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Post-Lab Questions 1. Use Table 1 in the Background section to record the approximate wavelength of light emitted for each metal in Data Table 2. 2. Convert each of the wavelengths in the Data Table from nanometers to meters. Record the wavelengths in meters in the ... CF#5607 Flame Test Kit SLK ...

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3.) The specific heat of a solution is 4.18]/(g*C °) and its density is 1.02g/mL. The solution is formed

by combing 25.0mL of solution A with 25.0mL of solution B, with each solution initially at 21.4 $^{\circ}$ C. The final temperate of the combined solutions is 25.3 $^{\circ}$ C. Calculate the heat of reaction, q rxn, assuming no heat loss to the calorimeter.

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