

Nuclear Physics

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1 Introduction

The fundamental constituents of matter was always a question for the great minds of each era. From the greek philosophers that tried to explain the matter using four fundamental ones, to the recent years that the standard model tries to explain what are the interactions between the fundamental particles that form up the world around us.

The physicists have found smaller and smaller particles that make up the matter around us. The atom, the substructure of the atom, electrons, nuclei, The substructure of nuclei, protons and neutrons, and the substructure of protons and neutrons which are quarks.

The two fundamental types of building blocks are the *leptons*, which include the electron and the neutrino, and the quarks. In the scattering experiments, these were found to be smaller than 10^{-18}m . They are possibly point-like particles. For comparison, protons are as large as 10^{-15}m .

By the end of 19th-century there were two known forces in the world. Gravity and electromagnetism, the first described all the observed motions in the night sky and the macroscopic objects around, while the latter explained the interactions causing chemical reactions. With the development of nuclear physics two new forces were introduced. The reason behind the absence of these forces was the short-range nature of them.

Thus the four fundamental interactions on which all physical interactions are based upon are:

- Gravitation
- Electromagnetic interaction
- Weak interaction
- Strong interaction

According to today's conceptions, we tend to assign a mediator for each interaction. Thus except from gravitation which is still a mystery in fundamental physics we have:

- Weak interactions mediators are **W-Boson** and **Z-Boson**.
- Strong interactions mediator is **Gluon**
- Electromagnetic interactions mediator is the photon.

The question to be raised is that why two electrons share electromagnetic interactions while a photon and an electron would not. This would raise the assumption that electrons should have a property that is only to the particles, which can share such interactions. Thus we would associate a property named charge for any interactions. For electromagnetic interactions an *electric* charge, for weak and strong we would also have weak and strong charges. The latter is also called colour charge or colour for short.

Callout — A particle is subjected to an interaction if and only if it carries the corresponding charge!

2 The Building Blocks of Matter

2.1 Atomic Theory and The Constituents of Matter

People have questioned the very nature of matter for a long time, whether it is possible to divide the matter infinite times and still be able to do so, or if it is made of a grain-like particles which cannot be divided into substructures. The question was an open one, though philosophers like Democritus or physicists like Boltzmann believed in an atomic (atomos means uncuttable in greek.) theory along the way.

The Philosophy: The idea was simple, matter is made up of discrete units that cannot be more divided into substructures. Democritus came out to be seen as the first person, who believed in such idea. But it was well

forgotten until the newly found works of Aristotle in 12th century which was exactly the opposite. That the matter was continuous and infinite and could be subdivided without limit.

Later on the discovery of major works describing atomist teachings gave more attention to the subject, but since the Epicureanism contradicted orthodox Christian teachings, the teachings were considered unacceptable by most philosophers. The French Catholic priest Pierre Gassendi revived the theory by some modification, that the God created the atoms.

John Dalton: At the end of 18th century there were two laws arising in chemistry, which were silently pointing towards the atomic nature of the matter.

1. The first one was the conservation of mass, which states that the total mass of the reactants must be equal to the total mass of the products.
2. Second one was the law of definite proportions. This law states that if a compound is broken down into its constituent chemical elements, then the masses of the constituents will always have the same proportions by weight, regardless of the quantity or source of the original substance.

Dalton proposed that each chemical element is composed of atoms of a single, unique type, and though they cannot be altered or destroyed by chemical means, they can combine to form more complex structures. This marked the first truly scientific theory of the atom, since Dalton reached his conclusions by experimentation and examination of the results in an empirical fashion.

Brownian Motion: In 1827, the British botanist Robert Brown observed that dust particles inside pollen grains floating in water constantly jiggled about for no apparent reason. In 1905, Albert Einstein theorized that this Brownian motion was caused by the water molecules continuously knocking the grains about, and developed a hypothetical mathematical model to describe it. The model was validated experimentally in 1908 by French physicist Jean Perrin, thus providing additional validation for particle theory (and by extension atomic theory).

Statistical Mechanics: In 1738, Daniel Bernoulli postulated that the pressure of gases and heat were both caused by the underlying motion of molecules. Nonetheless for the theory of ideal gases it was also necessary to postulate the existence of atoms.

James Clerk Maxwell used statistical mechanics in physics, a notion which Boltzmann and Rudolf Clausius used to expand and work on thermodynamics, and especially the law of entropy (2nd law of thermodynamics). All of statistical mechanics and the laws of heat, gas, and entropy were necessarily postulated upon the existence of atoms.

