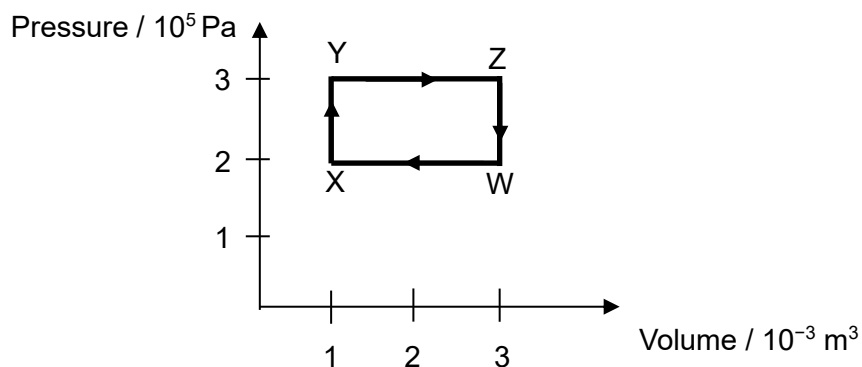


- 16** An ideal gas undergoes the cycle of pressure and volume changes  $W \rightarrow X \rightarrow Y \rightarrow Z \rightarrow W$ , as shown below. There is a decrease in internal energy of 600 J from W to X and an increase in internal energy of 150 J from X to Y.



Which of the following set of information is correct for one complete cycle?

	Net increase in internal energy / J	Net work done on gas / J	Net heat absorbed by gas / J
<b>A</b>	0	200	-200
<b>B</b>	0	-200	200
<b>C</b>	450	-200	650
<b>D</b>	-450	200	-650