

Ch 13 States Of Matter Answers

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Chapter 13 States of Matter 137 SECTION 13.1 THE NATURE OF GASES (pages 385–389) This section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the particles of a substance. Kinetic Theory and a Model for Gases (pages 385–386) 1.

Name Date Class STATES OF MATTER 13

CH. 13 STATES OF MATTER. The Nature of Gases 13.1. Kinetic Theory •Kinetic theory: all matter consists of tiny particles that are in constant motion. •Particles in a gas are considered to be small, hard spheres with insignificant volume •No attractive or repulsive forces exist between

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many particles colliding, the pressure can be substantial. In Chapter 14, you will learn how temperature, volume, and number of moles affect the pressure that a gas exerts. 388 Chapter 13 States of Matter EXAMPLE PROBLEM 13-1 Finding a Ratio of Diffusion Rates Ammonia has a molar mass of 17.0 g/mol; hydrogen chloride has a molar mass of 36.5 g/mol.

Chapter 13: States of Matter - Jayne Heier

Chapter 13 States of Matter - Chapter 13 "States of... Section 13.1 The Nature of Gases Equal pressures: 1 atm = 760 mm Hg = 101.3 kPa Sample 13.1, page 387 Most modern barometers do not contain mercury- too dangerous These are called aneroid barometers, and contain a sensitive metal diaphragm that responds to the number of collisions...

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Chapter 13 States of Matter notes. So $F/A = F'/A'$. A small force applied over a small area can give a bigger force over a bigger area like in a hydraulic system. The pressure you feel under water is due to the weight of the water over you. For a uniform fluid this depends on the density and the depth. $P = \rho hg$ where ρ is the density of the fluid.

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of matter. Whether your reading light comes from a fluorescent lamp or the sun, the light's source is plasma. Although plasma is still somewhat mysterious to humans, the plasma state of matter is the most common state in the universe. In this chapter, your exploration of the states of matter will go far beyond everyday, casual observations.

Chapter 13: States of Matter - dentonisd.org

Chapter 13- The States of Matter Gases- indefinite volume and shape, low density. Liquids- definite volume, indefinite shape, and high density. Solids- definite volume and shape, high density Solids and liquids have high densities because their molecules are close together.

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13 STUDY GUIDE FOR CONTENT MASTERY CHAPTER States of Matter Section 13.1 Gases In your textbook, read about the kinetic-molecular theory. Complete each statement. 1. The kinetic molecular theory describes the behavior of gases in terms of particles in 2. The kinetic-molecular theory makes the following assumptions. a.

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graph representing the relationships among the solid, liquid, and vapor states of a substance in a sealed container standard atmosphere pressure required to support 760 mm of mercury at 25°C

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Study Guide - Chapter 12 - States of Matter Section 12.1 Gases 1. motion 2. a. small b. forces c. random d. elastic; kinetic 3. $KE = \frac{1}{2}mv^2$ 4.

Study Guide - Chapter 12 - States of Matter

Match the intermolecular forces with their descriptions. 1. Weak forces between nonpolar molecules. 2. A type of one of the forces that is between hydrogen and a negatively charged particle.

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Chemistry Chapter 13: States Of Matter; Erin C. • 67 cards. Kinetic - Molecular Theory. All matter consists of tiny particles that are in constant motion. Kinetic - Molecular Theory. Used to explain the properties of matter in terms of the energy and movement of the particles and the forces acting between them. ...

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