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2) 2.50 moles of O<sub>2</sub> ANSWER: 56.0 L 3) 0.850 moles of N<sub>2</sub> ANSWER: 19.0 L 4) 3.75 moles of CO<sub>2</sub> ANSWER: 84.0 L 5) 0.70 mole of NH<sub>3</sub> ANSWER: 16 L 6) 12.0 g of H<sub>2</sub> ANSWER: 133 L 7) 125. g of O<sub>2</sub> ANSWER: 87.5 L 8) 56.0g of N<sub>2</sub> ANSWER: 44.8 L 9) 70. g of CO<sub>2</sub> ANSWER: 36. L 10) 20. g of NH<sub>3</sub> ANSWER: 26 L Chemistry IF8766 Page 51 Instructional Fair, Inc.

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### The Mole And Volume Worksheet Instructional Fair Inc

Chemistry IF8766 Olnstructional Fair, Inc. STOICHIOMETRY: MOLE-MOLE PROBLEMS I.  $N_2 + 3H_2$  Name How many moles of hydrogen are needed to completely react with two moles of nitrogen?  $2.0 + 3O_2$  How many moles of oxygen are produced by the decomposition of six moles of potassium chlorate?

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MIXED MOLE PROBLEMS - KEY 1. a) How many grams are there in  $1.5 \times 10^{25}$  molecules of CO? 1.110g 1 mol 44.0 g 6.0210molecules 1.510molecules CO 1mol 3 23 2 25! =!!!! b) What volume would the CO

### KEY - CP - Mixed Mole Problems

2 moles 3 moles  $X = 90$  g  $O_2$  1 mole  $O_2$  2 mole  $KClO_3$  122.5 g  $KClO_3 = 229.7$  g  $KClO_3$  32 g  $O_2$  3 moles  $O_2$  1 mole  $KClO_3$  I II III IV CHEMISTRY STOICHIOMETRY Mass (g,kg, etc.) Volume (L, mL, etc.) or # of Items (atoms,molec.) of Given reactant or product Moles of Given Mole Ratio from the Balanced Chemical Equation

### CHEMISTRY COMPUTING FORMULA MASS WORKSHEET - ISD 622

ANSWER: 0.99 mole 5) 53 g of  $CuSO_4 \cdot 5H_2O$  ANSWER: 0.21 mole Directions: determine the number of grams in each of the quantities below. 6) 2.8 moles of NaCl ANSWER: 160 g 7) 0.75 moles  $H_2SO_4$  ANSWER: 74 g 8) 2.75 moles of  $KMnO_4$  ANSWER: 435. g 9) 0.55 moles of KCl ANSWER: 41. g 10) 2.3 moles of  $CuSO_4 \cdot 5H_2O$  ANSWER: 570 g Chemistry IF8766 Page 50 Instructional Fair, Inc.

### Moles and Mass - Ms. Agostine's Chemistry Page

Directions: How many moles are in the number of molecules below? 6)  $6.02 \times 10^{23}$  molecules 7)  $1.204 \times 10^{24}$  molecules 8)  $1.5 \times 10^{20}$  molecules 9)  $3.4 \times 10^{26}$  molecules 10)  $7.5 \times 10^{19}$  molecules Chemistry IF8766 Page 52 Instructional Fair, Inc. Title: Microsoft Word - pg 52 - The Mole and Avagadro's Number.doc Author ...

### The Mole and Avagadro's Number - crazichemteacher.webs.com

Stoichiometry: Mixed Problems (KEY) 1)  $N_2 + 3H_2 \rightarrow 2NH_3$  What volume of  $NH_3$  at STP is produced if 25.0 of  $N_2$  is reacted with an excess of  $H_2$ ? 3 3 3 2 3 2 2 40.0L  $NH_3$  1mol  $NH_3$  22.4L  $NH_3$  1mol  $N_2$  2mol  $NH_3$  28.0g  $N_2$  25.0g  $N_2$  1mol  $N_2 \times \times \times =$  2)  $2KClO_3 \rightarrow 2KCl + 3O_2$  If 5.0g of  $KClO_3$  is decomposed, what volume of  $O_2$  is produced at STP? 2

### Stoichiometry: Mixed Problems (KEY)

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