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The Chemical Equations and Reactions chapter of this Holt McDougal Modern Chemistry Companion Course helps students learn the essential lessons associated with chemical equations and reactions.

### **Holt McDougal Modern Chemistry Chapter 8: Chemical ...**

Modern Chemistry 71 Chapter Test Name Class Date Chapter Test B, continued 17. A substance combines with oxygen, releasing a large amount of energy as heat and light, in a(n) . 18. The decomposition of a substance by an electric current is called . 19. A(n) orders the elements by the ease with which they undergo certain chemical reactions. 20.

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Modern Chemistry 68 Chapter Test Name Class Date Chapter Test A, continued \_\_\_\_ 21. Metal X replaces the ions of metal Y from solution, but it cannot replace the ions of metal Z from solution. The order these metals should have in the activity series (from top to bottom) is a. x, y, z b. z, x, y

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One way to test the salinity of a water sample is to add a few drops of silver nitrate solution with a known concentration. As the solutions of sodium chloride and silver nitrate mix, a precipitate of

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Holt McDougal Modern Chemistry 1 Chapter Test Assessment Chapter Test A Chapter: The Periodic Law Use the periodic table below to answer the questions in this Chapter Test. In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. \_\_\_\_ 1.

### **Assessment Chapter Test A - mrbijlani.weebly.com**

Modern Chemistry 47 Chapter Test Name Class Date Chapter Test A, continued \_\_\_\_ 7. The melting points of ionic compounds are higher than the melting points of molecular compounds because a. ionic substances tend to vaporize at room temperature. b. ionic substances are brittle. c. attractive forces between ions are greater than the attractive ...

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### **Solutions to Modern Chemistry (9780030367861) :: Free ...**

Modern Chemistry 84 Chapter Test Chapter: States of Matter In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. \_\_\_\_ 1. The kinetic-molecular theory of gases assumes that the particles of an ideal gas are separated by great distances. This implies that the gas

### **Assessment Chapter Test A - Kettering City School District**

Accelerated Chemistry Anderson - MCHS Modern Chemistry 7 Chemical Equations & Reactions CHAPTER 8 STUDY GUIDE Chemical Equations and Reactions SECTION 8-3 SHORT ANSWER Answer the following questions in the space provided. 1. List four metals that will not replace hydrogen in an acid.

### **CHAPTER 8 Chemical Equations and Reactions**

CHAPTER 1 REVIEW Matter and Change MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a homogeneous or heterogeneous substance. a. sugar d. plastic wrap b. iron filings e. cement sidewalk c. granola bar 2. For each type of investigation, select the most appropriate branch of chemistry from the following

### **mc06se cFMsr i-vi**

Study Sections 8.4–8.11. Focus on Goals 3–8 as you study. Strategy Table 8.2 provides a summary of all the types of reactions and equations. As with Assignment A, practice is the key. Answer The Equation-Classification Exercises (3–12) and Questions, Exercises, and Problems 2–14. Check your answers with those at the end of the chapter.

### **Chapter 8**

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CHAPTER 10 REVIEW States of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Match description on the right to the correct crystal type on the left. b ionic crystal (a) has mobile electrons in the crystal c covalent molecular crystal (b) is hard, brittle, and nonconducting a metallic crystal (c) typically has the lowest melting point of the four

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