

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

GCE Advanced Subsidiary Level and GCE Advanced Level

**MARK SCHEME for the May/June 2009 question paper
for the guidance of teachers**

9701 CHEMISTRY

9701/31

Paper 31 (Advanced Practical Skills 1), maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

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This is given by the expression $\frac{47.25}{\text{volume diluted}} \times \text{Examiner selected titre}$

Question	Sections	Indicative material	Mark	
1 (a)	PDO Layout	(i) Tabulates initial and final burette readings and volume added in each of the tables. <i>Do not award this mark if any final and initial burette readings are inverted or 50 is used as the initial burette reading.</i>	1	
	PDO Recording	(ii) Both burette readings in the dilution table and <u>final and initial</u> burette readings for all accurate titres in the titration table recorded to the nearest 0.05 cm ³ .	1	
	MMO Collection	(iii) Follows instructions: dilutes 47.00 cm ³ to 47.50 cm ³ and has any two titres within 0.20 cm ³	1	
	MMO Decisions	(iv) Has at least two uncorrected “accurate” titres within 0.1 cm ³ <i>Do not include any titre labelled “rough”/“trial” unless the candidate has ticked that value or used it in an expression when calculating the average in (b).</i>	1	
	MMO Quality	Accuracy (v) and (vi) Give (v) and (vi) if difference to Supervisor is 0.3 or less Give (vi) only for a difference of 0.3+ to 0.5 Give neither mark for a difference greater than 0.5 .	2	
(b)	ACE Interpretation	Candidate selects/calculates appropriate “average” from any uncorrected titre values within 0.20 cm ³ . <i>Candidate is permitted to use a titre labelled “rough” or “trial”.</i> <i>Titres to be used must be shown.</i> Where all titres are given to 1 decimal place the average should be calculated correct to 1 or 2 decimal places. Where any titre is recorded to 2 decimal places, the average should be calculated to 2 decimal places or rounded to the nearest 0.05 cm ³ .	1	

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[1]

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(c)	ACE Interpretation	<p>(i), (ii) and (iii)</p> <p>Award three marks if all steps are chemically correct. Withhold 1 mark for each chemical error – no negative marks. Count non-completed steps as chemical errors.</p> <p>step 1 $\frac{\text{titre}}{1000} \times 0.15$</p> <p>step 2 $\times \frac{1}{2}$</p> <p>step 3 $\times 2$</p> <p>step 4 $\times \frac{1000}{25}$</p> <p>step 5 $\times \frac{250}{\text{volume diluted}}$</p> <p>step 6 $\times 249.6$</p>	3	
	PDO Display	(iv) Working shown in at least three of the 5 steps	1	
		(v) Answers to 3 or 4 significant figures in final answer to each step attempted (minimum of three steps required)	1	[5]
(d)	ACE Interpretation	Explains that the maximum error is given by + 0.05 cm ³ on one burette reading and –0.05 cm ³ on the other burette reading, or Individual errors are in opposite directions.	1	[1]
(e)	ACE Interpretation	Calculates $\frac{0.1}{\text{titre}} \times 100 \%$ Answer must be correct to 2 or 3 decimal places.	1	[1]
[Total: 14]				

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Supervisor's Report

From the Supervisor's experimental results **round times to the nearest second** and calculate the average of (volume of sodium thiosulfate x time) for 50 cm³ and for 25 cm³ of sodium thiosulfate.

Candidate's scripts

From the candidate's experimental results **round times to the nearest second** and calculate (volume of sodium thiosulfate x time) as above.

Record values of (V x t) on script and use in assessing accuracy marks.

Question	Sections	Indicative material	Mark	
2 (a)	PDO Layout	(i) Tabulates all experimental readings: volumes of sodium thiosulfate and water, time and rate ($1/t$)	1	
	PDO Recording	(ii) Single table covering all three experiments <i>A single table has no repetition of column headings.</i>	1	
		(iii) Table has correct labels and units: e.g. /cm ³ , /cubic centimetres, or (cm ³), (cubic centimetres) or volume in cubic centimetres; Similarly for time (s or seconds but not sec) and rate (s ⁻¹ , rate (in) per second; $1/s$ etc.) At least two different units are required. <i>Where units have not been included in the column or row header there should be the appropriate unit for <u>each entry</u> in the table.</i>	1	
	MMO Collection	(iv) All times of reaction are recorded to the nearest second (no decimal places).	1	
	MMO Quality	(v) and (vi) Give (v) and (vi) if difference between candidate's (V x t) values (50 & 25 cm³ FA 1), is within 5% of the larger value. <i>Give (vi) only if the difference is > 5% but ≤10% of the larger value.</i>	2	
		(vii) and (viii) Compare the closer of the candidate's (V x t) values with the Supervisor's average Vt. Give (vii) and (viii) if difference is within 10% of the <u>Supervisor's value</u> . <i>Give (viii) only if the difference is > 10% but ≤20% of the <u>Supervisor's value</u>.</i>	2	
	MMO Decisions	(ix) Selects (10–15) or (35–40) cm ³ sodium thiosulfate and an appropriate volume of water to give a total volume of 50 cm ³ (or 55 cm ³ if the volume of acid is tabulated).	1	

