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Chapter 13 States Of Matter
Chapter 13: States of Matter. -As a result, gases fill their containers regardless of shape and volume of the containers -Uncontained gas can spread out into space without limit -particles travel in straight line paths until they collide with another particle or another object such as the wall of the container -particles change direction only...

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The electron density around each nucleus is, for a moment, greater in one region of each cloud. Each molecule forms a temporary dipole. 394 Chapter 13 States of Matter. When temporary dipoles are close together, a weak dispersion force exists between oppositely charged regions of the dipoles, as shown in Figure 13-8.

Chapter 13: States of Matter - Jayne Heier
Chapter 13 States of Matter137. SECTION 13.1 THE NATURE OF GASES (pages 385–389) This section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the particles of a substance. Kinetic Theory and a Model for Gases (pages 385–386)

Name Date Class STATES OF MATTER 13
Chapter 13 States of Matter - Chapter 13 “States of... Section 13.2 The Nature of Liquids OBJECTIVES: Describe the equilibrium between a liquid and its vapor. Section 13.2 The Nature of Liquids OBJECTIVES: Identify the conditions at which boiling occurs. Section 13.2 The Nature of Liquids Liquid particles are also in motion.

Chapter 13 States of Matter - Course Hero
Chapter 13 States of Matter notes. So F/A = F'/A' . A small force applied over a small area can give a bigger force over a bigger area like in a hydraulic system. The pressure you feel under water is due to the weight of the water over you. For a uniform fluid this depends on the density and the depth. P = ρ hg where is ρ the density of the fluid.

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13.1 The Fluid States 300 States of Matter FIGURE 13–1The ice cube, a solid, has a definite shape. But water, a fluid, takes the shape of its container.

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Chapter 13- The States of Matter Gases- indefinite volume and shape, low density. Liquids- definite volume, indefinite shape, and high density. Solids- definite volume and shape, high density Solids and liquids have high densities because their molecules are close together.

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13 STUDY GUIDE FOR CONTENT MASTERY CHAPTER States of Matter Section 13.1 Gases In your textbook, read about the kinetic-molecular theory. Complete each statement. I. The kinetic molecular theory describes the behavior of gases in terms of particles in 2. The kinetic-molecular theory makes the following assumptions. a.

CHAPTER 13 STATES OF MATTER.pdf
, The motion of the particles in this is rapid, constant, and random. , The SI unit of pressure is this, abbreviated with the letters (Pa) or (kPa). , The average kinetic energy of a particle in a substance is directly related to the substance's this. , At zero degrees Celsius and a pressure of 101.3 kPa (or 1 atmosphere) you have this term abbreviated with the letters (STP).</p>

Chapter 13: States of Matter Jeopardy Template
Chapter 13 States of Matter. What are the three assumptions of the kinetic theory as it applies to gases? the particles in a fas are considered to be small, hard spheres with an insignificant volume. The motion of the particles in a fas is rapid, constant, and random. All collisions between particles in a gas are perfectly elastic.

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Chapter 13 Concept Map: ... Most of the states of matter are pretty steady, but solids have two different type of solids. Notice how above, the graph says a solid is packed orderly? This is recognizing the crystal structure of a solid. Most solids are crystal, which means the particles are arranged in a repeating, 3D pattern.

Chapter 13: States of Matter - Chemistry by Anna
Chapter 13 - States of Matter Chapter 14 - Behavior of Gases Chapter 15 - Water and Aqueous Systems Chapter 16 - Solutions Chapter 17 -Thermochemistry Chapter 18 - Reaction Rates and Equilibrium Chapter 19 - Acids, Bases and Salts Chapter 20 - Oxidation-Reduction Reactions

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