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3. One major difference between the Bronsted-Lowry and the Arrhenius definitions of acids and bases is which of the following? In the Arrhenius definition, acids release protons while in the Bronsted-Lowry definition, they donate protons. Arrhenius said that bases donate electron pairs whereas ...

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your answer.CHAPTER 14 REVIEW Acids and Bases - Weebly Modern Chemistry 121 Acids and Bases CHAPTER 14 REVIEW Acids and Bases SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Answer the following questions according to the Brønsted-Lowry definitions of acids and bases: ____ a.

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3 is the Lewis acid because it accepts an electron pair in forming a covalent bond. It is not a Brønsted-Lowry acid, because it is not donating a proton. e. Which of the reactants is the Lewis base? Explain your answer. NF 3 is the Lewis base because it donates an electron pair in forming a covalent bond. 120 ACIDS AND BASES MODERN CHEMISTRY

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Modern Chemistry 121 Acids and Bases CHAPTER 14 REVIEW Acids and Bases SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Answer the following questions according to the Brønsted-Lowry definitions of acids and bases: ____ a. What is the conjugate base of H₂SO₃? ____ b.

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Modern Chemistry 129 Chapter Test Chapter: Acids and Bases PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

____ 1. Which of the following is an oxyacid? a. HCl b. H_2O c. H_2S d. H_2SO_4 ____ 2. Which acid is produced in the stomach? a. HNO_3 b. CH_3COOH c ...

Assessment Chapter Test B - clarkchargers.org

Modern Chemistry 136 Chapter Test Name Class Date Chapter Test A, continued ____ 14. A pH meter measures the pH of a solution by measuring the ... a. the color of the acid-base indicator should stay the same. b. the pH of the solution should change abruptly. c. the color of the acid-base indicator should change. d. Both (b) and (c)

Assessment Chapter Test A - Baumapedia

Arrhenius acid. b. Arrhenius base. c. Lewis acid. d. Lewis base. ____ 10. An Arrhenius acid in an aqueous solution a. attracts negatively charged anions. b. attracts positively charged cations. c. gives up one or more of its hydrogen ions to water molecules. d. forms ionic bonds with water molecules. ____ 11.

Assessment Chapter Test A - clarkchargers.org

CHAPTER 1 REVIEW Matter and Change MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a homogeneous or heterogeneous substance. a. sugar d. plastic wrap b. iron filings e. cement sidewalk c. granola bar 2. For each type of investigation, select the most appropriate branch of chemistry from the following

mc06se cFMsr i-vi

Modern Chemistry 127 Acid-Base Titration and pH CHAPTER 15 REVIEW Acid-Base Titration and pH SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. Below is a pH curve from an acid-base titration. On it are labeled three points: X, Y, and Z.

CHAPTER 15 REVIEW Acid-Base Titration and pH

Major topics: Arrhenius vs. Bronsted-Lowry definition of acids and bases, conjugate acid/base, acid dissociation constant (K_a), & strong vs weak acids

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