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Chem Quest 32 Answer Key

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(Base the answer to this question on the number of moles of propane that actually get

combusted—which is your answer to part a.) 12 moles. For every mole of propane that combusts 3 moles of CO₂ are produced, so the number of moles of CO₂ that can be produced when 4 moles of propane combusts = $4(3) = 12$.

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Justify your answer by citing experimental data from the data table. If the temperature or pressure is not constant are your statements in 9b and 9c correct? Justify your answer. ... ChemQuest 35 Information: Gas Pressure Critical Thinking Questions Information: Gas Laws Critical Thinking Questions ...

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14. Based on your answer to question 13, verify that this equation is true: This equation is called the "Combined Gas Law". Notice that it contains all of the equations combined into one! Again, the pressure times the volume divided by the Kelvin temperature equals 1.18 for every trial in the table. 15.

Information: Gas Pressure

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