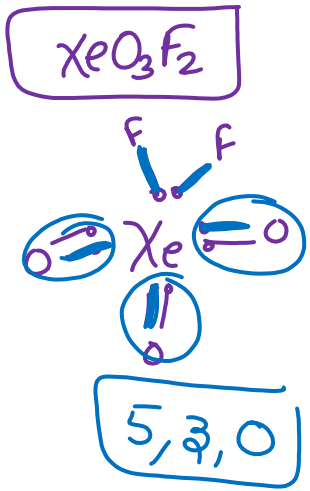
In XeO_3F_2, the number of bond pair(s), π-bond(s) and lone pair(s) on Xe atom respectively are:

A. 5, 3, 0

B. 5, 2, 0

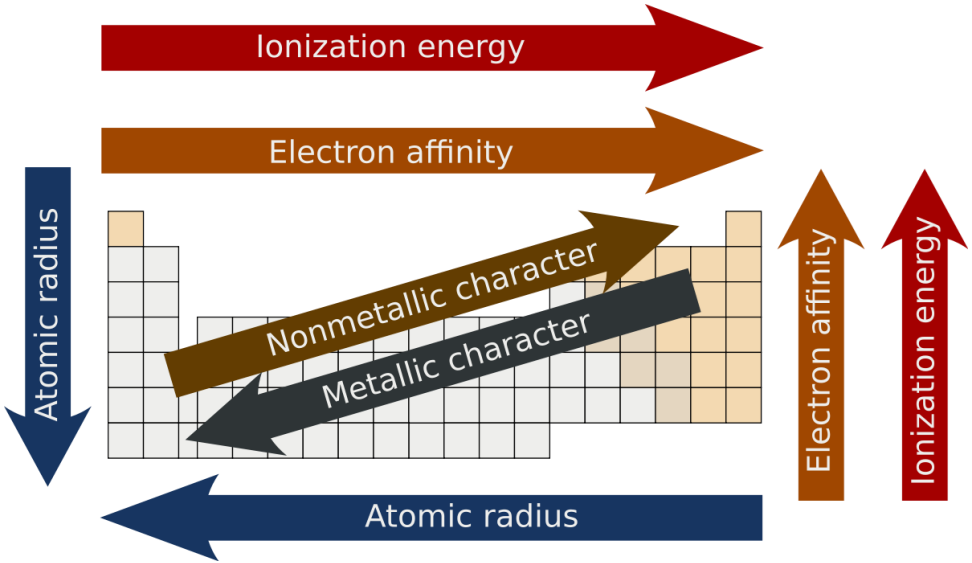
C. 4, 2, 2

D. 4, 4, 0

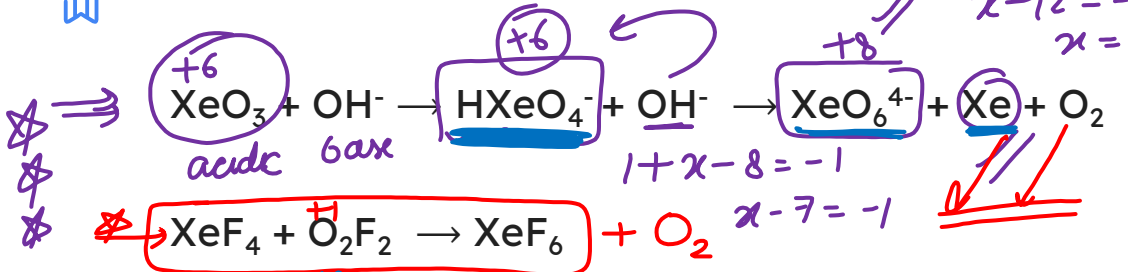


[April 15, 2018 (II)]

Reactions of Xe compounds



Reactions


+6
XeO3 + OH- → HXeO4- + OH- → XeO6 4- + Xe + O2
+6
+8
x-12=-4
x=8
acidic
base
1+x-8=-1
x-7=-1
XeF4 + O2F2 → XeF6 + O2
+1
+2
PF3 + XeF2 → PF5 + Xe
KF + XeF4 → K+ [XeF5]-
PF5 + XeF4 → [XeF3]+ [PF6]-
x-5=0
x=+5
PF6-
x-6=-1
x=+5

