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Chapter 16 Properties Of Solutions

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Chapter 16: Colligative Properties of Solutions 45 16-4. The mole fraction of $(\text{NH}_4)_2\text{SO}_4(\text{aq})$ is given by $x_{(\text{NH}_4)_2\text{SO}_4} = \frac{n_{(\text{NH}_4)_2\text{SO}_4}}{n_{(\text{NH}_4)_2\text{SO}_4} + n_{\text{H}_2\text{O}}}$ Because $(\text{NH}_4)_2\text{SO}_4(\text{aq})$ is a strong electrolyte, it dissociates completely into $\text{NH}_4^+(\text{aq})$ and $\text{SO}_4^{2-}(\text{aq})$ ions. Assume a one kilogram solution. The number of moles of ions in one ...

CHAPTER 16. Colligative Properties of Solutions

Properties of a solution that depend only on the number of particles. A nonvolatile substance is one that does. Look at Figure 16.13 on page 487. How is the decrease in vapor pressure of a solution with a nonvolatile solute. Assume 3 mol each of three different solutes have been added to three.

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Chapter 16 - Solutions - 16.1 Properties of Solutions - 16.1 Lesson Check - Page 524: 3 Answer Solute size, temperature, agitation, and surface area all affect the speed at which a substance will dissolve.

Chapter 16 - Solutions - 16.1 Properties of Solutions - 16 ...

Chapter 16 - Solutions - 16.1 Properties of Solutions - 16.1 Lesson Check - Page 524: 5 Answer The question is asking you to identify / recall factors that contribute to solubility, which is the mass of solute that dissolve in a given mass of solvent.

Chapter 16 - Solutions - 16.1 Properties of Solutions - 16 ...

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