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pH and pOH The pH of a solution indicates how acidic or basic that solution is. pH range of 0-7 acidic 7 neutral 7-14 basic Since $[H+]\cdot[OH-]=10-14$ at 25°C, if [H+] is known, the [OH-] can be calculated and vice versa. pH = $-\log[H+]$ So if [H]=10-6 M, pH = 6 pOH = $-\log[OH-]$ So if [OH]=10-8, pOH = 8 Together, pH + pOH = 14.

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pH and pOH CONTINUED 1. 0.01 M HCl [H+] = 0.01M pH = $-\log(0.01)$ = 2 2. 0.0010 M NaOH [OH-] = 0.0010M pOH = $-\log(0.010)$ = 3 \rightarrow pH = 11 3. 0.050 M NaOH

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Part 2: For each of the problems below, assume 100% dissociation. 1. A. Write the equation for the dissociation of hydrochloric acid. B. Find the pH of a 0.00476 M hydrochloric acid solution. 2. A.

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Calculating pH and pOH worksheet W 335 Everett Community College Tutoring Center Student Support Services Program 1) What is the pH of a 0.0235 M HCl solution? 2) What is the pOH of a 0.0235 M HCl solution? 3) What is the pH of a $6.50 \times 10-3$ M KOH solution? (Hint: this is a basic solution – concentration is of OH-)

Calculating pH and pOH worksheet

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