



Walter J. KILNER Aura Lens

https://en.wikipedia.org/wiki/Walter_John_Kilner

Walter John Kilner

Walter John Kilner, M.D. B.A., M.B. (Cantab.) M.R.C.P., etc. (1847–1920) was a medical electrician at St. Thomas Hospital, London. There, from 1879 to 1893, he was in charge of electrotherapy. He was also in private medical practice, in Ladbroke Grove, London.

He wrote papers on a range of subjects but is today best remembered for his late study *The Human Atmosphere*. In 1883 he became a Member of the Royal College of Physicians. In his spare time he was a keen chess player.

The Human Atmosphere

In 1911 Kilner published one of the first western medical studies of the "Human Atmosphere" or Aura, proposing its existence, nature and possible use in medical diagnosis and prognosis. In its conviction that the human energy field is an indicator of health and mood, Kilner's study resembles the later work of Harold Saxton Burr. However, while Burr relied upon voltmeter readings, Kilner, working before the advent of semiconductor technology, attempted to invent devices by which the naked eye might be trained to observe "auric" activity which, he hypothesised, was probably ultraviolet radiation, stating that the phenomena he saw were not affected by electromagnets.[1]

Glass slides or "Kilner Screens"[2] containing alcoholic solutions of variously coloured dyes, including a blue coal-tar dye called "dicyanin" were used as filters in "Kilner Goggles" which, together with lights, were held to train the eyes to perceive electromagnetic radiation outside the normal spectrum of visible light.[3] After being so trained, one could dispense with the apparatus. Kilner did not recommend merely viewing the subject through these lenses.

According to his study, Kilner and his associates were able, on many occasions, to perceive auric formations, which he called the Etheric Double, the Inner Aura and the Outer Aura, extending several inches from patients' naked bodies, and his book gave instructions by which the reader might construct and use similar goggles.[1]

Francis J. Rebman, a friend of Kilner supported his research in America,[4]

A drawback to Kilner's method was the scarcity and toxicity of the chemicals he recommended. Later, the biologist Oscar Bagnall[5] recommended substituting the dye pinacyanol (dissolved in triethanolamine) but this dye is also not easy to obtain. Carl Edwin Lindgren has stated that cobalt blue and purple glass may be substituted for the dyes used by Kilner and Bagnall.[6]

In 1920 a revised edition of his book was published under the title *The Human Aura*. Kilner's work was well-timed for the heyday of Theosophy and his findings were incorporated into Arthur E. Powell's book *The Etheric Double*. [7] Powell rightly made clear that Kilner had expressly differentiated between his own work and the clairvoyance and eastern systems of spiritualism.

Critical reception

In the British Medical Journal (BMJ) a review for Kilner's research stated that although Kilner contended the aura is a "purely physical phenomenon", evidence does not support this view. Scientists from the BMJ attempted to replicate Kilner's experiments but the results were negative. The review concluded that "Dr. Kilner has failed to convince us that his "aura" is more real than Macbeth's visionary dagger." [8]

American scholar J. Gordon Melton has written:

"Kilner's research was largely dismissed by later researchers on light and perception, and the results he reported were seen as artifacts of the observer's own optic process rather than reflective of any emanation being produced by the subject being observed. These findings did not prevent the marketing of Kilner's goggles, advertisements for which appeared in Esoteric periodicals as late as the 1970s." [9]

Skeptical investigator Joe Nickell has described Kilner's research as pseudoscience, noting that he "uncritically accepted the validity of non-existent N-rays and clairvoyant powers." [10]

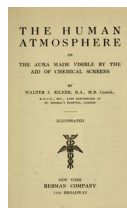
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http://www.mysticknowledge.org/The_Human_Atmosphere_By_Walter_J_Kilner.pdf

The Human Atmosphere by Walter J. Kilner



[PDF]

THE AURA OF HEALTHY PERSONS // THE ETHERIC DOUBLE // THE INNER AURA // OPTICAL PROBLEMS // THE EFFECTS OF THE DIFFERENT FORCES UPON THE AURA // COMPLEMENTARY COLOURS // THE OUTER AURA IN DISEASE // THE INNER AURA IN DISEASE // THE USE OF THE COMPLEMENTARY COLOURED BANDS IN DISEASE // THE AURA DURING PREGNANCY

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New Science and Invention in Pictures (1921)

The Human Aura by Hugo Gernsback

Dicyanin

SOME years ago, Dr. Kilner, late electrician at St. Thomas' Hospital, London, whose book on the human aura is well known, said that the discovery of a screen to make the aura visible was not accidental. He had been reading about the action of the N-rays upon phosphorescent sulfate of calcium and had experimented upon mechanical forces of certain bodily emanations.

