13 Thermal energy is supplied to 2.1 kg of ice at −4.5 °C. All the ice becomes water at 96 °C. There is no heat loss to the surroundings.

Specific heat capacity of water = 4190 J kg^{-1} K^{-1}
Specific latent heat of vaporization of water = 2260 kJ kg^{-1}
Specific heat capacity of ice = 2108 J kg^{-1} K^{-1}
Specific latent heat of fusion of water = 334 kJ kg^{-1}

How much thermal energy is supplied to the ice?

A 721 kJ **B** 865 kJ **C** 1.55 MJ **D** 1.57 MJ