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You are correct, stainless steel is a really poor conductor compared to most metals. This source lists it as $7.496 \times 10^{-7} \text{ } \mathrm{\Omega \cdot m}$ which is more than 40 times worse than copper.. The reason is that conductivity in metals is high is that metals form a crystal lattice where the outer shell electrons are shared and easily move through the lattice.

Why is stainless steel a poor conductor of electricity?

Diamond is a solid form of the element carbon with its atoms arranged in a crystal structure called diamond cubic. At room temperature and pressure, another solid form of carbon known as graphite is the chemically stable form, but diamond almost never converts to it. Diamond has the highest hardness and thermal conductivity of any natural material, properties that are utilized in major ...

Diamond - Wikipedia

In solid-state physics, the electronic band structure (or simply band structure) of a solid describes the range of energies an electron within the solid may have (called energy bands, allowed bands, or simply bands) and ranges of energy that it may not have (called band gaps or forbidden bands).. Band theory derives these bands and band gaps by examining the allowed quantum mechanical wave ...

Electronic band structure - Wikipedia

This paper reflects the achievements and the challenges ahead in the field of organic light emitting diodes (OLEDs). The primary intention of this paper is to study different organic materials synthesized so far and the OLEDs fabricated for solid-state lighting.

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