

- 17** A fixed amount of ideal gas undergoes the following processes:

Process 1: Positive work is done by gas at constant pressure

Process 2: Gas is allowed to expand while heat is supplied slowly to it

Process 3: Gas undergoes an adiabatic (i.e. no heat supplied or removed) expansion

How would the internal energy of the gas change for each of the processes?

	process 1	process 2	process 3
A	decrease	no change	decrease
B	decrease	increase	increase
C	increase	increase	decrease
D	increase	no change	decrease