2.2 Subatomic Particles

Electron: After the discovery of electric force and the existence of two types of charges, in 1859 Julius Plücker observed the radiation emitted from the cathode caused phosphorescent light. The first building block of the atom to be identified was the electron. Later on Arthur Schuster expanded upon the experiments by placing metal plates parallel to the cathode rays and applying an electric potential between the plates. The field deflected the rays toward the positively charged plate, providing further evidence that the rays carried negative charge. By measuring the amount of deflection for a given level of current, in 1890 Schuster was able to estimate the charge-to-mass ratio of the ray components.

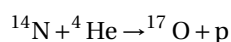
While studying naturally fluorescing minerals in 1896, the French physicist Henri Becquerel discovered that they emitted radiation without any exposure to an external energy source. These radioactive materials became the subject of much interest. J. J. Thomson performed experiments that indicated the particles were unique, rather than waves, atoms, or molecules as was believed earlier. The electron's charge was more carefully measured by the American Physicist Robert Millikan.

In the Standard Model of particle physics, electrons belong to the group of subatomic particles called leptons, which are believed to be fundamental or elementary particles with no substructure. Electrons have the lowest mass of any charged lepton and belong to the first-generation of fundamental particles. Leptons differ from the other basic constituent of matter, the quarks, by their lack of strong interaction. All members of the lepton group are fermions, because they all have half-odd integer spin; the electron has spin $\frac{1}{2}$.

Nuclei: In the Atom model of Thomson, the electrons, and an equivalent number of positively charged particles were uniformly distributed throughout the atom. The resulting atom was electrically neutral. Later on the Geiger-Marsden Experiment with Rutherford showed the image to be wrong. By shooting alpha particles to a thin golden foil. The explanation of the scattering data was a central Coulomb field caused by a massive, positively charged nucleus. Thus the atomic model had another step towards modeling reality more precisely.

Proton: Since the mass of a proton can be approximately of a hydrogen atom (the electron contributes really small to the total mass.) the concept of a hydrogen-like particle as a constituent of other atoms was developed over time. Because there were some indications such as the periodic table, the ratio of mass between different atoms, etc...

In 1927, Rutherford proved that the hydrogen nucleus is present in other nuclei, While bombarding nitrogen with α -particles, he observed positively charged particles with an unusually long range, which must have been ejected from the atom as well. From this he concluded that the nitrogen atom had been destroyed in these reactions and a light constituent of the nucleus had been ejected. He had already discovered similar long-ranged particles when bombarding hydrogen. From this he concluded that these particles were hydrogen nuclei which, therefore, had to be constituents of nitrogen as well:



Protons are spin $\frac{1}{2}$ fermions and are not elementary particles, which mean they have substructures, quarks that make up the proton are held together with strong interaction.

Neutron: The Neutron was also detected by bombarding nuclei with α -particles. Chadwick in 1932 found an appropriate experimental approach. He used the irradiation of beryllium with α -particles from a polonium source, and thereby established the neutron as a fundamental constituent of nuclei.

2.3 Nuclides

The Atomic and Mass Number: Knowing that the nuclei consists of positively charged particles (protons) the atomic number is the number of protons the nuclei has. Denoted by Z , the charge of the nucleus therefore is $Q = Ze$, the elementary charge $e \approx 1.6 \cdot 10^{-19}$, for neutral atoms this is also the number of electrons orbiting around the nucleus.

Callout — The Chemical properties of an atom is determined by its atomic number.

Callout — The classical method of determining the charge of the nucleus is the measurement of the characteristic X -rays of the atom to be studied. Moseley's law states that the energy of the K_α -line is proportional to $(Z - 1)^2$. Nowadays the detection of these characteristic X -rays are being used in material analysis.

In addition to the atomic number we define N as the number of neutrons, then the mass number is defined as:

$$A = Z + N \quad (1)$$

- Nuclides with the same mass number A are called isobars.
- Nuclides with the same atomic number Z are called isotopes.
- Nuclides with the same neutron number N are called isotones.

