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Periodic Law The physical/chemical properties of elements are periodic functions of their atomic numbers. Ionization Energy The energy it takes to remove one electron from a neutral atom Electronegativity the measure of the ability of an atom to attract electrons Electron Affinity Energy change that occurs when an electron is acquirred by a neutral atom [...]

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Modern Chemistry 47 Chemical Bonding CHAPTER 6 REVIEW Chemical Bonding SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. _____ In metals, the valence electrons are considered to be (a) attached to particular positive ions. (c) immobile. (b) shared by all surrounding atoms.(d) involved in covalent bonds. 2.

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mc06se cFMsr i-vi

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1. The charge on an ion is a. always positive. b. always negative. c. either positive or negative.

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Accelerated Chemistry Anderson - MCHS Modern Chemistry 1 Chemical Bonding CHAPTER 6 Chemical Bonding SECTION 1 Introduction to Chemical Bonding OBJECTIVES 1. Define Chemical bond. 2. Explain why most atoms form chemical bonds. 3. Describe ionic and covalent bonding. 4.

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