

Modern Chemistry Gases Section Quiz Answers

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Modern Chemistry Gases Section Quiz

Modern Chemistry 72 Quiz Section Quiz: Gases and Pressure In the space provided, write the letter of the term or phrase that best completes each sentence or best answers each question. ____ 1. What causes a gas to exert pressure? a. collisions b. density c. temperature d. elevation ____ 2. The SI unit for pressure is a. newton. b. mm Hg. c. pascal. d. liter. ____ 3.

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GASES Section Quiz: The Kinetic-Molecular Theory of Matter In the space provided, write the letter of the term or phrase that best completes each statement or answers each question. 1. A hypothetical gas that perfectly fits all the assumptions of the kinetic- molecular theory is known as a. a real gas. b. an ideal gas. c. an imaginary gas.

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Modern Chemistry Section Quiz The Gas Laws Answers

Modern Chemistry 74 Quiz Section Quiz: The Gas Laws In the space provided, write the letter of the term or phrase that best completes each sentence or best answers each question. ____ 1. For a fixed amount of gas at a constant temperature, the volume increases as the pressure a. remains steady. b. increases. c. decreases. d. fluctuates. ____ 2.

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182 Study Guide for An Introduction to Chemistry Section Goals and Introductions Section 11.1 Gases and Their Properties Goals To describe the particle nature of both real and ideal gases. To describe the properties of gases that can be used to explain their characteristics: volume, number of particles, temperature, and pressure.

Chapter 11 - Gases - An Introduction to Chemistry

States of Matter SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. Identify whether the descriptions below describe an ideal gas or a real gas. ideal gas a. The gas will not condense because the molecules do not attract each other. ideal gas b. Collisions between molecules are perfectly elastic.

10 States of Matter - Ms. Agostine's Chemistry Page

The common unit of pressure, symbolized mm Hg. Pressure (P) Defined as the force per unit area on a surface. Newton (N) The force that will increase the speed of a one-kilogram mass... The amount of force exerted per unit area of a surface. The SI unit for force; the force that will increase the speed... An instrument...

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$KE = (1/2) (mass) (velocity)^2$ Kinetic Molecular Theory particles of matter are always in motion 5 Assumptions of KMT 1) gases consist of a large number of tiny particles spread ap... Pressure volume = moles constant * temperature (K)... The force per unit area exerted by a gas on its container wall... The space occupied by an object equal to...

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CHEMISTRY NOTES - Chapter 10 States of Matter Goals : To gain an understanding of : 1.. The kinetic theory of matter. 2. Avogadro's hypothesis. 3. The nature of liquids, solids and gases. 4. Phase changes. NOTES: The kinetic theory of matter states that particles which make up all types of matter are in constant motion.

CHEMISTRY NOTES - Chapter 10 States of Matter

a. solid, liquid, gas b. liquid, solid, gas c. gas, liquid, solid d. gas, solid, liquid 3. Which is an example of a noncrystalline solid? a. emerald b. quartz c. ice d. glass 4. Which of the following is not a physical property of crystalline solids? a. definite shape b. definite melting point c. low rate of diffusion d. high compressibility 5.

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