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Modern Chemistry Chapter 7. compounds composed of two different elements; (compounds of 2 elements between metal (positive, s-block and p-block) and nonmetal (negative); metal is always listed first; change nonmetal to -ide ending)

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Holt McDougal Modern Chemistry Chapter 7: Chemical Formulas and Chemical Compounds Chapter Exam. Exam Instructions: Choose your answers to the questions and click 'Next' to see the next set of questions. You can skip questions if you would like and come back to them later with the yellow "Go To First Skipped Question" button.

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Modern Chemistry • CHAPTER 7 HOMEWORK 7-1 ... Answer the following questions in the space provided. 1. ... MODERN CHEMISTRY SECTION 7-1 REVIEW. 53.

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Modern Chemistry 33 Chapter Test Name Class Date Chapter Test B, continued 15. The energy state of an atom is called its ground state. 16. The number of waves that pass a point in one second is called. 17. When an electron drops from a higher-energy state to a lower-energy state, a(n) spectrum is produced. 18.

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The Chemical Formulas and Chemical Compounds chapter of this Holt McDougal Modern Chemistry Companion Course helps students learn the essential lessons associated with chemical compounds and ...

Holt McDougal Modern Chemistry Chapter 7: Chemical ...

Holt McDougal Modern Chemistry 1 Chapter Test Assessment Chapter Test A Chapter: The Periodic Law Use the periodic table below to answer the questions in this Chapter Test. In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. ____ 1.

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Modern Chemistry . Name Section Quiz, continued Class Date 6. How many atoms are present in one formula unit of barium acetate, c. 15 d. 16 7. What is the formula for the compound dinitrogen

tetroxide? a. d. NO 8. The term acid usually refers to a. a solution of the acid compound in water.

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CHAPTER 7 • HOMEWORK 7-1 VOCABULARY 1. false 2. true 3. false 4. false 5. true SKILL BUILDER
AgCl—In the compound, there is one atom of silver for each atom of chlorine. Ba(OH) 2 —The subscript 2 refers to everything inside the parentheses, so there are two hydroxide ions for

HOMEWORK 7-2 - Kettering City School District

Modern Chemistry • CHAPTER 7 HOMEWORK 7-3 (pp. 211–213) VOCABULARY Write the prefix that corresponds to each number. Try not to use the book. Number Prefix Number Prefix 16 27 38 49 510 GRAPHIC ORGANIZER Complete the chart by supplying the missing information. Name of Compound Formula of Compound CO 2 Nitrogen dioxide S 2 Cl 2 Carbon ...

HOMEWORK 7-3 - Kettering City School District

Modern Chemistry • CHAPTER 7 HOMEWORK 7-1 (pp. 203–206) VOCABULARY Write true or false for each statement. 1. A chemical formula indicates the relative number of molecules of each kind in a chemical compound. ____ 2. Monoatomic ions are ions formed from a single atom. ____ 3.

HOMEWORK 7-1 - Los Banos Unified School District

Modern Chemistry 92 Chapter Test Name Class Date Chapter Test B, continued 31. Calculate the energy released by freezing 13.8 g of a liquid. Its molar mass is 82.9 g/mol, and its molar enthalpy of fusion is 4.60 kJ/mol. PART VI Define the term equilibrium vapor pressure. Use it to explain the following two phenomena.

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CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Write formulas for the following compounds: CuCO 3 a. copper(II) carbonate Na 2SO 3 b. sodium sulfite (NH 4) 3PO 4 c. ammonium phosphate SnS 2 d. tin(IV) sulfide HNO 2 e. nitrous acid 2.

7 Chemical Formulas and Chemical Compounds

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Chemistry Chapter 7 Chemical Formulas and Compounds ...

Holt McDougal Modern Chemistry 3 Chapter Test Chapter Test B, continued 16. The measure of the ability of an atom in a chemical compound to attract electrons from another atom in the compound is called _____. 17. The energy required to remove one electron from an atom is called its _____. 18.

Assessment Chapter Test B - Wag & Paws

Modern Chemistry 137 Chapter Test Name Class Date Chapter Test A, continued ____ 19. In the figure on the previous page, the pH at the equivalence point a. is equal to 7.0. b. is greater than 7.0. c. is less than 7.0. d. cannot be determined from the data given. ____ 20. In the figure on the previous page, the volume of titration standard

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