The reaction in which the hybridisation of the underlined atom is affected is :

CB

gp 15

A. $\underline{H_3PO_2} \xrightarrow{\text{Disproportionation}} H_3PO_4 + PH_3$

gp 17/16

B. $H_2\underline{S}O_4 + NaCl \xrightarrow{420\ K} Na_2SO_4 + HCl$

gp 15 sp^3

C. $\underline{N}H_3 \xrightarrow{H^+} NH_4$ sp^3

sp^3 gp 18

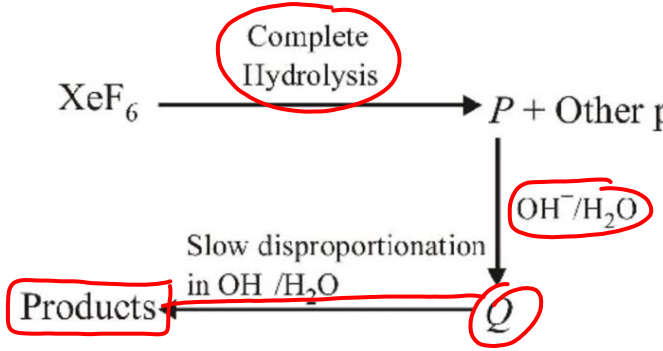
D. $\underline{Xe}F_4 + SbF_5 \longrightarrow [XeF_3]^+[SbF_6]^-$

JEE Main 2020

$=Xe:$ 4+2 sp^3d^2

$O=Xe=O$ 3+2 = sp^3d

Under ambient conditions, the total number of gases released as products in the final step of the reaction scheme shown below is



[Adv. 2014]

A. 0

B. 1

C. 2

D. 3

$XeF_6 \longrightarrow XeO_3$ (P) + other products

$XeO_3 \xrightarrow{OH^-/H_2O} HXeO_4^-$ (Q)

$HXeO_4^-$ (+6) $\xrightarrow{OH^-}$ $Xe + O_2 + XeO_6^{4-}$

g, g, 2

At 143 K, the reaction of XeF_4 with O_2F_2, produces a xenon compound *Y*. The total number of lone pair(s) of electron present on the whole molecule of *Y* is _____

19

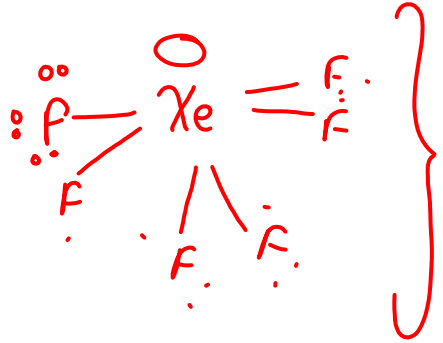
4

$XeF_4 + O_2F_2 \longrightarrow XeF_6 + O_2$

[Adv. 2019]

