Now the Gibbs energy differences between the beginning of each section and the highest activated state are considered.

In the figure, these are

$$\Delta^{\ddagger}G^{\circ}_{A} = G^{\circ}_{\ddagger_2} - G^{\circ}_{A}$$

$$\Delta^{\ddagger}G^{\circ}_{B} = G^{\circ}_{\ddagger_4} - G^{\circ}_{X_2}$$

$$\Delta^{\ddagger}G^{\circ}_{C} = G^{\circ}_{\ddagger_6} - G^{\circ}_{X_5}$$

See the figure

The figure shows that $\Delta^{\ddagger}G^{\circ}_{B}$ is the largest one. Therefore section B will control the overall reaction rate. However, the section B is not $X_2 \rightleftharpoons X_5$, it is originally,

$$X_2 \rightleftharpoons X_3 \rightleftharpoons X_4 \rightleftharpoons X_5 .$$

Among three steps. We have to find out which one is the rate determining step. We know that the rate controlled by the highest activated complex in the sequence. In the section B, it is $\ddagger_4$, therefore the rate determining step that controls the overall rate is $X_3 \longrightarrow X_4$