We define the binding energy as:

$$B(Z, A) = [ZM(^1\text{H}) + (A - Z)M_n - M(A, Z)] \cdot c^2 \quad (2)$$

where

$$\begin{aligned}
 M(^1H) &= M_p + m_e \\
 M_p &= 938.272 \text{ MeV}/c^2 = 1836.153 m_e \\
 M_n &= 939.565 \text{ MeV}/c^2 = 1838.684 m_e \\
 m_e &= 0.511 \text{ MeV}/c^2
 \end{aligned}$$

2.4 Parametrisation of Binding Energies

Semi-Empirical Mass Formula (SEMF): Apart from the lightest elements, the binding energy per nucleon for most nuclei is about 8–9 MeV. The parametrisation of nuclear masses as a function of A and Z , which is known as the *Weizsäcker formula*, or the semi empirical mass formula, is used to approximate the mass and various other properties of an atomic nucleus from its number of protons and neutrons. As the name suggests, it is based partly on theory and partly on empirical measurements. The formula represents the **liquid-drop model** proposed by George Gamow.

Liquid-Drop Model: The liquid-drop model was first proposed by George Gamow and further developed by Niels Bohr and John Archibald Wheeler. It treats the nucleus as a drop of incompressible fluid of very high density, held together by the nuclear force, there is similarity to the structure of a spherical liquid drop. While a crude model, the liquid-drop model accounts for the spherical shape of most nuclei and makes a rough prediction of binding energy.

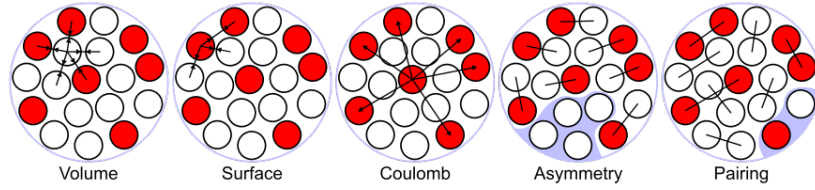


fig: 1 – Illustration of the terms of the semi-empirical mass formula in the liquid-drop model of the atomic nucleus.

1. **Volume Energy**, when an assembly of nucleons of the same size is packed together into the smallest volume, each interior nucleon has a certain number of other nucleons in contact with it. So, this nuclear energy is proportional to the volume.
2. **Surface Energy** corrects for the previous assumption made that every nucleon interacts with the same number of other nucleons. This term is negative and proportional to the surface area, and is therefore roughly equivalent to liquid surface tension.
3. **Coulomb Energy** the potential energy from each pair of protons. As this is a repulsive force, the binding energy is reduced.
4. **Asymmetry energy** (also called Pauli energy), which accounts for the Pauli exclusion principle. Unequal numbers of neutrons and protons imply filling higher energy levels for one type of particle, while leaving lower energy levels vacant for the other type.
5. **Pairing energy**, which accounts for the tendency of proton pairs and neutron pairs to occur. An even number of particles is more stable than an odd number due to spin coupling.

The Formula: Nuclear fusion and fission and other nuclear reactions, or measuring the mass of nucleus and comparing it with the mass of protons and neutrons it contains reveals that the mass of the nucleus as a whole is smaller than the sum of its constituents, This is because the binding energies between the nucleons are yet to be considered. From this we would write:

$$M = ZM_p + NM_n - \frac{E_B(N, Z)}{c^2} \quad (3)$$

The semi-empirical mass formula states the binding energy as:

$$E_B = a_V A - a_S A^{2/3} - a_C \frac{Z(Z-1)}{A^{1/3}} - a_A \frac{(N-Z)^2}{4A} + \delta(N, Z) \quad (4)$$

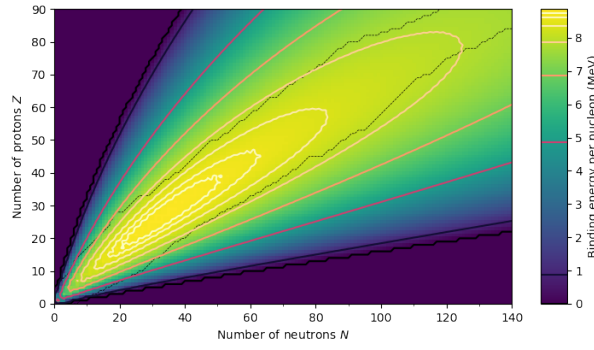


fig: 2 – The binding energy per nucleon (in MeV) shown as a function of the neutron number N and atomic number Z as given by the semi-empirical mass formula. A dashed line is included to show nuclides that have been discovered by experiment.

Volume Term: The term $a_V A$ is known as the volume term. The volume of the nucleus is proportional to A , so this term is proportional to the volume, hence the name.

