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Answer Multiplying volume (L) by molar concentration (mol/L) gives the number of moles in that volume of solution. If x is the concentration of gastric HCl (mol/L), $(0.010 \text{ L}) \times (0.0072 \text{ L})(0.1 \text{ mol/L}) \times 0.072 \text{ M}$ gastric HCl Given that $\text{pH} = -\log [\text{H}^+]$ and that HCl is a strong acid, $\text{pH} = -\log (7.2 \times 10^{-2}) = 1.15$.

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Lehninger Biochemistry Chapter 2/Problem #1 /#9

Chapter 2: Biochemistry Problems 7) Draw several molecules on the screen and calculate their formulas by hand. Check your work by clicking the "Calculate" button. Note that it is possible to draw molecules that the software cannot process properly. Some of these molecules are chemically possible, but their chemistry is beyond the scope

Chapter 2: Biochemistry Problems

6. A hydronium ion: A) has the structure H_3O^+ . B) is a hydrated hydrogen ion. C) is a hydrated proton. D) is the usual form of one of the dissociation products of water in solution.

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