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15. Use one of the methods in Model 3 that gave the correct answer for average atomic mass to calculate the average atomic mass for oxygen. Isotope information is provided below. Show all of your work and check your answer against the mass listed on the periodic table. Isoto e Natural Abundance on Earth (0/0) Atomic Mass (am u) - 16.00 160 170 180 99.76 O. 0.20 15.9949 16.9991 17.9992 + (. + . 00 n ST op Average Atomic Mass .

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Which isotope has an atomic mass closest to the average atomic mass listed on the periodic table? 24Mg isotope has the atomic mass closest to the average atomic mass b. Give a mathematical reason for your answer to part a. The average atomic mass is 24.305 amu. The mass of 24Mg is 23.985 amu. 17.

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Average Atomic Mass Answer Key. The isotope with a mass number of 355 is more abundant. The isotope with a mass number of 356 is more abundant. The isotopes have an equivalent abundance. An isotope with a mass number of 357 is more abundant.

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15. Use one of the methods in Model 3 that gave the correct answer for average atomic mass to calculate the average atomic mass for oxygen. Isotope information is provided below. Show all of your work and check your answer against the mass listed on the periodic table. Isotope Natural Abundance on Earth (%) Atomic Mass (amu) 160 99.76 15.9949

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CP Chemistry 2015-16 Kaspriskie Average Atomic Mass POGIL Activity. Isotopes Review. Fill in the following table. Isotope Symbol Atomic Number Mass Number Protons Neutrons 1940 9 19 9 10 918 19 40 19 21 S-31 16 31 16 15 Do all isotopes of an element have the same atomic number? Explain your answer. YES.

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the mass of a neutron) to determine the atomic mass of a particular atom. The atomic mass is the total number of protons + neutrons in the nucleus of the atom. The atomic mass is the average of all of the masses of the different isotopes of elements (isotopes will be discussed later). 3.

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Accelerated Chemistry: Average Atomic Mass Worksheet Calculate the average atomic mass for each element based on the natural abundance of its isotopes. 1, Find the average atomic mass for Li if 7.5% of Li atoms are 6Li with a mass of 6.0151223 amu and 92.5% are 7 Li with a mass of 7.0160041 amu. 0.075 + (7.0 = 2)

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14. Propose a possible way to calculate the average atomic mass of 100 magnesium atoms. Your answer may include a mathematical equation, but it is not required. Model 4 – Proposed Average Atomic Mass Calculations Key Questions 15. Complete the three proposed calculations for the average atomic mass of magnesium in Model 4. 16.

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