

Elements, their Atomic Number and Molar Mass

Element	Symbol	Atomic Number	Molar mass/ (g mol⁻¹)	Element	Symbol	Atomic Number	Molar mass/ (g mol⁻¹)
Actinium	Ac	89	227.03	Mercury	Hg	80	200.59
Aluminium	Al	13	26.98	Molybdenum	Mo	42	95.94
Americium	Am	95	(243)	Neodymium	Nd	60	144.24
Antimony	Sb	51	121.75	Neon	Ne	10	20.18
Argon	Ar	18	39.95	Neptunium	Np	93	(237.05)
Arsenic	As	33	74.92	Nickel	Ni	28	58.71
Astatine	At	85	210	Niobium	Nb	41	92.91
Barium	Ba	56	137.34	Nitrogen	N	7	14.0067
Berkelium	Bk	97	(247)	Nobelium	No	102	(259)
Beryllium	Be	4	9.01	Osmium	Os	76	190.2
Bismuth	Bi	83	208.98	Oxygen	O	8	16.00
Bohrium	Bh	107	(264)	Palladium	Pd	46	106.4
Boron	B	5	10.81	Phosphorus	P	15	30.97
Bromine	Br	35	79.91	Platinum	Pt	78	195.09
Cadmium	Cd	48	112.40	Plutonium	Pu	94	(244)
Caesium	Cs	55	132.91	Polonium	Po	84	210
Calcium	Ca	20	40.08	Potassium	K	19	39.10
Californium	Cf	98	251.08	Praseodymium	Pr	59	140.91
Carbon	C	6	12.01	Promethium	Pm	61	(145)
Cerium	Ce	58	140.12	Protactinium	Pa	91	231.04
Chlorine	Cl	17	35.45	Radium	Ra	88	(226)
Chromium	Cr	24	52.00	Radon	Rn	86	(222)
Cobalt	Co	27	58.93	Rhenium	Re	75	186.2
Copper	Cu	29	63.54	Rhodium	Rh	45	102.91
Curium	Cm	96	247.07	Rubidium	Rb	37	85.47
Dubnium	Db	105	(263)	Ruthenium	Ru	44	101.07
Dysprosium	Dy	66	162.50	Rutherfordium	Rf	104	(261)
Einsteinium	Es	99	(252)	Samarium	Sm	62	150.35
Erbium	Er	68	167.26	Scandium	Sc	21	44.96
Europium	Eu	63	151.96	Seaborgium	Sg	106	(266)
Fermium	Fm	100	(257.10)	Selenium	Se	34	78.96
Fluorine	F	9	19.00	Silicon	Si	14	28.08
Francium	Fr	87	(223)	Silver	Ag	47	107.87
Gadolinium	Gd	64	157.25	Sodium	Na	11	22.99
Gallium	Ga	31	69.72	Strontium	Sr	38	87.62
Germanium	Ge	32	72.61	Sulphur	S	16	32.06
Gold	Au	79	196.97	Tantalum	Ta	73	180.95
Hafnium	Hf	72	178.49	Technetium	Tc	43	(98.91)
Hassium	Hs	108	(269)	Tellurium	Te	52	127.60
Helium	He	2	4.00	Terbium	Tb	65	158.92
Holmium	Ho	67	164.93	Thallium	Tl	81	204.37
Hydrogen	H	1	1.0079	Thorium	Th	90	232.04
Indium	In	49	114.82	Thulium	Tm	69	168.93
Iodine	I	53	126.90	Tin	Sn	50	118.69
Iridium	Ir	77	192.2	Titanium	Ti	22	47.88
Iron	Fe	26	55.85	Tungsten	W	74	183.85
Krypton	Kr	36	83.80	Ununbium	Uub	112	(277)
Lanthanum	La	57	138.91	Ununnilium	Uun	110	(269)
Lawrencium	Lr	103	(262.1)	Unununium	Uuu	111	(272)
Lead	Pb	82	207.19	Uranium	U	92	238.03
Lithium	Li	3	6.94	Vanadium	V	23	50.94
Lutetium	Lu	71	174.96	Xenon	Xe	54	131.30
Magnesium	Mg	12	24.31	Ytterbium	Yb	70	173.04
Manganese	Mn	25	54.94	Yttrium	Y	39	88.91
Meitneium	Mt	109	(268)	Zinc	Zn	30	65.37
Mendelevium	Md	101	258.10	Zirconium	Zr	40	91.22

The value given in parenthesis is the molar mass of the isotope of largest known half-life.

