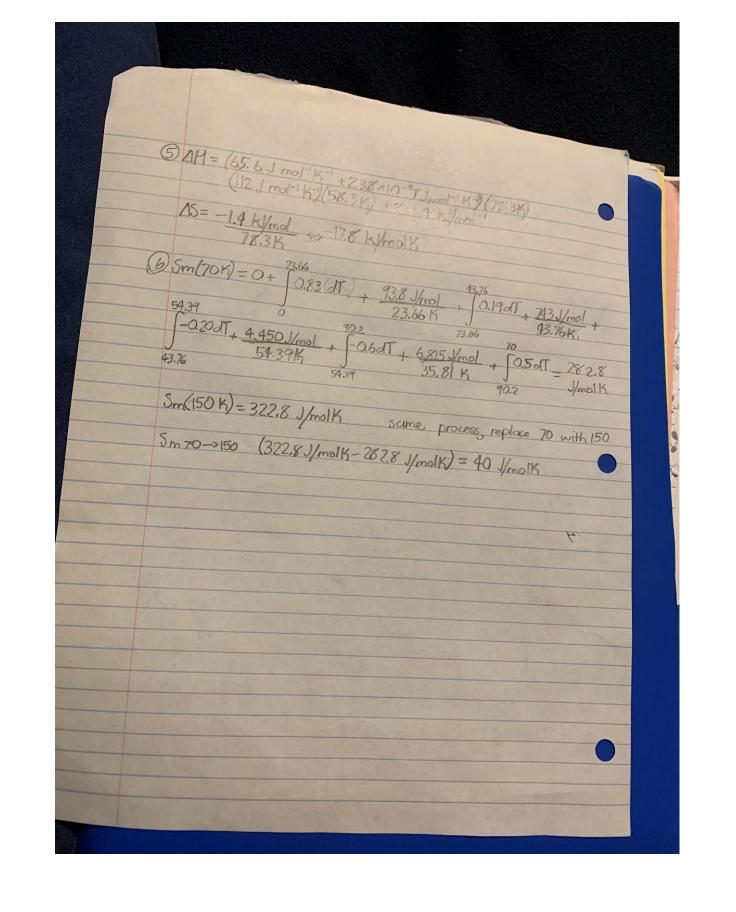
	"
Kevin Marry	
5.12 Using the Part that 5 Ts a State Tunchon 15=(35) dT, (35) dV	
15=(35) dT (85) dV Sentrope Tetro	
Tetemperature (35). CV	
Cv= heat capacity (under t val)	
The second law of Herman	
The second law of Hermodynamics is applied here. 5.13 The Dependence of 5 on T and P	
dS=(25) dT, (25) dP CP dT VBdP	
Cp = heat capacity (constant pressure)	
The second law of Hermodynamics is applied here.	
6.1 The Gibbs Erergy and the Helmholtz Freigy	-763 k
16= 14-TDS G=Gibs every H=enthalpy	1/mol
SA = AU-TAS A = Helmholtz erergy	2.30
oth equations involve reactions	16= Al
e sexond law of Hermodynamics is applied here.	
marginalis is applied here	

	5.2 The	
	M-LAPV A=U-15 G	of UHA
	M=UAPV A=U-15 G- M DU-TOS-POW	2000
	att = TOS+VOIP	dA
	121270 - 6 126	
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	law of Man	
	6.3 The Dependence of the Gibl	samies 5 applied here.
	6.3 The Dependence of the Gible Erergies on By and T	Telmholtz
	() J/T - 4 ()	TIPES (86)-11
	Sibbs-Helmholtz equation de	2 - 17
	2(1/T/P = H	Contract copy to
P	second low of How	/mol
	Trermod	dynamics is applied here. K
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@ Int Vo @ Ware = - (qub+ qud) The total work is O. The efficiency has to be less than I because a cyclic process can't convert heat to work with 100% efficiency. dS=CP dT_VBdP AHrm= (-1,364 KJ/mol+(-218 KJ/mol)+(-394 KJ/mol)-(-1,273 KJ/mol)=-763 KJ/mol 15° xn = (192 ymol K+ (161 J/mol K)+(213° J/mol K) - (209 J/mol K) = 357 J/mol K DSsur = (-763 KJ/mol) ≈ 2.56 KJ/mol K 15uni=357 J/molK-2.56 Ky/molK =-2,203 J/molk The reaction isn't sportsuneous because in the equation AG=SH-TAS, since AS is negative, and IASI>IAH, AG will be positive.



Questions