

## ***Mole Problems Answer Key***

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**Mole Problems Answer Key**

KEY Part 1: Molar Mass Use the periodic table to find the molar masses of the following. HCl K<sub>2</sub>CO<sub>3</sub> Ca(OH)<sub>2</sub> Na<sub>3</sub>PO<sub>4</sub> Part 2: Mole Conversions Work each of the following problems. SHOW ALL WORK. 1. How many atoms are in 6.2 moles of aluminum? ... Microsoft Word - 7-06a,b Mole Problems Wkst-Key.doc Author: Brent White

**2 3 2 3 2 1 2H<sub>3</sub>O<sub>2</sub> 3 2 ? f.u. NaCl = 3.55 mol NaCl 6.02 H ...**

View Stoichiometry Mole-Mole Problems Answer Key.pdf from ENGLISH 1201 at Mishawaka High School. i | i | i | STOICHIOMETRY: : Nome twp MOLE-MOLE PROBLEMS 1. N<sub>2</sub> + 3H<sub>2</sub> → 2NH<sub>3</sub> How many moles

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**Worksheet: Mole Problems Name\_\_\_\_\_ CHEMISTRY ...**

KEY 7.2 moles NO<sub>2</sub> Answer each of the following questions using the equation provided. BE SURE TO BALANCE EACH EQUATION BEFORE SOLVING ANY PROBLEMS. SHOW ALL WORK . 1. \_\_\_\_NO + \_\_\_\_O<sub>2</sub> → \_\_\_\_NO<sub>2</sub> a. 2 moles of NO will react with \_\_\_\_ mole(s) of O<sub>2</sub> to produce \_\_\_\_ mole(s) of NO<sub>2</sub> b. ? moles NO<sub>2</sub> = 3.6 moles O<sub>2</sub> × moles NO<sub>2</sub> moles O<sub>2</sub> = c.

**CHEMISTRY: A Study of Matter - Weebly**

Practice Problems: Moles (Answer Key) How many moles are in the following: a. 1.29 x 10<sup>24</sup> hydrogen atoms in HF 2.14 moles H atoms b. 7.36 x 10<sup>24</sup> free oxygen atoms 12.2 moles O atoms c. 3.28 x 10<sup>23</sup> Na atoms in salt (NaCl) 0.545 moles Na atoms; How many atoms are present in the following?

**Practice Problems: Moles (Answer Key)**

Answer Key. Stoichiometry: Mole-Mole Problems. N<sub>2</sub> + 3H<sub>2</sub> → 2NH<sub>3</sub>. How many moles of hydrogen are needed to completely react with 2.0 moles of nitrogen? 6.0 moles of hydrogen . 2. 2KClO<sub>3</sub> → 2KCl + 3O<sub>2</sub>. How many moles of oxygen are produced by the decomposition of 6.0 moles of potassium chlorate? 9.0 moles of oxygen .

**Stoichiometry: Mole-Mole Problems**

MIXED MOLE PROBLEMS - KEY 1. a) How many grams are there in 1.5 x 10<sup>25</sup> molecules of CO<sub>2</sub>? 1.110g 1 mol 44.0 g 6.0210molecules 1.510molecules CO<sub>1</sub>mol 3 23 2 25!=!!!! b) What volume would the CO

**KEY - CP - Mixed Mole Problems - newburyparkhighschool.net**

Mass and the Mole- Answer Key 1) How many moles are in 15 grams of lithium? 0.46 moles 2) How many grams are in 2.4 moles of sulfur? 77.0 grams 3) How many moles are in 22 grams of argon? 0.55 moles 4) How many grams are in 88.1 moles of magnesium? 2141 grams 5) How many moles

are in 2.3 grams of phosphorus? 0.074 moles

### **Molar Mass Worksheet Answer Key - River Dell Regional ...**

[ANSWER KEY] Chemistry I: Worksheet on All Kinds of Mole Problems SHOW ALL WORK!!! 1) What is the mass of 134 L of O<sub>2</sub> gas at STP?  $134 \text{ L O}_2 \times \frac{1 \text{ mol O}_2}{22.4 \text{ L}} \times 32.0 \text{ g O}_2 = 191 \text{ g O}_2$  2) How many molecules are there in 1.00 g of water?

### **[ANSWER KEY] Worksheet on All Kinds of Mole Problems SHOW ...**

Answer each of the following questions using the equation provided. BE SURE TO BALANCE EACH EQUATION BEFORE SOLVING ANY PROBLEMS. SHOW ALL WORK. 1.  $\text{___ NO} + \text{___ O}_2 \rightarrow \text{___ NO}_2$  a. 2 moles of NO will react with  $\text{___}$  mole(s) of O<sub>2</sub> to produce  $\text{___}$  mole(s) of NO<sub>2</sub> b. ? moles NO<sub>2</sub> = 3.6 moles O<sub>2</sub> c. moles NO<sub>2</sub> = c.

### **Worksheet: Mole/Mole Problems Name**

The ratio set up from data in the problem will almost always be the one with an unknown in it. Key point: the two ratios have to be set up with equivalent things in the same relative place in each ratio. A bit confusing? I will elaborate on this below. After setting up the proportion, you will cross-multiply and divide to get the answer.

### **ChemTeam: Stoichiometry: Mole-Mole Examples**

MOLE-MOLE PROBLEMS I.  $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$  Name How many moles of hydrogen are needed to completely react with two moles of nitrogen?  $2.0 \text{ mol N}_2 \times \frac{3 \text{ mol H}_2}{1 \text{ mol N}_2} = 6.0 \text{ mol H}_2$  2. How many moles of oxygen are produced by the decomposition of six moles of potassium chlorate?  $6 \text{ mol KClO}_3 \times \frac{3 \text{ mol O}_2}{2 \text{ mol KClO}_3} = 9 \text{ mol O}_2$  3.  $\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$  How many moles of hydrogen are produced from the reaction of three moles

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