APPENDIX II

Some Useful Conversion Factors

Common Unit of Mass and Weight

1 pound = 453.59 grams

1 pound = 453.59 grams = 0.45359 kilogram
1 kilogram = 1000 grams = 2.205 pounds
1 gram = 10 decigrams = 100 centigrams
= 1000 milligrams
1 gram = 6.022×10^{23} atomic mass units or u
1 atomic mass unit = 1.6606×10^{-24} gram
1 metric tonne = 1000 kilograms
= 2205 pounds

Common Unit of Volume

1 quart = 0.9463 litre

1 litre = 1.056 quarts

1 litre = 1 cubic decimetre = 1000 cubic centimetres = 0.001 cubic metre
1 millilitre = 1 cubic centimetre = 0.001 litre
= 1.056×10^{-3} quart
1 cubic foot = 28.316 litres = 29.902 quarts
= 7.475 gallons

Common Units of Energy

1 joule = 1×10^7 ergs

1 thermochemical calorie**
= 4.184 joules
= 4.184×10^7 ergs
= 4.129×10^{-2} litre-atmospheres
= 2.612×10^{19} electron volts
1 ergs = 1×10^{-7} joule = 2.3901×10^{-8} calorie
1 electron volt = 1.6022×10^{-19} joule
= 1.6022×10^{-12} erg
= 96.487 kJ/mol†
1 litre-atmosphere = 24.217 calories
= 101.32 joules
= 1.0132×10^9 ergs
1 British thermal unit = 1055.06 joules
= 1.05506×10^{10} ergs
= 252.2 calories

Common Units of Length

1 inch = 2.54 centimetres (exactly)

1 mile = 5280 feet = 1.609 kilometres
1 yard = 36 inches = 0.9144 metre
1 metre = 100 centimetres = 39.37 inches
= 3.281 feet
= 1.094 yards
1 kilometre = 1000 metres = 1094 yards
= 0.6215 mile
1 Angstrom = 1.0×10^{-8} centimetre
= 0.10 nanometre
= 1.0×10^{-10} metre
= 3.937×10^{-9} inch

Common Units of Force* and Pressure

1 atmosphere = 760 millimetres of mercury
= 1.013×10^5 pascals
= 14.70 pounds per square inch
1 bar = 10^5 pascals
1 torr = 1 millimetre of mercury
1 pascal = $1 \text{ kg}/\text{ms}^2 = 1 \text{ N/m}^2$

Temperature

SI Base Unit: Kelvin (K)

$$\begin{aligned} K &= -273.15^\circ C \\ K &= {}^\circ C + 273.15 \\ {}^\circ F &= {}^\circ C + 32 \\ {}^\circ C &= \frac{{}^\circ F - 32}{1.8} \end{aligned}$$

* Force: 1 newton (N) = 1 kg m/s^2 , i.e., the force that, when applied for 1 second, gives a 1-kilogram mass a velocity of 1 metre per second.

** The amount of heat required to raise the temperature of one gram of water from $14.5^\circ C$ to $15.5^\circ C$.

† Note that the other units are per particle and must be multiplied by 6.022×10^{23} to be strictly comparable.

