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AP® Chemistry 2013 Scoring Guidelines . The College Board . ... Answer the following questions involving the stoichiometry and thermodynamics of reactions containing aluminum species. 2  $\text{Al}_2\text{O}_3(\text{l}) + 3 \text{C}(\text{s}) \rightarrow 4 \text{Al}(\text{l}) + 3 \text{CO}_2(\text{g})$  An electrolytic cell produces 235 g of Al(l) according to the equation above.

### **AP Chemistry 2013 Scoring Guidelines - College Board**

2013 HSC Chemistry Marking Guidelines Section I, Part B Question 21 (a) Criteria Marks • Provides a reason for the use of acids as a food preservative related to properties of acids 2 • Provides a reason for the use of acids 1 Sample answer: Many microorganisms that cause spoilage of food cannot survive in low pH conditions. The

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FOR TEACHERS ONLY The University of the State of New York REGENTS HIGH SCHOOL EXAMINATION PHYSICAL SETTING/CHEMISTRY Thursday, January 24, 2013 — 1:15 to 4:15 p.m., only SCORING KEY AND RATING GUIDE

### **FOR TEACHERS ONLY - Regents Examinations**

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Young's modulus for a wire is  $Y$  MgL Constant A L L MgL AY L MgL AY Or, L L A Among the given options, the ratio of  $L/A$  is maximum for  $L = 50$  cm and  $D = 0.5$  mm.

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ANSWER KEY Unit 6 A Conversation 1 Conversation 2 Jack Charlotte Jack Charlotte Subject?  
Business and Marketing [accept either] Engineering Medicine Maths, Chemistry, etc. Year? 2nd 3rd  
4th Final [year] Going well or badly? [Really] well [Ss may also mention it's interesting] Not [very]  
well [Ss may mention the wrong subject]

**ANSWER KEY - Cengage**

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The AP Chemistry Exam covers a full-year introductory college course in chemistry with laboratory. The multiple-choice questions in Section I cover the breadth of the curriculum. Section II, the free-response part of the exam, includes three quantitative problems, one question on writing chemical reactions and predicting products, and two essays.

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©Modeling Instruction – AMTA 2013 1 U1 ws3 v2.0 Name Date Pd Chemistry – Unit 1 Worksheet 3  
Mass, Volume, and Density 1. Study the matter shown in Figure 1. Each dot represents a particle of matter. [Assume the particles are uniformly distributed throughout each object, and particles of the same size have the same mass.]

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