Unit 3 Modeling Atomic Structure

Mr. Maxwell

PACS

February 3, 2025

- Atomic Structure
 - Atomic Number
 - Mass Number
 - Hydrogen
 - Bohr Model
 - Energy Levels
 - The Periodic Table

Atomic Number

The is the number of

in the nucleus of an atom.

Atomic Number

The atomic number is the number of

in the nucleus of an atom.

Atomic Number

The atomic number is the number of protons in the nucleus of an atom.

The the total number of and in the nucleus of an atom.

The mass number the total number of of an atom.

and

in the nucleus

The mass number the total number of protons and of an atom.

in the nucleus

The mass number the total number of protons and neutrons in the nucleus of an atom.

Hydrogen

 ^{1}H

Hydrogen

 $^{1}\mathrm{H}$

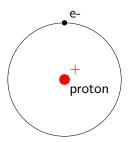
What does the 1 mean?

Hydrogen

 $^{1}\mathrm{H}$

What does the 1 mean?

1 is the total number of neutrons and protons.



 ${}^4_2{
m He}$



 $_2^4\mathrm{He}$

What does the 4 mean?



 ${}_{2}^{4}\mathrm{He}$

What does the 4 mean?

4 is the total number of neutrons and protons.

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What does the 2 mean?

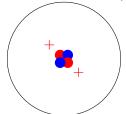
 ${}_{2}^{4}\mathrm{He}$

What does the 4 mean?

4 is the total number of neutrons and protons.

What does the 2 mean?

2 is the number of protons.



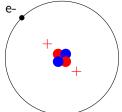
 ${}^4_2{
m He}$

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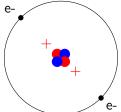
 ${}^4_2{\rm He}$

What does the 4 mean?

4 is the total number of neutrons and protons.

What does the 2 mean?

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 $^{7}_{3}\mathrm{Li}$

How many protons does Lithium have?

 $^{7}_{3}\mathrm{Li}$

How many protons does Lithium have? 3





 $^{7}_{3}\mathrm{Li}$



$$_3^7 \mathrm{Li}$$





 $^7_3\mathrm{Li}$

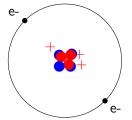


 $^7_3\mathrm{Li}$

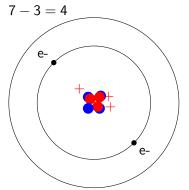
 $^{7}_{3}\mathrm{Li}$

 $^{7}_{3}\mathrm{Li}$

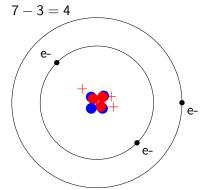
$$7 - 3 = 4$$



 $^7_3\mathrm{Li}$



$$^{7}_{3}\mathrm{Li}$$



Niels Bohr



The Bohr Model - Bohr proposed that an atom was a nucleus with electrons "orbiting" in different

Niels Bohr



The Bohr Model - Bohr proposed that an atom was a nucleus with electrons "orbiting" in different energy levels.

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The electrons closest to the nucleus have the further from away have energy.

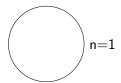
energy, while those

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The electrons closest to the nucleus have the lowest energy, while those further from away have higher energy.

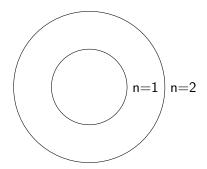
Energy Levels

The electrons closest to the nucleus have the lowest energy, while those further from away have higher energy.



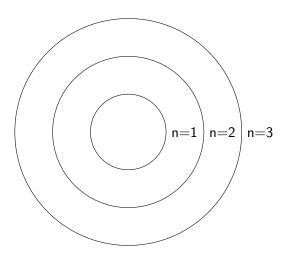
Energy Levels

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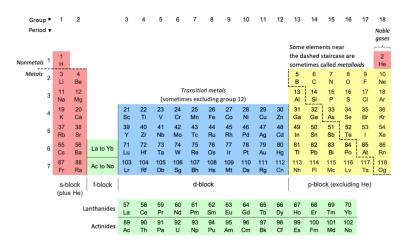


Energy Levels

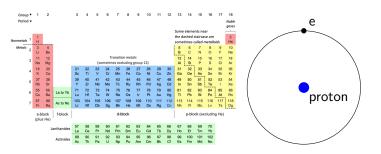
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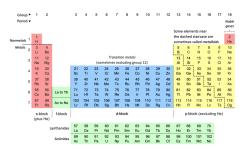


Energy Levels and the Periodic Table



Energy Level of Hydrogen













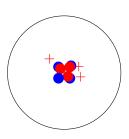




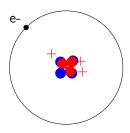


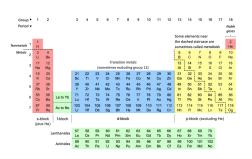


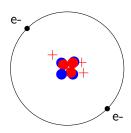


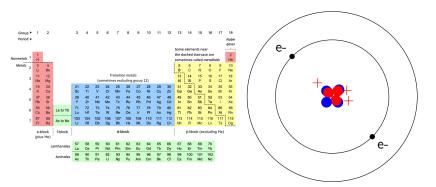


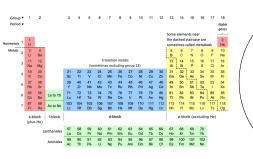


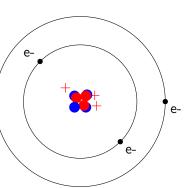


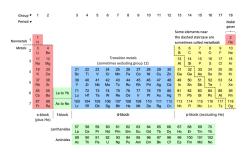


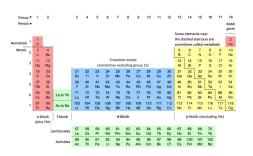




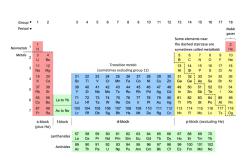


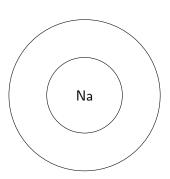


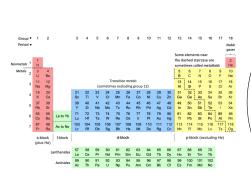


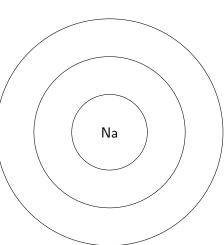




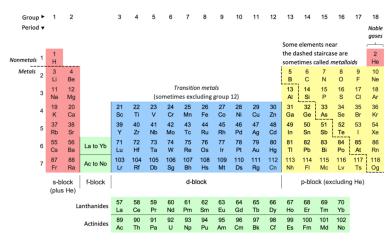






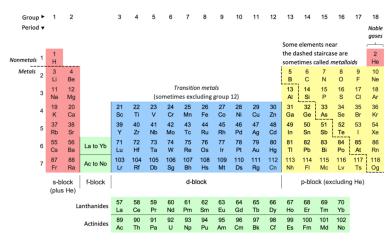


The Periodic Table



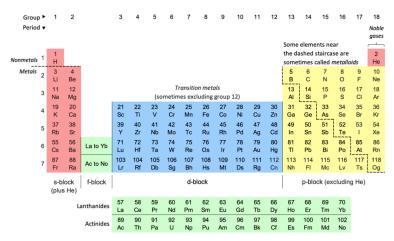
The Periodic Table has periods and groups.

The Periodic Table



The Periodic Table has 7 periods and groups.

The Periodic Table



The Periodic Table has 7 periods and 18 groups.