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Calcium oxide reacts with water in a combination reaction to produce calcium hydroxide: $\text{CaO(s)} + \text{H}_2\text{O(l)} \rightarrow \text{Ca(OH)}_2\text{(s)}$ In a particular experiment, a 3.00-g sample of CaO is reacted with excess water and 3.46 g of Ca(OH)_2 is recovered. What is the percent yield in this experiment?

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Express your answer with the appropriate units. Hint 1. How to approach the problem The pressure and the amount of air inside the balloon remain the same. Therefore this is a Charles's law problem. We can set up the following equation and solve it for : ANSWER: Now you can convert this value to degrees Celsius. ANSWER:

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