Standard potentials at 298 K in electrochemical order

Reduction half-reaction	<i>E</i> /V	Reduction half-reaction	<i>E</i> /V
$\text{H}_4\text{XeO}_6 + 2\text{H}^+ + 2\text{e}^- \rightarrow \text{XeO}_3 + 3\text{H}_2\text{O}$	+3.0	$\text{Cu}^+ + \text{e}^- \rightarrow \text{Cu}$	+0.52
$\text{F}_2 + 2\text{e}^- \rightarrow 2\text{F}^-$	+2.87	$\text{NiOOH} + \text{H}_2\text{O} + \text{e}^- \rightarrow \text{Ni}(\text{OH})_2 + \text{OH}^-$	+0.49
$\text{O}_3 + 2\text{H}^+ + 2\text{e}^- \rightarrow \text{O}_2 + \text{H}_2\text{O}$	+2.07	$\text{Ag}_2\text{CrO}_4 + 2\text{e}^- \rightarrow 2\text{Ag} + \text{CrO}_4^{2-}$	+0.45
$\text{S}_2\text{O}_8^{2-} + 2\text{e}^- \rightarrow 2\text{SO}_4^{2-}$	+2.05	$\text{O}_2 + 2\text{H}_2\text{O} + 4\text{e}^- \rightarrow 4\text{OH}^-$	+0.40
$\text{Ag}^+ + \text{e}^- \rightarrow \text{Ag}^+$	+1.98	$\text{ClO}_4^- + \text{H}_2\text{O} + 2\text{e}^- \rightarrow \text{ClO}_3^- + 2\text{OH}^-$	+0.36
$\text{Co}^{3+} + \text{e}^- \rightarrow \text{Co}^{2+}$	+1.81	$[\text{Fe}(\text{CN})_6]^{3-} + \text{e}^- \rightarrow [\text{Fe}(\text{CN})_6]^{4-}$	+0.36
$\text{H}_2\text{O}_2 + 2\text{H}^+ + 2\text{e}^- \rightarrow 2\text{H}_2\text{O}$	+1.78	$\text{Cu}^{2+} + 2\text{e}^- \rightarrow \text{Cu}$	+0.34
$\text{Au}^+ + \text{e}^- \rightarrow \text{Au}$	+1.69	$\text{Hg}_2\text{Cl}_2 + 2\text{e}^- \rightarrow 2\text{Hg} + 2\text{Cl}^-$	+0.27
$\text{Pb}^{4+} + 2\text{e}^- \rightarrow \text{Pb}^{2+}$	+1.67	$\text{AgCl} + \text{e}^- \rightarrow \text{Ag} + \text{Cl}^-$	+0.27
$2\text{HClO} + 2\text{H}^+ + 2\text{e}^- \rightarrow \text{Cl}_2 + 2\text{H}_2\text{O}$	+1.63	$\text{Bi}^{3+} + 3\text{e}^- \rightarrow \text{Bi}$	+0.20
$\text{Ce}^{4+} + \text{e}^- \rightarrow \text{Ce}^{3+}$	+1.61	$\text{SO}_4^{2-} + 4\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2\text{SO}_3 + \text{H}_2\text{O}$	+0.17
$2\text{HBrO} + 2\text{H}^+ + 2\text{e}^- \rightarrow \text{Br}_2 + 2\text{H}_2\text{O}$	+1.60	$\text{Cu}^{2+} + \text{e}^- \rightarrow \text{Cu}^+$	+0.16
$\text{MnO}_4^- + 8\text{H}^+ + 5\text{e}^- \rightarrow \text{Mn}^{2+} + 4\text{H}_2\text{O}$	+1.51	$\text{Sn}^{4+} + 2\text{e}^- \rightarrow \text{Sn}^{2+}$	+0.15
$\text{Mn}^{3+} + \text{e}^- \rightarrow \text{Mn}^{2+}$	+1.51	$\text{AgBr} + \text{e}^- \rightarrow \text{Ag} + \text{Br}^-$	+0.07
$\text{Au}^{3+} + 3\text{e}^- \rightarrow \text{Au}$	+1.40	$\text{Ti}^{4+} + \text{e}^- \rightarrow \text{Ti}^{3+}$	0.00
$\text{Cl}_2 + 2\text{e}^- \rightarrow 2\text{Cl}^-$	+1.36	$2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$	0.0 by definition
$\text{Cr}_2\text{O}_7^{2-} + 14\text{H}^+ + 6\text{e}^- \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$	+1.33	$\text{Fe}^{3+} + 3\text{e}^- \rightarrow \text{Fe}$	-0.04
$\text{O}_3 + \text{H}_2\text{O} + 2\text{e}^- \rightarrow \text{O}_2 + 2\text{OH}^-$	+1.24	$\text{O}_2 + \text{H}_2\text{O} + 2\text{e}^- \rightarrow \text{HO}_2^- + \text{OH}^-$	-0.08
