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In a chemical reaction, one or more reactants change into one or more products. Figure 11.1a shows the ingredients for making leavened bread— flour, salt, yeast, and water. Chemical reactions

take place when the ingredients are mixed together and heated in the oven. Figure 11.1b shows the product—a loaf of bread.

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This unit focuses on chemical properties and chemical reactions. Students analyze evidence and property changes that allow them to distinguish between chemical and physical reactions. In addition, students investigate the Law of Conservation of Mass as they look at how bonds are broken and formed in chemical reactions.

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1. Chemical equations and reactions<br />Chapter 8.1<br /> 2. Objectives:<br />List three observations that suggest that a chemical reaction has taken place.<br />List three requirements for a correctly written chemical equation.<br />Write a word equation and a formula equation for a given chemical reaction.<br />Balance a formula equation by inspection.<br /> 3.

### **Chapter 8.1 : Describing Chemical Reactions - SlideShare**

The production of a new substance, a gas, is evidence of a chemical change. Two new substances are produced in this reaction, oxygen gas and liquid water. You could describe this reaction by saying, "Hydrogen peroxide decomposes to form water and oxygen gas." You could also write a word equation.

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