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Worksheet - Entropy Entropy,  $S$ , is a measure of how dispersed or spread out a system's energy is among the available energy levels. The number of energy levels and how many are populated depend on a number of factors. As temperature is increased more of the excited state energy levels will be populated

**Worksheet - Entropy  $S$  dispersed energy levels Increasing ...**

The entropy change is negative - a decrease in entropy. Entropy is associated with the amount of disorder or number of degrees of freedom. In the liquid state the water molecules are able to move around randomly and so are very disordered (although not as disordered as in the gaseous state, when it is steam). When we freeze water to turn it into a solid, we remove the degrees of freedom ...

**Does the following reaction show an increase or decrease ...**

$\Delta S$  = change in entropy or randomness and Complete the table for the sign of  $\Delta G$ ; +, or undetermined, When conditions allow for an undetermined sign of  $\Delta G$ , temperature will decide spontaneity. Answer the questions below. The conditions in which  $\Delta G$  is always negative is when  $\Delta H$  is and  $\Delta S$  is  $\Delta H$   $\Delta S$  9. 3. ... Chemistry IF8766 ...

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