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CHAPTER 9 REVIEW Stoichiometry SECTION 2 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 4.5 mol The following equation represents a laboratory preparation for oxygen gas: 2KClO 3(s) → 2KCl(s) 3O 2(g) How many moles of O 2 form if 3.0 mol of KClO 3 are totally consumed?

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Chapter 9 – Stoichiometry Review #1 – #18, #31, & #38 Answers . 38. To ensure that all magnesium is converted to MgO, I would use pure oxygen, not air, to carry out the reaction, because Mg could react with NI in air to form Mg]N2. The pure oxygen should

1 – 18, 31, & 38 Answers
CHAPTER 9 REVIEW Stoichiometry SECTION 9-3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% If the actual yield of a reaction is 22 g and the theoretical yield is 25 g, calculate the percent yield. 2. 6.0 mol of N 2 are mixed with 12.0 mol of H 2 according to the following equation: N 2(g) 3H 2(g) → 2NH 3(g) N 2; 2.0 mol a.

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Chapter 9: Standard Review Worksheet 1. Answers will vary. An example is included below: 2H 2 O 2 (aq) 2H2 O(l) + O2 (g)This describes the decomposition reaction of hydrogen peroxide.

Chapter 9: Standard Review Worksheet
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