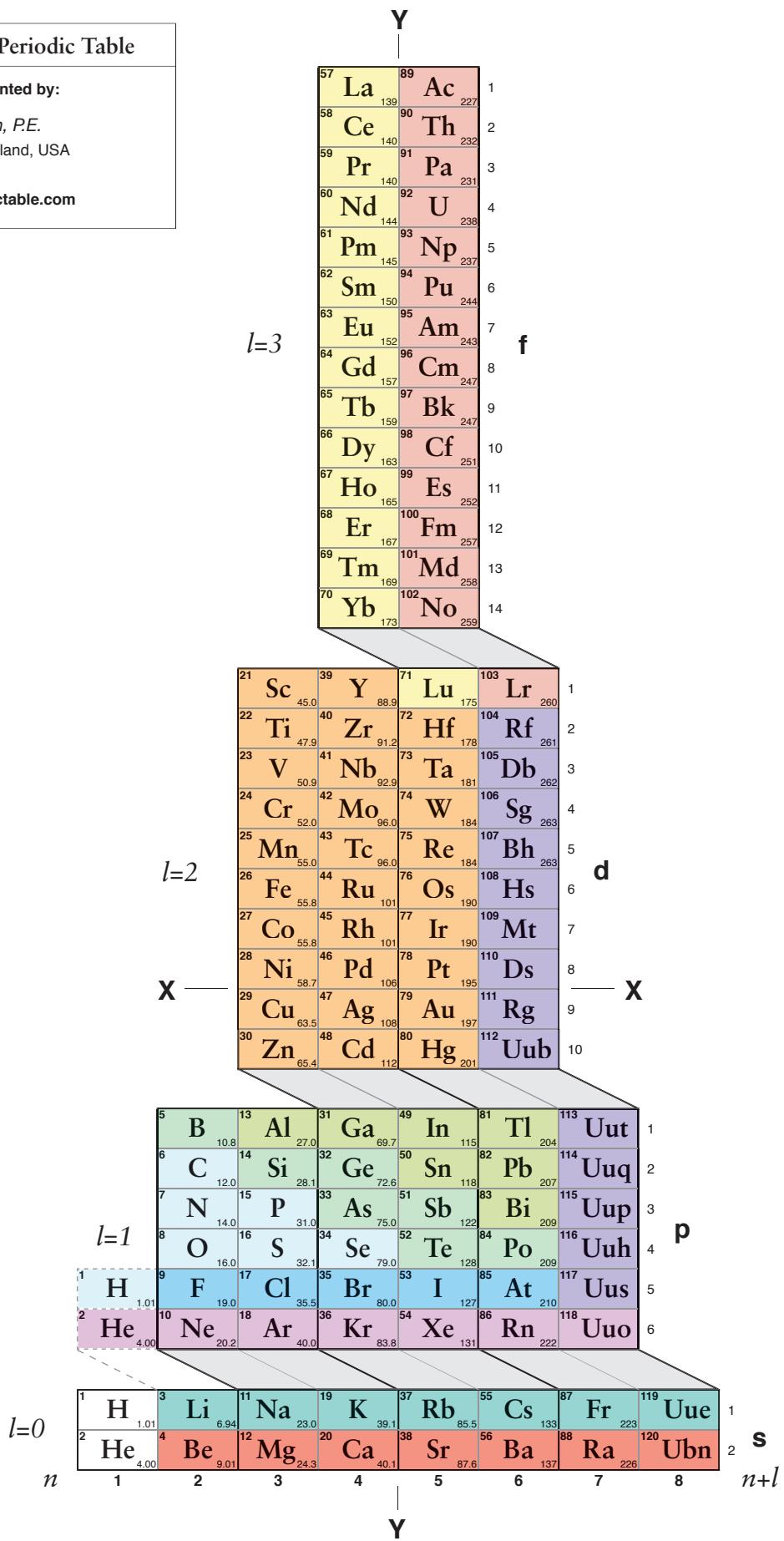


ADOMAH Periodic Table	
Presented by:	
Valery Tsimmerman, P.E.	
Howard County, Maryland, USA	
www.perfectperiodictable.com	



The diagram illustrates the Aufbau principle, showing the filling of atomic orbitals by increasing principal quantum number (n) and then by increasing azimuthal quantum number (l). The orbitals are represented as horizontal boxes.

- Left Column ($l=0$):** Contains 1s orbitals. The 1s orbital at $n=1$ is filled with 2 electrons. The 1s orbital at $n=2$ is filled with 2 electrons. The 1s orbital at $n=3$ is empty.
- Middle Column ($l=2$):** Contains 3d orbitals. The 3d orbital at $n=4$ is filled with 10 electrons. The 3d orbital at $n=5$ is filled with 10 electrons. The 3d orbital at $n=6$ is filled with 10 electrons. The 3d orbital at $n=7$ is empty.
- Right Column ($l=1$):** Contains 2p and 3p orbitals. The 2p orbital at $n=1$ is filled with 6 electrons. The 3p orbital at $n=2$ is empty. The 3p orbital at $n=3$ is filled with 6 electrons. The 4p orbital at $n=4$ is empty. The 5p orbital at $n=5$ is filled with 6 electrons. The 6p orbital at $n=6$ is empty. The 7p orbital at $n=7$ is filled with 6 electrons.
- Bottom Row ($l=3$):** Contains 4f and 5f orbitals. The 4f orbital at $n=6$ is filled with 14 electrons. The 5f orbital at $n=7$ is empty.

