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Chapter 10 Chemical Quantities Workbook

6.02 • 10^{23} representative particles of a substance molecule formula unit atom count the matter; measure the mass or weight; measure the volume They are questions about the amount of a substance and are similar to questions scientists ask. 0132525887_CHEM_WKBK_CH 10.indd 130 3/24/10 7:36:56 PM

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Chapter 10 - Chemical Quantities. Jennie L. Borders. Section 10.1 - The Mole: A Measurement of Matter. You often measure the amount of something by count, by mass, or by volume. A mole (mol) of a substance is 6.02×10^{23} representative particles of that substance.

Chapter 10 - Chemical Quantities

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Chemical Quantities Chapter 10 Answer Key

CHAPTER 10: Chemical Quantities BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

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Chapter 10 Chemical Quantities 243 Section Review Objectives • Convert the mass of a substance to the number of moles of a substance, and the number of moles of a substance to mass • Calculate the volume of a quantity of gas at STP Vocabulary • Avogadro's hypothesis • standard temperature and pressure (STP) • molar volume Key Equations • mass (grams) number of moles

Objectives Vocabulary Key Equations

92 Guided Reading and Study Workbook CHAPTER 10, Chemical Quantities (continued) 9. Complete the table about representative particles and moles. The Mass of a Mole of an Element (pages 293-294) 10. What is the atomic mass of an element? 11. Circle the letter of the phrase that completes this sentence correctly.

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296)

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