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Distinguish between the single- and double-replacement reactions indicated below by completing each, identifying each as to type, and explaining, in words, what has occurred in each.

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a. The only way to determine the products of a reaction is to carry out the reaction. b. All chemical reactions can be classified as one of five general types. c. Complete combustion has occurred when all the carbon in the product is in the form of carbon dioxide. d. A single reactant is the identifying characteristic of a decomposition ...

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Small whole numbers that are placed in front of the formulas i... Chemical Equation A representation of a chemical reaction Skeleton Equation A chemical equation that does not indicate the relative amount... Made from carbon and Hydrogen atoms vegetable oil oxygen, sulfur, nitrogen, phosphorus, o... Gasoline, medicines, shampoos,...

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chemical reactions review identify the type of reaction and balance the equation: 1. $\text{sb} + \text{i} \rightarrow \text{sbi}$ 2. $\text{li} + \text{h}_2\text{o} \rightarrow \text{lioh} + \text{h}_2$ 3. $\text{alcl}_3 + \text{al} + \text{cl}_2 \rightarrow \text{c}_6\text{h}_{12} + \text{o}_2$ 4. $\text{co}_2 + \text{h}_2\text{o} \rightarrow \text{alcl}_3 + \text{na}_2\text{co}_3$ 5. $\text{al}_2(\text{co}_3)_3 + \text{nacl} \rightarrow \text{hno}_3 + \text{ba}(\text{oh})_2$ 6. $\text{ba}(\text{no}_3)_2 + \text{h}_2\text{o} \rightarrow \text{al} + \text{pb}(\text{no}_3)_2$ 7. $\text{al}(\text{no}_3)_3 + \text{pb}$ identify the type of reaction & write a balanced ...

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Chapter 11: Chemical Reactions Study Guide. chemical equation. A representation of a chemical

reaction with reactants on the left, products on the right, and an arrow separating the two. skeleton equation. A chemical equation that does not indicate the relative amounts of the reactants and products.

Chapter 11: Chemical Reactions Study Guide

Combination Reactions The first type of reaction is the combination, or synthesis, reaction. A combination reaction is a chemical change in which two or more substances react to form a single new substance. As shown in Figure 11.5, magnesium metal and oxygen gas combine to form the compound magnesium oxide.

11.2 Types of Chemical Reactions - bleiker.weebly.com

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reactants products yields, gives, or reacts to produce false acetylene oxygen JKL carbon dioxide water mercury(II) oxide JKL mercury oxygen A chemical equation is a representation of a chemical reaction; the formulas of the reactants (on the left) are connected by an arrow with formulas of the products (on the right).

SECTION 11.1 DESCRIBING CHEMICAL REACTIONS (pages 321-329)

Chapter 11 • Stoichiometry 367 SStart-Up Activitiesstart-Up Activities LAUNNCH CH LLabab What evidence can you observe that a reaction is taking place? During a chemical reaction, reactants are consumed as new products are formed. Often, there are several telltale signs that a chemical reaction is taking place. Procedure 1.

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