

Chapter 10 Chemical Quantities Test A Answers

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CHAPTER 10: Chemical Quantities BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

CHAPTER 10: Chemical Quantities

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Chapter 10 (Chemical Quantities) Test Study Guide The mole is the SI unit used to measure the number of representative particles in a substance. A representative particle can be an atom, an ion, or a molecule, depending upon the way a substance commonly exists. Avogadro's number (6.02×10^{23}) is the number of representative particles of a substance present in 1 mole of that substance.

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Chapter 10 - Chemical Quantities. Jennie L. Borders. Section 10.1 - The Mole: A Measurement of Matter. You often measure the amount of something by count, by mass, or by volume. A mole (mol) of a substance is 6.02×10^{23} representative particles of that substance.

Chapter 10 - Chemical Quantities

Chapter Test B A. Matching Match each term in Column B with the correct description in Column A. Write the letter of the correct term on the line. ... CHEMICAL QUANTITIES 10 05_CTR_ch10 7/9/04 3:29 PM Page 256. Chapter 10 Review Package - Part 2. Chapter 10 Chemical Quantities. 257 ____ 12. What is the molar mass of C. 3. H. 8? a.

05 CTR ch10 7/9/04 3:29 PM Page 256 Chapter 10 Review ...

Chapter 10 "Chemical Quantities" Vocab. Tools. Copy this to my account; E-mail to a friend ... Use these activities to learn the vocabulary presented in this chapter. A B; percent composition: a description of the relative amounts of each element in a compound: empirical formula: the lowest whole- number ratio of the atoms of the elements in a ...

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