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Modern Chemistry 2 Stoichiometry CHAPTER 9 REVIEW Stoichiometry MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation: C ... Modern Chemistry 3 Stoichiometry MIXED REVIEW continued c. If 0.1 mol of N₂ combine with H₂, what must be true about the quantity of H

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CHAPTER 9 REVIEW Stoichiometry SECTION 9-3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% If the actual yield of a reaction is 22 g and the theoretical yield is 25 g, calculate the percent yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂ according to the following equation: N₂(g) + 3H₂(g) → 2NH₃(g) N₂; 2.0 mol a.

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Modern Chemistry • CHAPTER 9 HOMEWORK 9-1 (pp. 275–279) VOCABULARY Define. ... Modern Chemistry • CHAPTER 9 HOMEWORK 9-2 (pp. 280–282) VOCABULARY Complete each sentence. 1. The mole ratio is obtained from _____. 2. Theoretical stoichiometric calculations show _____. ...

HOMEWORK 9-1 - Los Banos Unified School District

CHAPTER 1 REVIEW Matter and Change MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a homogeneous or heterogeneous substance. a. sugar d. plastic wrap b. iron filings e. cement sidewalk c. granola bar 2. For each type of investigation, select the most appropriate branch of chemistry from the following

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14 Ions in Aqueous Solutions and Colligative Properties

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Let's try a real chemistry problem and use this concept to determine the limiting reagent and see how it determines how much product will be formed. Problem : Nitrogen gas reacts with hydrogen gas to produce gaseous ammonia (NH_3). If we react 25.0 g of nitrogen with 25.0 g of hydrogen, how much ammonia will be produced.

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Modern Chemistry 71 Chapter Test Name Class Date Chapter Test B, continued 17. A substance combines with oxygen, releasing a large amount of energy as heat and light, in a(n) . 18. The decomposition of a substance by an electric current is called . 19. A(n) orders the elements by the ease with which they undergo certain chemical reactions. 20.

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