

## *Net Ionic Equation Yahoo Answers*

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### Net Ionic Equation Yahoo Answers

Best Answer: The presence (or lack) of a driving force for reactions done in water can be evaluated using a net ionic equation. The total and net ionic equations show the real physical state of every component of the reaction. The compound is a strong electrolyte (a soluble ionic compound or a strong acid), then the compound is separated into the appropriate aqueous ions.

### Net Ionic Equations?!? | Yahoo Answers

The problem is: A gas forms when a solution of NaOH is mixed with a solution containing  $\text{Ca}^{2+}$ ,  $\text{Ba}^{2+}$ ,  $\text{K}^{+}$ , and  $\text{NH}_4^{+}$  ions. The mixture is heated. Write the net ionic equation. My question is: Can you please show how to write the net ionic equation? What is the purpose of heating this solution and any solution? HELP PLEASE!

### Write net ionic equation for...? | Yahoo Answers

Write a balanced net ionic equation for the following reaction. Enter the sum of all of the coefficients for the equation in the first answer box. Then enter the letters corresponding to the spectator ions listed below in the second box. (ex. BD) Calcium hydroxide(aq) + acetic acid(aq)  $\rightarrow$  A)  $\text{Ca}^{2+}$  B)  $\text{OH}^{-}$  C)  $\text{CH}_3\text{COOH}$  D)  $\text{H}_2\text{O}$  E)  $\text{CH}_3\text{COO}^{-}$  F) NONE

### Writing Net Ionic Equations? | Yahoo Answers

Which of the following statements are true for the reaction  $\text{Ca}(\text{OH})_2(\text{aq}) + \text{HCl}(\text{aq}) \rightarrow$  a)  $\text{CaCl}(\text{aq})$  is a product b) When balanced the coefficient of  $\text{HCl}(\text{aq})$  is 3 c) Ions cannot appear in the net ionic equation because they always cancel out on both sides of the equation d)  $\text{Ca}(\text{OH})_2(\text{aq})$  appears in the net ionic equation e)  $\text{H}_2\text{O}(\text{l})$  appears in the net ionic equation

### Net Ionic equation? | Yahoo Answers

Write the balanced formula, complete ionic, and net ionic equations for the reactions that occur when the following are mixed (include state symbols where needed). a. potassium hydroxide (aqueous) and nitric acid b. barium hydroxide (aq) and hydrochloric acid c. perchloric acid (aq) and solid iron(III) hydroxide

### Balanced Formulas and Net ionic? | Yahoo Answers

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### Net Ionic Equation..? | Yahoo Answers

If you're trying to find the net ionic equation you're going to need to know the states of those compounds. All solids, liquids, and gases stay the same when calculating the net ionic equation, only aqueous solutions break down into their ions, stating the number and charge.

### Net Ionic Equation Help? | Yahoo Answers

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### Net ionic equation for KOH + SnCl2 ???!!!!? | Yahoo Answers

Write balanced molecular, ionic, and net ionic equations for the reactions of tin (gives  $\text{Sn}^{2+}$ ) with a solution of hydrochloric acid to give hydrogen plus the metal ion in solution. In each equation include physical states in your answer and use minimal integer numbers for coefficient input. CAN SOMEONE PLEASE HELP

### Write balanced molecular, ionic, and net ionic equations ...

Net Ionic Equation Worksheet Answers Write balanced molecular, ionic, and net ionic equations (NIE) for each of the following reactions. Assume all reactions occur in aqueous solution. 1.  $2\text{NaCl}(\text{aq}) + \text{Pb}(\text{NO}_3)_2(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{NaNO}_3(\text{aq})$

### Net Ionic Equation Worksheet Answers

A Net Ionic Equation is a chemical equation for a reaction which lists only those species participating in the reaction. To write a Net Ionic Reaction, follow these 3 steps: 1) Start by simply ...

#### How do you write a net ionic equation - answers.com

5. (Points: 2) Determine the net ionic equation when the following two aqueous solutions are mixed together. If no reaction occurs, select that choice below.

#### 5. (Points: 2) Determine the net ionic equation when the ...

Complete and balance the following equations. Determine what ionic compound is the precipitate. Give the spectator ions and write out the net ionic equation.

#### Questions | Yahoo Answers

Net Ionic Equation Worksheet READ THIS: When two solutions of ionic compounds are mixed, a solid may form. This type of reaction is called a precipitation reaction, and the solid produced in the reaction is known as the precipitate.

#### Net Ionic Equation Worksheet Answers - hudson.k12.oh.us

PRACTICE PROBLEMS ON NET IONIC EQUATIONS page 1 of 3 Show the complete ionic and net ionic forms of the following equations. If all species are spectator ions, please indicate that no reaction takes place. Note: you need to make sure the original equation is balanced before proceeding!

#### PRACTICE PROBLEMS ON NET IONIC EQUATIONS

If the solvent is water, then you definitely gets sodium hydroxide forming too. once you're reacting a sodium salt like NaOH with HCl, a extra achievable occasion to apply for ionic equations, the molecular equation would be:  $\text{NaOH(aq)} + \text{HCl(aq)} \rightarrow \text{H}_2\text{O(l)} + \text{NaCl(aq)}$  This element shows each and all of the molecules as though they weren't ...

#### Net ionic equation question? | Yahoo Answers

Im stuck on this question. I have to find the net ionic equation for this reaction. Could someone work it out so I understand it.  $\text{AgCl(s)} + 2\text{NH}_3\text{(aq)} \rightarrow [\text{Ag(NH}_3)_2]\text{Cl(aq)}$

#### Net Ionic Equation..? | Yahoo Answers

Best Answer: Lithium sulfide is a somewhat unstable salt which easily reacts with the moisture in the air to produce  $\text{H}_2\text{S}$  gas. When dissolved in water it produces lithium ions, hydroxide ions, hydrogen sulfide gas and hydrogen sulfide ions  $\text{Li}_2\text{S(s)} + 2\text{H}_2\text{O} \rightarrow 2\text{LiOH(aq)} + \text{H}_2\text{S(aq)}$  The net ionic equation for the ...

#### Can you try to figure out what the net ionic equation for ...

Yahoo Philippines Answers ... What is the net ionic equation? 1- of the reaction of  $\text{MgCl}_2$  with  $\text{NaOH}$  I got this answer which is incorrect ... Total Ionic Equation/Net Ionic Equation? Answer Questions. What is the  $\Delta S_{\text{univ}}$  at 25.0°C for a reaction with a  $\Delta H_{\text{sys}}$  of 250.0 kJ and a  $\Delta S_{\text{sys}}$  of -80.0 J/K? Input your answer in units of J/K?

#### What is the net ionic equation? | Yahoo Answers

full ionic equation:  $3\text{Ba}^{2+}(\text{aq}) + 6\text{OH}^{-}(\text{aq}) + 6\text{H}^{+}(\text{aq}) + 2\text{PO}_4^{3-}(\text{aq}) \rightarrow \text{Ba}_3(\text{PO}_4)_2(\text{s}) + 6\text{H}_2\text{O(l)}$   
Now, the rules of changing the full ionic equation into a net ionic equation is to cancel out anything that is common to both sides of the full ionic equation. There is nothing to cancel. Therefore: Net ionic equation:

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