

- 19** An ideal gas undergoes an expansion in volume from $1.3 \times 10^{-4} \text{ m}^3$ to $3.6 \times 10^{-4} \text{ m}^3$ at a constant pressure of $1.3 \times 10^5 \text{ Pa}$. During the expansion, 24 J of heat is supplied to the gas.

What is the overall change in the internal energy of the gas?

- A** decrease of 54 J
- B** decrease of 6 J
- C** increase of 6 J
- D** increase of 54 J