$\text{O}_2 + 4\text{H}^+ + 4\text{e}^- \rightarrow 2\text{H}_2\text{O}$	+1.23	$\text{Pb}^{2+} + 2\text{e}^- \rightarrow \text{Pb}$	-0.13
$\text{ClO}_4^- + 2\text{H}^+ + 2\text{e}^- \rightarrow \text{ClO}_3^- + 2\text{H}_2\text{O}$	+1.23	$\text{In}^+ + \text{e}^- \rightarrow \text{In}$	-0.14
$\text{MnO}_2 + 4\text{H}^+ + 2\text{e}^- \rightarrow \text{Mn}^{2+} + 2\text{H}_2\text{O}$	+1.23	$\text{Sn}^{2+} + 2\text{e}^- \rightarrow \text{Sn}$	-0.14
$\text{Pt}^{2+} + 2\text{e}^- \rightarrow \text{Pt}$	+1.20	$\text{AgI} + \text{e}^- \rightarrow \text{Ag} + \text{I}^-$	-0.15
$\text{Br}_2 + 2\text{e}^- \rightarrow 2\text{Br}^-$	+1.09	$\text{Ni}^{2+} + 2\text{e}^- \rightarrow \text{Ni}$	-0.23
$\text{Pu}^{4+} + \text{e}^- \rightarrow \text{Pu}^{3+}$	+0.97	$\text{V}^{3+} + \text{e}^- \rightarrow \text{V}^{2+}$	-0.26
$\text{NO}_3^- + 4\text{H}^+ + 3\text{e}^- \rightarrow \text{NO} + 2\text{H}_2\text{O}$	+0.96	$\text{Co}^{2+} + 2\text{e}^- \rightarrow \text{Co}$	-0.28
$2\text{Hg}^{2+} + 2\text{e}^- \rightarrow \text{Hg}_2^{2+}$	+0.92	$\text{In}^{3+} + 3\text{e}^- \rightarrow \text{In}$	-0.34
$\text{ClO}^- + \text{H}_2\text{O} + 2\text{e}^- \rightarrow \text{Cl}^- + 2\text{OH}^-$	+0.89	$\text{Tl}^+ + \text{e}^- \rightarrow \text{Tl}$	-0.34
$\text{Hg}^{2+} + 2\text{e}^- \rightarrow \text{Hg}$	+0.86	$\text{PbSO}_4 + 2\text{e}^- \rightarrow \text{Pb} + \text{SO}_4^{2-}$	-0.36
$\text{NO}_3^- + 2\text{H}^+ + \text{e}^- \rightarrow \text{NO}_2 + \text{H}_2\text{O}$	+0.80	$\text{Ti}^{3+} + \text{e}^- \rightarrow \text{Ti}^{2+}$	-0.37
$\text{Ag}^+ + \text{e}^- \rightarrow \text{Ag}$	+0.80	$\text{Cd}^{2+} + 2\text{e}^- \rightarrow \text{Cd}$	-0.40
$\text{Hg}_2^{2+} + 2\text{e}^- \rightarrow 2\text{Hg}$	+0.79	$\text{In}^{2+} + \text{e}^- \rightarrow \text{In}^+$	-0.40
$\text{Fe}^{3+} + \text{e}^- \rightarrow \text{Fe}^{2+}$	+0.77	$\text{Cr}^{3+} + \text{e}^- \rightarrow \text{Cr}^{2+}$	-0.41
$\text{BrO}^- + \text{H}_2\text{O} + 2\text{e}^- \rightarrow \text{Br}^- + 2\text{OH}^-$	+0.76	$\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$	-0.44
$\text{Hg}_2\text{SO}_4 + 2\text{e}^- \rightarrow 2\text{Hg} + \text{SO}_4^{2-}$	+0.62	$\text{In}^{3+} + 2\text{e}^- \rightarrow \text{In}^+$	-0.44
$\text{MnO}_4^{2-} + 2\text{H}_2\text{O} + 2\text{e}^- \rightarrow \text{MnO}_2 + 4\text{OH}^-$	+0.60	$\text{S} + 2\text{e}^- \rightarrow \text{S}^{2-}$	-0.48
$\text{MnO}_4^- + \text{e}^- \rightarrow \text{MnO}_4^{2-}$	+0.56	$\text{In}^{3+} + \text{e}^- \rightarrow \text{In}^{2+}$	-0.49
$\text{I}_2 + 2\text{e}^- \rightarrow 2\text{I}^-$	+0.54	$\text{U}^{4+} + \text{e}^- \rightarrow \text{U}^{3+}$	-0.61
$\text{I}_3^- + 2\text{e}^- \rightarrow 3\text{I}^-$	+0.53	$\text{Cr}^{3+} + 3\text{e}^- \rightarrow \text{Cr}$	-0.74
		$\text{Zn}^{2+} + 2\text{e}^- \rightarrow \text{Zn}$	-0.76