Early in 1908 he thought certain dyes might help him, and fixed on the coal-tar dye dicyanin.

When the dye was obtained glass screens were made, the only satisfactory ones being glass cells filled with an alcoholic solution of dicyanin.

It was recognised very early that constantly looking through the dicyanin screens had an influence on the eyesight, an influence which is not altogether understood today. Everyone who uses the screens finds the eyesight markedly improved. Most were slightly presbyopic.

Modern commercial photographic plates were found comparatively insensitive beyond 450, but experiments at the Bureau of Standards, Washington, D. C., convinced the experimenters that with dicyanin it would be possible to get the spectra of bright stars as far as 480.

But the dye deteriorated quickly and there was ignorance of the spectral region in which dicyanin was most valuable. Dicyanin could not be depended on to preserve its useful properties indefinitely under ordinary conditions. Some change took place which destroyed its sensitizing value.

The fact that dicyanin does not sensitize the plates to green and yellow is apparent from the photographed spectra. The blue and red portions of the spectra are usually separated by a gap representing the insensitive organ. The addition of pinaverdol to the staining bath is an improvement.

Staining Solution	cc.
Distil. water	140
Ethyl alcohol	120
Dicyanin A (1:2000)	18
or Dicyanin 1:1000	7
Ammonia, 26 degrees	9
The human aura consists of a number of layer or strata one beyond the other extending out into space. There is the "etheric double," which entirely surrounds the body, conforming to its shape. Beyond this is the inner aura, 2 or 3 inches broad, and, beyond this again, the outer aura extending 5 or 6 inches, the outer part of which is termed the ultra-outer double.	
Dr. Kilner thinks the force from which the aura arises is generated within the body itself. He tells of details concerning the effect of electricity and chemicals on the inner aura, and the changes in shape and size of the aura generally as the result of nervous disease.	

<https://books.google.com/books?id=CTKViOzOOSUC&pg=PA491&lpg=PA491&dq=make+dicyanin:dye&source=bl&ots=PVoPJHAB9&sig=ACfU3U3P42U9yq3GrDIY6Kgj5cd0lGEVvw&hl=en&sa=X&ved=2ahUKFwjSeMGKMjnAhNS7Z4KHZFwAAU4ERDwATAJegQIRxAB#v=onepage&q=make%20dicyanin%20dye&f=false>

Application of Dicyanin to the Photography of Stellar Spectra
By Paul Willard Merrill

2. PHOTOGRAPHIC PROCESSES

The exact action of the **dye** in sensitizing plates is so little understood, and the subject is so complex, that in spite of the large amount of research upon it, any statements made at this time in regard to the best procedure must be regarded as provisional, as they are more than likely to be revised by subsequent experience. Similarly, certain precautions mentioned below may later be shown to be unnecessary.

Both **dicyanin** and a more recent product, **dicyanin A**, were tried at Harvard. The difference in the action of the two dyes is not very marked, **dicyanin A**, perhaps, being better for the very long waves. The dyes should be obtained as soon after manufacture as possible and should be kept cool and free from moisture. The stock solution should be made in the ratio 1:1000 or 1:2000

THE STOCK SOLUTION SHOULD BE MADE IN THE RATIO 1 : 1000 OR 1 : 2000 in the highest grade of absolute ethyl alcohol. The best results obtained by the writer have been with the following procedure:

Staining Solution	
	Cubic centimeters
Distilled water.....	140
Ethyl alcohol.....	120
Dicyanin A (1: 2000).....	18
Or dicyanin 1: 1000.....	7
Ammonia, 26 degrees.....	9

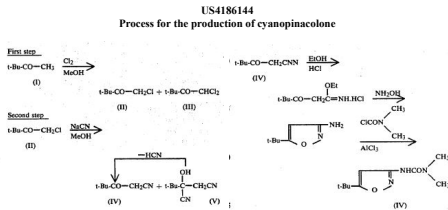
Directions for Sensitizing.—Mix the water and alcohol thoroughly and allow to stand. After adding the dye allow several minutes before the ammonia is put in and again an interval before staining. Bathe the plates 4¼ minutes at 20° C (68° F). Place for about 40 seconds in a bath of pure alcohol at the same temperature, remove,

492 *Bulletin of the Bureau of Standards* [Vol. 14

and dry in a current of air.⁹ Avoid high temperatures at any point of the process.

