

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

GCE Advanced Subsidiary Level and GCE Advanced Level

**MARK SCHEME for the October/November 2009 question paper
for the guidance of teachers**

9701 CHEMISTRY

9701/33

Paper 33 (Advanced Practical Skills), maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

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Page 2	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE A/AS LEVEL – October/November 2009	9701	33

Question	Sections	Indicative material	Mark	
1 (a)	PDO Layout MMO Collection	Three weighings recorded unambiguously (words needed) Mass used between 1.80 g and 2.00 g (for all experiments) Mass decreases after heating in a single experiment attempted (no repeat experiment allowed) (ignore heating to constant mass)	1 1 1	[3]
(b)	ACE Interpretation	Correct subtraction from experimental results in (a) for mass of water lost and for mass of residue (correct for any experiment attempted or for an average)	1	[1]
(c)	PDO Display ACE Interpretation	Shows working using values from (b) (can be ecf): $\frac{\text{mass of salt}}{151.9}$ and $\frac{\text{mass of water}}{18}$ or candidate's value of M_r Correct calculation of each value and must have between 2 and 4 sf and correct rounding (NO ecf from incorrect M_r but ecf allowed from masses used in first part)	1 1	[2]
(d)	ACE Interpretation PDO Display	Values from (c) shown in correct calculation or ratio or value of correct calculation showing dp and correct rounding for sf shown Value of x for fully correct answer given as an integer allow 0.5 to go up or down (ecf allowed) (If correct integer shown but no expression or calculation is shown award second mark only)	1 1	[2]

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Question	Sections	Indicative material	Mark	
(e) continued	PDO Display	(xi) Correct mean displayed to same number of decimal places as burette readings (If any burette reading is to 2 dp then mean to 2 dp; if two titres 0.05 apart then 0.025 or 0.075 is acceptable; similarly if titres recorded to 1 dp then 0.05; to no dp then to 0.5. If three titres used giving mean of 0.033... or 0.066... then allow mean to nearest 0.05)	1	[11]
(f)	ACE Interpretation	Maximum error for single reading of burette given as 0.05 cm ³ (even if readings recorded to 1 dp) Correct calculation to 2 sf (minimum) showing correct rounding of maximum % error $\frac{2 \times \text{candidate's error}}{\text{titre}} \times 100$ or allow $\frac{0.10}{\text{titre}} \times 100$ if 0.1 in first box.	1 1	[2]
(g)	ACE Conclusions	Mark (g) and (h) as a single unit (Mass loss is too high) Candidate suggests spitting or decomposition or overheating/heating too long (The latter will not gain any other mark unless a consequence of overheating is given.)	1	[1]
(h)	ACE Improvement	Use of lid or gentle heating (at first) – to reduce spitting/loss of solid or reduced temperature or heat for a shorter time – to prevent decomposition <u>Explains</u> that mass loss is too high or wtte (e.g. some solid spits out) and how modification will reduce this or a realisation that something other than water is lost.	1 1	[2]
Qn 1	Total			[24]

Page 5	Mark Scheme: Teachers' version	Syllabus	Paper
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Question	Sections	Indicative material	Mark	
FA 5 is $\text{NaNO}_2(\text{aq})$; FA 6 is $\text{NaCl}(\text{aq})$; FA 7 is KI; FA 8 is distilled water				
2 (a)	MMO Decisions	Chooses a named (dilute) strong acid to identify the nitrite (name could be in (b)). and Chooses silver nitrate or lead nitrate to identify the chloride and iodide ($\text{Ag}^+_{(\text{aq})}$ or aqueous solution of silver ions accepted, ditto lead ions)		[1]
(b)	MMO Collection	(i) Records a brown gas from FA 5 and acid and no reaction with second reagent. (ii) Records a colourless or blue solution on adding acid to FA 5 (iii) Records a white precipitate when Ag^+ or Pb^{2+} is added to FA 6 and no reaction with acid (iv) Records a yellow precipitate when Ag^+ or Pb^{2+} is added to FA 7 and no reaction with acid. If acid alone used and all four observations are correct (i) and (ii) and identity mark (v) can be awarded. If AgNO_3 or $\text{Pb}(\text{NO}_3)_2$ only used and all four observations are correct then can award (iii) and (iv) and identity mark (vi). If a white ppt is recorded with FA 5 and all other observations are correct award (iii) and (iv) but then deduct one observation mark for the FA 5 result. (Annotate the paper near the boxes.) Mark (vi) would be available to a candidate concluding both FA 5 and FA 6 are Cl^- and FA 7 is I^- . If a spurious reagent is chosen in (a) ignore and mark other (if correct) as single reagent. If three reagents are used, mark the observations for acid and the first of Ag^+ or Pb^{2+} mentioned in (a). Deduct one mark (from (i) – (iv)) if erroneous observation with FA 8 or test not done. A dash is similarly penalised (once, at the first opportunity)	1 1 1 1	[7]
	ACE Conclusions	(v) and (vi) are consequent on observations (v) Identifies nitrite in FA 5 Allow from effervescence or bubbling from FA 5 + acid only. (vi) Identifies chloride in FA 6 (only, unless single reagent as above) and iodide in FA 7	1 1	
	ACE Interpretation	(vii) Gives appropriate evidence for identification of all three ions (consequential on observations) (Brown gas with acid + FA 5 needed here.)	1	
(c)	MMO Decisions	Add (aqueous) ammonia to the precipitate formed with silver nitrate or Add (aqueous) lead nitrate or silver nitrate (to a fresh sample of) the solution Do not penalise lack of (aq) with an ion if already penalised in (a)	1	[1]
Qn 2	Total			[9]

Page 6	Mark Scheme: Teachers' version	Syllabus	Paper
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Question	Sections	Indicative material	Mark	
FA 9 is Pb(NO₃)₂, FA 10 is K₂CrO₄				
3 (a)	MMO Collection	White ppt formed in tests (i) and (ii) and yellow ppt in test (v) (no variation on white; ppts not soluble in excess acid) Solution changes from yellow to orange (both colours and 'solution' needed) in (iii) (Solution) changes to green or blue (blue/green, cyan, turquoise but no orange or orange–brown as final colour) in (iv) Ignore state	1 1 1	[3]
(b)	ACE Conclusions	FA 9 contains the cation Pb ²⁺ or lead II (all three ppts must be in (a) but ignore any PbCl ₂ ppt redissolving; no Ba ²⁺ unless no ppt with HCl) (White) ppt with both HCl and H ₂ SO ₄ given as evidence or yellow ppt with FA10 named as chromate or CrO ₄ ²⁻	1 1	[2]
(c)	ACE Conclusions	States that ethanol is being oxidised or is reducing agent or reduces FA10 Ethanol oxidised to ethanal or alcohol oxidised to aldehyde or chromate VI reduced to chromium III or dichromate reduced to chromium III	1 1	[2]
Qn 3	Total			[7]