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Chapter 8 Section 8.1 Describing Chemical Reactions. 5) Substances written on top of the arrow which indicates something supplied to the reaction; it does not take part in the reaction. Ex.) Triangle on top of arrow indicates heat being given off; temperature on top of an arrow indicates the temperature at which the reaction occurs; a substance on top of an arrow indicate the "catalyst".

Chapter 8 Section 8.1 Describing Chemical Reactions ...

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Chapter 8 Chemical Reactions. Adapted from notes by Stephen Cotton. 2. Section 8.1 Describing Chemical Change. OBJECTIVES: -Write equations describing chemical reactions, using appropriate symbols. -Write balanced chemical equations, when given the names or formulas of the reactants and products in a chemical reaction.

Section 8.1 Describing Chemical Change - Keweenaw

Section 8-1 Describing Chemical Reactions 1. 8-1 Describing Chemical Reactions. Chemical Equation - identifies and relates the amounts of reactants and products in a chemical reaction. Some Indicators of a Chemical Reaction. 1) Heat and light have been released 2) A gas is produced (you see bubbles)

Section 8-1 Describing Chemical Reactions - Kearsley

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Chapter 8.1 : Describing Chemical Reactions 1. Chemical equations and reactions
Chapter 8.1
 2.

Chapter 8.1 : Describing Chemical Reactions - SlideShare

Iron metal reacts with aqueous silver nitrate to produce aqueous iron(III) nitrate and silver metal. 11. Solid potassium nitrate yields solid potassium nitrite and oxygen gas. 12. Calcium metal reacts with chlorine gas to produce solid calcium chloride.

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Full Name Comment goes here. Chapter review answer key. 1. Chapter Review Answer Key (Unit 5)
1. Energy in the form of sound, light or heat (absorbed or released), production of a gas, formation of a precipitate, and a color change 3. A coefficient is a number that appears in front of a formula in a chemical equation.

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Chemical Reactions 193. To describe the burning of charcoal, you can substitute the reactants and products of the reaction into the word equation as follows. Carbon Oxygen → Carbon dioxide You can then simplify the word equation by writing the reactants and products as chemical formulas.

Section 7.1 7.1 Describing Reactions - PC\|MAC

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