

Baking Soda And Vinegar Stoichiometry Lab Answers

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Baking Soda And Vinegar Stoichiometry

Using the concept of stoichiometry, the amount of product that results from a chemical reaction can be predicted. Baking soda is a powdered chemical compound called sodium bicarbonate, and vinegar includes acetic acid.

Stoichiometry: Baking Soda and Vinegar Reactions

Vinegar and Baking Soda Stoichiometry Lab. Purpose: To predict the amount of Carbon Dioxide gas that should be produced in a chemical reaction; then calculate the amount of CO₂ released, the percent yield. Materials: Baking Soda (NaHCO₃), Vinegar (CH₃COOH), 2 beakers and electronic balance. Procedure: 1. Obtain and record the mass of 100 mL beaker.

Vinegar and Baking Soda Stoichiometry Lab

baking soda reacts comes into contact with the vinegar. 4. Shake the bottle gently until all the baking soda has reacted with the vinegar. Allow the solution to fizz up then slowly settle. Wait until the baking soda has dissolved completely into the vinegar, shown by no significant bubbling in the bottle. Keep the bottle sealed for Part 2.

Stoichiometry: Baking Soda and Vinegar Reactions

Using this technique, we discovered that the limiting reactant changed from vinegar to baking soda as the amount of baking soda used increased. With lighter masses, there was not enough baking soda to react with all of the vinegar, and with heavier masses, there was not enough vinegar to dissolve the baking soda.

Baking Soda and Vinegar Stoichiometry | The Chem Chapter

Dropping the baking soda into the vinegar in a plastic pouch prevents the reaction from occurring immediately once the baking soda is added. The plastic acts as a temporary barrier to mixing the two components, giving the experimentalist time to close and seal the bottle.

(PDF) Stoichiometry: Baking Soda and Vinegar Reactions ...

Stoichiometry Lab: Vinegar and Baking Soda Do Not Write On This Sheet Purpose: To predict the amount of Carbon Dioxide gas that should be produced in a chemical reaction; then calculate the % yield. $\text{CH}_3\text{COOH} + \text{NaHCO}_3 \rightarrow \text{NaCH}_3\text{COO} + \text{H}_2\text{O} + \text{CO}_2$ Materials: Baking Soda (NaHCO₃), Vinegar (CH₃COOH), and 2 plastic cups, scale. Procedure: Find and record the mass of cup A.

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Baking soda and vinegar were the reactants, and the products were carbon dioxide, water, and sodium acetate. After mixing these various chemicals we placed them on a hot plate and let them boil for about 20 minutes.

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