$XeF_6 \longrightarrow \frac{1}{2}[8+6] = 7$ → BP = 6, LP = 1

sp^3d^3



$3 \times 6 = 18$

$+ 1$

19

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Rapid Revision Batch

For JEE Main and Advanced 2022

17th November,2021

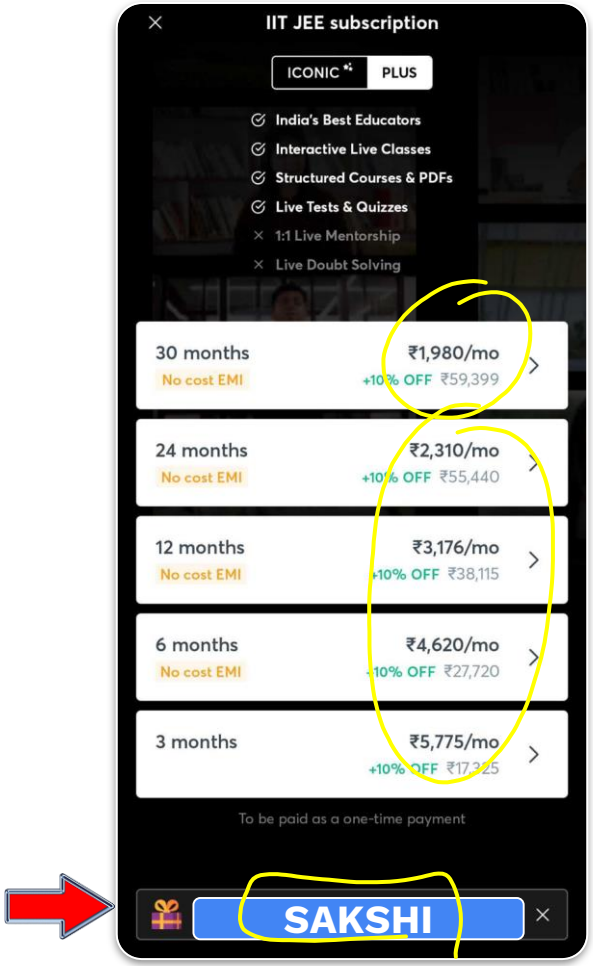

Nishant Vora


Praveen Kumar Pachuri


Mohammad Kashif Alam


Sakshi Ganotra


Ashish Bibyan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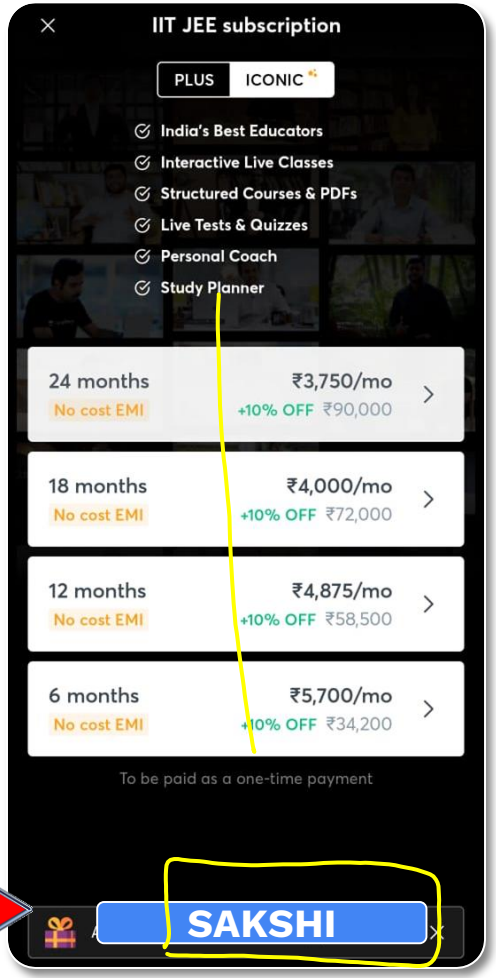
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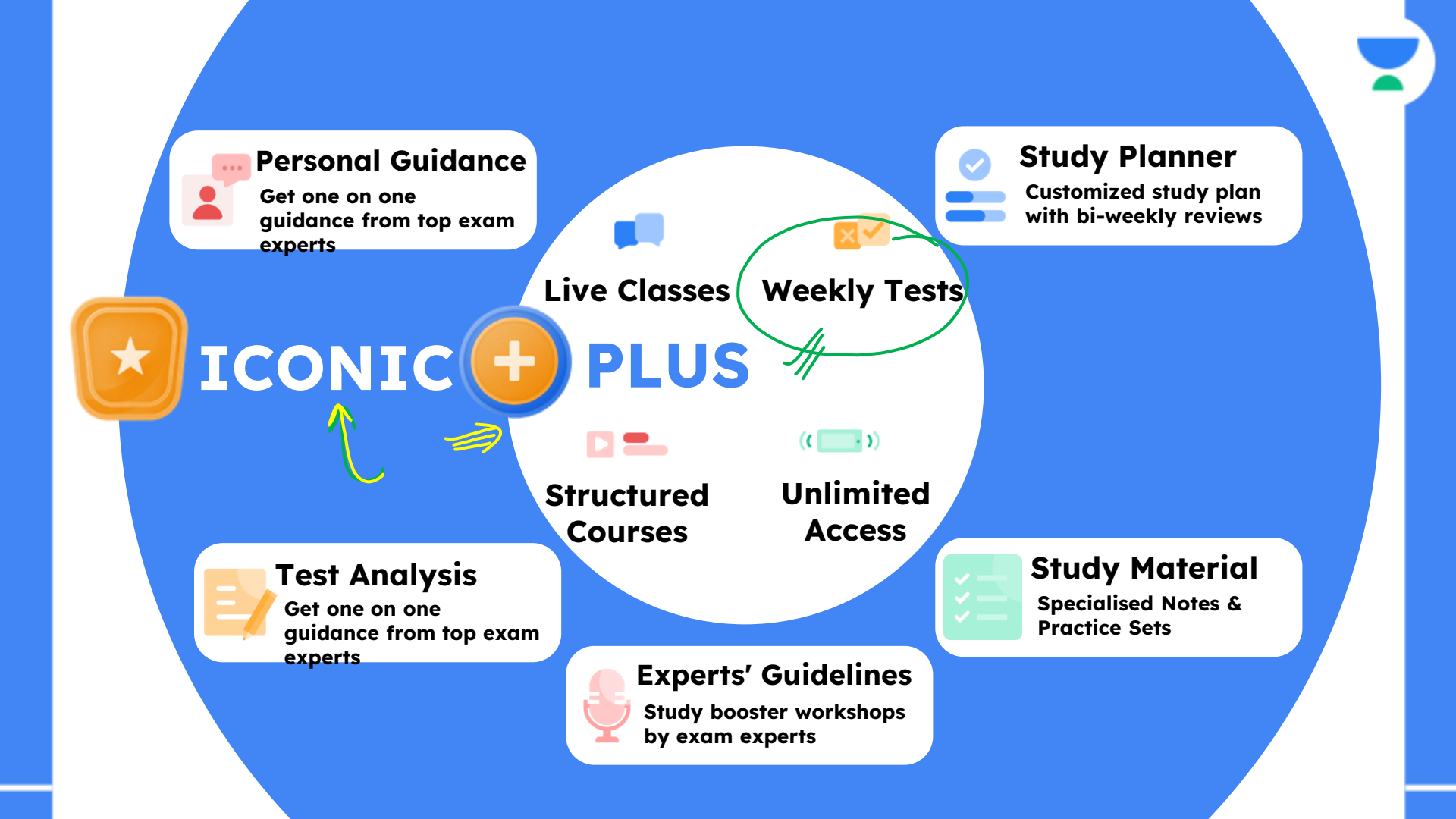
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Nov 13th | 8:00 PM

