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## ★ G.1 Ab Initio Philosophy

HF Hartree Fock theory is ~~unusually~~ useful in ab initio (from the beginning) philosophy because of ~~its~~ it can be used to understand more complicated theories. HF wave functions make it possible for wave functions to be used to test properties that are not energy

$$\begin{array}{c} \text{System energy in the HF limit} \\ \swarrow \\ E_{\text{corr}} = E - E_{\text{HF}} \\ \uparrow \quad \quad \uparrow \\ \text{Correlation energy, used} \quad \text{True energy} \\ \text{to supplement } E_{\text{HF}} \end{array}$$

## ★ G.2.1 Functional forms

STO =  
Slater-type  
orbital

GTO =  
Gaussian-type  
orbital

Slater-type orbitals are limited in ab initio HF theory. AO-like functions, ~~are favored~~ in the form of Gaussian functions are used in alternative to STO. The general functional form of a normalized GTO is expressed in atom-centred Cartesian coordinates.

$$\phi(x, y, z; a, i, j, k) = \left( \frac{2a}{\pi} \right)^{3/4} \left[ \frac{(8a)^{i+j+k} i! j! k!}{(2i)! (2j)! (2k)!} \right]^{1/2} x^i y^j z^k e^{-a(x^2+y^2+z^2)}$$

Exponent controlling width of GTO  
non-negative integers  
Dictates  
All 0  $\Rightarrow$  s-type GTO  
Index one is 1, p-type GTO  
Sum = 2, d-type GTO

### ★ 6.2.2 Contracted Gaussian Functions

Basis Functions  $\psi$  used for SCF calculations were linear combinations of GTOs used for the same purpose as (and as accurate) as an STO.

$$\psi(x,y,z; \{a\}, i, j, k) = \sum_{a=1}^M \underset{\substack{\uparrow \\ \text{coefficients}}}{c_a} \phi(x,y,z; a, i, j, k)$$

# of Gaussians  
↓  
M

### ★ 6.2.3 Single- $\zeta$ , Multiple- $\zeta$ , and Split-Valence

"Decontracting" basis sets increase their flexibility. It is ~~connecting~~ constructing 2 basis functions for each orbital, first is a contraction of the first 2 primitive Gaussians, and the second is the normalized, primitive third.

### ★ 6.2.4 Polarization Functions

~~Molecular orbitals~~

Molecular orbitals need more mathematical flexibility than atoms. AO-like GTOs add this flexibility in the form of basis functions corresponding to one quantum # of higher angular momentum than the valence orbitals.



### ★ 6.2.5 Diffuse Functions

Higher energy MOs are more diffuse. When an electron is unable to delocalize in a basis set, problems arise.

To solve this problem, standard basis sets are "augmented" w/ diffuse basis functions. The Pople Family + Dunning Family of basis sets uses such functions

### ★ 6.2.6. The HF Limit

<sup>These are</sup> Solutions of Equations w/ an infinite # of basis sets. More commonly, people ~~are~~ extrapolate the HF limit by using similar concepts for some basis-set incompleteness. There are potentially large errors w/ this approximation, so it is more of a formal standpoint

### ★ 6.2.7 Effective Core Potentials

~~Thereby~~ Replacing electrons w/ analytical functions that represents the combined-nuclear-electronic core to remaining electrons is a good way to use basis sets w/ large elements. ECP (Effective core potentials) is a nuclear point charge reduced in magnitude by the # of core electrons.

### ★ 6.2.8 Sources

Use websites that permits you to download large # of basis sets using different software packages + their reference information.

## 8 Density Functional Theory

### 8.1 Theoretical Motivation

#### ★ 8.1.1 Philosophy

We take advantage of our knowledge of quantum mechanics to ask about a particular observable.

$$\# \text{ of } e \rightarrow \mathcal{N} = \int p(r) dr$$

$$\left. \frac{\partial p(r)}{\partial r} \right|_{r=0} = -2Z_A p(r)$$

Diagram annotations:  
-  $\frac{\partial p(r)}{\partial r}$ : electron density  
-  $\left. \frac{\partial p(r)}{\partial r} \right|_{r=0}$ : spherically averaged density  
-  $r=0$ : radial distance from A  
-  $Z_A$ : atomic #  
-  $p(r)$ : electron density maximum  
-  $r=0$ : nucleus

#### ★ 8.1.2 Early Approximations

There is both kinetic + potential energy, and they can be found through several equations using several approximations. We can calculate exchange energy  $E_x$  using

$$E_x[p(r)] = -\frac{9\alpha}{8} \left(\frac{3}{\pi}\right)^{\frac{1}{3}} \int p^{\frac{4}{3}}(r) dr$$



## 8.2 Rigorous Foundation

### ★ 8.2.1 The Hohenberg-Kohn Existence Theorem

The first Hohenberg-Kohn theorem states that "the ground state of any interacting many particle system w/ a given fixed inner ~~part~~ particle interaction is a unique functional of the electron ~~dens~~ density."

I got over-zealous

$$E_{0,a} + E_{0,b} < \int [v_b(r) - v_a(r)] \rho_0(r) dr + \int [v_a(r) - v_b(r)] \rho_b(r) dr + E_{0,b} + E_{0,a}$$

$\downarrow$  diff  $e^-$

### ★ 8.2.2 The Hohenberg-Kohn Variational Theorem

To establish a dependence of the energy on the density in this theorem, it is employed in the ground-state density, which is sufficient to show that it determines the Hamiltonian operator

### ★ 8.3 Kohn-Sham Self-consistent Field Methodology

Kohn & Sham realized that things would be ~~be~~ simpler if only the Hamiltonian operator were for a non-interacting system of electrons

The Kohn-Sham one-electron operator is defined as

$$h_i^{KS} = \frac{1}{2} \nabla_i^2 - \sum_K \frac{Z_K}{|r_i - r_K|} + \int \frac{\rho(r')}{|r_i - r'|} dr' + V_{xc}$$

Function derivative  $\rightarrow V_{xc} = \frac{\delta E_{xc}}{\delta \rho}$