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Kinetic Theory - Solutions. 1) Estimate the fraction of molecular volume to the actual volume occupied by oxygen gas at STP. Take the diameter of an oxygen molecule to be 3 Å. Solution: Radius 'r' of an oxygen molecule,  $r = d/2 = (3/2) \text{ Å} = 1.5 \times 10^{-8} \text{ cm}$ . Volume occupied by 1 mole of oxygen gas at STP = 22.4 litres = 22400 cm<sup>3</sup>

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Chemistry Class 11 NCERT Solutions: Chapter 4 Chemical Bonding and Molecular Structure Part 1. The octet rule or the electronic theory of chemical bonding was developed by Kossel and Lewis. According to this rule, atoms can combine either by transfer of valence electrons from one atom to another or by sharing their valence electrons in order to attain the nearest noble gas configuration by having an octet in their valence shell.

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The law is a simple consequence of kinetic theory.] Answer Rate of diffusion of hydrogen,  $R_1 = 28.7 \text{ cm}^3 \text{ s}^{-1}$  Rate of diffusion of another gas,  $R_2 = 7.2 \text{ cm}^3 \text{ s}^{-1}$  According to Graham's Law of diffusion, we have: 32 g is the molecular mass of oxygen. Hence, the unknown gas is oxygen. 13.13.

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