The basis for this term is the strong nuclear force. The strong force affects both protons and neutrons, and as expected, this term is independent of Z . Because the number of pairs that can be taken from A particles is $A(A-1)/2$, one might expect a term proportional to A^2 . However, the strong force has a very limited range, and a given nucleon may only interact strongly with its nearest neighbors and next nearest neighbors. Therefore, the number of pairs of particles that actually interact is roughly proportional to A , giving the volume term its form.

Surface Term: The term $a_S A^{2/3}$ is known as the surface term. This term, also based on the strong force, is a correction to the volume term.

The volume term suggests that each nucleon interacts with a constant number of nucleons, independent of A . While this is very nearly true for nucleons deep within the nucleus, those nucleons on the surface of the nucleus have fewer nearest neighbors, justifying this correction. This can also be thought of as a surface-tension term, and indeed a similar mechanism creates surface tension in liquids.

If the volume of the nucleus is proportional to A , then the radius should be proportional to $A^{1/3}$ and the surface area to $A^{2/3}$. This explains why the surface term is proportional to $A^{2/3}$. It can also be deduced that a_S should have a similar order of magnitude to a_V .

Coulomb Term: The basis for this term is the electrostatic repulsion between the protons. The potential energy of sphere of uniform charge density is:

$$E = \frac{3}{5} \frac{1}{4\pi\epsilon_0} \frac{Q^2}{R}$$

Using an empirical nuclear radius of $R \approx r_0 A^{1/3}$ and $Q = Ze$ since the repulsion will only exist for more than one proton, Z^2 becomes $Z(Z-1)$:

$$E \approx \frac{3e^2 Z(Z-1)}{20\pi\epsilon_0 r_0 A^{1/3}} = a_C \frac{Z(Z-1)}{A^{1/3}} \quad (5)$$

Thus the term $a_C \frac{Z(Z-1)}{A^{1/3}}$ is known as the Coulomb or Electrostatic term.

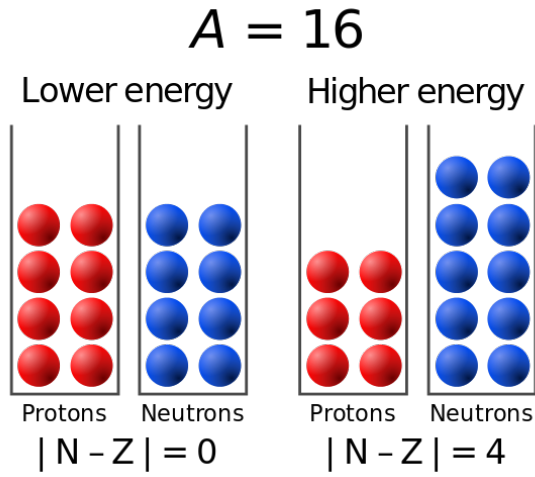


fig: 3 – Asymmetry Term

could move some particles from the neutron pool to the proton pool, in other words, change some neutrons into protons, we would significantly decrease the energy. The imbalance between the number of protons and neutrons causes the energy to be higher than it needs to be, for a given number of nucleons. This is the basis for the asymmetry term.

Pairing Term: The last term $\delta(A, Z)$ is known as the pairing term. This term captures the effect of spin coupling. It is given by:

$$\delta(A, Z) = \begin{cases} +\delta_0 & \text{for even } Z, N \\ 0 & \text{for odd } A \\ -\delta_0 & \text{for odd } Z, N \end{cases} \quad (6)$$

Constants The constants of this section are:

$$\begin{aligned} a_V &= 15.67 \text{ MeV}/c^2 \\ a_S &= 17.23 \text{ MeV}/c^2 \\ a_C &= 0.714 \text{ MeV}/c^2 \\ a_A &= 93.15 \text{ MeV}/c^2 \\ \delta_0 &= -11.2 \text{ MeV}/c^2 \end{aligned}$$

2.5 Isospin and The Charge Independence of the Nuclear Force

Protons and Neutrons not only have nearly equal masses, they also have similar nuclear interactions. This is particularly visible in the study of mirror nuclei.

Isospin: Isospin was introduced as a concept in 1932, well before the 1960s development of the quark model. The man who introduced it, Werner Heisenberg, did so to explain symmetries of the then newly discovered neutron (symbol n):

- The mass of the neutron and the proton (symbol p) are almost identical: they are nearly degenerate, and both are thus often called nucleons. Although the proton has a positive electric charge, and the neutron is neutral, they are almost identical in all other aspects.
- The mass of the neutron and the proton (symbol p) are almost identical: they are nearly degenerate, and both are thus often called nucleons. Although the proton has a positive electric charge, and the neutron is neutral, they are almost identical in all other aspects.

