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Chapter 1 Solutions. Engineering and Chemical Thermodynamics. Wyatt Tenhaeff Milo Koretsky. Department of Chemical Engineering Oregon State University. koretsm@engr.orst.edu 1.2 An approximate solution can be found if we combine Equations 1.4 and 1.5: $1.2 \text{ mV} \approx \frac{2.3 \text{ kT}}{e}$ $\approx \frac{2.3 \times 1.38 \times 10^{-23} \text{ J/K} \times 22 \text{ K}}{1.6 \times 10^{-19} \text{ C}}$ Assume the temperature is 22 C.

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The students examine heat-related phenomena of chemical thermodynamics using hot and cold packets, illustrating how thermodynamics work in the real world. ... For engineering, they should ask questions to define the problem to be solved and to elicit ideas that lead to the constraints and specifications for its solution. (NRC Framework 2012, p.

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