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Chapter 16 Solutions 167 SECTION 16.1 PROPERTIES OF SOLUTIONS (pages 471–477) This section identifies the factors that affect the solubility of a substance and determine the rate at which a solute dissolves. Solution Formation (pages 471–472) Look at Figure 16.1 on page 471 to help you answer Questions 1 and 2. 1.

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Chapter 16: Colligative Properties of Solutions 45 16-4. The mole fraction of $(NH_4)_2SO_4(aq)$ is given by $x(NH_4)_2SO_4 = \frac{n(NH_4)_2SO_4}{n(NH_4)_2SO_4 + n H_2O}$ Because $(NH_4)_2SO_4(aq)$ is a strong electrolyte, it dissociates completely into $NH_4^+(aq)$ and $SO_4^{2-}(aq)$ ions. Assume a one kilogram solution. The number of moles of ions in one ...

CHAPTER 16. Colligative Properties of Solutions
Chapter 16 Solutions 16.1 Properties of Solutions 16.2 Concentrations of Solutions 16.3 Colligative Properties of Solutions 16.4 Calculations Involving ... CHEMISTRY & YOU How do you think crystal-growing kits work? Use what you know about solubility and supersaturated solutions

Chapter 16
Check your answers with those at the end of the chapter. Workbook If your instructor recommends the Active Learning Workbook, do Questions, Exercises, and Problems 1–16. Chapter 16–Assignment B: Solution Concentrations Solutions are mixtures, not pure substances, so the composition of a solution is not fixed by a molecular formula.

Chapter 16
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Ch. 16 Textbook Answers 1. Henry's Law states that the solubility of a gas in a liquid is directly proportional to the pressure above the liquid. S 1 = S 2 0.16 g/L = 0.44 g/L p 1 p 2 104 kPa 288 kPa 2. Henry's Law states that the solubility of a gas in a liquid is directly proportional to the pressure above the liquid. S 1 = S 2 3.6 g/L = 9.5 g/L p 1 p 2 1 atm 2.6 atm 3. The factors that ...

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SAMPLE EXERCISE 16.4 Calculating [H+] for Pure Water Calculate the values of [H+] and [OH–] in a neutral solution at 25 °C. In an acid solution [H +] is greater than 1.0 ×10 –7 M ; in a basic solution [H +] is less than 1.0 ×0–7 M. PRACTICE EXERCISE Indicate whether solutions with each of the following ion concentrations are neutral,

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