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\ Chapter 11 Chemistry Study Guide. Chapter 11 Chemistry Study Guide. Chemical equation. An expression representing a chemical reaction. Catalyst. A substances that increases the rate of reaction by lowering the activation energy barrier; not used in the reaction. Coefficient.

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Chemistry 11 Final Exam Study Guide Chemistry 11 - Final Exam Study Guide Page 15 When Electronegativities of bonding atoms are the same (as they are in diatomic molecules) or close to the same, they SHARE electrons. Bonds formed when atoms share electrons are called Covalent Bonds .

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Stoichiometry is the tool for answering these questions. Stoichiometry The study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass. Recall from Chapter 3 that the law states that

Chapter 11: Stoichiometry

CHAPTER Section 11.1 continued In your textbook, read about mole ratios. Answer the questions about the following chemical reaction. sodium + iron(III) oxide \rightarrow sodium oxide + iron $6\text{Na(s)} + \text{Fe}_2\text{O}_3\text{(s)} \rightarrow 3\text{Na}_2\text{O(s)} + 2\text{Fe(s)}$ 15. What is a mole ratio? ... Study Guide Chemistry: Matter and Change Chapter 11 ...

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AP Chemistry – Solutions, Chapter 11/4b FAQ. Q: What are hydrophobic and hydrophilic? A: Hydrophobic is a “fear of water”. Nonpolar substances are typically hydrophobic. They are not attracted to water. Hydrophilic on the other hand are “water lovers”. These compounds are attracted to water, and therefore are commonly ionic or polar.

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