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(b)	ACE Interpretation	Candidate shows by calculation or by mathematical expression that $[\text{Na}_2\text{S}_2\text{O}_3] \propto \text{volume of Na}_2\text{S}_2\text{O}_3(\text{aq})$ in 50/55 cm^3 of solution. <i>Reference to $(^{50}/_{50}$ or $^{50}/_{55}$) and $(^{25}/_{50}$ or $^{25}/_{55}$)</i>	1	[1]
(c)	ACE Conclusions	Explains that rate is given by the inverse of time or is inversely proportional to time; or Rate $\propto \frac{1}{\text{time}}$ <i>Allow Rate = $\frac{1}{\text{time}}$</i>	1	[1]
(d)	ACE Interpretation	(i) Correctly evaluates all V_t values (using times given by candidate, including decimal places where appropriate – or (ii) gives an appropriate qualitative statement relating (rate or time) and concentration. <i>Award this mark if either is correct.</i>	1	[2]
	ACE Conclusions	Gives a quantitative description of relationship. <i>V_t values are required but do not have to be correctly evaluated. Where no pattern is obvious accept an appropriate statement to that effect.</i>	1	
(e)	ACE Improvements	Explains that volumes of reactants or concentration (of thiosulfate and acid) must be kept constant and describes <u>how</u> the temperature will be varied.	1	[1]
				[Total: 14]

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Question	Sections	Indicative material	Mark	
FA 5 is $K_2CrO_4(aq)$; FA 6 is $NaNO_2(aq)$; FA 7 is $Pb(NO_3)_2(aq)$, FA 8 is $MgSO_4(aq)$				
3 (a)	MMO Collection	(i) Records no reaction, no change or no precipitate on adding NaOH and $NH_3(aq)$ to FA 5 and FA 6 .	1	
		(ii) Records white ppt soluble (in excess NaOH) and white ppt insoluble (in excess NH_3) with FA 7	1	
		(iii) Records white ppt insoluble (in excess for both NaOH and NH_3) with FA 8	1	
	ACE Conclusion	(iv) Conclusion is marked consequentially from the observations for a single cation and a pair of cations. <i>Mg^{2+}/magnesium from white ppt insoluble in an excess of $NaOH(aq)$ and in an excess of $NH_3(aq)$</i> <i>Ca^{2+}/calcium from white ppt insoluble in an excess of $NaOH(aq)$ no ppt in $NH_3(aq)$</i> <i>Pb^{2+}/Al^{3+} from white ppt soluble in an excess of $NaOH(aq)$ and insoluble in an excess of $NH_3(aq)$</i> <i>Ba^{2+}/NH_4^+ from no ppt with $NaOH(aq)$ or $NH_3(aq)$</i> FA 6 only	1	
(b)	MMO Decisions	(Warms) with NaOH and Al(s) and records appropriate test for ammonia. Gas must be tested in at least one test. <i>This is a mark for the method not the observation.</i>	1	[4]
	ACE Conclusion	Must have indication that the test has been performed with FA 6 , FA 7 and FA 8 . In awarding the conclusion mark, assume, in this section only , that a blank box indicates no reaction (no ammonia detected). Award this mark for any of the following: (i) a conclusion, from correct observations, that FA 6 and FA 7 contain nitrate or nitrite (ii) correct observations for NH_3 – only with FA 6 and FA 7 , but no conclusion given (iii) a statement that NH_3 is evolved – only with FA 6 and FA 7 (iv) observation that red litmus turns blue (gas not needed) – only with FA 6 and FA 7	1	
[2]				

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(c)	MMO Collection	(i) Observes a change in colour (from yellow) to yellow/orange or orange (solution), no ppt , with FA 5 and a white ppt with FA 7 .	1	[4]
	ACE Conclusions	(ii) Observes a brown gas formed with only FA 6 .	1	
		Mark (iii) and (iv) consequentially to observations		
		(iii) Give this mark for one conclusion providing it is supported by an acceptable explanation.	2	
		(iv) Give this mark for two further conclusions supported by acceptable explanations.		
		Minimum acceptable supporting evidence: CrO ₄ ²⁻ from yellow soln or soln turning orange in acid NO ₂ ⁻ from brown gas or from effervescence/fizzing/bubbling with acid, if named soln has yielded ammonia or an alkaline gas in (b) NO ₃ ⁻ no brown gas etc with acid, but ammonia evolved in (b) Pb ²⁺ white ppt with HCl if Pb ²⁺ in (a) (iv) Al ³⁺ no white ppt with HCl if Al ³⁺ in (a) (iv)		
(d)	MMO Collection	Mixes FA 5 and FA 7 and observes a yellow ppt. If this section has not been attempted, the correct <u>observation</u> on mixing FA 5 and FA 7 can be carried forward from the conclusions in (c) .	1	[2]
	ACE Conclusions	Concludes that FA 5 contains CrO ₄ ²⁻ and FA 7 contains Pb ²⁺ providing the ions have been previously mentioned in (a) or (c) .	1	
[Total: 12]				