This behavior is not unlike the electron, where there are two possible states based on their spin. Other properties of the particle are conserved in this case. Heisenberg introduced the concept of another conserved

Asymmetry Term: The term $a_A \frac{(N-Z)^2}{A}$ is known as the asymmetry term (or Pauli term). The Theoretical justification for this term is more complex. The pauli exclusion principle states that no two identical fermions can occupy exactly the same quantum state in an atom. At a given energy level, there are only finitely many quantum states available for particles. What this means in the nucleus is that as more particles are added, these particles must occupy higher energy levels, increasing the total energy of the nucleus. Note that this effect is not based on any of the fundamental forces, only the Pauli exclusion principle.

Protons and neutrons, being distinct types of particles, occupy different quantum states. One can think of two different "pools" of states – one for protons and one for neutrons. Now, for example, if there are significantly more neutrons than protons in a nucleus, some of the neutrons will be higher in energy than the available states in the proton pool. If we

quantity that would cause the proton to turn into a neutron and vice versa. In 1937, Eugene Wigner introduced the term "isospin" to indicate how the new quantity is similar to spin in behavior, but otherwise unrelated. Protons and neutrons were then grouped together as nucleons because they both have nearly the same mass and interact in nearly the same way, if the (much weaker) electromagnetic interaction is neglected. In particle physics, the near mass-degeneracy of the neutron and proton points to an approximate symmetry of the Hamiltonian describing the strong interactions. It was thus convenient to treat them as being different states of the same particle.

Although the neutron does in fact have a slightly higher mass due to isospin breaking (this is now understood to be due to the difference in the masses of the up and down quarks and the effects of the electromagnetic interaction), the appearance of an approximate symmetry is useful even if it does not exactly hold; the small symmetry breakings can be described by a perturbation theory, which gives rise to slight differences between the near-degenerate states.

Callout — Isospin is similar to, but should not be confused with weak isospin. Briefly, weak isospin is the gauge symmetry of the weak interaction which connects quark and lepton doublets of left-handed particles in all generations; for example, up and down quarks, top and bottom quarks, electrons and electron neutrinos. By contrast (strong) isospin connects only up and down quarks, acts on both chiralities (left and right) and is a global (not a gauge) symmetry.

The symmetry between protons and neutrons may therefore be described by this formalism. The proton and Neutron are treated as two states of the nucleon which form a doublet:

$$\text{Nucleon } I: 1/2 \quad \begin{cases} \text{proton:} & I_3 = +1/2 \\ \text{neutron:} & I_3 = -1/2 \end{cases} \quad (7)$$

Formally, isospin is treated as a quantum mechanical angular momentum:

$$I_3^{\text{nucleus}} = \sum I_3^{\text{nucleons}} = \frac{Z - N}{2} \quad (8)$$

3 Nuclear Stability

There are only a small number of nuclei, which are stable, the other are unstable and would decay in various ways. As the number of protons increases there would be more electric repulsion between the protons that would lower the binding energy in total. The Iron and Nickel isotopes possess the maximum binding energy per nucleon, thus they are the most stable nuclides. For still heavier masses nuclei become unstable to fission and decay spontaneously into two or more lighter nuclei. Because of released energy in the process the mass of the original atom must be larger than the sum of the masses of the daughter atoms. For a two body decay thus, the inequality below must apply:

$$M(A, Z) > M(A - A', Z - Z') + M(A', Z') \quad (9)$$

- An isotope is said to be stable if its lifetime is considerably larger than the age of the solar system.
- It is very often the case that one of the daughter nuclei is a ${}^4\text{He}$ nucleus. This decay mode is called α -decay and the nucleus of Hydrogen is called α -particle.
- If a heavy nucleus decays into two similarly massive daughter nuclei we speak of spontaneous fission. The probability of spontaneous fission exceeds that of α -decay only for nuclei with $Z \leq 110$ and is a fairly unimportant process for the naturally occurring heavy elements.

Radioactive Decay: nuclear decay is the process in which an (unstable) nucleus loses energy by radiating particles, or light. The decaying nucleus is called the parent radionuclide, and the process produces at least one daughter nuclide, unless it decays light. There are six types for decay:

1. **Alpha Decay** occurs when the nucleus ejects an alpha particle (helium nucleus).
2. **Beta Decay** occurs in two ways;
 - beta-minus decay, when the nucleus emits an electron and an antineutrino in a process that changes a neutron to a proton.