The bath can be used repeatedly if 3 or 4 c c of ammonia is added each time. The dye and all solutions should be protected from light. In packing the plates after staining do not allow the film to come in contact with any solid object, even glass or another film. The stained plates deteriorate rather quickly; if placed in an ice box, they will keep fairly well for a few days, but it is advisable to sensitize them as shortly before using as convenient. Use any clear working developer ¹⁰ a few degrees cooler than usual.

The fact that dicyanin does not sensitize the plates to green and yellow light is made apparent by the appearance of the photographed spectra. As the illustrations show, the blue and red portions of the spectra are usually separated by a gap representing the insensitive region. A more nearly uniform spectrum can undoubtedly be obtained by the addition of pinaverdol ¹¹ to the staining bath.



<https://mindmachine.ru/catalog/en/Stalker/>

Glasses PranaVision Stalker

Glasses vision of aura and subtle world.

Adjustment of the spectrum can be done in two ways:

1. Loosening the screws on the side of the eyepieces, to carry out the rotation of the eyepieces of glasses with handles.
2. Fixing the eyepiece in the selected position, make small movements of the head in the up-down direction.

The differences between models and Econom Stalker type lenses - they work on the uptake and lighter for street lighting. The main distinguishing feature of the model Stalker is the ability to adjust the spectrum of the blocking of the incoming light. The ability to adjust the spectrum was made possible by changing the angle of incidence to the surface with special reflective filters...

The ability to adjust the spectrum from blue to violet and red, while maintaining the bandwidth of the optical filter in the ultraviolet region, gives spectra almost all the range of dyes used for seeing auras and energies of the subtle world. For example, one of the most effective dyes dicyanin, corresponds to the 90° angle of incidence on the filter surface.

[https://hatch.kookscience.com/wiki/Dicyanin_\(synthetic_dye\)](https://hatch.kookscience.com/wiki/Dicyanin_(synthetic_dye))

Dicyanin (synthetic dye)

Dicyanin is a synthetic blue dye derived from coal tar that was used for sensitizing photographic plates, first manufactured in the early twentieth century by the dye works of Meister, Lucius & Brining at Höchst (near Frankfurt), Germany. It became associated with auric research thanks to the experiments of Walter J. Kilner, who used the dye (which he also referred to as specturamine) in the manufacture of his Kilner screens.

Preparation

In Dyestuffs & Coal-Tar Products (1915), the authors — Beacall, Martin, et. al. — relate that the dicyanin is a quinoline dye of the cyanine type, reporting it is "prepared by action of KOH [potassium hydroxide] + atmospheric O [oxygen] on α - γ -Dimethylquinolinium salts," for intended use as a sensitizer of "silver bromide gelatine plates up to the red line α , with a strong minimal effect between E and F," specifically noting it is "used as a red sensitizer for scientific work, but the prepared plates are not very sensitive." [1]

Later, in Constitution of the Dicyanines (1924), Mills and Odams demonstrated dicyanine is a carbocyanin, giving a formula of 2,4-dimethylquinoline iodoethylate and sodium methylate in methanol[2] (for 1,1'-diethyl-2,4'-carbocyanine iodide).[3]

Studies

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The Synthesis of Dicyanine A

By S. Palkin

<https://www.docdroid.net/CLKqg27/dicyanine-a-mikeska1920.pdf>

Synthesis of Photosensitizing Dyes (II), Dicyanine A\

A. Mikeska, et al.

<https://www.docdroid.net/mGviJux/tetramethylquinolines-mikeska1920.pdf>

TetraMethyl-Quinolines

L.A., Mikeska, E. Adams

Cyanopinacolone & Cyanogen Preparation Patents

CA1093097A

PROCESS FOR THE PRODUCTION OF CYANOPINACOLONE

Abstract

Cyanopinacolone is produced by reacting pinacolone with about 1.0 to about 1.2 molar equivalents of chlorine in methanol and further reacting the resulting monochloropinacolone with about 1.0 to about 1.2 molar equivalents of an alkali metal cyanide in methanol.

