

Chapter 11 Review Gases Answers

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Chapter 11 Review Gases Answers

Review Previous Concepts Gases CHAPTER 11 1. What is a mole? What information does the molar mass of a substance relate about the substance? 2. What is the kinetic-molecular theory of matter? Key Concepts LOUISIANA STANDARDS LA.PS.29 Predict the properties of a gas based on gas laws. (e.g., temperature, pressure, volume). (PS-H-C7) 340 CHAPTER 11

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CHAPTER 11 REVIEW Gases SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. State whether the pressure of a fixed mass of gas will increase, decrease, or stay the same in the following circumstances:

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Chapter 11 181 Chapter 11 - Gases Review Skills 11.1 Gases and Their Properties Ideal Gases ... Work all of the selected problems at the end of the chapter, and check your answers with the solutions provided in this chapter of the study guide. Ask for help if you need it.

Chapter 11 - Gases - An Introduction to Chemistry

Modern Chemistry 95 Gases CHAPTER 11 REVIEW Gases SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. State whether the pressure of a fixed mass of gas will increase, decrease, or stay the same in the following circumstances: ____ a. temperature increases, volume stays the same

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CHAPTER . 11 . REVIEW . Gases . SHORT ANSWER . Answer the following questions in the space provided. 1. Consider the following data table: Approximate pressure (kPa) Altitude above sea level (km) 100 . o (sea level) 50 5.5 (peak of Mt. Kilimanjaro) 25 11 Get cruising altitude} < 0.1 22 (ozone layer) ~.

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CHAPTER 11 REVIEW Gases SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. c The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L. (b) divided by the mass of 1 mol. (d) divided by 22.4 L. 2. c For the expression $V \propto nRT/P$, which of the following will cause the volume to increase?

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Gases and Gas Laws Learn with flashcards, games, and more — for free. ... 26 terms. jkinsler12. Modern Chemistry Chapter 11. Gases and Gas Laws. STUDY. PLAY. Ideal Gas Law. $PV=nRT$. P. Pressure. V. Volume. n. Moles. R. Constant. T. Temperature. Boyle's Law. The volume of a fixed mass of gas varies inversely with the pressure at constant ...

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Modern Chemistry Chapter 11 GASES Section 1 Gases & Pressure Pressure & Force Pressure (P) is defined as the force per unit of area on a surface. $\text{pressure} = \text{force} \div \text{area}$ $P = F/A$ Atmospheric pressure (atm) is the pressure exerted on an object due to the weight of the column of the air above it in the atmosphere.

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Temperature of gas depends on average kinetic energy of particles Ch 10 pg 329-330 Parameters for an ideal gas Kinetic molecular theory only applies to "ideal" gases. Most gases are only nearly ideal if pressure is not too high and temperature is not too low

Gases - Los Angeles County High School for the Arts

Holt McDougal Modern Chemistry Chapter 11: Gases Chapter Exam Instructions. Choose your answers to the questions and click 'Next' to see the next set of questions.

Holt McDougal Modern Chemistry Chapter 11: Gases ...

CHAPTER 11 REVIEW Molecular Composition of Gases MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. The average speed of a gas molecule is most directly related to the . (a) polarity of the molecule (b) pressure of the gas (c) temperature of the gas (d) number of moles in the sample 2.

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