STANDARD TABLE SHOWN IN RESPECT TO FILLING OF ORBITALS

JANET TABLE SHOWN IN RESPECT TO FILLING OF ORBITALS

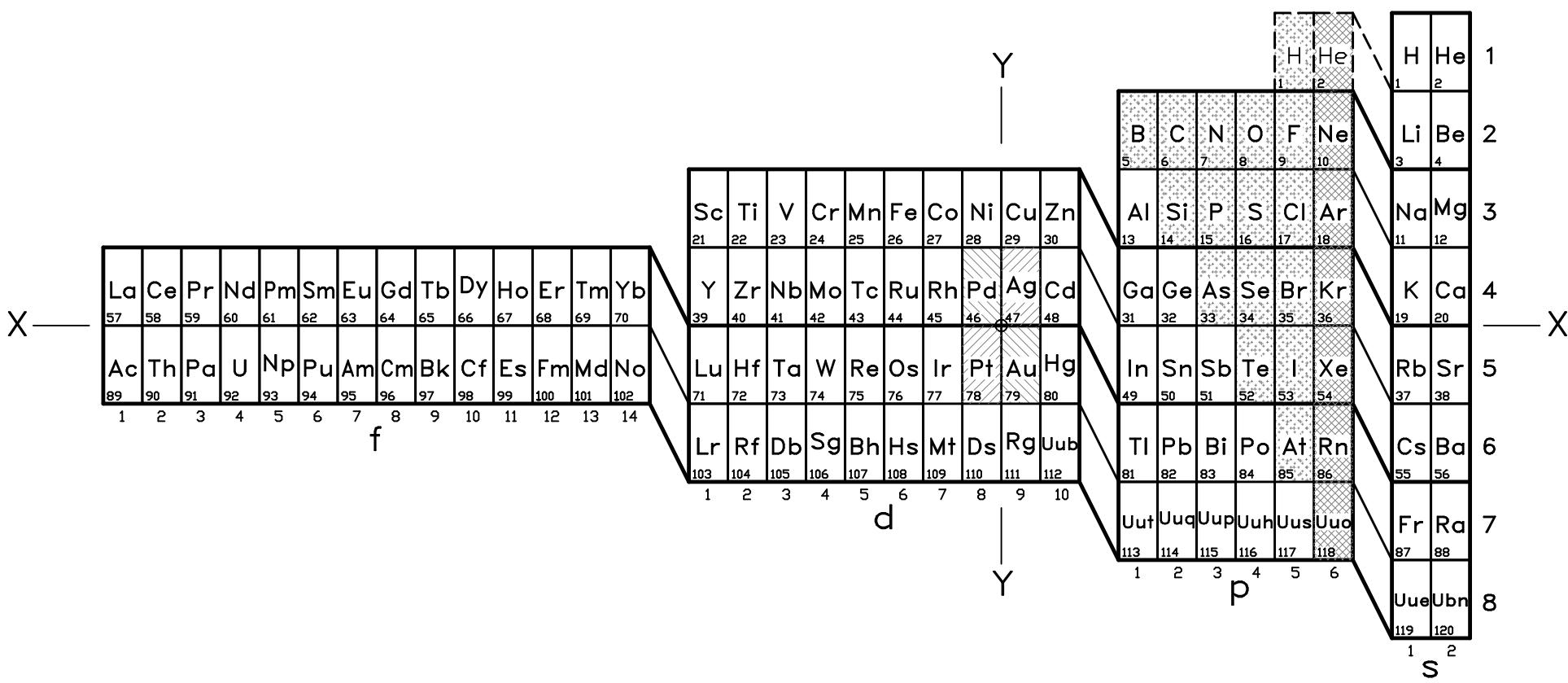
The diagram illustrates the Aufbau principle, mapping atomic orbitals onto a series of nested shells. The shells are labeled by the principal quantum number n (from 1 to 8) on the right. Within each shell, subshells are labeled by the azimuthal quantum number ℓ (from 0 to 3). The subshells and their corresponding atomic orbitals are:

- $\ell=0$: 1s (one orbital)
- $\ell=1$: 2s (one orbital), 2p (three orbitals)
- $\ell=2$: 3s (one orbital), 3p (three orbitals), 3d (five orbitals)
- $\ell=3$: 4s (one orbital), 4p (three orbitals), 4d (five orbitals), 4f (seven orbitals)

The shells are represented as nested rectangles, and the subshells within each shell are also nested rectangles. The angular momentum quantum number m_ℓ is implicitly defined by the orientation of the subshell boxes.

ADOMAH TABLE (MODIFIED FOR COMPARISON)
SHOWN IN RESPECT TO FILLING OF ORBITALS

Organized by: Valery Tsimerman, P.E., ORAH Constructive Technologies, Inc.
Howard County, Maryland, USA
Tel: (410) 442-4658, Fax: (410) 442-5726
email: orahct@comcast.net



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