		Group																																	
		I	II																	III	IV	V	VI	VII	VIII										
Period	1	1 H																								2 He									
	2	3 Li	4 Be																	5 B	6 C	7 N	8 O	9 F	10 Ne										
	3	11 Na	12 Mg																	13 Al	14 Si	15 P	16 S	17 Cl	18 Ar										
	4	19 K	20 Ca																	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr
	5	37 Rb	38 Sr																	39 Y	40 Zr	41 Nb	42 Mo	43 Tc	44 Ru	45 Rh	46 Pd	47 Ag	48 Cd	49 In	50 Sn	51 Sb	52 Te	53 I	54 Xe
	6	55 Cs	56 Ba	57 La	58 Ce	59 Pr	60 Nd	61 Pm	62 Sm	63 Eu	64 Gd	65 Tb	66 Dy	67 Ho	68 Er	69 Tm	70 Yb	71 Lu	72 Hf	73 Ta	74 W	75 Re	76 Os	77 Ir	78 Pt	79 Au	80 Hg	81 Tl	82 Pb	83 Bi	84 Po	85 At	86 Rn		
	7	87 Fr	88 Ra	89 Ac	90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr	104 Rf	105 Db	106 Sg	107 Bh	108 Hs	109 Mt	110 Ds	111 Rg	112 Cn	113 Nh	114 Fl	115 Uup	116 Lv	117 Uus	118 Uuo		

fig: 4 – Radioactivity is characteristic of elements with large atomic numbers. Elements with at least one stable isotope are shown in light blue. Green shows elements of which the most stable isotope has a half-life measured in millions of years. Yellow and orange are progressively less stable, with half-lives in thousands or hundreds of years, down toward one day. Red and purple show highly and extremely radioactive elements where the most stable isotopes exhibit half-lives measured on the order of one day and much less.

- beta-plus decay, when the nucleus emits a positron and a neutrino in a process that changes a proton to a neutron, also known as positron emission.
3. In **Gamma Decay** a radioactive nucleus first decays by the emission of an alpha or beta particle. The daughter nucleus that results is usually left in an excited state and it can decay to a lower energy state by emitting a gamma ray photon.
 4. In **Neutron Emission**, extremely neutron-rich nuclei, formed due to other types of decay or after many successive neutron captures, occasionally lose energy by way of neutron emission, resulting in a change from one isotope to another of the same element.
 5. In **Electron Capture**, the nucleus may capture an orbiting electron, causing a proton to convert into a neutron. A neutrino and a gamma ray are subsequently emitted.
 6. In **Cluster Decay** and nuclear fission, a nucleus heavier than an alpha particle is emitted.

Decay Constant: The radioactive decay law states that the probability per unit time that a nucleus will decay is constant, independent of time. This constant is called the decay constant and is denoted by λ .

The rate of nuclear decay is also measured in terms of half-lives. The half-life is the amount of time that a given nuclei would lose half of its radioactivity.

The decay constant is related to lifetime and half-life by:

$$\tau = \frac{1}{\lambda} \quad \text{and} \quad t_{1/2} = \frac{\ln 2}{\lambda} \quad (10)$$

The measurement of the decay constants of radioactive nuclei is based upon finding the activity (the number of decays per unit time)

$$A = -\frac{dN}{dt} = \lambda N \quad (11)$$

where N is the number of radioactive nuclei in the sample. The unit of activity is defined to become:

$$1\text{Bq [Becquerel]} = 1\text{decay/s} \quad (12)$$

For short-lived nuclides, the fall-off over time of the activity

$$A(t) = \lambda N(t) = \lambda N_0 e^{-\lambda t} \quad (13)$$

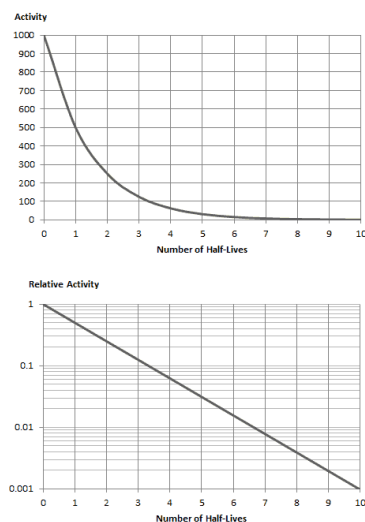


fig: 5 – Decay constant determines the rate of decay. Decay constant is denoted by λ , “lambda”. This constant probability may vary greatly between different types of nuclei, leading to the many different observed decay rates.

may be measured using fast electronic counters. This method of measuring is not suitable for lifetimes larger than about a year. For longer-lived nuclei both the number of nuclei in the sample and the activity must be measured in order to obtain the decay constant from eq[11]

3.1 Beta Decay

A Beta decay is a type of radioactive decay, on which one of the nucleons (proton or neutrons) transforms into the other by decaying a beta particle (positron or electron respectively) and a quasi-massless particle (neutrino or antineutrino respectively). The two types of beta decay is known as beta minus (for electron decay) and beta plus (for positron decay).

