

## *Can A Solution With Undissolved Solute Be Supersaturated*

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### **Can A Solution With Undissolved**

Supersaturated Solutions: A saturated solution contains the maximum mass of dissolved solute in a given volume of solvent at a particular temperature, and it may or may not contain undissolved solute.

### **Can a solution with undissolved solute be supersaturated ...**

So, in order to form SATURATED SOLUTIONs you may obtain undissolved solute. In order to form SUPERSATURATED SOLUTIONs you have to use FREEZE APPARATUS to avoid any EQUILIBRIUM PHENOMENA, e.g. the crystallization-germs : if precipitation occurs, you will obtain not SUPERSATURATED SYSTEMs. I hope this could be clear.

### **Can a solution with undissolved solute be supersaturated ...**

In this question, we discuss whether a solution with undissolved solute can be supersaturated by considering the state of equilibrium in solutions with undissolved solute. When a solution has all of a particular solute that can possibly be dissolved already dissolved into the solution, we call that solution saturated .

### **Solved: Can a solution with undissolved solute be ...**

A solution that is super saturated has more solute than can be dissolved under normal conditions. Such solutions are generally difficult to make because the solute quickly forms crystals.

### **Can a solution with undissolved solute be supersaturated?**

Chapter 15 ChemPhys. STUDY. PLAY. ... Can a solution with undissolved solute be supersaturated? Explain. No, if there were undissolved solute, the excess solute would come out of a supersaturated solution. What is the effect of pressure on the solubility of gases in liquids?

### **Chapter 15 ChemPhys Flashcards | Quizlet**

Undissolved definition, to make a solution of, as by mixing with a liquid; pass into solution: to dissolve salt in water. See more.

### **Undissolved | Definition of Undissolved at Dictionary.com**

A solution is saturated when no more solute can be dissolved at the current temperature. A saturated solution is one in which the dissolved and undissolved solutes are in equilibrium.

### **What is a Solution? - Edinformatics**

We can prepare a homogeneous saturated solution by adding excess solute (in this case, greater than 35.9 g of NaCl) to the solvent (water), stirring until the maximum possible amount of solute has dissolved, and then removing undissolved solute by filtration.

### **13.2: Saturated Solutions and Solubility - Chemistry ...**

For a saturated aqueous solution in equilibrium with a molecular (covalent) solute which is a solid, molecules of solute dissolved in water are in equilibrium with undissolved molecules of solute and we can write:

### **Chemistry Tutorial Solubility and Le Chateliers Principle**

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## **Can A Solution With Undissolved Solute Be**

## Supersaturated

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