

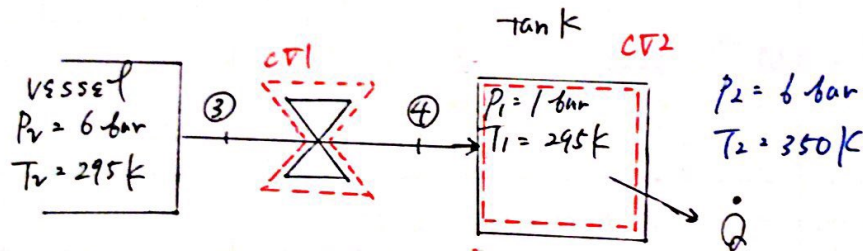
GIVENrigid tank $V = 2.0 \text{ m}^3$ >> initially $P_1 = 1 \text{ bar}$, $T_1 = 295 \text{ K}$ (Air)>> valve opened from vessel of $P_v = 6 \text{ bar}$, $T_v = 295 \text{ K}$ >> final of tank $P_2 = 6 \text{ bar}$, $T_2 = 350 \text{ K}$ FIND

Q

ASSUMP open sys, uniform flow, ideal gas, $\Delta KE = \Delta PE = 0$, lumped mass, incompressible fluid, constant specific heat, no work, SFEQN

$$\frac{dm}{dt}_{\text{sys}} = \sum \dot{m}_i - \sum \dot{m}_e, \quad \frac{dE}{dt}_{\text{sys}} = \dot{Q} - \dot{W} + \sum \dot{m}_i (h + \text{pe} + \text{ke}) - \sum \dot{m}_e (h + \text{pe} + \text{ke})$$

$$PV = n R_{\text{air}} T \quad (R_{\text{air}} = 0.287 \text{ kJ/kg}\cdot\text{K}), \quad \Delta E = \Delta U + \Delta KE + \Delta PE$$

FFDSOLN

first of all $m_1 = \frac{P_1 V}{R_{\text{air}} T_1} = \frac{(1 \times 10^5 \text{ Pa})(2.0 \text{ m}^3)}{(0.287 \text{ kJ/kg}\cdot\text{K})(295 \text{ K})} \approx 2.362 \text{ kg}$

$$m_2 = \frac{P_2 V}{R_{\text{air}} T_2} = \frac{(6 \times 10^5 \text{ Pa})(2.0 \text{ m}^3)}{(0.287 \text{ kJ/kg}\cdot\text{K})(350 \text{ K})} \approx 11.95 \text{ kg}$$

<CV1>

because this is a valve

$$\frac{dE}{dt}_{\text{sys}} = \dot{Q} + \sum \dot{m}_i h - \sum \dot{m}_e h \Leftrightarrow 0 = \dot{m}_3 h_3 - \dot{m}_4 h_4$$

$$\frac{dm}{dt}_{\text{sys}} = \sum \dot{m}_i - \sum \dot{m}_e \Leftrightarrow \dot{m}_i = \dot{m}_e \Leftrightarrow \dot{m}_3 = \dot{m}_4$$

So, \dot{m} into tank is steady

<CV2>

$$\frac{dU}{dt} = \dot{Q} + \dot{m} h \Leftrightarrow \frac{dU}{dt} = \frac{dQ}{dt} + h_i \frac{dm}{dt}$$

$$\int_{U_1}^{U_2} dU = \int_0^t \dot{Q} dt + h_i \int_{m_1}^{m_2} dm$$

$$U_2 - U_1 = Q_{12} + h_i (m_2 - m_1)$$

$$m_2 u_2 - m_1 u_1 = Q_{12} + h_i (m_2 - m_1)$$

$$\therefore Q_{12} = m_2 u_2 - m_1 u_1 - h_i (m_2 - m_1)$$

since the air temp from the vessel is

$$T_v = 295 \text{ K}$$

$$h_i = h|_{T=295\text{K}} = 295.1 \text{ kJ/kg} \quad (\text{from table})$$

$$\text{and } u_1 = 210.5 \text{ kJ/kg}$$

$$u_2 = 250.0 \text{ kJ/kg} \quad \text{from table}$$

$$\begin{aligned} \therefore Q_{12} &= (11.95 \text{ kg})(250.0 \text{ kJ/kg}) - (2.362 \text{ kg})(210.5 \text{ kJ/kg}) \\ &\quad - (295.1 \text{ kJ/kg})(11.95 \text{ kg} - 2.362 \text{ kg}) \\ &\approx -339.1 \text{ kJ} \end{aligned}$$

$$Q_{12} = -339 \text{ kJ}$$