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How many electrons can fit in each of the following energy levels: 1 st energy level = 2 because only 2 can fit in an s sublevel and the first energy level only has an s sublevel. 2 nd energy level = 8; 2 in the 2s sublevel and 6 in the 2p sublevel giving a total of 8. 3 rd energy level = 18; 2 in the 3s sublevel, 6 in the 3p sublevel, and 10 ...

### **34 ChemQuest 11 Name: - Yumpu**

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### **10 ChemQuest 3 Name: - Yumpu**

10. Given the following symbols and your answers to the above questions, which of the following equations is correct. Make sure you have the correct answer before proceeding to the next questions!  $\Delta T$  = temperature change of a substance  $C_p$  = specific heat capacity  $m$  = mass of the substance  $q$  = amount of heat energy applied to the substance

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Ga = 21s 2s22p63s23p64s23d104p1 10. Write the electron configurations for N, P and As and note their similarities. N 2= 1s 2s22p3 P 2= 1s 2s22p63s23p3 As = 1s 2s 2p63s23p64s23d104p3 11. Complete the following statement: Elements ending in p1 are in column number 13 or IIIA of

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