Nov 20th | 8:00 PM

SAKSHI

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Unacademy Champions League

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All India Level Test

Win Prizes worth 5 Crores

Scholarship test in Aptitude format

TEST	DATES	QUESTIONS	TIME
GROUP STAGE	14th & 21st November	25	1-2 PM
FINALS	28th November	75	1-4 PM

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GROUP STAGE - Prizes & Scholarships

RANK	PRIZES* EACH	SCHOLARSHIP %
1-10	₹20,000	100%
11-30	₹10,000	50%
31-50	₹6,000	25%
50-200	--	15%

FINALS - Prizes & Scholarships

RANK	PRIZES* EACH	SCHOLARSHIP %
1	₹1,00,000	100%
2 & 3	₹60,000	100%
4-10	₹50,000	100%
11-30	--	50%
31-50	--	25%
51-200	--	15%

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*Amazon Gift Vouchers

UNACADEMY COMBAT

IIT-JEE MEGA COMBAT

21 NOVEMBER | 11 AM

CLASS 11 | CLASS 12 | CLASS 13

60 MINUTES | 21 QUESTIONS

unacademy

UPTO 100% SCHOLARSHIP WORTH

₹10 CRORE*+

*T&C applied

Rank 1-3: 100% Scholarship AV worth 3000*

Rank 4-10: 75% Scholarship AV worth 2000*

Rank 11-50: 50% Scholarship

Rank 51-150: 25% Scholarship

Top 50 Rankers and 500 Lucky Participants to win SURPRISE GIFTS !!!

I-PAD Mini

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ANDROID APP AND WEB

LIVE Test

*T&C applied

Thank You!