

## *Nature Of Gases Section Review Answer Key*

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### **Nature Of Gases Section Review**

The theory assumes that the volume occupied by a gas is mostly \_\_\_\_\_, that the particles of gas are relatively \_\_\_\_\_, move \_\_\_\_\_ of each other, and are in constant \_\_\_\_\_ motion.

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Page 317 THE NATURE OF GASES 13 ... THE NATURE OF GASES 13.1 ... Section Review Objectives ...  
that are introduced in this section. Each blank can be completed with a term, short  
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13.1 ... Section Review Objectives ... Answer the following in the space provided. 22. A gas is at a  
pressure of 4.30 atm. SECTION 13.1 THE NATURE OF GASES SECTION 13.2 THE NATURE OF ...

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Name Date Class. THE NATURE OF GASES Section Review. Objectives □ Describe the assumptions of  
the kinetic theory as it applies to gases □ Interpret gas pressure in terms of kinetic theory □ Define  
the relationship between Kelvin temperature and average kinetic energy.

### **THE NATURE OF GASES - Quia**

that are introduced in this section. Each blank can be completed with a term, short phrase, or  
number. Liquids are much than gases. Liquids and solids 1. are known as states of matter. The  
conversion of a liquid 2. to a gas or vapor is called . When a liquid that is not 3. changes to a gas,  
the process is called evaporation. A 4.

### **05 CTR ch13 7/12/04 8:12 AM Page 317 THE NATURE OF GASES 13**

Chapter 13 States of Matter137. SECTION 13.1 THE NATURE OF GASES (pages 385–389) This  
section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure  
and explains how temperature is related to the kinetic energy of the particles of a substance.  
Kinetic Theory and a Model for Gases (pages 385–386)

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The Nature of Gases 13.1. because during an evaporation, particles with highest kinetic energy tends to escape first and the particles left have lower kinetic energy making it cool.

**The Nature of Gases 13.1 Flashcards | Quizlet**

gases are similar, and one way in which they are different. Both liquids and gases can flow, so they can take the shape of their container. The molecules in a liquid have intermolecular attractions that are not present in gases. Therefore, liquids have a definite volume and will not simply fill their container.

**Chapter 13**

Chapter 13- The States of Matter 13.1- The Nature of Gases section 13.2 the nature of liquids (pages 390-395) This section describes a model for liquids in terms of kinetic energy and the attractive forces between the particles in a liquid.

**13 1 The Nature Of Gases Section Review Answers Pearson ...**

Section 13.2 Review ... According to the kinetic theory, both the particles in gases and the particles in liquids have kinetic energy. This energy allows the particles in gases and ... Ethanol 1.63 5.85 18.04 47.02 108.34 225.75. Section 13.2 The Nature of Liquids.

**13.2 The Nature of Liquids - Henry County Schools / Overview**

Chapter 13.2 States of Matter, The nature of liquids. The particles in a liquid have an attraction for each other. The kinetic energy in them cannot overcome this attraction ... Vapor pressure is the pressure exerted by a gas above a liquid in a sealed container. More heat = more vapor pressure. Past boiling point is more vapor pressure.

**Chapter 13.2 States of Matter, The nature of liquids**

THE PROPERTIES OF GASES 14.1 Section Review Objectives why gases are easier to compress than solids or liquids are Describe the three factors that affect gas pressure Vocabulary compressibility Part A Completion Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section.

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