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Calcium oxide reacts with water in a combination reaction to produce calcium hydroxide: $\text{CaO(s)} + \text{H}_2\text{O(l)} \rightarrow \text{Ca(OH)}_2\text{(s)}$ In a particular experiment, a 3.00-g sample of CaO is reacted with excess water and 3.46 g of Ca(OH)₂ is recovered. What is the percent yield in this experiment?

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