1 The Clausius-Clapeyron equation for measuring relative vapour pressure of a gas $\frac{P}{P_o}$, at a temperature T is as follows:

$$\frac{P}{P_o} = e^{-\frac{\Delta H}{R}\left(\frac{1}{T}-\frac{1}{T_o}\right)}$$

where ΔH is the enthalpy of vaporization, R is the molar gas constant, P and P_o are the final and initial pressures and T_o is the initial temperature of a gas.

Which of the following has the same units as ΔH?

A energy per mole

B molar gas constant

C temperature

D no units

2 A student made a series of measurements of the diameter d of a wire using four micrometer