

# SPECTROSCOPY

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*Every honest researcher I know admits he's just a professional amateur.  
He's doing whatever he's doing for the first time. That makes him an  
amateur. He has sense enough to know that he's going to have a lot of  
trouble, so that makes him a professional.*

— Charles F. Kettering (1876-1958) (Holder of 186 patents)

## ACKNOWLEDGEMENTS

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I express my sincere gratitude to our instructors, Dr. K. S. Viswanathan and Dr. K. R. Shamasundar, for bringing the subject to life and helping us discover, in depth, the science behind the procedures.

I also thank Vivek Sagar (MS11017) for his contribution to this report as my lab-partner, who made the task of performing experiments immensely comfortable and productive at the same time.



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## Part I

### EXPERIMENTS





## INTRODUCTION TO INFRA RED AND ULTRAVIOLET - VISIBLE SPECTROSCOPY

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August 9, 2012

### 1.1 OBJECTIVE

To

1. understand the method of Fourier Transformed Infrared and Ultraviolet spectroscopy
2. understand the preparation substances for analysis
3. handle and operate the Spectroscopes provided

### 1.2 THEORY

#### 1.2.1 Basic Concept

To find the presence of elements and/or compounds within a given substance, we can use spectroscopy techniques, specially when their concentrations are small and they satisfy certain requirements. The essential idea behind this measurement comes from the fact that elements/compounds absorb lights of certain frequencies to get to a higher energy state. These frequencies are mostly discrete as they correspond to quantized energy levels. This energy could be absorbed for, say, changing the vibrational energy (IR Spectroscopy) or for exciting an electron in the substance to a higher energy level (UV-vis Spectroscopy). We note here that these quantized energy levels are properties of individual substances and are, for most practical purposes, unique.

For the analysis to be possible, the first condition is that the substance must *absorb* light incident to it. Granted this, we can obtain an absorption spectrum for the given substance, which behaves like a fingerprint of the substance. This can thus be used to not only identify the compound, but also to quantify it. For identification, in the simplest case, we simply need to observe the frequency corresponding to the peaks in the absorption spectrum and match it with the known/expected substance(s). Quantification harnesses a rather “obvious” law, termed *Beer-Lambert’s Law*. In the simplest form, the law



How much absorption, well, the limit comes from the sensitivity of the experimental setup and concentration of substance given.

quantifies the intuitive notion; higher the concentration of the analyte, higher is the absorption. The relation is given as

$$T = \frac{I}{I_0} = 10^{-\alpha l} = 10^{-\epsilon lc} \quad (1)$$

where  $I$  is intensity of incident light,  $I_0$  is intensity of transmitted light,  $\epsilon$  is molar absorptivity,  $l$  is the optical path length, and  $c$  is molar concentration.

### 1.2.2 Infra Red Spectroscopy

Infra Red spectroscopy usually deals with energies of the level that cause change in vibrational energies. The wavelength ranges from  $2.5\mu\text{m}$  to  $25\mu\text{m}$  ( $4000\text{cm}^{-1}$  to  $400\text{cm}^{-1}$ ). These energies are characteristic for different bonds which is how, using the spectrum, we can identify (and quantify) the bonds present and thus the compound.

The way the spectroscope works for Infra Red, is rather interesting and ingenious. The setup uses a Michelson interferometer.

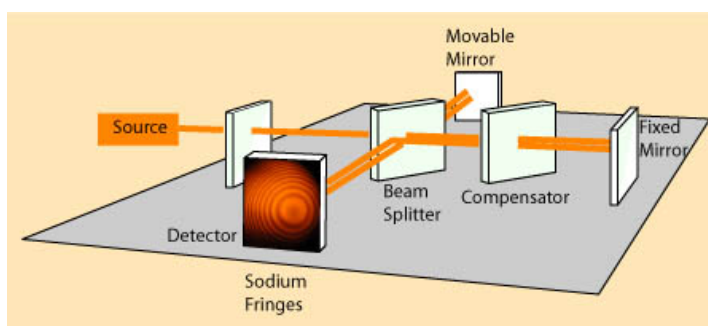




Figure 1: Michelson Interferometer [? ]

