Chapter 9 Review Stoichiometry Section 3 Answer Key

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CHAPTER 9 REVIEW Stoichiometry SECTION 2 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 4.5 mol The following equation represents a laboratory preparation for oxygen gas: $2KClO_3(s) \rightarrow 2KCl(s)$ $3O_2(g)$ How many moles of O_2 form if 3.0 mol of $KClO_3$ are totally consumed? 2. 200 g Given the ...

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Stoichiometry. SECTION 1. SHORT ANSWER Answer the following questions in the space provided. 1. _____ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. ... CHAPTER 9 REVIEW ...

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