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M a s t e r i n g C h e m i s t r y C h a p t e r 1 9 P a g e | 1 Chapter 19 Mastering Chemistry Homework Spontaneity and Reversibility: Part A For a single substance at atmospheric pressure, classify the following as describing a spontaneous process, a nonspontaneous process, or an equilibrium system. Part B Determine whether each of these processes is reversible or irreversible.

Chem 152 Chapter 19 Mastering Chemistry Homework ...

Finally, calculate the volume of the suspension required from the required mass that we calculated: $(72.64 \text{ milligrams}) / (100 \text{ mg} / 5.0 \text{ mL}) = 3.632 \text{ mL}$ should be the answer. Note: 100mg/5.0 mL is the density of the suspension, so it makes sense to divide the mass required by the density to get the volume required.

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