

Behavior Of Gases Core Teaching Resources Answers

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Behavior Of Gases Core Teaching

Gas and Momentum. The particles in a gas have momentum, the product of their mass and velocity (mv). The pressure exerted by a gas in a given volume is proportional to the average kinetic energy of its molecules ($\frac{1}{2}mv^2$). The mass of the particles is constant, so any change in momentum is due to a change in velocity.

14.1 Properties of Gases 14 - Grandview Independent School ...

When the temperature of a sample of steam increases from 100°C to 200°C , the average kinetic energy of its particles doubles. This section explains the relationships among the volume, pressure, and temperature of gases as described by Boyle's law, Charles's law, Gay-Lussac's law, and the combined gas law.

SECTION 14.1 PROPERTIES OF GASES(pages 413-417)

exerted by a mixture of gases is equal to the sum of the partial pressures of the component gases. c. d. the contribution each gas in a mixture makes to the total pressure e. A gas tends to move to an area of lower concentration until the concentration is uniform throughout. f. $P_1 T_1 V_1 P_2 T_2 V_2$ g. $P V n R T h$.

05 CTR ch14 7/12/04 8:13 AM Page 361 THE BEHAVIOR OF GASES 14

In answering the question below, please be sure to include BOTH your work and your answer. When 3.56 grams of a compound is vaporized at 245°C and 1.20 atm, the resulting gas occupies a volume of 344 mL.

Behavior of Gases (College) - Help Teaching

This Behavior of Gases Worksheet is suitable for 9th - 12th Grade. In this gas behavior worksheet, students answer 5 questions about gases, how pressure, altitude and temperature effect the behavior of gases and about Gay-Lussac's Law.

Behavior of Gases Worksheet for 9th - 12th Grade | Lesson ...

A mixture of gases at a total pressure of 145.0 kPa contains N_2 , CO_2 , and O_2 . 14.4 The partial pressure of the N_2 is 28.5 kPa, and the partial pressure of the CO_2 is 76.0 kPa. What is the partial pressure of the O_2 ? 5 14.4 4 14.3 3 14.1 2 14.1 1 14.1 05_CTR_ch14 7/12/04 8:13 AM Page 360

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Chemistry Ch 14B - Name Date Class THE BEHAVIOR OF GASES... The volume of a fixed mass of gas is directly proportional to its Kelvin temperature if the pressure is kept constant. The rate of effusion of a gas is inversely proportional to the square root of its molar mass. For a given mass of gas at constant temperature,...

Chemistry Ch 14B - Name Date Class THE BEHAVIOR OF GASES ...

As the temperature of an enclosed gas decreases, the pressure 9 Part B True-False. Classifr each of these statements as always true, AT; sometimes true, ST; or never true, NT 10. According to kinetic theory, the volume of the particles in a gas is 348 Core Teaching Resources 11.

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Postulates of the Kinetic Molecular Theory. The behavior of gases is explained by what scientists call the Kinetic Molecular Theory. According to this theory, all matter is made of constantly moving atoms or molecules. Because of their mass and velocity, they possess kinetic energy, ($K.E. = \frac{1}{2}mv^2$).

The Theories and Behavior of Gas | Owlcation

(At a higher temperature, the behavior of a gas is closer to the ideal; so the curve might not dip below the straight line.) Use Visuals. Figure 14.15 Point out that the straight line represents ideal conditions. For an ideal gas, by definition $PV = nRT$, and the ratio of PV to nRT is 1.

14.3 Ideal Gases - Grandview Independent School District

Chapter 14 The Behavior of Gases 347 Name _____ Date _____ Class _____ THE PROPERTIES OF GASES 14.1

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