Solutions Lesson 1 Check for Understanding

Part ONE: **Terminology** - Define the following terms

a. Solubility:

b. Soluble:

c. Insoluble:

d. Miscible:

e. Immiscible :

f. Electrolyte:

g. Nonelectrolyte:

Part TWO: **Determining Solubility:**

1. Name or give the chemical formula for each of the following compounds.
2. State whether they are soluble (will dissolve) or insoluble (will not dissolve) in solution. Use the Solubility Rules from your Reference Table.

Chemical Formula	Name	Solubility
$Ba(OH)_2$		
	Iron (II) Carbonate	
NaOH		
$RbNO_3$		
	Cesium Sulfate	
$MgSO_4$		
Hg_2SO_4		

Part THREE:**Electrolytes:**

Classify the following compounds:

1st : Determine the Type of compound *(Ionic/Salt/Acid/Base/Covalent)* - ***State all that apply*** -

2nd : Based on your answer to step 1, decide if an electrolyte or nonelectrolyte

3rd: If an electrolyte and ionic, show its dissociation (breaking up) in water

Compound	Ionic / Salt / Acid / Base / Covalent	Electrolyte	Nonelectrolyte	Dissociation
NaCl	Ionic & Salt	✔		$\rightarrow Na^{1+} + Cl^{1-}$
$MgSO_4$				
$C_3H_5(OH)_3$				
HCl				
$C_6H_{12}O_6$				
NaOH				
NH_3				