

### **U3 Lesson 3 Check for Understanding**

#### **Naming & FW for Covalent Compounds**

**Part ONE: Covalent Compounds** *(No roman numerals in name and no need to check charges for formula)*

Provide the chemical name:

1. CO \_\_\_\_\_
2. CO<sub>2</sub> \_\_\_\_\_
3. NO \_\_\_\_\_
4. NO<sub>2</sub> \_\_\_\_\_
5. SF<sub>6</sub> \_\_\_\_\_
6. SiF<sub>4</sub> \_\_\_\_\_
7. N<sub>2</sub>S<sub>3</sub> \_\_\_\_\_
8. B<sub>2</sub>H<sub>6</sub> \_\_\_\_\_
9. SO<sub>2</sub> \_\_\_\_\_
10. CH<sub>4</sub> \_\_\_\_\_
11. SO<sub>3</sub> \_\_\_\_\_
12. N<sub>2</sub>O<sub>4</sub> \_\_\_\_\_

Provide the chemical formula:

13. Boron trichloride \_\_\_\_\_
14. Nitrogen monoxide \_\_\_\_\_
15. Dinitrogen monoxide \_\_\_\_\_
16. Dinitrogen pentoxide \_\_\_\_\_
17. Sulfur hexachloride \_\_\_\_\_
18. Carbon monoxide \_\_\_\_\_
19. Carbon disulfide \_\_\_\_\_
20. Oxygen difluoride \_\_\_\_\_
21. Dinitrogen tetrahydride \_\_\_\_\_
22. Silicon tetrahydride \_\_\_\_\_
23. carbon tetrachloride \_\_\_\_\_
24. trisulfur nonabromide \_\_\_\_\_

### U3 Lesson 4 Check for Understanding

#### Mixed Practice

#### Part TWO: Mixed Review

Determine whether the substance is ionic type A (I -A) ionic type B (I -B) or covalent (molecular) (C). If Ionic, you will need to decide whether you need to put a roman numeral in the name and always check charges in the formula. If Covalent, no need to use roman numerals in the name and not need to check charges in the formula.

I /C	Provide the chemical name:	I /C	Provide the chemical formula:
_____ 1. CuO	_____	_____ 26. Phosphorus trichloride	_____
_____ 2. SrO	_____	_____ 27. Chlorine monofluoride	_____
_____ 3. B <sub>2</sub> O <sub>3</sub>	_____	_____ 28. Copper(II) chloride	_____
_____ 4. TiCl <sub>4</sub>	_____	_____ 29. Copper(I) sulfide	_____
_____ 5. K <sub>2</sub> S	_____	_____ 30. Calcium nitride	_____
_____ 6. OF <sub>2</sub>	_____	_____ 31. Carbon tetrabromide	_____
_____ 7. NH <sub>3</sub>	_____	_____ 32. Lithium oxide	_____
_____ 8. VF <sub>5</sub>	_____	_____ 33. Potassium chloride	_____
_____ 9. CuCl	_____	_____ 34. Titanium(IV) bromide	_____
_____ 10. MnO <sub>2</sub>	_____	_____ 35. Magnesium sulfide	_____
_____ 11. MgO	_____	_____ 36. Manganese(II) nitride	_____
_____ 12. B <sub>2</sub> H <sub>6</sub>	_____	_____ 37. Calcium bromate	_____
_____ 13. Li <sub>2</sub> Te	_____	_____ 38. Sodium chloride	_____
_____ 14. Fe(NO <sub>3</sub> ) <sub>3</sub>	_____	_____ 39. Trinitrogen dioxide	_____
_____ 15. CaSO <sub>4</sub>	_____	_____ 40. Lithium phosphate	_____
_____ 16. NaCl	_____	_____ 41. Ammonium chloride	_____
_____ 17. K <sub>2</sub> SO <sub>4</sub>	_____	_____ 42. Copper(II) chlorite	_____
_____ 18. CO <sub>2</sub>	_____	_____ 43. Nitrogen monoxide	_____
_____ 19. SF <sub>6</sub>	_____	_____ 44. Iron(II) iodide	_____
_____ 20. KClO	_____	_____ 45. Calcium phosphate	_____
_____ 21. N <sub>2</sub> O <sub>5</sub>	_____	_____ 46. Dinitrogen dioxide	_____
_____ 22. IF <sub>5</sub>	_____	_____ 47. Magnesium oxide	_____
_____ 23. Co(MnO <sub>4</sub> ) <sub>2</sub>	_____	_____ 48. Iron(III) chromate	_____
_____ 24. Sn(SO <sub>4</sub> ) <sub>2</sub>	_____	_____ 49. Sulfur dioxide	_____
_____ 25. FrCl	_____	_____ 50. Aluminum iodate	_____