Acids,Bases and Salts

Learning Points

1. Definitions in terms of furnishing of H+ and OH– ions
2. General properties of acids and bases
3. Examples and uses
4. Chemical reactions of acids and bases
    - Reaction with metals
    - Reaction with Metal carbonates
    - Reaction of metal oxides with acids
    - Reaction of non-metallic oxides with bases
5. Neutralization reaction
6. concept of pH scale
7. importance of pH in everyday life
8. Chemicals from common salt
    - Sodium Hydroxide,
    - Bleaching powder
    - Baking soda
    - Washing soda
    - Plaster of Paris.

Introduction

An acid is any hydrogen-containing substance that is capable of donating a proton (hydrogen ion) to another substance. A base is a molecule or ion able to accept a hydrogen ion from an acid. Acidic substances are usually identified by their sour taste. Salt is formed by the reaction of acid and base. Salt is a neutral substance.