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•covalent bond: a bond formed by the sharing of electrons between atoms •molecule: a neutral group of atoms joined by covalent bonds •diatomic molecule: a molecule consisting of two atoms  
Glossary Terms

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Chapter 8 Covalent Bonding and Molecular Structure 8-4 H<sub>2</sub> molecule. More sophisticated descriptions of chemical bonding will be discussed in Chapter 9. 8.3 Lewis Structures OWL Opening Exploration 8.X One of the most important tools chemists use to predict the properties of a chemical species is its Lewis structure.

### Chapter 8: Covalent Bonding and Molecular Structure

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single covalent bond. two atoms held together by sharing one pair of electrons. structural formula. represents the covalent bonds as dashes and shows the arrangement of covalently bonded atoms. unshared pair. a pair of valence electrons that is not shared between atoms, also known as a lone pair or non-bonding pair.

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Chemical Bonding SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. Use the concept of potential energy to describe how a covalent bond forms between two atoms. As the atoms involved in the formation of a covalent bond approach each other, the electron-proton attraction is stronger than the electron-electron and ...

**6 Chemical Bonding - effinghamsschools.com**

In this lesson students delve into covalent bonding through taking notes and practicing with whiteboards and individual practice. Students have already learned the basics of covalent bonds through the introductory lessons of Bonding Inquiry and Ionic, Covalent, and Metallic Bonds.. This lesson aligns with NGSS Performance Expectation: HS-PS1-2: Construct and revise an explanation for the ...

**Covalent Answer Key - BetterLesson**

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**Covalent Bonding Answer Key - Printable Worksheets**

In covalent bonds, electron sharing usually occurs so that atoms attain the electron configurations of noble gases. • For example, a single hydrogen atom has one electron.

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