

# Specific Heat Capacity of Metals

## Period 11

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### 1 Objective

The objective of this experiment is to measure the specific heat capacity of three different samples of metal and to compare those with the accepted values. The samples consist of aluminum, zinc and copper.

### 2 Definitions

**Heat** Heat is the measure of the internal kinetic energy of a substance.

**Temperature** Temperature is a measure of the kinetic energy of a particle. It is the degree or intensity of heat in a substance. Celcius is a unit of temperature. One degree Celcius represents the temperature change of one gram of water when  $2.39 \times 10^{-5}$  Joules of heat is added to it.

**Specific Heat Capacity** The specific heat capacity is the energy transferred to one kilogram of substance causing its temperature to increase by one degree Celcius.Homer (2014)

**Thermal Equilibrium** Thermal equilibrium is a condition where two substances in physical contact with each other exchange no net heat energy. Substances in thermal equilibrium are at the same temperature.

### 3 Theory

The change in the internal energy of an object or substance is equal to the product of the mass and the specific heat capacity and the change in temperature.

$$\Delta U = mC_p\Delta T$$

When water and the metal samples are in thermal equilibrium the change in heat of the water is equal in magnitude to the change in heat of the metal.

$$\Delta U_{metal} = \Delta U_{water}$$

From this relationship we may derive a formula for the specific heat capacity of the metal sample given the mass of metal, mass of water, change in temperature of the water, change in temperature of the metal and the specific heat capacity of water.

$$m_{metal}C_{metal}\Delta T_{metal} = m_{water}C_{water}\Delta T_{water}$$

$$C_{metal} = \frac{m_{water}}{m_{metal}} \frac{\Delta T_{water}}{\Delta T_{metal}} C_{water}$$

## 4 Materials

- Kettle
- Short bar, long bar, cube
- styrofoam cups
- graduated cylinder
- scale
- thermometer
- tongs
- flask of water

## 5 Method

- a. Weigh the samples and record
- b. Measure 350 ml of water in graduated cylinder and transfer to styrofoam cup
- c. Measure the initial temperature of the water
- d. Boil water and add metal samples to kettle
- e. Use tongs to transfer a sample to the cup with water
- f. Place thermometer in cup, cover it, stir and record equilibrium temperature
- g. Repeat steps b-f for each sample

## 6 Data

Metal	Mass Metal	Temp Water Initial	Temp Final
cube	90.6 g	350 g	20.5 Celcius 24.5 Celcius
short bar	64.1 g	350 g	20.9 Celcius 22.5 Celcius
long bar	203.0 g	300 g	20.8 Celcius 24.8 Celcius

Table 1: Experimental data

Material	Specific Heat Capacity
Water	4180 J/kg·°C
Aluminum	900 J/kg·°C
Zinc	380 J/kg·°C
Copper	387 J/kg·°C
Iron	452 J/kg·°C
Steel	452 J/kg·°C
Lead	128 J/kg·°C
Silver	230 J/kg·°C

Table 2: Known specific heat capacities

## 7 Example Calculations

This is the calculation for the specific heat capacity of copper.

$$C_{metal} = \frac{m_{water}}{m_{metal}} \frac{\Delta T_{water}}{\Delta T_{metal}} C_{water}$$
$$\Delta T_{water} = 24.5 - 20.5 = 4.0 \text{ Celcius}$$
$$\Delta T_{metal} = 100 - 24.5 = 75.5 \text{ Celcius}$$
$$C_{metal} = \frac{0.350 \text{ kg}}{0.0906 \text{ kg}} \frac{4.0 \text{ Celcius}}{75.5 \text{ Celcius}} 4180 \text{ J/kg} \cdot ^\circ \text{C} = 855 \text{ J/kg} \cdot ^\circ \text{C}$$

The percent error is calculated as follows.

$$Error = \frac{900 - 855}{900} = 5\%$$

## 8 Results

Material	Measured $C_p$	Percent Error
Aluminum	900 J/kg $\cdot$ $^\circ$ C	5.0%
Zinc	363 J/kg $\cdot$ $^\circ$ C	0.8%
Copper	404 J/kg $\cdot$ $^\circ$ C	4.4%

Table 3: Calculated specific heat capacities

## 9 Discussion of Error

For the error, I have some situation will loss the heat, for example, when we put this metal out of the water, we have no way to control the heat lost, we know we didn't have enough thing to have save this heat, so even we try put these metal in the water as quick as possible, we still will loss the heats. And the paper cup is the same reason, because it not the heat saver cup, it can't have a good effort to save the heat, that's why it have the error

## 10 Conclusion

Heat is very difficult to measure and control, and that's make the error bigger than other kind of data, but it still work with the research, because we know that the "error is normal in the science" because science need to face the trouble and try to fix it by other way, like use many data, or try to complete our measures as better as possible. we can't have a prefect reason but we can try our best to have a good balance with the error.

## References

Homer, J. (2014). *Physics*. Oxford, 3rd edition.