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6) How many grams are in 11.9 moles of chromium? 7) How many moles are in 9.8 grams of calcium? 8) How many grams are in 238 moles of arsenic? Solve the following: 9) How many grams are in 4.5 moles of sodium fluoride, NaF? (molar mass of NaF is $22.99 + 19.00 = 41.99$ g/ mole) $4.5 \text{ moles} \times 41.99 \text{ grams} = 188.955 \text{ g NaF} =$

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The mole is the unit for amount of substance. The number of particles in a substance can be found using the Avogadro constant. The mass of product depends upon the mass of limiting reactant.

The mole - Higher - bbc.com

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Parabolic Calculator - 101science.com

10.1 Equation Stoichiometry 369 The ratio of moles of P_4O_{10} to moles of P (which came from the subscripts in the chemical formula, P_4O_{10}) provided the key conversion factor that allowed us to convert from units of phosphorus to units of tetraphosphorus decoxide. Now let's assume that you have been transferred to the division responsible for the

Chapter 10 Chemical Calculations and equations - Mark Bishop

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Ideal gases and the ideal gas law: $pV = nRT$ - Main Menu

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