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At the completion of this episode's lesson(s), you should be able to: • Distinguish between saturated, unsaturated, and supersaturated solutions and between dilute and concentrated solutions. • Identify factors that affect solubility.

Chemistry 1002: Solubility | Georgia Public Broadcasting

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When you can answer "yes" to a question, ... Likewise with chemistry, ... • solutions conduct electricity • normally crystalline solids at room temperature.

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7. How many liters of solution can be produced from 2.5 moles of solute if a 2.0 M solution is needed? $2.0\text{ M} = \frac{2.5\text{ moles}}{\text{liters of solution}}$ liters of solution = $1.25\text{ L} = 1.3\text{ L}$ 8. What would be the concentration of a solution formed when 1.00 g of NaCl are dissolved in water to make 100.0 mL of solution? $? \text{ mol} = \frac{1.00\text{ g NaCl}}{58.5\text{ g/mol}}$ $\times 1\text{ mol NaCl}$ 58.5 g ...

Molarity: Molarity = 1. 2. - Central Bucks School District

Chemistry & Physics. Chemistry 201: Physical and Chemical Properties and Changes. Season 1 Episode 201 | 27m Matter: Physical and Chemical Properties and Changes Physical and chemical properties, and physical and chemical changes are defined, and students learn to identify each.

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