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Chapter 13 States Of Matter

Chapter 13: States of Matter. -As a result, gases fill their containers regardless of shape and volume of the containers -Uncontained gas can spread out into space without limit -particles travel in straight line paths until they collide with another particle or another object such as the wall of the container -particles change direction only...

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The electron density around each nucleus is, for a moment, greater in one region of each cloud. Each molecule forms a temporary dipole. 394 Chapter 13 States of Matter. When temporary dipoles are close together, a weak dispersion force exists between oppositely charged regions of the dipoles, as shown in Figure 13-8.

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Section 13.1 The Nature of Gases Kinetic refers to motion The energy an object has because of it's motion is called kinetic energy The kinetic theory states that the tiny particles in all forms of matter are in constant motion! Section 13.1 The Nature of Gases Three basic assumptions of the kinetic theory as it applies to gases: #1.

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Chapter 13 States of Matter notes. So $F/A = F'/A'$. A small force applied over a small area can give a bigger force over a bigger area like in a hydraulic system. The pressure you feel under water is due to the weight of the water over you. For a uniform fluid this depends on the density and the depth. $P = \rho hg$ where is ρ the density of the fluid.

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CHAPTER 13, States of Matter (continued) SECTION 13.3 THE NATURE OF SOLIDS (pages 396–399) This section describes the highly organized structures of solids, distinguishes between a crystal lattice and a unit cell, and explains how allotropes of an element differ.

Name Date Class STATES OF MATTER 13

13.1 The Fluid States 300 States of Matter FIGURE 13–1 The ice cube, a solid, has a definite shape. But water, a fluid, takes the shape of its container.

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Chapter 13- The States of Matter Gases- indefinite volume and shape, low density. Liquids- definite volume, indefinite shape, and high density. Solids- definite volume and shape, high density Solids and liquids have high densities because their molecules are close together.

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13 STUDY GUIDE FOR CONTENT MASTERY CHAPTER States of Matter Section 13.1 Gases In your

textbook, read about the kinetic-molecular theory. Complete each statement. 1. The kinetic molecular theory describes the behavior of gases in terms of particles in 2. The kinetic-molecular theory makes the following assumptions. a.

CHAPTER 13 STATES OF MATTER.pdf

Chapter 13 States of Matter What are the three assumptions of the kinetic theory as it applies to gases? the particles in a gas are considered to be small, hard spheres with an insignificant volume. The motion of the particles in a gas is rapid, constant, and random.

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