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$P_{ideal} = 1.0004$, $P_1 = nRT/(V - nb)$ $P_1 = 1.0029$, $P_2 = (a \cdot n^2)/V^2$ $P_2 = 0.0129$, $P_{waals} = P_1 +$
 P_2 $P_{waals} = 1.0158$ The ideal gas law predicts a pressure of 1.0004 atmospheres, while the van
der Waals model predicts 1.0158 atmospheres.

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end disp(x) T4.5-1 The script le is angle = input('Enter an angle in degrees.') switch angle case 45
disp('Angle is in first quadrant') case -45 disp('Angle is in second quadrant') case 135 disp('Angle
is in third quadrant') case -135 disp('Angle is in fourth quadrant') otherwise disp('Quadrant is
unknown.') end T4.7-1 The 0.75 in the matrix should be replaced with 0.70.

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