(continued)

APPENDIX III CONTINUED

Reduction half-reaction	E /V	Reduction half-reaction	E /V
$\text{Cd}(\text{OH})_2 + 2\text{e}^- \longrightarrow \text{Cd} + 2\text{OH}^-$	-0.81	$\text{La}^{3+} + 3\text{e}^- \longrightarrow \text{La}$	-2.52
$2\text{H}_2\text{O} + 2\text{e}^- \longrightarrow \text{H}_2 + 2\text{OH}^-$	-0.83	$\text{Na}^+ + \text{e}^- \longrightarrow \text{Na}$	-2.71
$\text{Cr}^{2+} + 2\text{e}^- \longrightarrow \text{Cr}$	-0.91	$\text{Ca}^{2+} + 2\text{e}^- \longrightarrow \text{Ca}$	-2.87
$\text{Mn}^{2+} + 2\text{e}^- \longrightarrow \text{Mn}$	-1.18	$\text{Sr}^{2+} + 2\text{e}^- \longrightarrow \text{Sr}$	-2.89
$\text{V}^{2+} + 2\text{e}^- \longrightarrow \text{V}$	-1.19	$\text{Ba}^{2+} + 2\text{e}^- \longrightarrow \text{Ba}$	-2.91
$\text{Ti}^{2+} + 2\text{e}^- \longrightarrow \text{Ti}$	-1.63	$\text{Ra}^{2+} + 2\text{e}^- \longrightarrow \text{Ra}$	-2.92
$\text{Al}^{3+} + 3\text{e}^- \longrightarrow \text{Al}$	-1.66	$\text{Cs}^+ + \text{e}^- \longrightarrow \text{Cs}$	-2.92
$\text{U}^{3+} + 3\text{e}^- \longrightarrow \text{U}$	-1.79	$\text{Rb}^+ + \text{e}^- \longrightarrow \text{Rb}$	-2.93
$\text{Sc}^{3+} + 3\text{e}^- \longrightarrow \text{Sc}$	-2.09	$\text{K}^+ + \text{e}^- \longrightarrow \text{K}$	-2.93
$\text{Mg}^{2+} + 2\text{e}^- \longrightarrow \text{Mg}$	-2.36	$\text{Li}^+ + \text{e}^- \longrightarrow \text{Li}$	-3.05
$\text{Ce}^{3+} + 3\text{e}^- \longrightarrow \text{Ce}$	-2.48		

Logarithms

Sometimes, a numerical expression may involve multiplication, division or rational powers of large numbers. For such calculations, logarithms are very useful. They help us in making difficult calculations easy. In Chemistry, logarithm values are required in solving problems of chemical kinetics, thermodynamics, electrochemistry, etc. We shall first introduce this concept, and discuss the laws, which will have to be followed in working with logarithms, and then apply this technique to a number of problems to show how it makes difficult calculations simple.

We know that

$$2^3 = 8, 3^2 = 9, 5^3 = 125, 7^0 = 1$$

In general, for a positive real number a , and a rational number m , let $a^m = b$,

where b is a real number. In other words

the m^{th} power of base a is b .

Another way of stating the same fact is

logarithm of b to base a is m .

