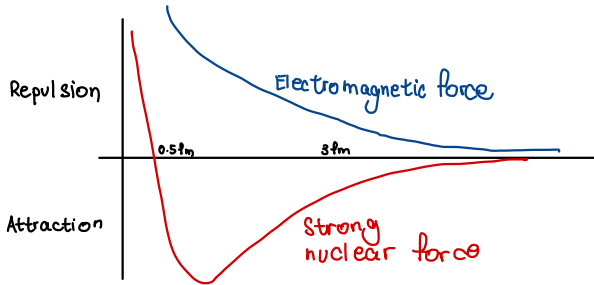


# Stable and Unstable Nuclei

## Forces in the nucleus:

- **Electromagnetic force** causes protons in the nucleus to repel
- **Gravitational force** causes nucleons (protons & neutrons) to attract each other
- **Strong nuclear force** to combat the electromagnetic force to hold the nucleus together.



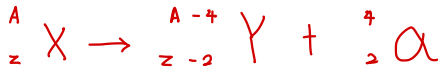
- The nuclear force is repulsive up to 0.5 fm to prevent it crushing the nucleus
- Past 0.5 fm, the strong nuclear force becomes attractive

## What makes the nucleus unstable?

- The range of the strong nuclear force is only a few femtometres. It struggles to hold together very large nuclei, which makes them unstable.
- Unstable nuclei will emit particles to become more stable - known as nuclear decay.

## Alpha decay

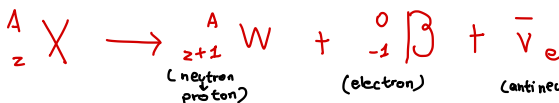
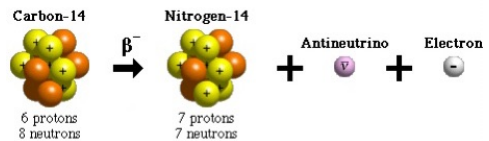
- Happens in very big atoms - as the nuclei of these atoms are too big for the strong nuclear force to keep them stable
- An alpha particle is emitted:  ${}^4_2\alpha$  (contains 2 protons and 2 neutrons - helium nucleus)



(element negatively charged?)

## Beta-minus decay ( $\beta^-$ )

- Happens in isotopes that are "neutron rich"
- An electron & an antineutrino particle are emitted
- **One neutron is changed into a proton**
- This leaves the element positively charged  $\downarrow$  1 less electron



this is needed due to lepton number conservation

## Positron emission / Beta-plus decay ( $\beta^+$ )

- Takes place in an unstable nucleus with too many protons - "proton-rich".
- A proton changes into a neutron.
- The positron ( $\beta^+$ ) is the antiparticle of the electron, so it carries a positive charge.



- PET scanning involves a positron-emitting isotope.

## Gamma decay/radiation

- It is electromagnetic radiation emitted by an unstable nucleus.
- It is emitted by a nucleus with too much energy.
- Gamma radiation has no mass and no charge. It can pass through thick metal plates.