Description: Beta decay is a consequence of the weak force, which is characterized by relatively lengthy decay times. Nucleons are composed of up quarks and down quarks, and the weak force allows a quark to change its flavour by emission of a W boson leading to creation of an electron/antineutrino or positron/neutrino pair. For example, a neutron, composed of two down quarks and an up quark, decays to a proton composed of a down quark and two up quarks.

Electron capture is sometimes included as a type of beta decay, because the basic nuclear process, mediated by the weak force, is the same. In electron capture, an inner atomic electron is captured by a proton in the nucleus, transforming it into a neutron, and an electron neutrino is released.

Beta decay conserves a quantum number known as the lepton number, or the number of electrons and their associated neutrinos (other leptons are the muon and tau particles). These particles have lepton number +1, while their antiparticles have lepton number -1. Since a proton or neutron has lepton number zero, β^+ decay (a positron, or antielectron) must be accompanied with an electron neutrino, while β^- decay (an electron) must be accompanied by an electron antineutrino.

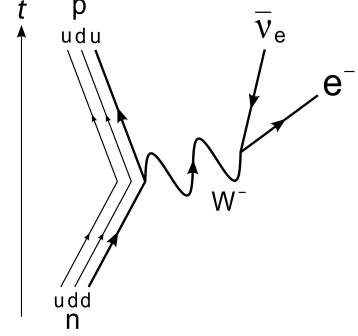


fig: 6 – Feynman diagram of the β^- decay

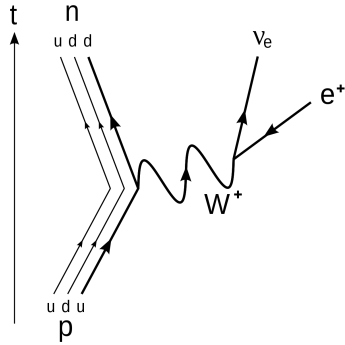


fig: 7 – Feynman diagram of the β^+ decay

Reaction and Phenomenology The β^- decay reaction is written as:

$${}^A_Z X_N \rightarrow {}^A_{Z+1} X'_{N-1} + e^- + \bar{\nu} \quad (14)$$

The underlying interaction as shown in figure 6 is:

$$n \rightarrow p + e^- + \bar{\nu} \quad (15)$$

The β^+ decay reaction however is:

$${}^A_Z X_N \rightarrow {}^A_{Z-1} X'_{N+1} + e^+ + \nu \quad (16)$$

with the underlying reaction as shown in figure 7:

$$p \rightarrow n + e^+ + \nu \quad (17)$$

Let us consider nuclei with equal mass number A we can write the SEMF as:

$$M(A, Z) = \alpha \cdot A - \beta \cdot Z + \gamma \cdot Z^2 + \delta(N, Z) \quad (18)$$

where

$$\alpha = M_N - a_V + a_S A^{-1/3} + \frac{a_A}{4}$$

$$\beta = a_A + a_C + M_N - M_P - m_e$$

$$\gamma = \frac{a_C}{A^{1/3}} + \frac{a_A}{A}$$

The smallest binding energy means the most stable nucleus as well. Therefore by taking the derivative and setting equal to zero we find the most stable atomic number Z :

$$\begin{aligned}\frac{dM}{dZ} &= 0 \\ \rightarrow -\beta + 2\gamma Z &= 0\end{aligned}$$

Thus the most stable atomic number Z for any given mass number is:

$$Z_{\text{most stable}} = \frac{\beta}{2\gamma} = \frac{a_A + a_C + M_N - M_P - m_e}{2\left(\frac{a_C}{A^{1/3}} + \frac{a_A}{A}\right)} = \frac{66.2787A}{130.462 + A^{2/3}} \quad (19)$$

Beta Decay in Odd Mass Nuclei Let's find the most stable atom with atomic mass $A = 101$ using eq[18]. The result is $Z = 44$. Isobars with more neutrons, such as $^{101}_{42}\text{Mo}$ and $^{101}_{43}\text{Tc}$, would have beta minus decay.