If for a positive real number a , $a \neq 1$

$$a^m = b,$$

we say that m is the logarithm of b to the base a .

We write this as $\log_a^b = m$,

“log” being the abbreviation of the word “logarithm”.

Thus, we have

$$\log_2 8 = 3, \quad \text{Since } 2^3 = 8$$

$$\log_3 9 = 2, \quad \text{Since } 3^2 = 9$$

$$\log \frac{125}{5} = 3, \quad \text{Since } 5^3 = 125$$

$$\log_7 1 = 0, \quad \text{Since } 7^0 = 1$$

Laws of Logarithms

In the following discussion, we shall take logarithms to any base a , ($a > 0$ and $a \neq 1$)

First Law: $\log_a(mn) = \log_a m + \log_a n$

Proof: Suppose that $\log_a m = x$ and $\log_a n = y$

Then $a^x = m$, $a^y = n$

Hence $mn = a^x \cdot a^y = a^{x+y}$

It now follows from the definition of logarithms that

$$\log_a(mn) = x + y = \log_a m + \log_a n$$

Second Law: $\log_a \left(\frac{m}{n} \right) = \log_a m - \log_a n$

Proof: Let $\log_a m = x$, $\log_a n = y$

Then $a^x = m$, $a^y = n$

$$\text{Hence } \frac{m}{n} = \frac{a^x}{a^y} = a^{x-y}$$

Therefore

$$\log_a \left(\frac{m}{n} \right) = x - y = \log_a m - \log_a n$$

Third Law : $\log_a(m^n) = n \log_a m$

Proof : As before, if $\log_a m = x$, then $a^x = m$

$$\text{Then } m^n = (a^x)^n = a^{nx}$$

giving $\log_a(m^n) = nx = n \log_a m$

Thus according to First Law: "the log of the product of two numbers is equal to the sum of their logs. Similarly, the Second Law says: the log of the ratio of two numbers is the difference of their logs. Thus, the use of these laws converts a problem of multiplication / division into a problem of addition/ subtraction, which are far easier to perform than multiplication/division. That is why logarithms are so useful in all numerical computations.

Logarithms to Base 10

Because number 10 is the base of writing numbers, it is very convenient to use logarithms to the base 10. Some examples are:

$\log_{10} 10 = 1$,	since $10^1 = 10$
$\log_{10} 100 = 2$,	since $10^2 = 100$
$\log_{10} 10000 = 4$,	since $10^4 = 10000$
$\log_{10} 0.01 = -2$,	since $10^{-2} = 0.01$
$\log_{10} 0.001 = -3$,	since $10^{-3} = 0.001$
and $\log_{10} 1 = 0$	since $10^0 = 1$

The above results indicate that if n is an integral power of 10, i.e., 1 followed by several zeros or 1 preceded by several zeros immediately to the right of the decimal point, then $\log n$ can be easily found.

If n is not an integral power of 10, then it is not easy to calculate $\log n$. But mathematicians have made tables from which we can read off approximate value of the logarithm of any positive number between 1 and 10. And these are sufficient for us to calculate the logarithm of any number expressed in decimal form. For this purpose, we always express the given decimal as the product of an integral power of 10 and a number between 1 and 10.

Standard Form of Decimal

We can express any number in decimal form, as the product of (i) an integral power of 10, and (ii) a number between 1 and 10. Here are some examples:

(i) 25.2 lies between 10 and 100

$$25.2 = \frac{25.2}{10} \times 10 = 2.52 \times 10^1$$

(ii) 1038.4 lies between 1000 and 10000.

$$\therefore 1038.4 = \frac{1038.4}{1000} \times 10^3 = 1.0384 \times 10^3$$

(iii) 0.005 lies between 0.001 and 0.01

$$\therefore 0.005 = (0.005 \times 1000) \times 10^{-3} = 5.0 \times 10^{-3}$$

(iv) 0.00025 lies between 0.0001 and 0.001

$$\therefore 0.00025 = (0.00025 \times 10000) \times 10^{-4} = 2.5 \times 10^{-4}$$

In each case, we divide or multiply the decimal by a power of 10, to bring one non-zero digit to the left of the decimal point, and do the reverse operation by the same power of 10, indicated separately.