US5017356A

Preparation of cyanogen from glyoxime

Abstract

Cyanogen is prepared from glyoxime by reaction with acetic anhydride to form N,N'-diacetoxyglyoxime which is then pyrolyzed to cyanogen.

NL7105148A

Cyanogen preparation

Abstract

Cyanogen is produced by catalytic oxidation of HCN according to BE 764361 by oxidation with H2O2 in the presence of CuSO4 with or without ferric sulphate. The ferric sulphate may be prepared in situ by the reaction of H2O2 with ferrous sulphate. The Cu++ or Fe+++ ions may be recovered by precipitation with alkaline or alkaline-earth carbonates in the cyanogen-free reaction-soln. The cyanogen-free catalytic soln. may be recycled to the reaction step after the water introduced with the H2O2 has been distilled off.

BE766248R

Cyanogen preparation

Abstract

Cyanogen is produced by catalytic oxidation of HCN according to BE 764361 by oxidation with H2O2 in the presence of CuSO4 with or without ferric sulphate. The ferric sulphate may be prepared in situ by the reaction of H2O2 with ferrous sulphate. The Cu++ or Fe+++ ions may be recovered by precipitation with alkaline or alkaline-earth carbonates in the cyanogen-free reaction-soln. The cyanogen-free catalytic soln. may be recycled to the reaction step after the water introduced with the H2O2 has been distilled off.

US4503025A

Process for preparing dicyanogen

Abstract

Dicyanogen is prepared by reacting trimethylsilyl cyanide with cyanogen halides in the presence of Lewis acids.

JPH07315831A

METHOD FOR PREPARING AMMONIA-FREE CYANOGEN GAS

Abstract
PURPOSE:To surely trap ammonia and to obtain an ammonia-free cyanogen gas by a simple procedure, by treating a cyanogen-containing mixed gas blended with an ammonia gas by an acidic ion exchange resin. CONSTITUTION:A cyanogen-containing mixed gas blended with an ammonia gas is led to a container charged with an acidic ion exchange resin and passed through the container so that the mixed ammonia gas is caught by the ion exchange resin to give an ammonia-free cyanogen gas. This method is applicable to the case of $\text{-}^{11}\text{-C}$ labeled cyanide-containing gas in which the cyanogen-containing mixed gas is formed from $\text{-}^{11}\text{-C}$ labeled methane.

GB579785A
Production of cyanogen

Abstract
Cyanogen is obtained by reacting hydrocyanic acid with chlorine or cyanogen chloride in the vapour phase at an elevated temperature. The reaction may be effected in an empty glass or quartz tube having smooth walls at a temperature not lower than 600-700 $\text{^{\circ}C}$., or in a reaction chamber packed with inert solids, but preferably, the reaction mixture is passed over a surface-active material at a temperature of 200-1,000 $\text{^{\circ}C}$. Suitable surface-active materials are active carbon or silica gel. If the reaction is conducted under essentially adiabatic conditions, the gases are fed to the reaction vessel at temperatures suitably below the final reaction temperature, yet high enough to initiate the reaction. Preferably when operating adiabatically the reaction mixture is diluted with partially cooled off gas from the reaction chamber, thus recycling part of the product. The hydrocyanic acid may be employed in slight stoichiometrical excess. The offgas is cooled, scrubbed with water to remove hydrochloric acid as a concentrated solution and dried. Finally, the cyanogen may be purified from its slight content of hydrocyanic acid by rectification.

GB1306528A
PRODUCTION OF CYANOGEN

Abstract
Cyanogen is obtained by initially reacting HCN and Cl 2 on a surface active catalyst (i.e. carbon) and then treating the gas mixture still containing HCN with 1Å0 to 1Å1 mols of Cl 2 per mol of HCN on a surface active catalyst at 195-800 $\text{^{\circ}C}$. for 10 seconds to 0Å01 second. In a preferred embodiment the initial stage is carried out in two steps. In the first step the HCN and Cl 2 are reacted in a mole ratio of 1 : 1 or less and a residence time of greater than 10 \times 3 second and the residual amount of HCN required to use up the excess Cl 2 is directly introduced to the second reactor together with the gas mixture from the first reactor.

