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### **Chem 152 Chapter 19 Mastering Chemistry Homework ...**

Finally, calculate the volume of the suspension required from the required mass that we calculated:  $(72.64 \text{ milligrams}) / (100 \text{ mg} / 5.0 \text{ mL}) = 3.632 \text{ mL}$  should be the answer. Note: 100mg/5.0 mL is the density of the suspension, so it makes sense to divide the mass required by the density to get the volume required.

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