 Coherent means phase locked and Monochromatic means single frequency

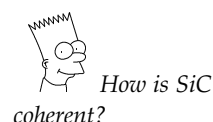
Before complicating things, let us assume that the source of light is coherent and monochromatic. Now in the [Figure 1](#), we assume the source of light to be the light transmitted through the sample to be analysed. Say the interference at the detector is, at the given configuration, constructive. If we move the movable mirror by  $\lambda/4$ , where  $\lambda$  is the wavelength of the light, then the detector will receive a dark, destructive interference. If we plot the intensity at the detector as a function of displacement of the moveable mirror, we will, in this case, receive a sine wave.

 Note the fact that here, the measurement is simultaneous!

Now let us crank it up a notch. Let us consider the light to still be coherent, but not monochromatic. Let the source of light contain all the transmitted frequencies. Now if the intensity is plotted against the displacement of the moveable mirror, we will get a superposition of sine waves, and we already know how to decompose them to find individual frequencies using *Fourier Transformation*. This essentially

gives us the spectrum, with wavenumber (dimensions of one over distance) on one axis, and intensity on the other.

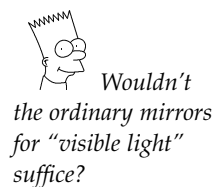
The source of infra-red light is a Silicon Carbide rod, that is heated to produce the desired radiations. The detector is capacitive, or so we were told.



### 1.2.3 Ultraviolet-Visible Spectroscopy

Here the wavelengths range from 200nm to 900nm. This range is often in the range of differences of electronic energy levels. The light source is a Tungsten-Neon lamp. In this technique, usually movable diffraction grating is used to split the transmitted light into its component frequencies, "one" of which goes to an intensity detector. Which component is detected depends on the orientation of the diffraction grating. Using this information, an intensity vs. frequency plot is generated which is the desired spectrum.

It must be noted here that the process of moving the orientation of the grating, requires time, which means that the measurement for various frequencies' intensities is not simultaneous, which in certain cases might not be suitable. One solution to such a problem is the use of multiple (arrayed) detectors. Yet it doesn't explain why a Michelson interferometer can't be used in this case. The answer to this comes from the fact that the wavelength in this case is much smaller, thereby complicating the task for preparing reflectors.



## 1.3 EXPERIMENTAL DETAILS

For this session, both IR and UV-visible spectroscopy techniques were demonstrated to us in groups of two. For IR spectroscopy, we used KBr as the base and Benzoic Acid as the analyte. For UV-vis, we used water as the base and  $\text{KMnO}_4$  as the analyte.

### 1.3.1 Infrared

#### *Procedure*

The general procedure is as follows:

- A. Weighed 100mg of KBr and ground it well.
- B. Transferred it the dye (palletizing module).
- C. Converted it into a pellet using the Hydraulic Press (about 8 tons of pressure).
- D. Did IR spectroscopy.
- E. Weighed 100mg of KBr and with it, 2mg of sample (in this case Benzoid Acid) and repeated from [item B](#).

Details about the IR spectroscopy are as follows:

- A. Ran the background to get rid of noise.
- B. Did the accumulation 16 times (set the option).
- C. Subtracted the first spectroscopy from the second

*Observation*

Detected the carboxylic group (peaks for OH and C = O)

1.3.2 UV-Vis

*Procedure*

The general procedure is as follows:

- A. Took a suitable amount of water.
- B. Did UV-Vis spectroscopy.
- C. Again took water and added a suitable amount of  $\text{KMnO}_4$  and mixed it well.
- D. Did UV-Vis spectroscopy.

Details about the UV-vis spectroscopy is corresponding to the previous section.

*Observation*

Peaks were detected in the range 520nm-550nm, complimentary to Purple as expected.

1.4 ACKNOWLEDGEMENTS

I thank the PhD students, Ms. Shruti and Ms. Shilpa, who helped us with the session and performed the experiment. They also walked us through the procedure for using the apparatus.

## Part II

### THE SHOWCASE

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