CN206566534U
Neighbour's cyanogen stills

Abstract
The utility model discloses a neighbour's cyanogen stills, including barrel, agitator, thermowell, cover and speed reducer, be equipped with the cover directly over the barrel, the cover passes through sealing device with the barrel to be connected, and the agitator upper end links to each other with the speed reducer, and the lower extreme passes through the cover tromppl and extends to go into inside the barrel, thermowell open position extends to go into inside the barrel in barrel upper end, thermowell, then, the utility model discloses stills distillation heat transfer even speed is fast, and distillation efficiency is high, from having saved the resource to a certain extent, and be difficult to the scale deposit in the cauldron, the time of not only having protected the cauldron body but also having avoided the clearance to cause is extravagant.

GB845471A
Method of preparing cyanogen

Abstract
A gaseous mixture of HCN and Cl2 is ignited in the presence of H2 at at least 1200 DEG C. to form cyanogen which is subsequently recovered. Cl2, HCN and H2 in molar ratio 1:1.5-2.0:0.25-2.0 respectively, preheated if desired to 200 DEG C.-600 DEG C. are supplied either separately or variously premixed e.g. HCN and Cl2 may be premixed to a burner, where they p are ignited either by an oxy-hydrogen torch, an electric spark or a heated Pt wire, and burn at a temperature of at least 1200 DEG C. usually 1200-1800 DEG C. Combustion products containing cyanogen formed in a reaction time of 0.001-0.00001 seconds are removed from the burner and passed say through a water scrubber to remove acidic and other impurities, cyanogen being recovered as the resultant purified gas. The combustion may be carried out in the presence of an inert gas e.g. N2 or He and preheating of the separate gases or gaseous mixture is obtained either by heat exchange with the combustion products leaving the burner or by independent electrical heating.

JP2000178779A
PRODUCTION OF CYANOGEN

Abstract
PROBLEM TO BE SOLVED: To improve a conversion rate by electrolytically oxidizing cyan ion in a cyan ion-containing solution buffered to a specific pH range. SOLUTION: The cyan-containing solution is buffered with a buffer such as phosphoric acid/boric acid/acetic acid/sodium hydroxide or the like to pH 4-7. As the cyan ion source, a metallic cyanide such as sodium cyanide generating cyan ion by dissolving in water is used and the cyan ion concentration in the solution is controlled to 0.01-10 mol%. A redox active metallic ion such as Cu ion, iron ion is further incorporated in the cyan-containing solution. A titanium electrode coated with a noble metal oxide or various kinds of carbon electrodes are used as the electrode material of the anode and the cathode. The electrolytic reaction is performed usually under a condition of a fixed current and a fixed voltage at 0.1-500 mA/cm2 current density and room temp. to 100 deg.C. Gaseous nitrogen or the like is passed through the electrolyte and the electrolytic cell as many small bubbles and the generated cyanogen is carried to the outside of the system.

GB1338025A
PRODUCTION OF CYANOGEN

Abstract
Cyanogen is produced by catalytic oxidation of HCN with H 2 O 2 in the presence of CuBr 2 , CuCl 2 , Cu(CN) 2 , Cu(NO 3) 2 , or CuSO 4 . Fe3+ salts may also be present, with anions corresponding to the Cu salt (Fe 2 (SO 4) 3 may be formed in situ from FeSO 4). The HCN may be used as a gas, liquid or aqueous solution, and the H 2 O 2 as a 3-90% aqueous solution or as an organic solution. The temperature may be 0-100 $\text{^{\circ}C}$. and the pressure 1-5 atm. An organic solvent, e.g. tetramethylene sulphone, ethyl acetate or n-propyl acetate, may be present. The process may be operated continuously, the catalyst solution being distilled to remove water and recycled, any HCN and CN evolved during distillation also being recycled. Alternatively, the solution, optionally after distillation, may be treated with alkali metal or alkaline earth metal carbonate, e.g. CaCo 3 , to precipitate Cu and Fe carbonates which may then be converted, e.g. into sulphates with H 2 SO 4 , and recycled. CNCl is formed as by-product in Examples 6 and 7.

US3302996A
Process for preparing cyanogen

CA186179A
CYANOGEN PRODUCTION

US2075046A
Photographic emulsion containing supersensitized pinacyanol
