

## *Mole Study Guide For Content Mastery Answer*

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Mole the basic unit in the International System of Units (SI), representing the amount of a substance expressed in grams containing as many atoms, molecules, or ions as the number of atoms in 12 grams of carbon-12 (which is Avogadro's number, or  $6.022 \times 10^{23}$ ).

### **Moles Study Guide Chemistry Flashcards | Quizlet**

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Conversions. Fri 1/18. 20. Chapter 10 Study Guide (9-15). Tues 1/22. 10. The mole is the SI unit to measure the amount of a substance. 12g of carbon-12 is set  $2.50 \text{ mol} \times 6.02 \times 10^{23} \text{ atoms} = 1 \text{ mol}$ . = Section 11.1 Measuring Matter. Counting The Mole – Chapter 11 Study Guide for Content Mastery. Read your.

### **Chapter 10 Study Guide The Mole Section 10 1 Measuring Matter**

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