# Chapter 9 Stoichiometry Answers Section 2

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#### **Chapter 9 Stoichiometry Answers Section**

CHAPTER 9 REVIEW Stoichiometry SECTION 2 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 4.5 mol The following equation represents a laboratory preparation for oxygen gas:  $2KCIO 3(s) \rightarrow 2KCI(s) 3O 2(g)$  How many moles of O 2 form if 3.0 mol of KCIO 3 are totally consumed?

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Chapter 9 – Stoichiometry Name \_\_\_\_\_ Answers to text questions and worksheets Section 1 Formative Assessment (page 285) 1. the branch of chemistry that deals with the mass relationships of elements in compounds and the mass relationships between reactants and products in a chemical reaction 2. a.

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Chapter 9 Section 9.1: Team Learning Worksheet 1. An individual coefficient does not tell us anything. What is important is the ratio between the reactants and products. For example, suppose we were going to make cookies and a recipe told us to use two eggs, some butter, some flour (etc.) and we would make some cookies. The fact

# Chapter 9 Section 9.1: Team Learning Worksheet

CHEMISTRY NOTES – Chapter 9 Stoichiometry Goals: To gain an understanding of: 1. Stoichiometry. 2. Limiting reagents and percent yield. NOTES: Stoichiometry is the calculation of chemical quantities from balanced equations. The four quantities involved in stoichiometric calculations are:

# **CHEMISTRY NOTES - Chapter 9 Stoichiometry**

Chapter menu Resources Chapter 9 Section 1 Introduction to Stoichiometry Stoichiometry Definition
• Composition stoichiometry deals with the mass relationships of elements in compounds. •
Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction.

#### **Section 1 Introduction to Chapter 9 Stoichiometry**

Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a chemical reaction. Students had an introduction to composition stoichiometry in Chapter 3 and will now move on to some more difficult problems.

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Section 9.2 Chemical Calculations XOBJECTIVES: • Construct mole ratios from balanced chemical equations, and apply these ratios in mole-mole stoichiometric calculations. • Calculate stoichiometric quantities from balanced chemical equations using units of moles, mass, representative particles, and volumes of gases at STP. Mole - Mole Problem

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# **Chapter 9 Section 3 Stoichiometry Answers**

CHAPTER 9 Stoichiometry Stoichiometry comes from the Greek ... SECTION 9-1 O BJECTIVES Define stoichiometry. Describe the importance of the mole ratio in stoichio-metric calculations. Write a mole ratio relating ... 276 CHAPTER 9 mass of amount of amount of mass of given substance ...

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CHAPTER 9 REVIEW Stoichiometry SECTION 9-3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% If the actual yield of a reaction is 22 g and the theoretical yield is 25 g, calculate the percent yield. 2. 6.0 mol of N 2 are mixed with 12.0 mol of H 2 according to the following equation: N 2(g) 3H 2(g)  $\rightarrow$  2NH 3(g) N 2; 2.0 mol a.

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