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Chapter 14 The Behavior of Gases147 SECTION 14.1 PROPERTIES OF GASES(pages 413–417) This section uses kinetic theory to explain the properties of gases. This section also explains how gas pressure is affected by the amount of gas, its volume, and its temperature.

SECTION 14.1 PROPERTIES OF GASES(pages 413–417)
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418 Chapter 14 Gases CHAPTER 14 What You'll Learn You will use gas laws to cal-culate how pressure, tem-perature, volume, and number of moles of a gas will change when one or more of these variables is altered. You will compare properties of real and ideal gases. You will apply the gas laws and Avogadro's principle to chemical equations ...

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Chapter 14 The Ideal Gas Law and Its Applications. Gases Revisited Properties of Gases (Section 4.1) Gases may be compressed. Gases expand to fill their containers uniformly. All gases have low densities compared with ... Chapter14 4th ed Cracolice super custom white.key Author:

Chapter 14 The Ideal Gas Law and Its Applications
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Chapter 14
Gas flows from a region of lower pressure to a region of higher pressure. Adding air to an object will cause the object to inflate. Four variables are used to describe a gas, P, V, T, and n, where n = number of moles. Chapter 14 The Behavior of Gases 347

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Chapter 13 Gases 1. Solids and liquids have essentially fixed volumes and are not able to be compressed easily. Gases have volumes that depend on their conditions, and can be compressed or expanded by changes in those conditions. Although the particles of matter in solids are

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Chapter 14–Assignment E: Volume–Volume Gas Stoichiometry Chapter 14 concludes with a section on converting between volumes of gases reacting and produced in a chemical reaction. The main idea in this section is: 1) The ratio of volumes of gases in a reaction is the same as the ratio of moles,

Chapter 14
Relationships among P, V, T and the Number of Gas Particles Answer the following questions in complete sentences. Chapter 13 - Gases - An Introduction to Chemistry

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CHAPTER 14 THE BEHAVIOR OF GASES ANSWER KEY
Chapter 13: Standard Review Worksheet 1. ... constant pressure for a fixed amount of gas): V = (constant) _ T Answer Key. Avogadro's law states that the volume of a gas sample is proportional to the number of ... 14. The total pressure in a mixture of gases is the sum of the individual partial pressures of the

Chapter 13: Standard Review Worksheet
Chapter 14 SCREEN-CAST. Please watch the videos and ... PRACTICE QUESTIONS WITH AN ANSWER KEY AT THE END ALSO AP CHEMISTRY REVIEW PROBLEMS FOR EXAM ON KINETICS ... Find the reaction rate of the production of oxygen gas between 2 and 6 seconds. 1c.) Identify the reactant and product based on the graph. 2.) How does temperature affect reaction rate?

Chapter 14 - CHEMISTRY - Google Sites
Chapter 14 The Behavior of Gases 347 ... Answer the following in the space provided. 20. ... Key Equation • Ideal gas law: P V n R T or PV nRT Part ACompletion Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short

05 CTR ch14 7/12/04 8:13 AM Page 347 THE PROPERTIES OF ...
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414 Chapter 14 For: Links on Gases Visit: www.SciLinks.org Web Code: cdn-1141 Kinetic theory can explain why gases are compressed more easily than liquids or solids. Gases are easily compressed because of the space between the particles in a gas. Remember that the volume of the particles in a gas is small compared to the overall volume of the gas.

14.1 Properties of Gases 14 - Grandview Independent School ...
THE GAS LAWS 14.2 Section Review Objectives Describe the relationship among the temperature, volume, and pressure of a gas Use the combined gas law to solve problems Key Part B True-False Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT. ... Chapter J 4 The Behavior of Gases 349 .

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