

## *Chapter 10 Chemical Quantities Assessment Answers*

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### **Chapter 10 Chemical Quantities Assessment**

Chapter 10 – Chemical Quantities. Jennie L. Borders. Section 10.1 – The Mole: A Measurement of Matter. You often measure the amount of something by count, by mass, or by volume. ... Section 10.1 Assessment. Describe the relationship between Avogadro's number and one mole of any substance.

### **Chapter 10 - Chemical Quantities**

Chapter 10 - Chemical Quantities - 10 Assessment - Page 342: 121 Answer An atom is the smallest particle of an element that retains its identity in a chemical reaction, whereas a molecule is a group of atoms joined together by covalent bonds.

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CHAPTER 10: Chemical Quantities BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

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Chapter 9 Chemical Quantities 1. Although we define mass as the "amount of matter in a substance," the units in which we measure mass are a human invention. Atoms and molecules react on an individual particle-by-particle basis, and we have to count individual particles when doing chemical calculations. 2.

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its chemical formula and can be used to convert from mass to moles of that ... 320 Chapter 10 • The Mole Section 110.10.1 Measuring Matter MAIN Idea Chemists use the mole to count atoms, molecules, ... mole quantities of water, copper, and salt are shown, each with a different

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