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According to rate law,  $\text{rate} = k [X]^2$  If  $[X]$  is increased to 3 times, then the new rate is  $\text{rate}' = k [3X]^2$   
 $\text{rate}' = 9 k [X]^2 = 9 \text{ rate}$  Thus, rate of reaction becomes 9 times and hence rate of formation of Y increases 9-times.

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NCERT Solutions for Class 12 Chemistry Chapter 4 Chemical Kinetics. Class XII Chemistry Chapter 4 Chemical Kinetics NCERT Solution is given below. For the reaction  $R \rightarrow P$ , the concentration of a reactant changes from 0.03 M to 0.02 M in 25 minutes. Calculate the average rate of reaction using units of time both in minutes and seconds.

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95 Chemical Kinetics Rate of appearance of P Increase in concentration of P  $P = \text{Time taken } t$  (4.2) Since,  $\Delta[R]$  is a negative quantity (as concentration of reactants is decreasing), it is multiplied with  $-1$  to make the rate of the reaction a positive quantity. Equations (4.1) and (4.2) given above represent the average rate of a reaction,  $r_{av}$ .

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CHAPTER 13: CHEMICAL KINETICS 343 From the first set of data:  $3.20 \times 10^{-1} \text{ M/s} = k(1.50 \text{ M})$   $k = 0.213 \text{ s}^{-1}$  What would be the value of  $k$  if you had used the second or third set of data? Should  $k$  be constant? 13.18 Strategy: We are given a set of concentrations and rate data and asked to determine the order of the reaction and the initial rate for specific concentrations of X and Y.

**CHAPTER 13 CHEMICAL KINETICS - kau**

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