

**17** A fixed amount of ideal gas undergoes the following processes:

Process 1: Positive work is done by gas at constant pressure

Process 2: Gas is allowed to expand while heat is supplied slowly to it

Process 3: Gas undergoes an adiabatic (i.e. no heat supplied or removed) expansion

How would the internal energy of the gas change for each of the processes?

	process 1	process 2	process 3
<b>A</b>	decrease	no change	decrease
<b>B</b>	decrease	increase	increase
<b>C</b>	increase	increase	decrease
<b>D</b>	increase	no change	decrease

