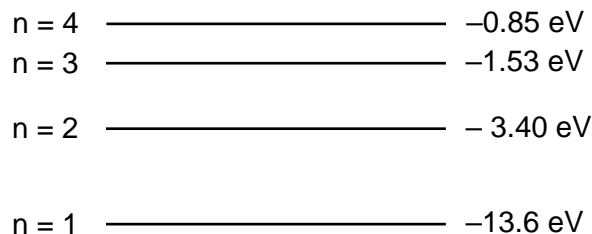The figure shows the four lowest energy levels of a hydrogen atom. It is known that the wavelength of visible light ranges from 400 nm to 700 nm.

n = 4	———————————	−0.85 eV
n = 3	———————————	−1.53 eV
n = 2	———————————	− 3.40 eV
n = 1	———————————	−13.6 eV

If electrons having kinetic energy of 12.5 eV are used to bombard cool hydrogen gas, how many spectral lines in the visible region can be obtained subsequently?

A 0 **B** 1 **C** 2 **D** 3