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3. What is the molarity of a solution of HNO₃ that contains 12.6 grams HNO₃ in 1.0 L of solution? ?
 $\text{mol HNO}_3 = 12.6 \text{ g HNO}_3 \times 1 \text{ mol HNO}_3 / 63.0 \text{ g HNO}_3 = 0.200 \text{ mol HNO}_3$
 $M = 0.200 \text{ mol HNO}_3 / 1.0 \text{ L} = 0.200 \text{ M}$
4. How many grams of potassium nitrate are required to prepare 0.250 L of a 0.700 M solution?
 $0.700 \text{ M} = \text{moles of solute} / 0.250 \text{ L}$
 $\text{moles of solute} = 0.700 \text{ M} \times 0.250 \text{ L} = 0.175 \text{ mol}$
 $\text{grams of solute} = 0.175 \text{ mol} \times 101 \text{ g/mol} = 17.7 \text{ g}$

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Molarity Problems Worksheet $M = \frac{n}{V}$ - n = # moles V - V must be in liters (change if necessary) - Use M or mol/L as unit for molarity 1. What is the molarity of a 0.30 liter solution containing 0.50 moles of NaCl?

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