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A pH indicator is a halochromic chemical compound added in small amounts to a solution so the pH (acidity or basicity) of the solution can be determined visually. Hence, a pH indicator is a chemical detector for hydronium ions (H_3O^+) or hydrogen ions (H^+) in the Arrhenius model. Normally, the indicator causes the color of the solution to change depending on the pH.

pH indicator - Wikipedia

HgS is dimorphic with two crystal forms: red cinnabar ($\alpha\text{-HgS}$, trigonal, hP6, P3221), is the form in which mercury is most commonly found in nature.; black, metacinnabar ($\beta\text{-HgS}$), is less common in nature and adopts the zinc blende (T 2 d-F 4 3m) crystal structure. Crystals of red, $\alpha\text{-HgS}$, are optically active. This is caused by the Hg-S helices in the structure.

Mercury sulfide - Wikipedia

During exercise, the muscles use up oxygen as they convert the chemical energy in glucose to mechanical energy. This O_2 comes from hemoglobin in the blood. CO_2 and H^+ are produced during the breakdown of glucose, and they are removed from the muscles via the blood. The production and removal of CO_2 and H^+ , together with the use and transport of O_2 , cause chemical changes in the blood.

pH Buffers in the Blood

Chemical Changes in the Blood During Exercise. In previous tutorials ("Hemoglobin and the Heme Group: Metal Complexes in the Blood for Oxygen Transport", "Iron Use and Storage in the Body: Ferritin and Molecular Representations", "Maintaining the Body's Chemistry: Dialysis in the Kidneys") you learned about the daily maintenance required in the blood for normal everyday activities such as ...

pH Buffers in the Blood - chemistry.wustl.edu

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