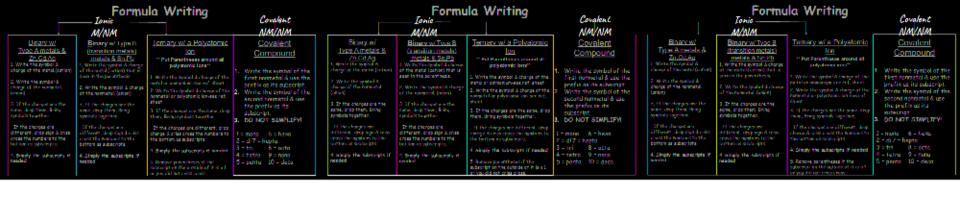
	Ionic Na M/NM	ming	Covalent NM/NM		Ionie No M/NM	aming	Covalent NM/NM		Tonic No M/VM	aming	Covalent NM/NM
element (metal) cation  2. Change the ending of the 2nd element	Type 4 but have provertisees beside the metal name to show the charge and a roman hammer. I have been been been been been been been be	as Type A, <u>BUT</u> look up the name of the pulyatumic ion on the reference sheet. If Type B metal don't forget (Roman Numeral) of metal's charge		Risary w/ Israe Ametais & Zn.Gd.An  I. Write the nome of 1st element (metal) -cation  7. Change the ending of the 2nd element (nometal) -anion to *-ide*	Binary will Tube B Ithanistion mistakin	Ternary will a Polystomic Ion Do the exact sone thing as type A, BUT look up the name of the polystomic ion on the reference sheet.  If Type is metal don't forget (Bunan Nameral) of metal's change "Nate amonorium NI-1," is treated like a metal so it may come first in the compound "NI-1," in th	Binary Covalant Compound Use profiles to represent the substraint # unit he that FXEPT — In NOT use MONO on the flight name of left nomental Z. Write the profix and name of the nomental with the LIDE sading 1 - nome   2 - not 1 - not 2 - not 3 -	Bloom will Type Americals & Zondo Ag  1. Write the name of 1st element (metal) -cution  2. Change the ending of the 2nd element (metal) -union to "-ide"	metals & Sin.Eb  Do the exoct searching as Type A but leave per eithered beside the metal frame to show the shorty as a nation remains Sink Way  L. Uharross the automorph.  2. 2º the exist charge is convert, then got unreast to find the metals charge 3. 2º the arisen charge has been mediated (an incorrect) their metals charge 3. 2º the arisen charge has been mediated (an incorrect) their ministry both is subscripted by the	Iernary w/ at Holyatomic Ion to the swort some thing as Typic A, BLT link up the name of the polyatomic ion on the rester since the same of the polyatomic ion on the rester since shared forget (Roman Nameral) of metal's charge ""Note ammonium NH <sub>4</sub> , is threated like a metal so it may come first in the compound."""	Binary Considert Composing Use prefers to represent the superies to represent the superies if on both nones EXCEPT - De NOT use MONIS on the first roome- 1 White the prefix and tenne of Set nomestal 2, White the prefix and future of the Zni instead with the IDE ending 1 - main 6 - 1 - 1 - 1 - 1 - 1 - 1 - 1 - 1 - 1 -



	6	congress:				Енатрко				Sumples	
	Iron (III) chloride	Magnoskim sulfato	Carbon monorade	Calcium bromido	hor(II)dilande	Magnesium suffide	Carbon monovido	Celdum bromide	Iron(II)chlorido	Magnesiani salate	Carbon monodde
Aluminum cuido	Coppor(!) oxido	Coball (II) Trydroude	Diansonic trioxide	Aluminum colde	Copper(i) adde	Cobalt (8) hydroxida	Diamenic Incolde	Alameian oxide	Copper(f) and	Coball (II) hydrocide	Diametric broade
Ultium fluoride	Critel(II) nilitie	Sodium nitrato	Safer howards	Litran feorde	Colorl(I) intride	Sortium nitrole	Suffur howards	Lithkon Suorido	Coboli(II) nitrido	Sodium nitrate	Safer hexcelde
Caldium nitride	Ylinen (I) wode	Ammonium chloride	Dinitrogon tetrahydrido	Calcium intaide	Yttrken (f) oxido	Ammonium chloride	Dinilrogen leitvityddde	Caldum nitride		Ammonium chlorido	Umbogen ktralijdisde
Massium sulfide	Tin(IV) sulfide	Ammonium carbonato	Othyrimgen monoxide	Potassium suffide	In(IV) salide	Ammonium carbonale	Dhydrogen monoada	Poloskim suilide	Tin(IV) sulfide	Annonen cabonile	Dihydragen monoxide

## Practice on your own using the U3 L9 CFU part TWO

You have already completed part ONE, now complete part TWO which is comprised of all naming and formula writing.

If you need the worksheet, click here:

https://docs.google.com/document/d/1TK-ts3\_DMeSYDmvTr3ukJkN-IvMfpbTAqiR

zxltdbFE/edit?usp=sharing

## Part TWO: Mixed Review

Determine whether the substance is ionic type A (I -A) ionic type B (I -B)or covalent (molecular) (C). If Ionic, you will need to decide whether you need to put a roman numeral in the name and always check charges in the formula. If Covalent, no need to use roman numerals in the name and not need to check charges in the formula.

I/C	Pro	vide the chemical name:	I/C	Provide the chemical formula:
	1. CuO			26. Phosphorus trichloride
	2. SrO			27. Chlorine monofluoride
	3. $B_2O_3$			28. Copper(II) chloride
	4. TiCl <sub>4</sub>			29. Copper(I) sulfide
	5. K <sub>2</sub> S			30. Calcium nitride
	6. OF <sub>2</sub>			31. Carbon tetrabromide
	7. NH <sub>3</sub>			32. Lithium oxide
	8. VF <sub>5</sub>			33. Potassium chloride
	9. CuCl			34. Titanium(IV) bromide
	$10.~\mathrm{MnO}_2$			35. Magnesium sulfide

