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Modern Chemistry 171 Nuclearchemistry CHAPTER 21 REVIEW Nuclear Chemistry SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. _____ The nuclear equation is an example of an equation that represents (a) alpha emission. (b) beta emission. (c) positron emission. (d) electron capture. 2.

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Chapter 21 Nuclear Chemistry Review. nucleons. Means nucleus (made up of protons and neutrons) Three main types of radiation. alpha – beta – gamma. Sheet of paper/clothing can stop this kind of radation (least penetrating) Alpha. Type of radation that can be stopped by a piece of wood.

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Chapter 21: Nuclear Chemistry 21.1: The Nature of Nuclear Reactions Nucleons: - the particles that make up a nucleus of an atom (protons,(1 1 p + or 1 1 H) and neutrons, (1 0 n)). Isotopes: - atoms that have different mass number but the same atomic number or number of protons.

Chapter 21 Nuclear Chemistry Notes (answers)
Chapter 21–Assignment C: Summary and Review You may think of nuclear chemistry as an untamed jungle, but there are rules to help you find the trails, just as you found the rules and trails in ordinary chemical reactions. For example, natural radioactivity has only three possible forms, as described below:

Chapter 21
How It Works: Identify the lessons in the Holt McDougal Nuclear Chemistry chapter with which you need help. Find the corresponding video lessons within this companion course chapter.

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Chapter 22 Review: Nuclear Chemistry. History Just because it is interesting. Marie Curie: Henri Becquerel (Physics professor, 1852 ... Nuclear particles involved in nuclear reactions ... effect is mass same, protons up one thus neutrons down one. Example: 42 20 Ca g 0-1 e + 42 21 Sc; Alpha emission (occurs when the nucleus is too large ...

Chapter 22 Review: Nuclear Chemistry
Chapter 21 - Nuclear Chemistry. History and Discovery of Radioactivity. The Discovery of Radioactivity (1896) ... Nuclear Chemistry Nuclear reaction – process that alters the ... Review of Nuclear Structure Every atom of an element has the same number of

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In this lecture I'll teach you about radiotracers in medical diagnostics, and I'll teach you how to perform rudimentary calculations using Einstein's energy-mass equation: E = mc2.

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CHAPTER 21 REVIEW NUCLEAR CHEMISTRY ANSWERS
(c) Products from a nuclear fission of a uranium atom such as 90Sr and 137Cs are highly radioactive and decay by emission of beta particles. (d) Nuclear fusion requires large amounts of energy to get started, whereas nuclear fission can occur spontaneously, although both processes release energy.

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