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mole of water is the water molecule, the representative particle in a mole of copper is the copper atom, and the representative particle in a mole of sodium chloride is the formula unit. 310 Chapter 11 The Mole Figure 11-2 The amount of each substance shown is 6.02 1023 or one mole of representative particles. The representative particle for

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Study Guide - Chapter 10 - The Mole Sectio n 10.1 Measuring Matter 1. pair 2. 5 3. dozen 4. gross 5. 200 6. ream 7. 6,000,000,000 8. 0.5 mol 9. 6.02 1023 10. four moles 11. 6.02 10 Cu atoms23 1 mol Cu 12. 4 23 4 1 mol CH 6.02 10 molecules CH 13. 23 1 mol Xe 6.02 10 molecules Xe 14. 23 2 2 6.02 10 molecules F 1 mol F Section 10.2 Mass and the ...

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Chemistry: Matter and Change Study Guide for Content Mastery Answer Key Name Date Class 62 Chemistry: Matter and Change • Chapter 11 Study Guide for Content Mastery Section 11.2 Mass and the Mole In your textbook, read about the mass of a mole. For each statement below, write true or false. 1.

Study Guide for Content Mastery - Teacher Edition

368 Chapter 11 • Stoichiometry Section 111.11.1 Objectives Describe the types of relationships indicated by a balanced chemical equation. State the mole ratios from a balanced chemical equation. Review Vocabulary reactant: the starting substance in a chemical reaction New Vocabulary stoichiometry mole ratio Defining Stoichiometry

Chapter 11: Stoichiometry

study the entire chapter online ... 322 Chapter 10 • The Mole Converting Between Moles and Particles Suppose you buy three-and-one-half dozen roses and want to know how many roses you have. Recall what you have learned about conver- ... Thus, 2.11×10 24 molecules of sucrose is 3.50 mol of sucrose.

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each problem, write in the space provided the mole ratio that can be used to solve the problem. Complete the equation by writing the correct value on the line provided.

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Section 11.1 What is stoichiometry? In your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false. _____ 1. The study of the quantitative

relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry.

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