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6) How many grams are in 11.9 moles of chromium? 7) How many moles are in 9.8 grams of calcium? 8) How many grams are in 238 moles of arsenic? Solve the following: 9) How many grams are in 4.5 moles of sodium fluoride, NaF? (molar mass of NaF is $22.99 + 19.00 = 41.99$ g/ mole) $4.5 \text{ moles} \times 41.99 \text{ grams} = 188.955 \text{ g NaF} =$

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The mole is the unit for amount of substance. The number of particles in a substance can be found using the Avogadro constant. The mass of product depends upon the mass of limiting reactant.

The mole - Higher - bbc.com

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Gas Laws - Purdue University College of Science Welcome

10.1 Equation Stoichiometry 369 The ratio of moles of P 4O 10 to moles of P (which came from the subscripts in the chemical formula, P 4O 10) provided the key conversion factor that allowed us to convert from units of phosphorus to units of tetraphosphorus decoxide. Now let's assume that you have been transferred to the division responsible for the

Chapter 10 ChemiCal alCulations and equations - Mark Bishop

This page looks at the assumptions which are made in the Kinetic Theory about ideal gases, and takes an introductory look at the Ideal Gas Law: $pV = nRT$. This is intended only as an introduction suitable for chemistry students at about UK A level standard (for 16 - 18 year olds), and so there is no ...

Ideal gases and the ideal gas law: $pV = nRT$ - Main Menu

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So although for simple calculation purposes we use the value 22.4 dm³ for all gases, you can see that it isn't exactly true. Even at ordinary temperatures and pressures, real gases can deviate slightly from the ideal value.

REAL GASES - chemguide

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