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Chapter 18 Review Chemical Equilibrium

Modern Chemistry 149 Chemical Equilibrium CHAPTER 18 REVIEW Chemical Equilibrium SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. Match the solution type on the right to the corresponding relationship between the ion product and the K_{sp} for that solution, listed on the left. ____ The ion product exceeds the K_{sp} .

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b. chemical equilibrium. d. both a and b. 8. According to collision theory, in order for a chemical reaction to occur, the reactant atoms must: a. make contact with each other. b. have a minimum level of kinetic energy. ... Exam Review Chapter 18-Equilibrium ...

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sometimes referred to as the chemical-equilibrium expression. The H_2 , I_2 , HI Equilibrium System Consider the reaction between H_2 and I_2 vapor in a sealed flask at an elevated temperature. The rate of reaction can be followed by observing the rate at which the violet color of the iodine vapor diminishes, as $[C]^x[D]^y[A]^n[B]^m$ 556 CHAPTER 18

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CHAPTER 18 REVIEW Chemical Equilibrium MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Consider the following equilibrium equation: $N_2(g) + 2O_2(g) \rightleftharpoons 2NO_2(g)$; $\Delta H = 33 \text{ kJ/mol}$ At equilibrium, which reaction is favored when

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560 Chapter 18 Chemical Equilibrium H 2 Time Concentration NH_3 N_2 Figure 18-1 The concentrations of the reac-tants (H_2 and N_2) decrease at first while the concentration of the product (NH_3) increases. Then, before the reactants are used up, all concentrations become constant.

Chapter 18: Chemical Equilibrium - Jayne Heier

1 Chapter 18...Reaction Rates and Equilibrium RATES OF REACTION Speeds of chemical reactions can be extremely fast or extremely slow. The rate is a measure of the speed of any change that occurs within an interval of time (which could range from fractions of a second

Chapter 18 Reaction Rates and Equilibrium

CHAPTER 17 EQUILIBRIUM: THE EXTENT OF CHEMICAL REACTIONS ... The equilibrium constant for the reverse reaction is the inverse of the original constant. When a ... 17.18 The concentration of solids and pure liquids do not change, so their concentration terms are not written in the

CHAPTER 17 EQUILIBRIUM: THE EXTENT OF CHEMICAL REACTIONS

CHAPTER 8 REVIEW. Chemical Equations and Reactions. ... CHAPTER 18 REVIEW. Chemical Equilibrium. In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. 16.A chemical reaction that is at equilibrium always has.

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