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Kinetic Theory of Gases 1) Indefinite shape and volume 2) Made up of particles in rapid, constant, random motion 3) Temperature changes greatly affect the volume of a gas 4) All collisions between particles in a gas are perfectly elastic - no loss of energy 5) Point masses - no volume or diameter - so [...]

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236 Study Guide for An Introduction to Chemistry Section Goals and Introductions Section 14.1 Collision Theory: A Model for the Reaction Process Goals To describe a model, called collision theory, that helps us to visualize the process of many chemical reactions. To use collision theory to explain why not all collisions between possible reactants lead

Chapter 14 - The Process of Chemical Reactions

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220 Study Guide for An Introduction to Chemistry Exercises Key Exercise 14.1-Using Electronegativities: Classify the following bonds as nonpolar covalent, polar covalent, or ionic. If a bond is polar covalent, identify which atom has the partial negative charge and which has the partial positive charge.

Chapter 14 Liquids: Condensation, Evaporation, and Dynamic ...

Chapter 14-Assignment E: Volume-Volume Gas Stoichiometry Chapter 14 concludes with a section on converting between volumes of gases reacting and produced in a chemical reaction. The main idea in this section is: 1) The ratio of volumes of gases in a reaction is the same as the ratio of moles,

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