

## ***Mole Ratios Packet Answer Key***

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let's consider the equation  $\text{CH}_4 + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$  write the molar ratios for  $\text{CH}_4/\text{H}_2\text{O}$ ,  $\text{H}_2\text{O}/\text{H}_2$ ,  $\text{H}_2/\text{CH}_4$   
ii. if you use mole of  $\text{CH}_4$  how many moles of  $\text{H}_2$  could be

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Mole ratios are used as conversion factors between products and reactants in stoichiometry calculations.. For example, in the reaction  $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{g})$  The mole ratio between  $\text{O}_2$  and  $\text{H}_2\text{O}$  is  $\#(1 \text{ mol O}_2)/(2 \text{ mol H}_2\text{O})\#$ .. The mole ratio between  $\text{H}_2$  and  $\text{H}_2\text{O}$  is  $\#(2 \text{ mol H}_2)/(2 \text{ mol H}_2\text{O})\#$ .. Example: How many moles of  $\text{O}_2$  are required to form 5.00 moles of  $\text{H}_2\text{O}$ ?

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#### **PracticePacket((Unit6: Moles(&Stoichiometry**

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#### **Moles Answer Key - HelpTeaching.com**

Moles & Stoichiometry Answers Key Questions & Exercises 1. The atomic weight of carbon is 12.0107 u, so a mole of carbon has a mass of 12.0107 g. Why doesn't a mole of carbon weigh 12 g? The atomic weight refers to the weighted average of masses of the isotopes comprising a naturally occurring sample of carbon.

#### **Moles & Stoichiometry Answers Key Questions & Exercises**

W/ answers Website Upload Big Numbers and Chemistry At the most fundamental level, the chemist needs a unit that describes a very large quantity. One of the most well-known numbers in the study of chemistry is number of units in a mole. The number of

#### **Unit 6: Reactions and Stoichiometry**

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The mole ratio can be used to "convert" one substance in an equation to another using dimensional ... Key Questions 1. Without using numbers, explain the significance if the mole ratio term works out to 1/1 for any given ... POGIL 04 - Stoichiometry 4 - Reactants, Products, and the Mole Ratio -

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**POGIL 04 - Stoichiometry 4 - Reactants, Products, and the ...**

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This is called Stoichiometry- (stoi-key-ah-meh-tree) Another way of looking at it is using the mole ratio from the balanced equation and information about one compound in the reaction to determine information about another compound in the equation. A mole ratio is a conversion factor derived from the coefficients of a balanced chemical equation

**POGIL- Stoichiometry - Oakmont High School**

The molecular weight of O<sub>2</sub> is 32 Pogil the mole answer key. 00 u, so a mole of O<sub>2</sub> would have a mass of 32. 00 g and would contain 6. 022 x 10<sup>23</sup> O. 2 molecules. Each O<sub>2</sub> molecule is composed of two oxygen atoms, so one mole of O<sub>2</sub> contains two moles of oxygen atoms Pogil the mole answer key. 3. The mole is sometimes described as “the chemist’s dozen”.

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