Thus, any positive decimal can be written in the form

$$n = m \times 10^p$$

where p is an integer (positive, zero or negative) and $1 \leq m < 10$. This is called the "standard form of n ".

Working Rule

1. Move the decimal point to the left, or to the right, as may be necessary, to bring one non-zero digit to the left of decimal point.
2. (i) If you move p places to the left, multiply by 10^p .
 (ii) If you move p places to the right, multiply by 10^{-p} .
 (iii) If you do not move the decimal point at all, multiply by 10^0 .
 (iv) Write the new decimal obtained by the power of 10 (of step 2) to obtain the standard form of the given decimal.

Characteristic and Mantissa

Consider the standard form of n

$$n = m \times 10^p, \text{ where } 1 \leq m < 10$$

Taking logarithms to the base 10 and using the laws of logarithms

$$\begin{aligned} \log n &= \log m + \log 10^p \\ &= \log m + p \log 10 \\ &= p + \log m \end{aligned}$$

Here p is an integer and as $1 \leq m < 10$, so $0 \leq \log m < 1$, i.e., m lies between 0 and 1. When $\log n$ has been expressed as $p + \log m$, where p is an integer and $0 \leq \log m < 1$, we say that p is the "characteristic" of $\log n$ and that $\log m$ is the "mantissa of $\log n$ ". Note that characteristic is always an integer – positive, negative or zero, and mantissa is never negative and is always less than 1. If we can find the characteristics and the mantissa of $\log n$, we have to just add them to get $\log n$.

Thus to find $\log n$, all we have to do is as follows:

1. Put n in the standard form, say

$$n = m \times 10^p, 1 \leq m < 10$$

2. Read off the characteristic p of $\log n$ from this expression (exponent of 10).

3. Look up $\log m$ from tables, which is being explained below.

4. Write $\log n = p + \log m$

If the characteristic p of a number n is say, 2 and the mantissa is .4133, then we have $\log n = 2 + .4133$ which we can write as 2.4133. If, however, the characteristic p of a number m is say -2 and the mantissa is .4123, then we have $\log m = -2 + .4123$. We cannot write this as -2.4123. (Why?) In order to avoid this confusion we write $\bar{2}$ for -2 and thus we write $\log m = \bar{2}.4123$.

Now let us explain how to use the table of logarithms to find mantissas. A table is appended at the end of this Appendix.

Observe that in the table, every row starts with a two digit number, 10, 11, 12, ..., 97, 98, 99. Every column is headed by a one-digit number, 0, 1, 2, ..., 9. On the right, we have the section called "Mean differences" which has 9 columns headed by 1, 2, ..., 9.

0	1	2	3	4	5	6	7	8	9	1	2	3	4	5	6	7	8	9
..
61	7853	7860	7868	7875	7882	7889	7896	7803	7810	7817	1	1	2	3	4	4	5	6
62	7924	7931	7935	7945	7954	7959	7966	7973	7980	7987	1	1	2	3	4	5	6	6
63	7993	8000	8007	8014	8021	8028	8035	8041	8048	8055	1	1	2	3	3	4	5	6
..

Now suppose we wish to find $\log(6.234)$. Then look into the row starting with 62. In this row, look at the number in the column headed by 3. The number is 7945. This means that

$$\log(6.230) = 0.7945^*$$

But we want $\log(6.234)$. So our answer will be a little more than 0.7945. How much more? We look this up in the section on Mean differences. Since our fourth digit is 4, look under the column headed by 4 in the Mean difference section (in the row 62). We see the number 3 there. So add 3 to 7945. We get 7948. So we finally have

$$\log(6.234) = 0.7948.$$

Take another example. To find $\log(8.127)$, we look in the row 81 under column 2, and we find 9096. We continue in the same row and see that the mean difference under 7 is 4. Adding this to 9096, and we get 9100. So, $\log(8.127) = 0.9100$.

Finding N when $\log N$ is given

We have so far discussed the procedure for finding $\log n$ when a positive number n given. We now turn to its converse i.e., to find n when $\log n$ is given and give a method for this purpose. If $\log n = t$, we sometimes say $n = \text{antilog } t$. Therefore our task is given t , find its antilog. For this, we use the ready-made antilog tables.

