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### **How Sublime Exploring and Measuring the Triple Point of ...**

Sublimation is the transition from the solid phase to the gas phase without passing through an intermediate liquid phase. This endothermic phase transition occurs at temperatures and pressures below the triple point. The term only applies to physical changes of state and not to the transformation of a solid into a gas during a chemical reaction.

### **Sublimation Definition (Phase Transition in Chemistry)**

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9. (Chemistry) to change or cause to change directly from a solid to a vapour or gas without first melting: to sublime iodine; many mercury salts sublime when heated. 10. (Chemistry) to undergo or cause to undergo this process followed by a reverse change directly from a vapour to a solid: to sublime iodine onto glass.

### **Sublime - definition of sublime by The Free Dictionary**

'The sublime transformed itself into feminine agency, the ability to occupy a space in an active way so as not to be utterly overwhelmed by the sublime effect of nature.' 'While Turner was a complex and eclectic artist, much of his work is suffused with a Romantic sense of nature's sublime power and wonder.'

### **sublime | Definition of sublime in English by Oxford ...**

Sublimation (phase transition) The pressure referred to is the partial pressure of the substance, not the total (e.g. atmospheric) pressure of the entire system. So, all solids that possess an appreciable vapour pressure at a certain temperature usually can sublime in air (e.g. water ice just below 0 °C).

### **Sublimation (phase transition) - Wikipedia**

Sublime vs. Sublimate. At first glance, the question of whether sublime and sublimate are related might seem like an easy one to answer, as they appear to come from the same source. However, the most common senses in which each of these words is used today are dissimilar enough to give pause.

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