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Electron Energy and Light 3 11. Consider the hydrogen spectrum in Model 2. a. Which color of light has the most energy? b. Which color of light has the least energy? 12. Does a gas discharge tube filled with boron emit the same wavelengths of light as a tube filled with hydrogen? Use evidence from Model 2 to support your answer. 13.

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Electron Energy And Light. STUDY. PLAY. violet. high energy, high frequency, short wavelengths. red. low energy, low frequency, long wave lengths. ... for an electron to move from an energy level far from the nucleus to an energy level close to the nucleus it would need to _____ energy. lose. electron moves up.

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LIGHT QUESTIONS and ANSWERS 1. What does it mean to say an atom is excited? An electron or electrons have more energy than their ground state, which is where they would be if they had not gained energy in some way. 2. How does the energy of the photon that is emitted by an electron jumping between two

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Emission of light from an atom occurs when an electron _____. ... a. a quantum mechanical model c. the Schrodinger equation b. direction with respect to the nucleus d. the electron energy sublevel Short Answer. 63. Give the electron configuration for a neutral atom of beryllium. ... Chapter 5 Practice Test ...

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