Suppose $\log n = 2.5372$.

To find n , first take just the mantissa of $\log n$. In this case it is .5372. (Make sure it is positive.) Now take up antilog of this number in the antilog table which is to be used exactly like the log table. In the antilog table, the entry under column 7 in the row .53 is 3443 and the mean difference for the last digit 2 in that row is 2, so the table gives 3445. Hence,

$$\text{antilog}(.5372) = 3.445$$

Now since $\log n = 2.5372$, the characteristic of $\log n$ is 2. So the standard form of n is given by

$$n = 3.445 \times 10^2$$

$$\text{or } n = 344.5$$

Illustration 1:

If $\log x = 1.0712$, find x .

Solution: We find that the number corresponding to 0712 is 1179. Since characteristic of $\log x$ is 1, we have

$$\begin{aligned}x &= 1.179 \times 10^1 \\&= 11.79\end{aligned}$$

Illustration 2:

If $\log x = \underline{2}.1352$, find x .

Solution: From antilog tables, we find that the number corresponding to 1352 is 1366. Since the characteristic is $\underline{2}$ i.e., -2, so

$$x = 1.366 \times 10^{-2} = 0.01366$$

Use of Logarithms in Numerical Calculations

Illustration 1:

Find 6.3×1.29

Solution: Let $x = 6.3 \times 1.29$

Then $\log x = \log(6.3 \times 1.29) = \log 6.3 + \log 1.29$

Now,

$$\log 6.3 = 0.7993$$

$$\log 1.29 = 0.1106$$

$$\therefore \log x = 0.9099,$$

* It should, however, be noted that the values given in the table are not exact. They are only approximate values, although we use the sign of equality which may give the impression that they are exact values. The same convention will be followed in respect of antilogarithm of a number.

Taking antilog

$$x = 8.127$$

Illustration 2:

$$\text{Find } \frac{(1.23)^{1.5}}{11.2 \times 23.5}$$

$$\text{Solution: Let } x = \frac{(1.23)^{\frac{3}{2}}}{11.2 \times 23.5}$$

$$\text{Then } \log x = \log \frac{(1.23)^{\frac{3}{2}}}{11.2 \times 23.5}$$

$$= \frac{3}{2} \log 1.23 - \log (11.2 \times 23.5)$$

$$= \frac{3}{2} \log 1.23 - \log 11.2 - \log 23.5$$

Now,

$$\log 1.23 = 0.0899$$

$$\frac{3}{2} \log 1.23 = 0.13485$$

$$\log 11.2 = 1.0492$$

$$\log 23.5 = 1.3711$$

$$\log x = 0.13485 - 1.0492 - 1.3711$$

$$= \overline{3.71455}$$

$$\therefore x = 0.005183$$

Illustration 3:

$$\text{Find } \sqrt{\frac{(71.24)^5 \times \sqrt{56}}{(2.3)^7 \times \sqrt{21}}}$$

$$\text{Solution: Let } x = \sqrt{\frac{(71.24)^5 \times \sqrt{56}}{(2.3)^7 \times \sqrt{21}}}$$

$$\text{Then } \log x = \frac{1}{2} \log \left[\frac{(71.24)^5 \times \sqrt{56}}{(2.3)^7 \times \sqrt{21}} \right]$$

$$= \frac{1}{2} [\log (71.24)^5 + \log \sqrt{56} - \log (2.3)^7 - \log \sqrt{21}]$$

$$= \frac{5}{2} \log 71.24 + \frac{1}{4} \log 56 - \frac{7}{2} \log 2.3 - \frac{1}{4} \log 21$$

Now, using log tables

$$\log 71.24 = 1.8527$$

$$\log 56 = 1.7482$$

$$\log 2.3 = 0.3617$$

$$\log 21 = 1.3222$$

$$\therefore \log x = \frac{5}{2} \log (1.8527) + \frac{1}{4} (1.7482) - \frac{7}{2} (0.3617) - \frac{1}{4} (1.3222)$$
$$= 3.4723$$

$$\therefore x = 2967$$

