

## *Buffer Solution Definition Chemistry*

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### **Buffer Solution Definition Chemistry**

Buffer Definition. A buffer is a solution containing either a weak acid and its salt or a weak base and its salt, which is resistant to changes in pH. In other words, a buffer is an aqueous solution of either a weak acid and its conjugate base or a weak base and its conjugate acid. Buffers are used to maintain a stable pH in a solution,...

### **Buffer Definition in Chemistry and Biology - ThoughtCo**

Buffer solution. A buffer solution (more precisely, pH buffer or hydrogen ion buffer) is an aqueous solution consisting of a mixture of a weak acid and its conjugate base, or vice versa. Its pH changes very little when a small amount of strong acid or base is added to it. Buffer solutions are used as a means...

### **Buffer solution - Wikipedia**

Definition of Buffers (acid-base) A buffer solution can be made by mixing a weak acid with one of its salts OR mixing a weak base with one of its salts. A more technical way of saying this is that a buffer solution consists of a mixture of a weak acid and its conjugate base OR a weak base and its conjugate acid.

### **Definition of buffers\_acid\_base - Chemistry Dictionary**

For example, a mixture of acetic acid and sodium acetate acts as a buffer solution with a pH of about 4.75. Alkaline buffers, on the other hand, have a pH above 7 and contain a weak base and one of its salts. For example, a mixture of ammonium chloride and ammonium hydroxide acts as a buffer solution with a pH of about 9.25.

### **Buffer Solutions: Definition, Types, Preparation, Examples ...**

Definition of Buffers. A solution which tends to resist changes in pH is called buffer solution. Buffer solutions are the solutions that resist changes in the concentration of hydronium ion and hydroxide ion (and therefore pH) when adding low amounts of acid or base, or when diluting the solution.

### **Buffer Solution : Definition Examples & Applications ...**

buffer solution one that resists appreciable change in its hydrogen ion concentration (pH) when acid or alkali is added to it. colloid solution (colloidal solution) imprecise term for colloid (def. 3). hyperbaric solution one having a greater specific gravity than a standard of reference.

### **Buffers (chemistry) | definition of Buffers (chemistry) by ...**

buffer - (chemistry) an ionic compound that resists changes in its pH. starting buffer - buffer solution at the start of a reaction. PBS, phosphate buffer solution - a solution containing a phosphate buffer.

### **Buffer (chemistry) - definition of Buffer (chemistry) by ...**

Alkaline buffer solutions are commonly made from a weak base and one of its salts. A frequently used example is a mixture of ammonia solution and ammonium chloride solution. If these were mixed in equal molar proportions, the solution would have a pH of 9.25.

### **7. Buffer Solutions - Chemistry LibreTexts**

Definition A buffer solution is one which resists changes in pH when small quantities of an acid or an alkali are added to it. Acidic buffer solutions An acidic buffer solution is simply one which has a pH less than 7. Acidic buffer solutions are commonly made from a weak acid and one of its salts - often a sodium salt.

### **BUFFER SOLUTIONS - chemguide**

Buffer: Buffer, in chemistry, solution usually containing an acid and a base, or a salt, that tends to maintain a constant hydrogen ion concentration. Ions are atoms or molecules that have lost or gained one or more electrons. An example of a common buffer is a solution of acetic acid ( $\text{CH}_3\text{COOH}$ ) and sodium.

**Buffer | chemistry | Britannica.com**

- [Voiceover] Buffer solutions resist changes in pH and so let's think about a solution of a weak acid and its conjugate base. So here we have HA which is our generic weak acid and so the conjugate base would be A<sup>-</sup>. A buffer solution needs to have substantial amounts of both present and that's what ...

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