

H3-2022: Extra Practice Question Set 1

1. The unit cell of a crystal, with a face-centered cubic (fcc) structure, is shown in Figure 1. The lattice constant is $a = 5.5 \text{ \AA}$ and there is one atom at each lattice point.

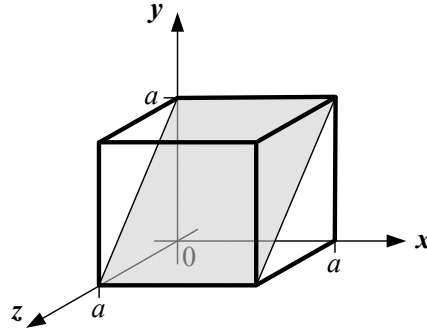


Figure 1

- (a) Write down the Miller indices of the shaded plane.
- (b) Calculate the surface density of atoms (in units of atoms/cm²) on the (010) plane.

[Ans: (a) (011), (b) $6.612 \times 10^{14} \text{ cm}^{-2}$]

2. An intrinsic semiconductor with a bandgap energy of 2.1 eV is uniformly illuminated with light. If the wavelength of light is 0.7 \mu m , will excess electron-hole pairs be produced in the semiconductor? Justify your answer.
3. Figure 3 shows a shaded plane in a body-centered cubic (bcc) crystal structure with a lattice constant of a . There is one atom at each lattice point of the crystal structure.

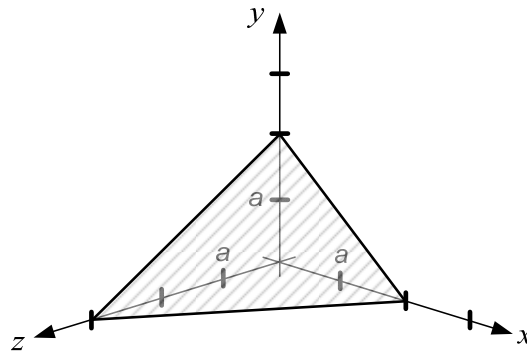


Figure 3

- (a) Determine the Miller indices of the shaded plane.
- (b) The surface density of atoms on the (111) plane of the crystal is $2.506 \times 10^{14} \text{ cm}^{-2}$. Find the value of a .
- (c) Calculate the shortest distance between any two atoms in this crystal structure.

[Ans: (a) (3 3 2), (b) 4.8 \AA , (c) 4.157 \AA]

4. Figure 4 shows the Si unit cell with a lattice constant of a . The surface density of atoms on the (010) plane is $6.783 \times 10^{14} \text{ cm}^{-2}$.

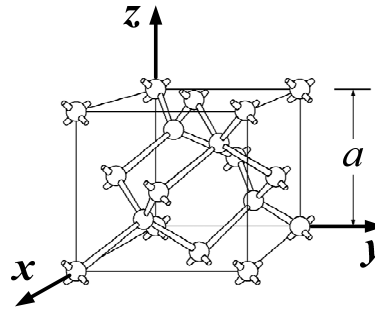


Figure 4

- Determine the value of a .
- Sketch the (110) plane and indicate clearly the position of all the atoms on the plane.
- Hence calculate the surface density of atoms on the (110) plane.

[Ans: (a) 5.43 \AA , (c) $9.59 \times 10^{14} \text{ cm}^{-2}$]

5. The position of atoms on the (111) plane of a hypothetical semiconductor is shown in Figure 5. The surface density of atoms on the plane is $1.45 \times 10^{15} \text{ cm}^{-2}$.

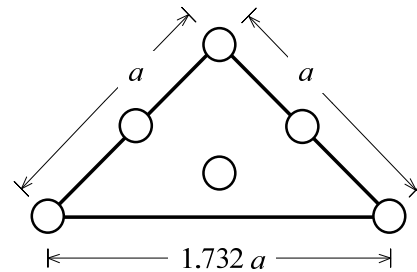


Figure 5

- Determine the number of atoms in the plane.
- What is the value of a ?

[Ans: (a) 2.5, (b) 6.392 \AA]

6. Figure 6 shows the conduction band of two semiconductor samples. Which sample has a smaller electron effective mass? Justify your choice.

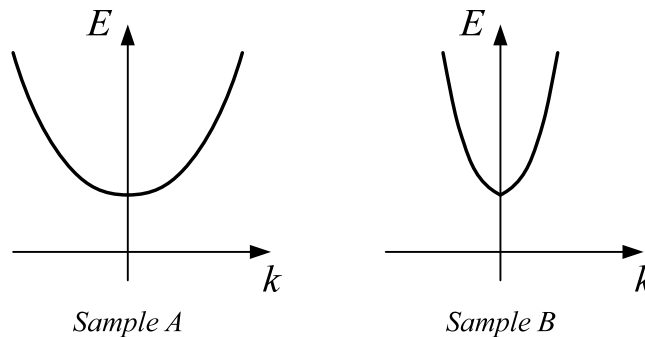


Figure 6