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Chapter 13 States Of Matter

The Kelvin temperature is directly proportional to the average kinetic energy -as the temp rises the air particles speed up and increase in kinetic energy which cause the raft to expand -as the temperature drops the air particles slow down resulting in a decrease in kinetic energy, causing the raft to deflate to its original YOU MIGHT ALSO LIKE...

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The electron density around each nucleus is, for a moment, greater in one region of each cloud. Each molecule forms a temporary dipole. 394 Chapter 13 States of Matter. When temporary dipoles are close together, a weak dispersion force exists between oppositely charged regions of the dipoles, as shown in Figure 13-8.

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Section 13.1 The Nature of Gases Kinetic refers to motion The energy an object has because of it's motion is called kinetic energy The kinetic theory states that the tiny particles in all forms of matter are in constant motion! Section 13.1 The Nature of Gases Three basic assumptions of the kinetic theory as it applies to gases: #1.

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Chapter 13 States of Matter notes. So $F/A = F'/A'$. A small force applied over a small area can give a bigger force over a bigger area like in a hydraulic system. The pressure you feel under water is due to the weight of the water over you. For a uniform fluid this depends on the density and the depth. $P = \rho hg$ where is ρ the density of the fluid.

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CHAPTER 13, States of Matter (continued) SECTION 13.3 THE NATURE OF SOLIDS (pages 396–399) This section describes the highly organized structures of solids, distinguishes between a crystal lattice and a unit cell, and explains how allotropes of an element differ.

Name Date Class STATES OF MATTER 13

lamp or the sun, the light's source is plasma.

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Chapter 13- The States of Matter Gases- indefinite volume and shape, low density. Liquids- definite volume, indefinite shape, and high density. Solids- definite volume and shape, high density Solids and liquids have high densities because their molecules are close together.

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13 STUDY GUIDE FOR CONTENT MASTERY CHAPTER States of Matter Section 13.1 Gases In your textbook, read about the kinetic-molecular theory. Complete each statement. I. The kinetic

molecular theory describes the behavior of gases in terms of particles in 2. The kinetic-molecular theory makes the following assumptions. a.

CHAPTER 13 STATES OF MATTER.pdf

Chapter 13 States of Matter What are the three assumptions of the kinetic theory as it applies to gases? the particles in a gas are considered to be small, hard spheres with an insignificant volume. The motion of the particles in a gas is rapid, constant, and random.

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Chapter 13 Concept Map: ... Most of the states of matter are pretty steady, but solids have two different type of solids. Notice how above, the graph says a solid is packed orderly? This is recognizing the crystal structure of a solid. Most solids are crystal, which means the particles are arranged in a repeating, 3D pattern.

Chapter 13: States of Matter - Chemistry by Anna

Chapter 13 - States of Matter Chapter 14 - Behavior of Gases Chapter 15 - Water and Aqueous Systems Chapter 16 - Solutions Chapter 17 - Thermochemistry Chapter 18 - Reaction Rates and Equilibrium Chapter 19 - Acids, Bases and Salts Chapter 20 - Oxidation-Reduction Reactions

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