

Balancing Chemical Equations Yahoo Answers

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Yahoo Answers Sign in Mail ☼ ... Balancing Chemical Equations? How would one balance the chemical equation: $\text{N}_2 + \text{H}_2 \Rightarrow \text{NH}_3$ Please provide steps and show how you solved the answer. I always choose best answer. Follow . 4 answers 4.

Balancing Chemical Equations? | Yahoo Answers

I am in 7th grade and we are briefly touching on balancing chemical equations. I don't really get it. Can someone please explain balancing chemical equations. Thanks P.S. I kind of need my answer ASAP!!! Thanks

Balancing Chemical Equations? | Yahoo Answers

Yahoo New Zealand Answers ... Balancing Chemical Equations? Write a balanced chemical equation for the reaction of iron with oxygen forming iron oxide (Fe_2O_3). If anyone can help me answer this it will be greatly appreciated! Follow . 2 answers 2.

Balancing Chemical Equations? | Yahoo Answers

Best Answer: $2\text{H}_2 + \text{O}_2 = 2\text{H}_2\text{O}$ To balance the equations, look at each ion: $\text{H}_2 + \text{O}_2 = \text{H}_2\text{O}$ (unbalanced) Looking at the left side, you see that there are 2 oxygen ions (O_2 consists of O----O bond, so there are 2 atoms), and when you look at the right side, you see that H_2O has only one oxygen ion, the digit "2" tells us about 2 ions of hydrogen, but there's only 1 O ion.

Balancing chemical equations? | Yahoo Answers

Yahoo Canada Answers ... Help with balancing chemical equations please!? How do you balance chemical equations? Answer Questions. How effective is benzodiazepines, antihistamines, cocaine and alcohol combined for suicide?

Balancing chemical equations help ?!? | Yahoo Answers

Best Answer: The idea behind balancing an equation is to make sure there are equal amounts of element on each side of the equation. Then, you add coefficients to each item to multiply the amounts and get even numbers of each element on each side. I believe my chem teacher told us to do it in this order ...

balancing chemical equations? | Yahoo Answers

$\text{H}_2 + \text{Br}_2 = \text{HBr}$ & HBr You have to make sure that there is the same amount of elements on each side of the equation. so for this equation it makes 2 molecules of HBr because there are 2 of each element on the other side of the equation.

Balancing Chemical Equations!? | Yahoo Answers

Yahoo Malaysia Answers ... When balancing an equation, you must have the same number of molecules of a particular substance on each side of the equation. In your example, left as it is there are 2 hydrogen atoms on one side (reactant) and 2 on the other side (product). ... Help with balancing chemical equations please!? How do you balance ...

Balancing chemical equations? | Yahoo Answers

Yahoo En Español Respuestas ... Help with balancing chemical equations please!? How do you balance chemical equations? Responder preguntas ¿Como varia la densidad del vino respecto al tiempo de fermentación? ¿Dado 1000 pies³ de un gas a 15°C y 763 mmHg. Calcular el volumen a -6°C y 420 mmHg?

balancing chemical equations? | Yahoo Respuestas

Give the balanced equation for each of the following. (Use the lowest possible coefficients. Include states-of-matter under the given conditions in your answer.) (a) The combustion of ethanol ($\text{C}_2\text{H}_5\text{OH}$) forms carbon dioxide and water vapor. A combustion reaction refers to a reaction of a substance with oxygen gas. b.) Aqueous solutions of lead(II) nitrate and sodium phosphate are mixed ...

Balancing chemical equations? | Yahoo Answers

Best Answer: Basically, it is a math problem when you know the formulas of the products and educts. Taking the first problem: for each S on the product side, there have to be 2 H. Already there in the unbalanced form. However, on the product side we find 3 O, on the educt side we find 2 O. This means we ...

Balancing chemical equations.....?!? | Yahoo Answers

When balancing an equation, you must have the same number of molecules of a particular substance on each side of the equation. In your example, left as it is there are 2 hydrogen atoms on one side (reactant) and 2 on the other side (product).

Balancing chemical equations? | Yahoo Answers

Best Answer: $\text{Na}_2\text{SO}_4 + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O} + \text{SO}_2 + \text{O}_2$ This is your complete reaction For balancing the equation you should check if al the elements of the reactants are equal in magnitude to the same elements in the products example: Checking for Na Rectant - Na_2SO_4 - $\rightarrow \text{Na}_2$ Product - $\text{NaCl} \rightarrow \text{Na}$ which means Na ...

Balancing chemical equations? | Yahoo Answers

Best Answer: Do 1 element at a time If you end up with fractional coefficients let them stay for time being then multiply both sides by the denminator So start with the C's 3 C's on this side so put 3 CO_2 on other side Now you have 7 O on right side (6 from 3 CO_2 and 1 from H_2O) There are 8 H's on left side ...

Balancing Chemical Equations? | Yahoo Answers

Best Answer: I think you mean Fe_2O_3 , since FeO_3 does not exist. Thus, the equation should be $\text{Fe}_2\text{O}_3 + \text{CO} \rightarrow \text{Fe} + \text{CO}_2$ First, check how many of each type of atom you have on either side of the reaction symbol. On the left, you have 2 Fe, 4 O and 1 C. On the right, you have 1 Fe, 2 O and 1 C. O and Fe are thus ...

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