

Chemistry If8766 Significant Figures Answer Key

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Chemistry If8766 Significant Figures Answer

Example: 100.0 has 4 significant figures. 100 has 1 significant figure. 1 4) Zeros in the beginning of a number whose only function. 1 is to place the decimal point are not significant. , Example: 0.0025 has 2 significant figures. 5) Zeros following a decimal significant figure are significant.

SIGNIFICANT FIGURES Name - Dearborn Public Schools

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Chemistry If8766 Worksheets - Printable Worksheets

Name: _____. Calculations using Significant Figures. When multiplying and dividing, limit and round to the least number of significant figures for any of the factors. Example 1: $23.0 \text{ cm} \times 432 \text{ cm} \times 19 \text{ cm} = 188,784 \text{ cm}^3$ The answer is expressed as $190,000 \text{ cm}^3$ since 19 cm has only two significant figures.

Calculations using Significant Figures

Arkansas State University Department of Chemistry and Physics. 2. Determine the number of significant figures in each of the following: a) 3427 b) 0.00456 c) 123,453 d) 172 e) 0.000984 f) 0.502 g) 3100.0×102 h) 0.0114×104 i) 107.2 j) 0.0000455 k) 2205.2 l) 30.0×10^{-2} m) 0.982×10^{-3} n) 0.0473 o) 650.502 p) 3.03×10^{-1} q) 20.4×10^5 r) 1.29 s)...

Worksheet 1 - Scientific notation/ significant figures

Losing Significant Figures. Sometimes significant figures are 'lost' while performing calculations. For

example, if you find the mass of a beaker to be 53.110 g, add water to the beaker and find the mass of the beaker plus water to be 53.987 g, the mass of the water is. $53.987 - 53.110 \text{ g} = 0.877 \text{ g}$.

Calculations Using Significant Figures (Sig Figs)

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