Callout — Energetically β^- decay is possible whenever the mass of the daughter atom $M(A, Z + 1)$ is smaller than the mass of its isobaric neighbors.

$$M(A, Z) > M(A, Z + 1)$$

We consider here the mass of the whole atom and not just that of the nucleus alone and so the rest mass of the electron created in the decay is automatically taken into account

The reverse is possible for $^{101}_{45}\text{Rh}$ and $^{101}_{46}\text{Pd}$



Callout — Since the mass of a free neutron is larger than the proton mass, this process is only possible inside a nucleus. By contrast, neutrons outside nuclei can and do decay (They would do a β^- decay since neutrons are not stable particles.).

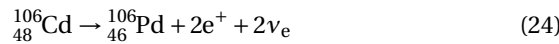
Energetically, beta plus decay is possible whenever the following relationship between masses is satisfied:

$$M(A, Z) > M(A, Z - 1) + 2m_e$$

This relation takes into account the creation of a positron and the existence of an excess electron in the parent atom.

Beta Decay in even mass Nuclei: For even mass nuclei there exist two parabolas, one for the odd-odd nuclei and one for the even-even nuclei. Often there's more than one β -stable isobar, especially in the range $A > 70$. For the case of $A = 106$ in figure 8. There exists two parabolas, but also to β -stable atoms.

$^{106}_{46}\text{Pd}$ and $^{106}_{48}\text{Cd}$ isobars are on the lower parabolas, and $^{106}_{46}\text{Pd}$ is the stablest. $^{106}_{48}\text{Cd}$ is β -stable, since its two odd-odd neighbours both lie above it. The conversion of $^{106}_{48}\text{Cd}$ is thus only possible through a double β -decay into $^{106}_{46}\text{Pd}$



The probability for such process is so small that $^{106}_{48}\text{Cd}$ may be considered to be a stable nuclei. Odd-Odd nuclei always have at least one more strongly bound even-even neighbour. Thus they are unstable. The only exceptions are light nuclei such as ^2_1H , ^6_3Li , $^{10}_5\text{B}$ and $^{14}_7\text{N}$, which are stable to β -decay.

Callout — Some odd-odd nuclei can undergo both beta minus and beta plus decays. for example $^{40}_{19}\text{K}$ and $^{64}_{29}\text{Cu}$.

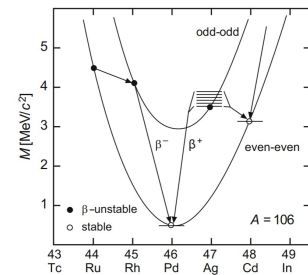


fig: 8 – The Mass parabola of $A = 106$ which shows two parabola because of even-even and odd-odd parity term.

3.2 Electron Capture

Electron capture or K-electron capture is a process in which the nuclei would transform one proton into a neutron by capturing one of atoms electrons. The process is as below:



This reaction occurs mainly in heavy nuclei where the nuclear radii are larger and the electron orbits are more compact. Usually the electrons that are captured are from the innermost (the "K") shell since such electrons are closest to the nucleus and their radial wave function has a maximum at the center of the nucleus.

Electron-capture reactions compete with β^+ -decay. The following condition is a consequence of energy conversion.

$$M(A, Z) > M(A, Z - 1) + \epsilon \quad (26)$$

3.3 Alpha Decay

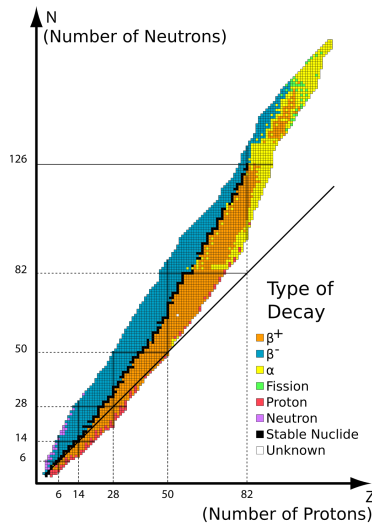


fig: 9 – Graph of isotopes by type of nuclear decay.

Orange and blue nuclides are unstable, with the black squares between these regions representing stable nuclides. The unbroken line passing below many of the nuclides represents the theoretical position on the graph of nuclides for which proton number is the same as neutron number. The graph shows that elements with more than 20 protons must have more neutrons than protons, in order to be stable.