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Chapter 10 The Mole Assessment

CHAPTER 10 SOLUTIONS MANUAL The MoleThe Mole Solutions Manual Chemistry: Matter and Change • Chapter 10 161 Section 10.1 Measuring Matter page 320–324 Practice Problems pages 323–324 1. Zinc (Zn) is used to form a corrosion-inhibiting surface on galvanized steel. Determine the number of Zn atoms in 2.50 mol of Zn. 2.50 mol Zn

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The mole The mole, abbreviated mol, is the SI base unit used to measure the amount of a substance. A mole is defined as the number of carbon atoms in exactly 12 g of pure carbon-12. Through years of experimentation, it has been established that a mole of anything contains 6.0221367×10^{23} particles.

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mole of water is the water molecule, the representative particle in a mole of copper is the copper atom, and the representative particle in a mole of sodium chloride is the formula unit. 310 Chapter 11 The Mole Figure 11-2 The amount of each substance shown is 6.02 $\times 10^{23}$ or one mole of representative particles. The representative particle for

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10.5, 10.7; Assessment 10.1 • Go Online, Section 10.1 10.1 FOCUS Objectives 10.1.1 Describe methods of measuring the amount of something. 10.1.2 Define Avogadro's number as it relates to a mole of a substance. 10.1.3 Distinguish between the atomic mass of an element and its molar mass. 10.1.4 Describe how the mass of a mole of a compound ...

10.1 The Mole: A Measurement of Matter 10

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