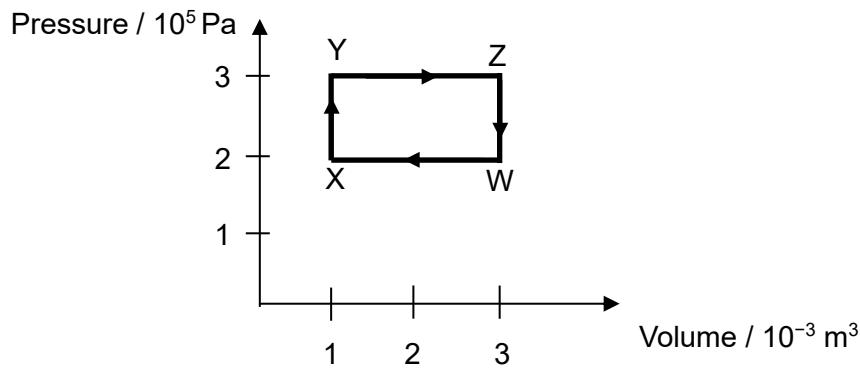


- 16 An ideal gas undergoes the cycle of pressure and volume changes $W \rightarrow X \rightarrow Y \rightarrow Z \rightarrow W$, as shown below. There is a decrease in internal energy of 600 J from W to X and an increase in internal energy of 150 J from X to Y .



Which of the following set of information is correct for one complete cycle?

	Net increase in internal energy / J	Net work done on gas / J	Net heat absorbed by gas / J
A	0	200	-200
B	0	-200	200
C	450	-200	650
D	-450	200	-650