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HEAT AND ITS MEASUREMENT Name _____ ~at (or energy) can be measured in units of calories or joules. When there is a ,,,,nperature change (ΔT), heat (Q) can be calculated using this formula: $Q = \text{mass} \times L \times \text{specific heat capacity}$ ($L \Delta T = \text{Final Temp} - \text{Initial Temp}$) During a phase change, we use this formula:

HEAT AND ITS MEASUREMENT - Mr. Kay - Chemistry

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specific heat of water and its mass. $C_{\text{water}} = C_{s,\text{water}} \cdot m_{\text{water}}$ which leads to $q_{\text{cal}} = m_{\text{water}} C_{s,\text{water}} \Delta T + C_{\text{hardware}} \Delta T$ The specific heat for water (an intensive property) times the mass of the water (extensive property) equals the heat capacity of the water (an extensive property). Always check and see what you are given in a problem.

Dr. McCord Calorimetry - Chemistry 301

Physical vs. Chemical Changes In a physical change, the original substance still exists; it has only changed in form, so some of its physical and chemical properties are still the same.

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